Additional Information

2.2.1 The Institution assesses the learning levels of the students and organizes special programme for Advanced and Slow learners

Details of Enclosures

I. Representative Programmes conducted for Slow Learners

II. Representative Programmes conducted for Advanced Learners

III. Other Programmes

A list of representative activities conducted for slow learners is shown below

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Slow learners

Students with varied degree of learning capabilities and diversified back ground get admission in the college. Therefore the level of challenge faced by them is also different. The college is empathetic to this aspect and takes maximum steps to optimize the learning outcomes of the students. In this direction the college has undertaken a proactive step of identifying slow learners and advanced learners.

Students who scored 40% or <40% marks in the university examinations are categorized as slow learners and those scored >70% are advanced learners.

Department of chemistry

A sample copy of slow learners and advanced learners belonging to chemistry department is presented here. Complete details are available in college website.

B.Sc: Batch:2016-2019

Programme code	Semester wise number of slow learners					
	SEMESTER -II	SEMESTER-III	SEMESTER -IV	SEMESTER-V		SEMESTER-VI
				Paper-5	Paper-6	
MPC -E.M&TM	14	17	14	6	1	3
BBC-E.M	3	11	12	NIL	2	3
MZC-E.M	3	7	5	4	2	NIL
CBZ-TM	8	7	8	8	6	4
Total	28	42	39	18	11	10

B.Sc: Batch:2017-2020

Programme code	Semester wise number of slow learners					
	SEMESTER-II	SEMESTER-III	SEMESTER-IV	SEMESTER-V		SEMESTER-VI
				Paper-5	Paper-6	
MPC -E.M&TM	5	5	4	NIL	1	NIL
BBC-E.M	4	3	3	3	4	2
MZC-E.M	11	9	4	2	2	NIL
CBZ-TM	6	7	5	NIL	NIL	NIL
Total	26	24	16	5	7	2

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

BATCH-2016-2019

Programme code	SEMESTER WISE NUMBER FOR SLOW LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	SEMESTER-III (BASED ON II SEMESTER)	SEMESTER-IV (BASED ON III SEMESTER)	SEMESTER-V (BASED ON IV SEMESTER)	SEMESTER-VI (BASED ON V SEMESTER) PAPER-5	SEMESTER-VI (BASED ON V SEMESTER) PAPER-6
MPCS	2	3	2	1	1	NIL
MSCS	9	6	4	NIL	NIL	NIL
Total	11	9	6	1	1	0

BATCH-2017-2020

Programme code	SEMESTER WISE NUMBER FOR SLOW LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	SEMESTER-III (BASED ON II SEMESTER)	SEMESTER-IV (BASED ON III SEMESTER)	SEMESTER-V (BASED ON IV SEMESTER)	SEMESTER-VI (BASED ON V SEMESTER) PAPER-5	SEMESTER-VI (BASED ON V SEMESTER) PAPER-6
MPCS	NIL	NIL	NIL	NIL	NIL	NIL
MSCS	1	1	3	NIL	NIL	NIL
Total	1	1	3	0	0	0

BATCH-2018-2021

Programme code	SEMESTER WISE NUMBER FOR SLOW LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	SEMESTER-III (BASED ON II SEMESTER)	SEMESTER-IV (BASED ON III SEMESTER)	SEMESTER-V (BASED ON IV SEMESTER)	SEMESTER-VI (BASED ON V SEMESTER) PAPER-5	SEMESTER-VI (BASED ON V SEMESTER) PAPER-6
MPCS	NIL	NIL	NIL	1	NIL	NIL
MSCS	NIL	NIL	1	NIL	NIL	NIL
Total	0	0	1	1	0	0

1. Remedial Coaching

Slow learners require additional encouragement and support for better academic performance. The college regularly conducts remedial coaching for such students. This led to improved performance of the students in the university examinations.

Special programmes conducted for slow learners

Phone: (0877) 2264599 E-mail: sgsartscollegettd@gmail.com Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

ODD SEMESTERS

CIRCULAR

Date: 11-06-2019

All the HODs of the department are here by instructed to plan remedial classes for slow learners of III & V semesters in order to improve their learning capability and for further improvement in the forthcoming university examinations. The Time –Table for the conduct of remedial classes for slow learners is enclosed. The classes may be conducted out of the college hours i.e. before 10.00A.M / after4.00P.M

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE Co-ordinator NAAC TIRUPATI

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

NOTICE

Date: 14-06-2019

SGSAC/Chem dept/Remedial classes- 2019/HOD/01

All the slow learners in chemistry of III&V semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 17-06-2019 in Room numbers 225& 230 from 4.00Pm to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.

ACADEMIC YEAR 2019-20

II B.Sc, IIISEMESTER

S.No	HALL Ticket Number	Student Name	GROUP& Medium
1	NIL	NIL	MPC EM & TM
1	319008007	T. Goutham	MZC-EM
2	319008026	S.Soma Sekar	
1	318008054	G.Suman	CBZ-TM
2	318008050	P.SAI	
3	318008038	S. Balaji Naik	
4	318008055	E.G. Suri	
5	319008259	P.Narasimha	BBC-EM
6	319008269	D.Sasi Kumar	
7	319008248	M.Chandra	
8	319008242	S.Choudaiah	
9	319008254	S.Manjunadh	
10	319008261	M.Naveen	
11	319008280	T,Vijay	
12	319008247	K.Jagadeswara	
13	319008250	T.Jayaram	

14	319008257	V.Nagendra Prasad	
15	319008246	B.Hari Babu	
16	319008262	A.Pedda Obulesu	
17	319008274	T.SREEKANTH REDDY	
III BSC SEMESTER-V PAPER -5			
1	318008159	M.Kusuma Priya	BBC-EM
2	318008146	S.Gangadhar	
3	318008163	G,.Mahesh	
4	318008284	T.Akash	MZC-EM
5	318008265	K.Vamsi	
1	NIL	NIL	CBZ-TM
1	NIL	NIL	MPC-EM & TM
III BSC SEMESTER-V PAPER -6			
1	318008161	C.Madhu Mohan	BBC-EM
2	318008160	C.Lokesh	
3	318008159	M.Kusuma Priya	
4	318008163	G,.Mahesh	
5	318008284	T.Akash	MZC-EM
6	318008283	M.Sreedhar	
7	318008201	Chittaturu Gnana Sekhar	MPC-EM&TM
1	NIL	NIL	CBZ-TM

H.O.D of the Department

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

To

1. All teaching staff of chemistry
2. Department for circulation.
3. Department notice board

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

S.G.S ARTS COLLEGE,TIRUCHANOORROAD,TIRUPATI

Institutional Time -Table for Remedial classes (Odd semester)

Programme name : B.A, B.Sc, B.Com SEMESTER : III & V

TIME: 4.00PM-5.30PM **YEAR : II & III**

DAY	SUBJECT	SEMESTER	ROOM NO
MONDAY	CHEMISTRY	III	225
		V	230
	STATISTICS	III	221
		V	222
	COMPUTER SCIENCE	III	115
		V	116
	POLITICS	III	211
		V	212
	COMMERCE	III	307
		V	308
TUESDAY	CHEMISTRY	III	225
		V	230
	STATISTICS	III	221
		V	222
	COMPUTER SCIENCE	III	115
		V	116

	POLITICS	III	211
		V	212
	COMMERCE	III	307
		V	308
WEDNESDAY	MATHEMATICS	III	225
		V	230
	ZOOLOGY	III	Z LAB-2
		V	129
	BIOTECHNOLOGY	III	315
		V	316
	HISTORY	III	211
		V	212
	ACCOUNTS	III	307
		V	308
THURSDAY	MATHEMATICS	III	225
		V	230
	ZOOLOGY	III	Z LAB-2
		V	129
	BIOTECHNOLOGY	III	315
		V	316
	HISTORY	III	211
		V	212
	ACCOUNTS	III	307
		V	308
FRIDAY	PHYSICS	III	225
		V	230
	BOTONY	III	B LAB-2
		V	129

	MICROBIOLOGY	III	314
		V	315
	ECONOMICS	III	211
		V	212
	TELUGU	III	204
		V	205
SATURDAY	PHYSICS	III	225
		V	230
	BOTONY	III	B LAB-2
		V	129
	MICROBIOLOGY	III	314
		V	315
	ECONOMICS	III	211
		V	212
	TELUGU	III	205
		V	206

S.G.S. ARTS COLLEGE ★ TIRUPATI ★ Co-ordinator NAAC

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S.G.S. ARTS COLLEGE
TIRUPATI

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc., Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2016-17			
SEMESTERS - III & V		**DATE : 27/06/2016 to 31/09/2016**	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

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S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2017-18			
SEMESTERS - III & V		**DATE : 19/06/2017 to 03/10/2017**	
Name of the Lecturer	**MONDAY**	**TUESDAY**	**TIME**
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19			
SEMESTERS - III & V		**DATE : 11/6/2018 to 11/9/2018**	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
Dr. V. Venkata Lakshmi	III - SEMESTER	III - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - III & V		**DATE : 17/06/2019 to 29/10/2019**	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
Dr. V. Venkata Lakshmi	III - SEMSTER	III - SEMSTER	4.00 TO 5.30
Dr. P. SUGUNA	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

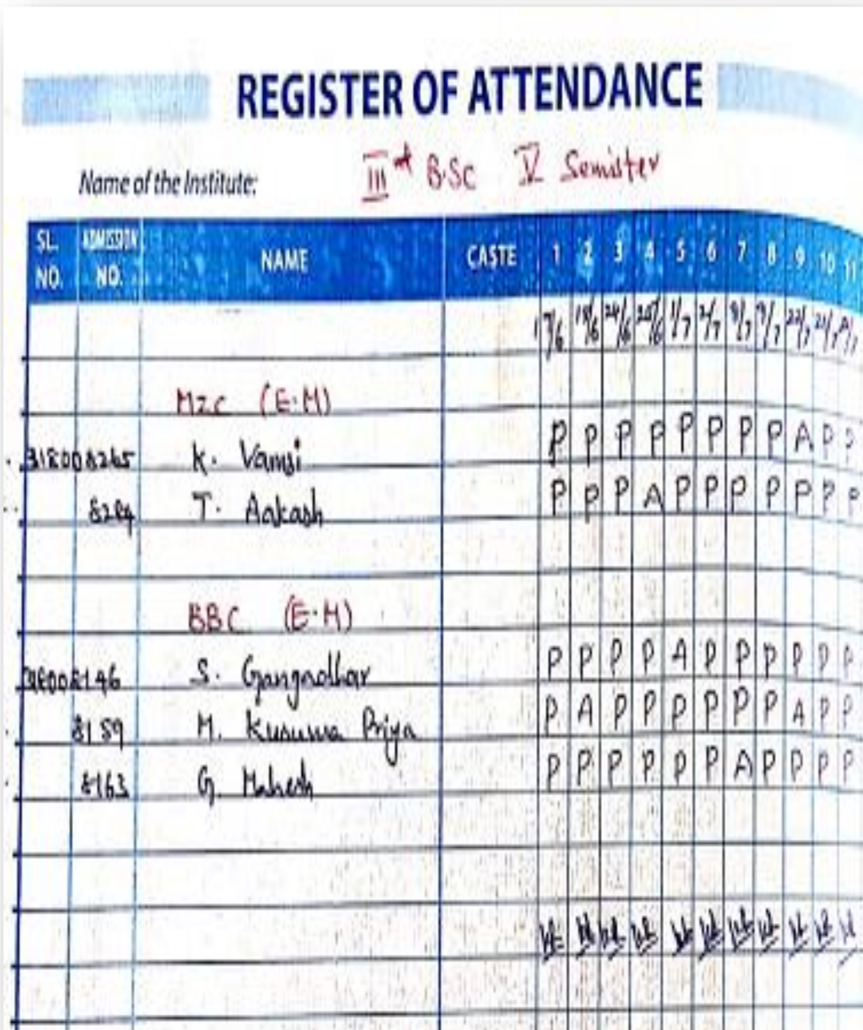
TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with chemistry combination

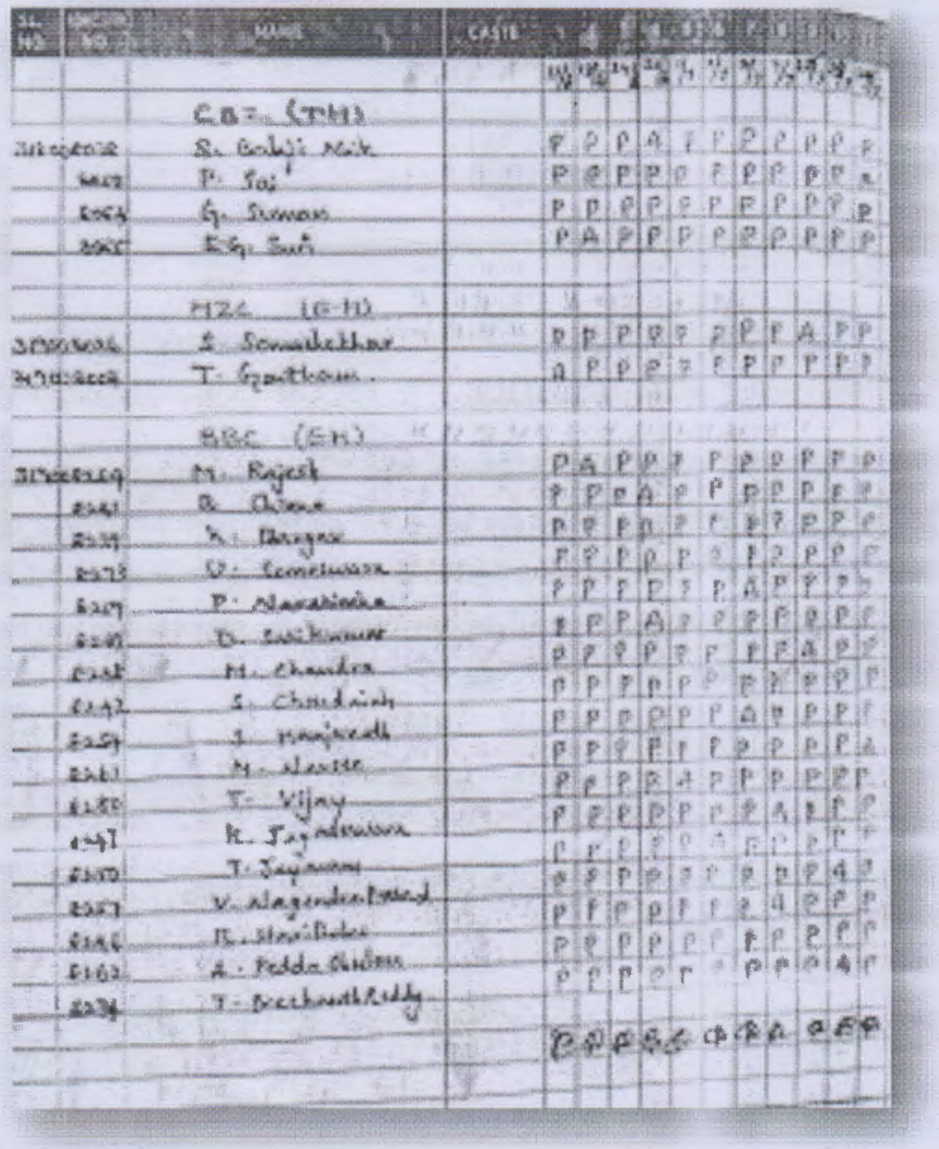
Semester: **III&V** **Date:** **17-06-2019 To29-10-2019**



REGISTER OF ATTENDANCE

Name of the Institute: III rd B.Sc V Semister

SL. NO.	ADMISSION NO.	NAME	CASTE	1	2	3	4	5	6	7	8	9	10	11
				17/6	18/6	24/6	26/6	1/7	2/7	9/7	[illegible]	[illegible]	[illegible]	[illegible]
		MZC (E.M)												
	3180008265	K. Vamsi		P	P	P	P	P	P	P	P	A	P	P
	8264	T. Aakash		P	P	P	A	P	P	P	P	P	P	P
		BBC (E.M)												
	[illegible]8146	S. Gangadhar		P	P	P	P	A	P	P	P	P	P	P
	8159	M. Kusuma Priya		P	A	P	P	P	P	P	P	A	P	P
	8163	G. Mahesh		P	P	P	P	P	P	A	P	P	P	P



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TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

EVEN SEMESTERS

CIRCULAR

Date: 02-12-2019

All the HODs of the department are here by instructed to plan remedial classes for slow learners of II , IV & VI semesters in order to improve their learning capability and for further improvement in the forthcoming university examinations. The Time –Table for the conduct of remedial classes for slow learners is enclosed. The classes may be conducted out of the college hours i.e. before 10.00A.M / after 4.00 P.M

S.G.S. ARTS COLLEGE ★ TIRUPATI ★ Co-ordinator NAAC

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TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
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HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

NOTICE

Date: 06-12-2019

SGSAC/Chem dept/Remedial classes- 2019/HOD/02

All the slow learners in chemistry of II, IV&VI semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 09-12-2019 in Room numbers 224,225 & 230 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory for the students mentioned below. Stringent action will be taken against the students who will be absent for the remedial classes.

List of Slow Learners

LIST OF SLOW LEARNERS - 2019-2020			
I B.Sc. , II SEMESTER			
S.No	Hall Ticket No.	Name of the student	Group & Medium
1	320008027	A. Venakatarjun	MPC (EM&TM)
2	320008028	A. Chandu Prakash	
3	320008002	B. Dinesh Naik	CBZ (TM)
4	320008003	C. Lakshmi	
5	320008004	C. Harinath	
6	320008006	C. Chandra	
7	320008008	D. Bala	
8	320008009	D. Munilatha	
9	320008010	E. Sree	

10	320008011	Gowrayya	
11	320008012	G. Rajasekhar	
12	320008014	K. Usha	
13	320008015	K. Swathi	
14	320008017	K. Madhu	
15	320008018	K. Naveen Naik	
16	320008020	M. Hareesh	
17	320008212	A. Giri Ramesh	MZC (EM)
18	320008249	S. Reena	
19	320008261	B. Govardhan	BBC (EM)
20	320008262	V. Jhansi Laskhmi	
21	320008265	G. Ramakanth	
22	320008286	M. Siva Kumar	
23	320008289	P. Sunil	
24	320008293	T. Umamaheswar	
25	320008294	S. Muni Sindhuja	
26	320008297	S.R.P. Jenny	
27	320008302	V. Nethesh Kumar	
II B.Sc. , IV SEMESTER			
S.no	Hall Ticket No.	Name of the student	Group & Medium
1	319008036	T. Asha	CBZ (TM)
2	319008044	K. Himabindu	
3	319008047	A. Manendra	
4	319008239	K. Bhargav	BBC (EM)
5	319008241	B. Chinna	
6	319008242	B. Haribabu	
7	319008246	K. Jagadeeswara	
8	319008247	M. Chandra	
9	319008248	T. Jayaram	

10	319008250	S. Manjunath	
11	319008254	V. Nagendra Prasad	
12	319008257	P. Narasimha	
13	319008259	M Naveen	
14	319008261	A. Pedda Obulesu	
15	319008262	D. Sasi Kumar	
16	319008269	U. Someswara	
17	319008273	T. Sreekanth Reddy	
18	319008274.	T. Vikay	

III B.Sc. , VI SEMESTER

S.no	Hall Ticket No.	Name of the student	Group & Medium
1	318008146	S. Gangadar	MPC (EM&TM)
2	318008159	M. Kusuma Priya	
3	318008163	G. Mahesh	
4	318008161	C. Madhu Mohan	
5	318008265	K. Vamsi	MZC (EM)
6	318008201	C. Gnana Sekahr	
7	318008240	P. Sivakumar Reddy	

H.O.D of the Department

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

To

1. All teaching staff of chemistry Department for circulation.
2. Department notice board

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

INSTITUTIONAL TIME - TABLE FOR REMEDIAL CLASS (Even semester)

S.G.S ARTS COLLEGE,TIRUPATI			
COURSE : B.A, B.Sc, B.Com		SEMESTER:II , IV &VI	
TIME:4.00PM-5.30PM		**YEAR: I,II& III**	
DAY	**SUBJECT**	**SEMESTER**	**ROOM NO**
MONDAY	CHEMISTRY	II	224
		IV	225
		VI	230
	STATISTICS	II	220
		IV	221
		VI	222
	COMPUTER SCIE4NCE	II	114
		IV	115
		VI	116
	POLITICAL SCIENCE	IV	210
		VI	212
		II	306
	COMMERCE	IV	307
		VI	308
		II	224
TUESDAY	CHEMISTRY	II	224
		IV	225
		VI	230

	STATISTICS	II	220
		IV	221
		VI	222
	COMPUTER SCIENCE	II	114
		IV	115
		VI	116
	POLITICAL SCIENCE	II	210
		IV	211
		VI	212
	COMMERCE	II	306
		IV	307
		VI	308
WEDNESDAY	MATHEMATICS	II	224
		IV	225
		VI	230
	ZOOLOGY	II	Z LAB
		IV	ZLAB-2
		VI	129
	BIOTECHNOLOGY	II	LAB
		IV	315
		VI	316
	HISTORY	II	210
		IV	211
		VI	212
	ACCOUNTS	II	306
		IV	307
		VI	308
THURSDAY	MATHEMATICS	II	224
		IV	225
		VI	230
	ZOOLOGY	II	LAB
		IV	Z LAB-2

		VI	129
	BIOTECHNOLOGY	II	LAB
		IV	315
		VI	316
	HISTORY	II	210
		IV	211
		VI	212
	ACCOUNTS	II	306
		IV	307
		VI	308
FRIDAY	PHYSICS	II	224
		IV	225
		VI	230
	BOTONY	II	B-LAB
		IV	B-LAB-2
		VI	129
	MICROBIOLOGY	II	LAB
		IV	314
		VI	315
	ECONOMICS	II	210
		IV	211
		VI	212
	TELUGU	II	203
		IV	204
		VI	205
SATURDAY	PHYSICS	II	224
		IV	225
		VI	230
	BOTONY	II	B-LAB
		IV	B LAB-2
		VI	129
	MICROBIOLOGY	II	LAB

		IV	314
		VI	315
	ECONOMICS	II	210
		IV	211
		VI	212
	TELUGU	II	203
		IV	205
		VI	206

S.G.S. ARTS COLLEGE ★ TIRUPATI ★ Co-ordinator NAAC

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA **Phone No: 9490728655**
M.Sc.,Mphil.,(Ph.D) **Mail.Id : parameswara1964@gmail.com**
HOD OF CHEMISTRY

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2016-17			
SEMESTERS - II, IV & VI		**DATE : 23/01/2017 to 11/04/2017**	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2017-18			
SEMESTERS - II, IV & VI -		DATE : 06/12/2017 to 03/04/2018	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19			
SEMESTERS - II, IV & VI	**DATE : 12/11/2018 to 11/02/2018**		
Name Of The Lecturer	**MONDAY**	**TUESDAY**	
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF CHEMISTRY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - II, IV & VI	DATE : 09/12/2019 to 03/03/2020		
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
SRI D. PARAMESWARA	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. V. Venkata Lakshmi	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. P. SUGUNA	II - SEMESTER	II - SEMESTER	4.00 TO 5.30
Smt. K. REVATHI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with chemistry combination

Semester: II, IV&VI **Date: 09-12-2019 to 03-03-2020**

Name of the Institute: II BSc IV Sem

SL. NO.	ADMISSION NO.	NAME	CASTE	1	2	3	4	5	6	7	8	9	10	11
				7/12	10/12	14/12	17/12	28/12	24/12	30/12	3/12	7/1	8/1	21/1
		CBZ (TM)												
31900	8036	T. Asha		P	P	P	P	P	P	P	P	P	P	A
	8044	K. Hima bindhu		P	P	P	P	A	P	P	P	P	P	P
	8047	A Manendra		P	A	P	P	P	P	P	P	A	P	P
		BBC (EM)												
	8239	K. Baraghav		P	P	P	P	P	P	A	P	P	P	P
	8241	B. Chinna		P	P	P	A	P	P	P	P	P	P	P
	8242	S. Chowdaiah		P	A	P	P	P	P	P	P	P	A	P
	8246	B. Haribabu		P	P	P	P	P	A	P	P	P	P	P
	8247	K. Jagadeeswara.		P	P	P	A	P	P	P	P	P	P	P
	8248	M. Chandra		A	P	P	P	P	P	P	P	P	P	A
	8250	T. Jayaram		P	P	P	P	A	P	P	P	P	P	P
	8254	S. Manjunadh		P	P	P	P	P	P	P	A	P	P	P
	8257	V. Nagendra Prasad		P	P	A	P	P	P	P	P	P	P	P
	8259	P. Narasimha		A	P	P	P	P	P	P	P	P	P	P
	8261	M. Naveen		P	P	P	P	P	A	P	A	P	P	P
	8262	A. Pedda Obulesu		P	P	P	P	P	P	A	P	P	P	P
	8269	D. Sai Kumar		P	P	A	P	P	P	P	P	P	P	P
	8273	U. Someswara		P	P	P	P	P	A	P	P	P	P	P
	8274	T. Sreekanth Reddy		P	P	P	P	P	P	A	P	P	P	P
	8280	T. Vijay		P	P	P	P	P	P	P	A	P	P	P

SL. NO.	ADMISSION NO.	NAME	CASTE	1	2	3	4	5	6	7	8	9
				7/12	10/12	14/12	17/12	33/12	24/12	34/12	9/12	1/12
		MPC (EM & TM)										
31800	8146	S. Gangadar		P	P	P	P	A	P	P	P	P
	8159	M. Kusuma priya		P	P	P	P	P	P	A	P	P
	8163	G. Mahesh		P	P	P	A	P	P	P	P	P
	8161	C. Madhu Mohan		P	A	P	P	P	P	P	P	A
		MPC (EM)										
	8265	K. Vamsi		P	P	A	P	P	P	P	P	A
		BBC (EM)										
	8201	C. Gnanasekhar		P	P	A	P	P	P	P	P	P
	8240	P. Sivakumar Reddy		P	P	P	P	P	A	A	P	P

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Students engaged in Study hours conducted for slow learners by the Department of Chemistry.

Students engaged in peer learning

Department of Microbiology

1. Remedial Coaching

Slow learners require additional encouragement and support for better academic performance. The college regularly conducts remedial coaching for such students. This led to improved performance of the students in the university examinations.

Special programmes conducted for slow learners

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Sri V. VENKATALAKSHMI
M.Sc.,Mphil.,(Ph.D)
HOD OF MICROBIOLOGY

ODD SEMESTERS
CIRCULAR

Date: 11-06-2019

As Per the instructions of the principal we are planning to conduct remedial classes for slow learners to improve their performance in the university examinations. The classes may be conducted out of the college hours i.e., before 10 A.M /after 4 P.M . The time tabel of remedial classes for slow learners is enclosed.

(HOD)

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF MICROBIOLOGY

Dr V. Venkatalakshmi **Phone No: 9492075161**
M.Sc..,(Ph.D)
HOD OF MICROBIOLOGY

NOTICE

Date: 25-06-2019

SGSAC/Microbiology dept/Remedial classes- 2019/HOD/01

All the slow learners in Microbiology of III&V semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 28-06-2019 in Room numbers 314& 315 from 4.00Pm to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.

III - SEMESTER :

S. No	Hall Ticket No.	Name of the student	Group & Medium
1	0319008001	A Balaji	MZC(EM)
2	0319008005	D Ganesh	
3	0319008014	E Naveen	
4	0319008018	G Pranay kumar	
5	0319008026	D somasekhar	
6	0319008027	P subramanyam	
7	0319008030	G swarna latha	
8	0319008034	B vijay kumar	
9	0319008035	N vinod kumar	
10	0319008033	N venkatesh	

V - SEMESTER :

S. No	Hall Ticket No.	Name of the student	Group & Medium
1	0318008258	B anwarlal hussain	MZC (EM)
2	0318008267	Sk mazeed babu	

H.O.D of the Department

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517502

To

1. All teaching staff of microbiology
2. Department for circulation.
3. Department notice board

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF MICROBIOLOGY

Dr V. Venkatalakshmi **Phone No: 9492075161**
M.Sc.,(Ph.D)
HOD OF MICROBIOLOGY

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF MICROBIOLOGY			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - III & V		DATE : 28/06/2019to 31/10/2019	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
Dr. K. SRIDEVI	III - SEMSTER	III - SEMSTER	4.00 TO 5.30
Dr. K. SRIDEVI	V - SEMESTER	V - SEMESTER	4.00 TO 5.30

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI
DEPARTMENT OF MICROBIOLOGY

Dr. V. VENKATALAKSHMI **Phone No: 9492075161**
M.Sc.,(Ph.D)
HOD OF MICROBIOLOGY

NOTICE

Date: 03-12-2019

SGS AC /Microbiology dept/Remedial classes- 2019/HOD/01

All the slow learners in Microbiology of II, IV&VI semesters as enclosed in the list are here by intimated that remedial classes are planned to conduct from 06-12-2019 in Room numbers lab, 314& 315 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory for the students mentioned below. Stringent action will be taken against the students who will be absent for the remedial classes.

H.O.D of the Department

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

To

1. All teaching staff of microbiology Department for circulation.
2. Department notice board

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF MICROBIOLOGY

Sri V. Venkata Lakshmi
M.Sc.,Mphil.,(Ph.D)
HOD OF MICROBIOLOGY

Phone No: 9492075161
Mail.ID: parameswara1964@gmail.com

Departmental Time-Table for remedial classes

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI			
DEPARTMENT OF Microbiology			
REMEDIAL CLASSES			
TIME-TABLE FOR THE ACADEMIC YEAR - 2019-20			
SEMESTERS - II, IV & VI		DATE : 06/12/2019 to 07/03/2020	
Name Of The Lecturer	**MONDAY**	**TUESDAY**	**TIME**
SRI K. SRIDEVI	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	VI - SEMESTER	VI - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	II - SEMESTER	II - SEMESTER	4.00 TO 5.30
Dr. K. SRIDEVI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

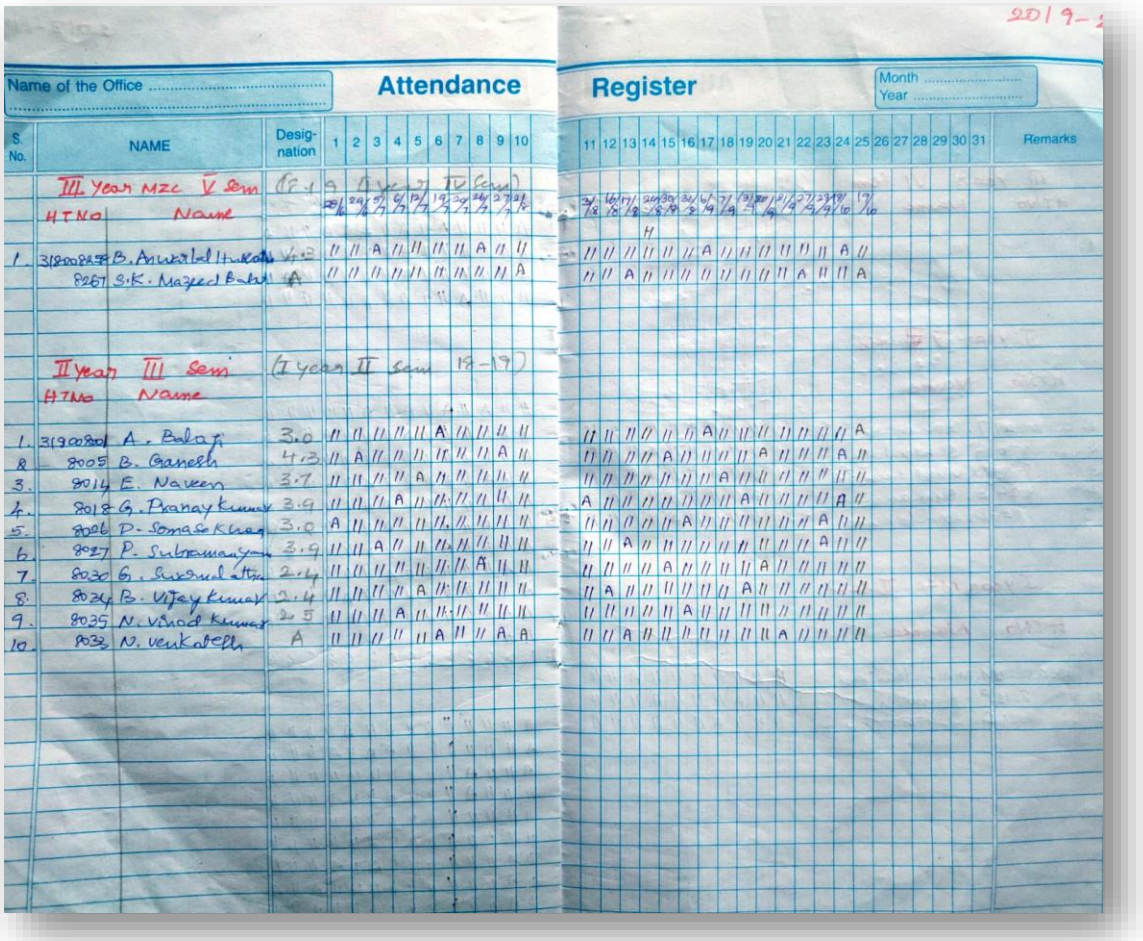
TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with Microbiology combination

Semester: III&V **Date: 28-06-2019 To 19-10-2019**



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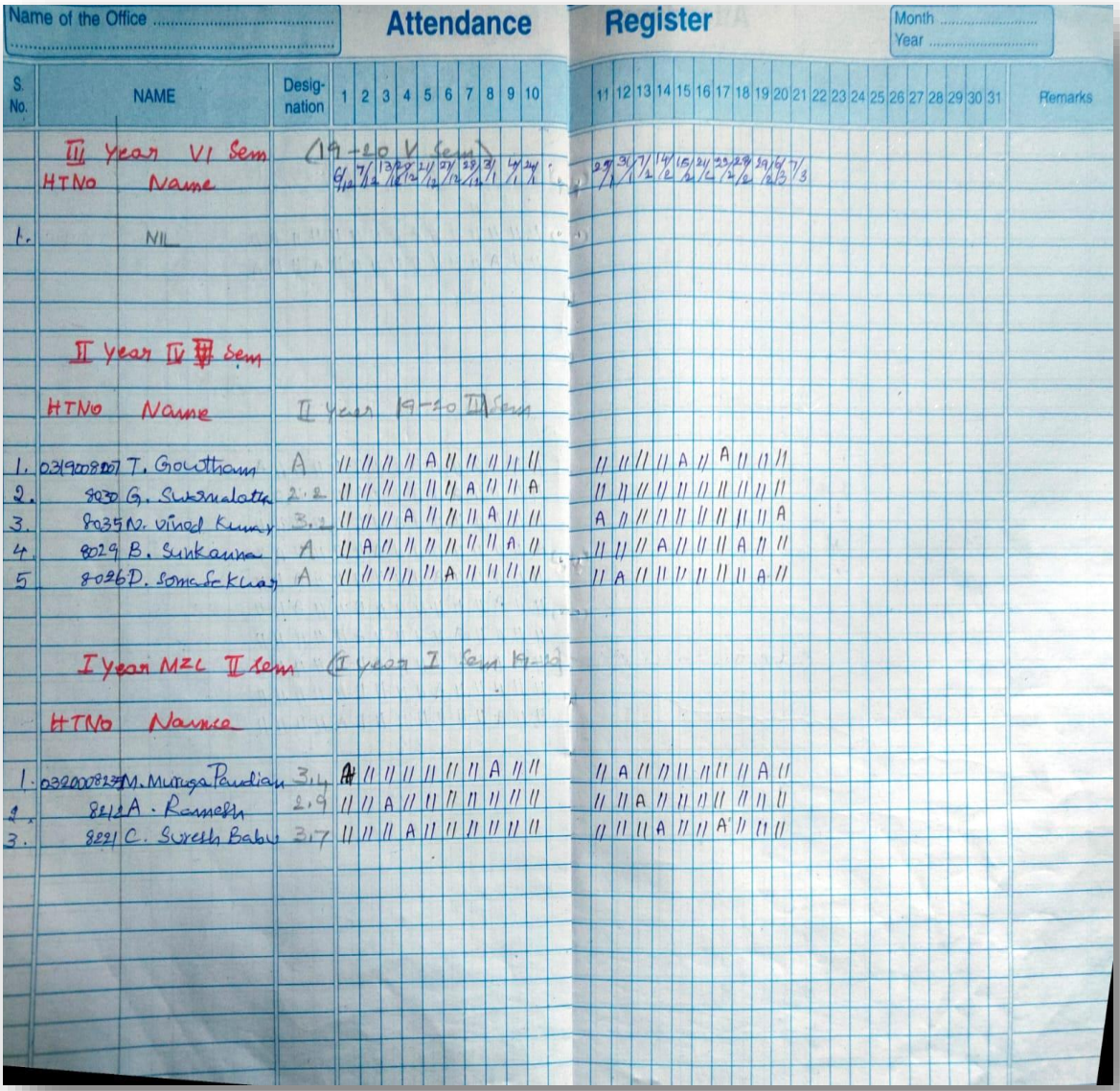
TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Attendance list of slow learners of all programmes with Microbiology combination

Semester: II, IV&VI **Date: 09-12-2019 to 03-03-2020**



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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

LIST OF SLOW LEARNERS

II B.Sc. , III SEMESTER

S.no	Hall Ticket No.	Name of the student	Group&Medium
1	0318008075	SANAMBATLA HIMAVANTH	MSCs-EM

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web:sgsac.edu.in

"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S. ARTSCOLLEGE, TIRUPATI
AIDED (FOUNDED IN 1969)

TIME-TABLE FOR THE ACADEMIC YEAR - 2018-19

DATE : 11/6/2018 to 11/9/2018

SEMESTERS - III

NAME OF THE LECTURER	MONDAY	TUESDAY	TIME
B.TRIVENI N.JAYALAKSHMI	III - SEMESTER	III - SEMESTER	4.00 TO 5.30

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S.ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

III & V SEMESTERS ATTENDANCE REGISTER FOR SLOW LEARNERS :: 2018-2019

S.No	Name of the Student	11/06	12/6	18/6	19/6	25/6	26/6	2/07	5/07	9/07	10/7	16/7	23/7	24/7	30/7	31/7	6/8	7/8	13/8	14/8	20/8	21/8	27/8
II YEAR III SEMESTER MSCS[EM]																							
1	SANAMBATLA HIMAVANTH	/	/	/	A	/	/	/	A	/	/	/	/	/	A	/	/	/	/	/	/	A	/

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

PP044

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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

CLASS TEST MARK SHEET FOR SLOW LEARNERS

II B.Sc. III SEMESTER

S.no	Hall Ticket No.	Name of the student	Group& Medium	TEST-1 (25M)	TEST-2 (25M)
1	0318008075	SANAMBATLA HIMAVANTH	MSCs-EM	20	22

HEAD
DEPARTMENT OF COMPUTER SCI...
S.G.S. ARTS COLLEG...
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS · (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	SANAMBATLA HIMAVANTH	SLOW LEARNERS	II YEAR MSCS-EM	42	44

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web:sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

Date:- 08/11/2018

NOTICE

All the slow learners in Computer Science of IV Semester as enclosed in the list are here by intimated that remedial classes are planned to conduct from 12/11/2018 in Room numbers 114 from 4.00P.M to 5.30 P.M. They are further instructed that attendance is mandatory. Stringent action will be taken against the students who will be absent for the remedial classes.

H.O.D. of the Department

DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

To

1. All teaching staff of Computer Science Department for circulation.
2. Department notice board

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web:sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

LIST OF SLOW LEARNERS

II B.Sc. , IV SEMESTER

S.no	Hall Ticket No.	Name of the student	Group&Medium
1	0318008118	G VAMSI KRISHNA	MSCs-EM
2	0318008099	KESANI SAI KUMAR	
3	0318008075	SANAMBATLA HIMAVANTH	

HEAD
DEPARTMENT OF COMPUTER SCIENCE
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TIRUPATI

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"OM NAMO VENKATESAYA"
TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S. ARTSCOLLEGE, TIRUPATI
AIDED (FOUNDED IN 1969)

TIME-TABLE FOR THE ACADEMIC YEAR :: 2018-19

DATE : 12/11/2018 to 11/02/2018

SEMESTER – IV

NAME OF THE LECTURER	MONDAY	TUESDAY	TIME
B.TRIVENI N.JAYALAKSHMI	IV - SEMESTER	IV - SEMESTER	4.00 TO 5.30

DEPARTMENT OF
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S.ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

II, IV & VI SEMESTERS ATTENDANCE REGISTER FOR SLOW LEARNERS :: 2018-2019

S.No	Name of the Student	12/11	13/11	19/11	20/11	26/11	27/11	3/12	4/12	10/12	11/12	17/12	31/12	7/01	21/01	22/01	28/01	29/01	4/02	5/02	11/02
II YEAR IV SEMESTER MSCS[EM]																					
1	G VAMSI KRISHNA	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
2	KESANI SAI KUMAR	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A
3	SANAMBATLA HIMAVANTH	A	P	P	A	P	P	P	A	A	P	P	P	A	P	P	P	P	P	P	A

HEAD
DEPARTMENT OF COMPUTER SCIENCE
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TIRUPATI

PP050

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TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

CLASS TEST MARK SHEET FOR SLOW LEARNERS

II B.Sc. , IV SEMESTER

S.no	Hall Ticket No.	Name of the student	Group& Medium	TEST-1 (25M)	TEST-2 (25M)
1	0318008118	G VAMSI KRISHNA	II YEAR IV SEM MSCs-EM	20	21
2	0318008099	KESANI SAI KUMAR		18	20
3	0318008075	SANAMBATLA HIMAVANTH		15	AB

DEPARTMENT OF COMPUTER ...
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING

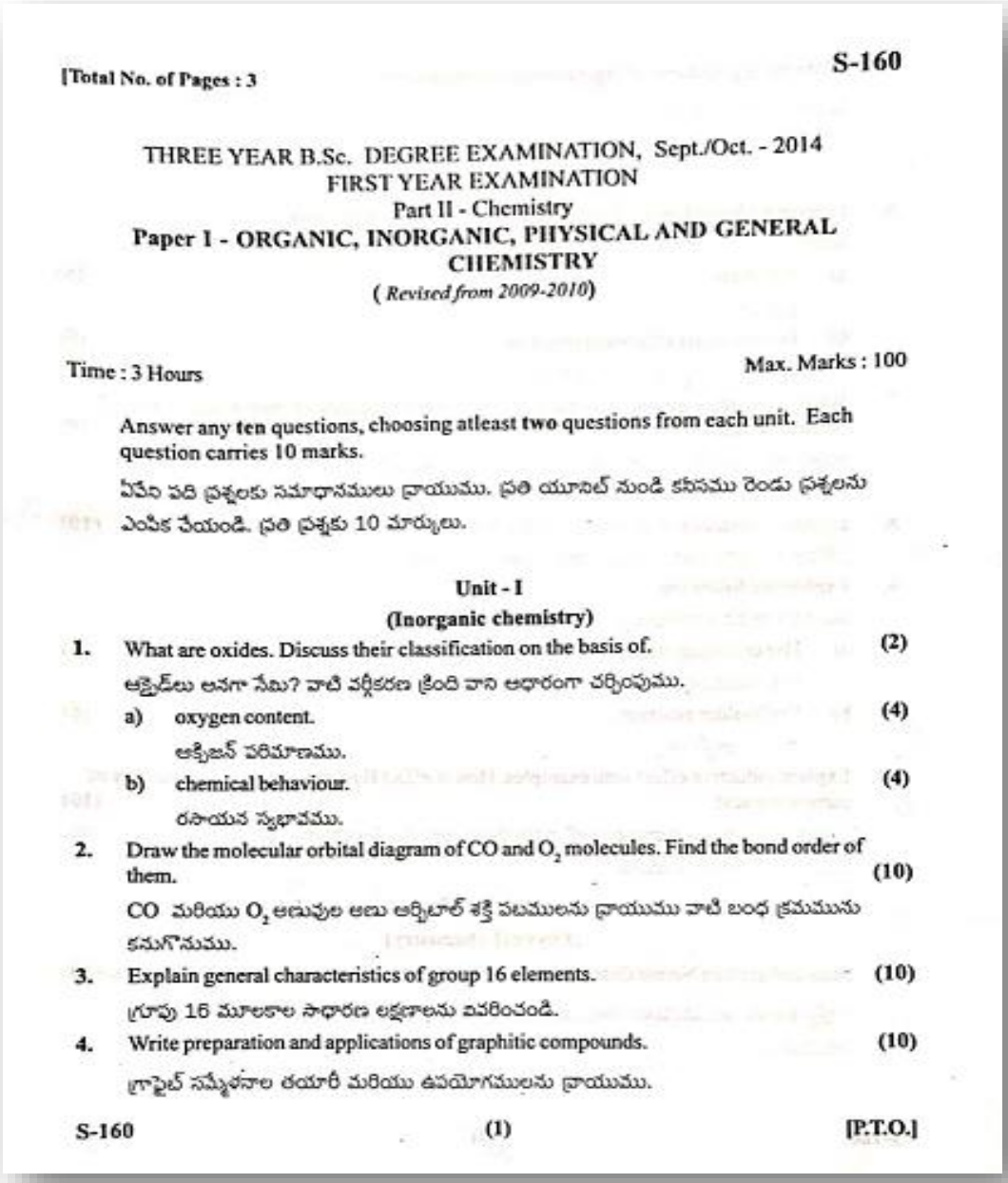
LIST OF SLOW LEARNERS

S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group& Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	G VAMSI KRISHNA	SLOW LEARNERS	II YEAR IV SEM MSCS	20	66
2	KESANI SAI KUMAR			18	41
3	SANAMBATLA HIMAVANTH			22	AB

HEAD
DEPARTMENT OF COMPUTER SC
S.G.S. ARTS COLLEGE
TIRUPATI

2. Rubrics

A rubric is an evaluation tool which promotes consistent application of learning expectations and learning standards in the class room. The college conduct a good number of classes to solve previous examination question papers. This programme helps the slow learners to augment their subject knowledge and ensures better scores in the university examinations.



[Total No. of Pages : 3

S-160

THREE YEAR B.Sc. DEGREE EXAMINATION, Sept./Oct. - 2014
FIRST YEAR EXAMINATION
Part II - Chemistry
Paper I - ORGANIC, INORGANIC, PHYSICAL AND GENERAL CHEMISTRY
(Revised from 2009-2010)

Time : 3 Hours

Max. Marks : 100

Answer any **ten** questions, choosing atleast **two** questions from each unit. Each question carries 10 marks.

ఏవేని పది ప్రశ్నలకు సమాధానములు వ్రాయుము. ప్రతి యూనిట్ నుండి కనీసము రెండు ప్రశ్నలను ఎంపిక చేయండి. ప్రతి ప్రశ్నకు 10 మార్కులు.

Unit - I
(Inorganic chemistry)

1. What are oxides. Discuss their classification on the basis of. (2)
ఆక్సైడ్‌లు అనగా నేమి? వాటి వర్గీకరణ క్రింది వాని ఆధారంగా చర్చింపుము.
   a) oxygen content. (4)
   ఆక్సిజన్ పరిమాణము.
   b) chemical behaviour. (4)
   రసాయన స్వభావము.
2. Draw the molecular orbital diagram of CO and O_2 molecules. Find the bond order of them. (10)
CO మరియు O_2 అణువుల అణు ఆర్బిటాల్ శక్తి పటములను వ్రాయుము వాటి బంధ క్రమమును కనుగొనుము.
3. Explain general characteristics of group 16 elements. (10)
గ్రూపు 16 మూలకాల సాధారణ లక్షణాలను వివరించండి.
4. Write preparation and applications of graphitic compounds. (10)
గ్రాఫైట్ సమ్మేళనాల తయారీ మరియు ఉపయోగములను వ్రాయుము.

S-160 (1) [P.T.O.]

5. Write the applications of organometallic compounds. (10)

సేంద్రీయ లోహ సమ్మేళనాలు యొక్క ఉపయోగాలను సమీకరణాలతో వ్రాయుము.

Unit - II

(Organic Chemistry)

6. Explain with mechanism for the following reactions of benzene.

బెంజీన్ పొందే ఈ క్రింది చర్యలను విధాన పూర్వకముగా వివరింపుము.

a) Nitration (5)

నైట్రేషన్

b) Friedel crafts alkylation reaction. (5)

ఫ్రీడల్ - క్రాఫ్ట్ ఆల్కైలేషన్ చర్యలు.

7. What is meant by mesomeric effect? How it explains acidity of phenol and acidity of carboxylic acid. (10)

మీసోమెరిక్ ప్రభావము అనగా నేమి? ఫీనాల్ యొక్క ఆమ్లత్వమును మరియు కార్బాక్సిలిక్ ఆమ్లాల ఆమ్లత్వమును, దీని ద్వార ఎలా వివరించవచ్చును.

8. Explain orientation of aromatic substitution. (10)

ఎరోమాటిక్ ప్రతిక్షేపణము యొక్క స్థాన నిర్దేశకతను వివరింపుము.

9. Explain the following.

ఈ క్రింది వాటిని వివరింపుము.

a) Hyper conjugation. (5)

అతి సంయుగ్మము.

b) Diels-older reaction. (5)

డీల్స్ - ఆల్డర్ చర్య

10. Explain inductive effect with examples. How it effect Basicity of amines and acidity of carboxylic acid. (10)

ప్రేరేపక ప్రభావమును ఉదాహరణలతో వివరింపుము అది ఎమైన్‌ల క్షారతను మరియు ఆమ్లాల ఆమ్లతను ఏవిధంగా ప్రభావితం చేయును?

Unit - III

(Physical chemistry)

11. State and explain Nernst distribution Law. Explain any four applications of it (10)

నెర్న్స్ట్ వివరణ నియమమును తెల్పి, వివరింపుము. దాని యొక్క నాలుగు అనువర్తనములను తెలుపుము.

12. Deduce the values of critical constants in terms of vander waals constants. (10)

వాండర్ వాల్ స్థిరాంకముల పరంగా సంధిగ్ధ స్థిరాంకములను ఉత్పాదించు విధానము తెలుపుము.

13. What is critical solution temperature? Explain the phenol water system. (10)

సంధిగ్ధ ద్రావణ ఉష్ణోగ్రత అనగా నేమి? ఫినాల్ – నీటి వ్యవస్థను వివరింపుము.

14. What is Joule Thomson effect? Explain. (10)

జౌల్ – థామ్సన్ ప్రభావము అనగా నేమి? వివరింపుము.

15. Derive Bragg's equation. (10)

బ్రాగ్ సమీకరణమును ఉత్పాదింపుము.

Unit - IV

(General chemistry)

16. What is meant by R, S - configuration? Explain with examples. (10)

R, S విన్యాసము అనగా నేమి? ఉదాహరణలతో వివరింపుము.

17. Write the difference between physical and chemical adsorption. (10)

అనురూపాత్మక సాదృశ్యము అనగా నేమి? n - బ్యూటేన్ యొక్క అనురూపాత్మక సాదృశ్యాన్ని వివరింపుము.

18. Explain the following

ఈ క్రింది వాటిని వివరింపుము.

a) Compton effect (5)

కాంప్టన్ ఫలితము

b) Photo electric effect. (5)

ఫోటో ఎలక్ట్రిక్ ప్రభావము.

19. What is conformational isomerism? Explain the conformational isomerism of n-butane. (10)

అనురూపాత్మక సాదృశ్యము అనగా నేమి? n - బ్యూటేన్ యొక్క అనురూపాత్మక సాదృశ్యాన్ని వివరింపుము.

20. What is meant by a colloid? Explain optical and electrical properties of colloids. (10)

కొల్లాయడు అనగానేమి? దృశ్య మరియు విద్యుత్ ధర్మాలను వివరింపుము.

THREE YEAR B.Sc. DEGREE EXAMINATION — MARCH/APRIL 2014

THIRD YEAR EXAMINATION

Part II — Chemistry

Paper III — INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY

(Revised from 2010-2011)

Time : 3 hours Max. Marks : 100

Answer any TEN questions, choosing not less than THREE from each Section.

All questions carry equal marks.

ప్రతి సెక్షన్ నుండి మూడు ప్రశ్నలకు తక్కువ కాకుండా పది ప్రశ్నలకు సమాధానములు వ్రాయుము.

అన్ని ప్రశ్నలకు మార్కులు సమానము.

(Marks: $10 \times 10 = 100$)

SECTION – I

సెక్షన్ – I

(INORGANIC CHEMISTRY)

What is crystal field theory? Discuss splitting of d-orbitals in case of tetrahedral and octahedral complexes. (10)

స్ఫటిక క్షేత్ర సిద్ధాంతమును తెలుపుము? అష్ట వలక మరియు చతుర్ముఖ సమ్మేళనాలలో d - ఆర్బిటాల్ విభజనను వివరింపుము.

Explain the determination of composition of a complex by mole-ratio method. (10)

సంశ్లిష్ట సమ్మేళనాల సంఘటనమును మోల్-నిష్పత్తి పద్ధతి ద్వారా ఎట్లు నిర్ణయించెదవో తెల్పుము.

Explain SN^1 and SN^2 reaction mechanism in coordination compounds. (10)

సంశ్లిష్ట సమ్మేళనాలలో SN^1 మరియు SN^2 చర్యల చర్యా విధానమును వ్రాయుము.

Explain electronic absorption spectrum of $[Ti(H_2O)_6]^{3+}$. (10)

$[Ti(H_2O)_6]^{3+}$ అయాన్ యొక్క ఎలక్ట్రానిక్ శోషణ వర్ణపటమును వివరింపుము.

(a) Explain biological significance of Na^+ and K^+ ions. (5)

జీవ వ్యవస్థలో Na^+ మరియు K^+ అయాన్ల ప్రాముఖ్యతను తెలుపుము.

(b) Explain biological significance of copper. (5)

జీవ వ్యవస్థలో కాపర్ యొక్క ప్రాముఖ్యతను తెలుపుము.

Explain the geometrical isomerism in complexes. (10)

సంశ్లిష్టాలలోని క్షేత్ర సాదృశ్యమును వివరించండి.

SECTION — II

సెక్షన్ - II

(ORGANIC CHEMISTRY)

7. What are amines? Discuss the action of nitrous acid on primary, secondary and tertiary amines. (10)

ఎమైన్స్ అంటే ఏమిటి? ప్రైమరీ, సెకండరీ మరియు టెర్షరీ ఎమైన్లలో నైట్రస్ ఆమ్లం చర్య ఏ విధంగా ఉంటుందో తెలపండి.

8. What are heterocylic compounds? Write any two methods of preparation of pyrrole. Explain its acid character. (10)

విజాతీయ వలయ సమ్మేళనాలు అనగా నేమి? పిర్రోల్ తయారుచేయుటకు రెండు పద్ధతులను వ్రాయండి. పిర్రోల్ ఆమ్ల గుణత్వాన్ని వివరించండి.

9. What are Amino Acids? Discuss their classification and explain (10)

(a) Zwitter ion structure and

(b) Isoelectric point.

ఎమైనో ఆమ్లాలు అనగా నేమి? వాటి వర్గీకరణ గూర్చి వ్రాసి

(a) జ్విట్టర్ ఆయాన్ నిర్మాణము మరియు

(b) సమ విద్యుత్ స్థానములను గురించి వివరించండి.

10. Write the structural elucidation of glucose. (10)

గ్లూకోజ్ నిర్మాణమును గురించి వివరించండి.

11. Write notes on the following :

(a) Hoffman's bromamide reaction with mechanism (5)

హాఫ్‌మన్ బ్రోమమైడ్ చర్య మరియు చర్యా విధానము

(b) Carbylamine reaction. (5)

కార్బైల్ ఎమైన్ చర్యలను గూర్చి గూర్చి క్లుప్తముగా వివరించండి

12. Write the preparation of pyridine and its nucleophilic substitution reactions.

పిరిడీన్ తయారు చేయు పద్ధతిని తెల్పి, దాని న్యూక్లియోఫిలిక్ ప్రతిక్షేపణ చర్యలను వ్రాయుము.

SECTION – III

(PHYSICAL CHEMISTRY)

13. Derive an equation for the rate constant of a first order reaction and write its characteristics. (10)

ప్రథమ క్రమాంక చర్య రేటు స్థిరాంక సమీకరణమును ఉత్పాదించుము మరియు దాని లక్షణాలను తెలుపుము.

14. Give the various definitions for the first law of thermodynamics. Define C_p and C_v and derive the relation between them. (10)

ఉష్ణ గతిక శాస్త్ర ప్రథమ నియమాన్ని తెలిపే వివిధ నిర్వచనాలను వ్రాయుము. Cp మరియు Cv లను నిర్వచించి వాటి మధ్యగల సంబంధమును రాబట్టుము.

15. State and explain Einstein's Law of photochemical equivalence. (10)

ఐన్‌స్టీన్ కాంతి రసాయన తుల్యతా నియమమును నిర్వచించి వివరించుము.

16. Derive and explain Kirchoff's equation and explain the Carnot cycle. (10)

కిర్‌ఛాఫ్ సమీకరణమును ఉత్పాదించి వివరింపుము. మరియు కార్నాట్ వలయమును గురించి వివరింపుము.

17. State and explain quantum yield. Explain high quantum yield in $H_2 - Cl_2$ reaction. (10)

క్వాంటమ్ ప్రాప్తి గూర్చి తెల్పి వివరించుము. $H_2 - Cl_2$ చర్య యొక్క అధిక క్వాంటమ్ ప్రాప్తి గూర్చి వివరించండి.

18. Describe various methods for determining the order of a chemical reaction. (10)

చర్య క్రమాంకాలను కనుగొను పద్ధతులను గూర్చి వ్రాయుము.

(23051)

THREE YEAR B.Sc. DEGREE EXAMINATION, APRIL 2014

SECOND YEAR EXAMINATION

Part II – Chemistry

Paper II — INORGANIC, ORGANIC, PHYSICAL AND GENERAL CHEMISTRY

Time : 3 Hours Max. Marks : 100

Answer any TEN questions choosing not less than TWO from each Section.

All question carry equal marks.

ప్రతి సెక్షన్ నుండి కనిసము రెండు ప్రశ్నలను ఎంచుకొను ఏ వది ప్రశ్నలకైనా సమాధానములు వ్రాయుము.

అన్ని ప్రశ్నలకు మార్కులు సమానము.

(Marks : 10 × 10 marks = 100 marks)

SECTION – I

సెక్షన్ – I

(INORGANIC CHEMISTRY)

(మూలక రసాయన శాస్త్రము)

Explain the colour and magnetic behaviour of transition elements.

పరివర్తన మూలకాల రంగు మరియు అయస్కాంత ధర్మములను వివరింపుము.

Give an account of the following

(a) Lanthanide contraction

(b) Oxidation states of *f* –block elements.

ఈ క్రింది వాటిపై వ్యాఖ్యానింపుము.

(a) లాంథనైడ్ సంకోచము.

(b) *f* –బ్లాకు మూలకల ఆక్సీకరణ స్థితులు.

Explain the free electron theory and give any two of its limitations.

స్వేచ్ఛా ఎలక్ట్రాన్ సిద్ధాంతమును వివరింపుము. ఆ సిద్ధాంతమును లోపాలను ఏవేని రెండింటిని తెలుపుము.

Explain the band theory of conductors, iusulators and semiconductors.

వాహకాలు , అర్ధ వాహకాలు మరియు నిరోధకాలను లోహాల పట్టీ సిద్ధాంతపరంగా వివరించండి.

[P.T.O]

5. Discuss the structures and shapes of $[Ni(Co)_4]$ and $[Fe(Co)_5]$.

$[Ni(Co)_4]$ మరియు $[Fe(Co)_5]$ ల నిర్మాణాలను, ఆకారములను వివరింపుము.

SECTION – II

సెక్షన్ – II

(ORGANIC CHEMISTRY)

(కర్బన రసాయన శాస్త్రము)

6. Give the mechanism of SN^2 reaction. Explain the stereochemistry of SN^2 reactions example.

SN^2 చర్యల చర్యా విధానమును తెలుపుము. SN^2 చర్యల త్రిమితీయ రసాయన శాస్త్రమును సోదాహ వివరింపుము.

7. Write a note on the following

(a) Synthesis of alcohols fran grignard reagents

(b) Pinacol–pinacolone rearrangement

ఈ క్రింది వాటిపై చర్యా వ్యాఖ్యను వ్రాయుము.

(a) గ్రిగ్నాండ్ కారకాల నుండి ఆల్కహాల్ సంశ్లేషణ

(b) పినకాల్– పినకోలోన్ పునఃఅమరికా

Write the mechanisms of the following reactions.

(a) Aldol condensation

(b) Perkin reaction.

ఈ క్రింది వాటికి చర్యా విధానాలను వివరింపుము.

(a) ఆల్డాల్ సంఘనన చర్య

(b) పెర్కిన్ చర్య.

Explain the following reactions.

(a) Huns –Diecker reaction

(b) HVZ–reaction

ఈ క్రింది చర్యలను వివరింపుము.

(a) హాన్స్–డికర్స్ చర్య

(b) HVZ–చర్య.

How do you synthesise the following from aceto acetic ester.

(a) *n*-Butyric acid
(b) Succinic acid
(c) Crotonic acid
(d) Heterocyclic compound.

అసిటో, అసిటిక్ నుండి ఈ క్రింది వాటిని ఎలా విశ్లేషిస్తారు.

(a) ఎన్-బుట్రిక్ ఆసిడ్
(b) సుసినిక్ ఆసిడ్
(c) క్రొటోనిక్ ఆసిడ్
(d) హెటిరో సైకిలిక్ సమ్మేళనం.

SECTION – III

సెక్షన్ – III

(PHYSICAL CHEMISTRY)

(భౌతిక రసాయన శాస్త్రము)

11. Define phase rule. Draw and explain the phase diagram of Nacl – H_2O system.

ప్రావస్థా నియమాన్ని నిర్వచించుము. Nacl – H_2O వ్యవస్థ యొక్క ప్రావస్థా చిత్రాన్ని గీసి వివరించుము.

12. Define osmotic pressure. How will you determine the osmotic pressure of a solutic experimentally?

ద్రవాభిసరణ పీడనము అనగా నేమి? ప్రయోగ పూర్వకంగా ఒక ద్రావణం యొక్క ద్రవాభిసరణ పీడనాన్ని ఎ నిర్ణయిస్తారు.

13. State and explain Kohlrauschi's law. Briefly explain any two of its applications.

కోల్‌రాష్ నియమాన్ని తెలిపి, ఏవేని రెండు అనువర్తనాలను వివరించుము.

14. Write notes on the following

(a) Ostwald's dilution law
(b) Calomel electrode.

ఈ క్రింది వాటిని వివరించుము.

(a) ఆస్వాల్డ్ విలీన ద్రావణాల నియమము
(b) కాలొమెల్ ఎలక్ట్రోడ్.

15. Define EMF of a cell. How EMF of a cell can be experimentally measured.

ఘటక EMF అనగా నేమి? ఘటక EMF ను ప్రయోగపూర్వకంగా ఎలా కనుగొంటారు.

SECTION - IV

సెక్షన్ - IV

(GENERAL CHEMISTRY)

(సాధారణ రసాయన శాస్త్రము)

16. Define symmetry. Explain the different elements of symmetry with examples.

సౌష్ఠవత అనగా నేమి? వివిధ సౌష్ఠవాంశాలను ఉదాహరణలతో వివరింపుము.

17. Briefly explain the principle involved in complexo metric titration's.

సంశ్లిష్టాత్మక అంశ మాపనాల సిద్ధాంతాన్ని క్లుప్తంగా వివరింపుము.

18. Define and explain "Significant figures" and its importance.

సార్థక సంఖ్యలను నిర్వచించి వివరింపుము. వాటి ప్రత్యేకతను వివరింపుము.

19. What is asymmetric synthesis? Give any two methods of asymmetric synthesis.

అసౌష్ఠవ సంశ్లేషణ అనగా నేమి? అసౌష్ఠవ సంశ్లేషణ పద్ధతులను ఏవేని రెండింటిని తెలుపుము.

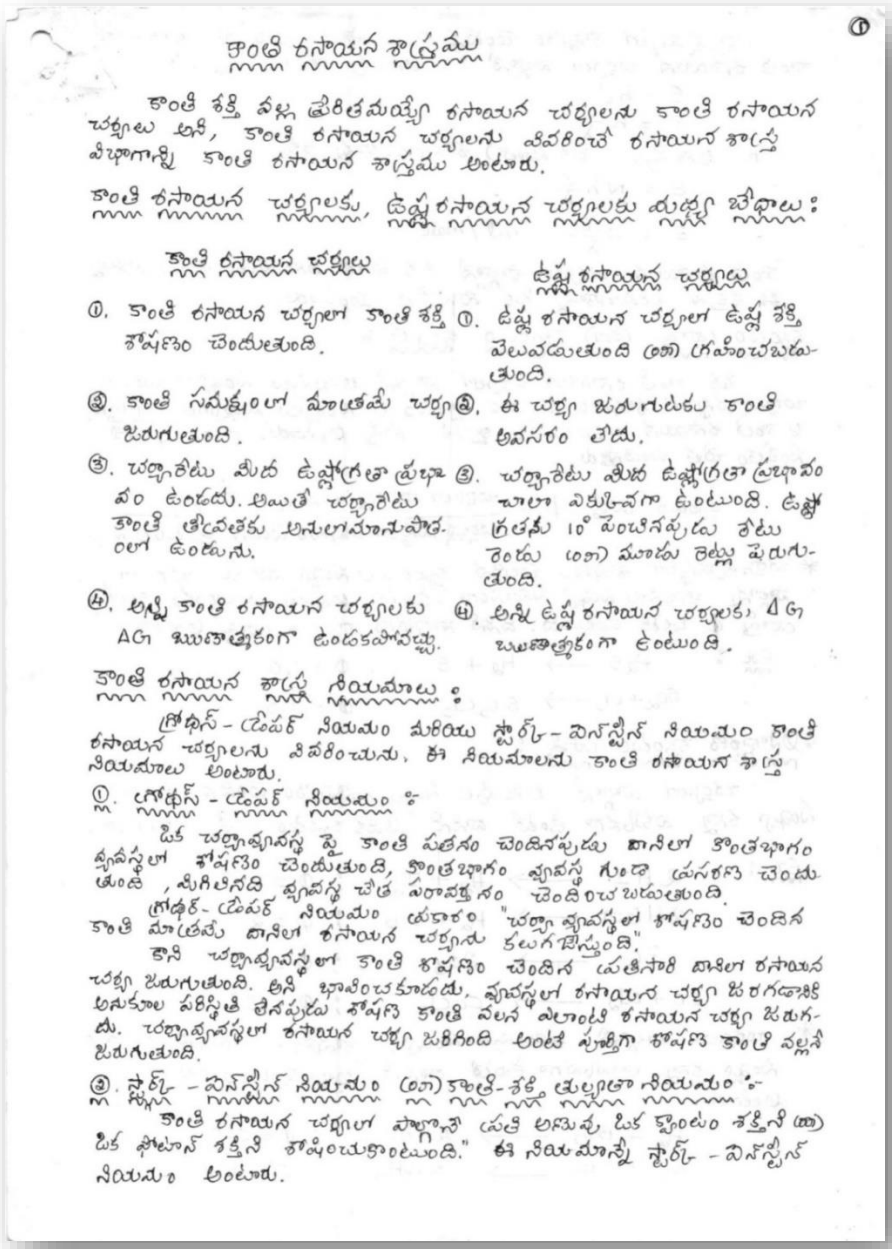
20. Write a note on acid base indicators.

ఆమ్లక్షార సూచికలను వివరింపుము.

3. Simplified study material

Slow learners are those students' whose learning pace is slower than their peers. The respective4 class teachers provide simplified study notes to enhance the confidence in the subject and augment their learning peace.

Photochemistry



కాంతి రసాయన శాస్త్రము ①

కాంతి శక్తి వల్ల ప్రేరితమయ్యే రసాయన చర్యలను కాంతి రసాయన చర్యలు అని, కాంతి రసాయన చర్యలను వివరించే రసాయన శాస్త్ర విభాగాన్ని కాంతి రసాయన శాస్త్రము అంటారు.

కాంతి రసాయన చర్యలకు, ఉష్ణ రసాయన చర్యలకు మధ్య భేదాలు:

కాంతి రసాయన చర్యలు	ఉష్ణ రసాయన చర్యలు
①. కాంతి రసాయన చర్యలలో కాంతి శక్తి శోషణం చెందుతుంది.	①. ఉష్ణ రసాయన చర్యలలో ఉష్ణ శక్తి వెలువడుతుంది లేదా గ్రహించబడుతుంది.
②. కాంతి సమక్షంలో మాత్రమే చర్య జరుగుతుంది.	②. ఈ చర్య జరుగుటకు కాంతి అవసరం లేదు.
③. చర్యావేగం మీద ఉష్ణోగ్రతా ప్రభావం ఉండదు. అయితే చర్యావేగం కాంతి తీవ్రతకు అనులోమానుపాతంలో ఉండును.	③. చర్యావేగం మీద ఉష్ణోగ్రతా ప్రభావం చాలా ఎక్కువగా ఉంటుంది. ఉష్ణోగ్రత 10° పెంచినప్పుడు వేగం రెండు లేదా మూడు రెట్లు పెరుగుతుంది.
④. అన్ని కాంతి రసాయన చర్యలకు A_G ఋణాత్మకంగా ఉండకపోవచ్చు.	④. అన్ని ఉష్ణ రసాయన చర్యలకు ΔG ఋణాత్మకంగా ఉంటుంది.

కాంతి రసాయన శాస్త్ర నియమాలు:

గ్రోథస్ - డ్రేపర్ నియమం మరియు స్టార్క్ - ఐన్‌స్టీన్ నియమం కాంతి రసాయన చర్యలను వివరించును, ఈ నియమాలను కాంతి రసాయన శాస్త్ర నియమాలు అంటారు.

① గ్రోథస్ - డ్రేపర్ నియమం :

ఒక చర్యావ్యవస్థ పై కాంతి పతనం చెందినప్పుడు పాక్షికంగా కొంతభాగం వ్యవస్థలో శోషణం చెందుతుంది, కొంతభాగం వ్యవస్థ గుండా ప్రసరణ చెందుతుంది, మిగిలినది వ్యవస్థ చేత పరావర్తనం చెందించబడుతుంది.

గ్రోథర్ - డ్రేపర్ నియమం ప్రకారం "చర్యా వ్యవస్థలో శోషణం చెందిన కాంతి మాత్రమే పాక్షికంగా రసాయన చర్యను కలుగజేస్తుంది."

కానీ చర్యావ్యవస్థలో కాంతి శోషణం చెందిన ప్రతిసారి పాక్షికంగా రసాయన చర్య జరుగుతుంది. అని భావించకూడదు, వ్యవస్థలో రసాయన చర్య జరగడానికి అనుకూల పరిస్థితి లేనప్పుడు శోషణ కాంతి వలన ఎలాంటి రసాయన చర్య జరుగదు. చర్యావ్యవస్థలో రసాయన చర్య జరిగింది అంటే పూర్తిగా శోషణ కాంతి వల్లనే జరుగుతుంది.

② స్టార్క్ - ఐన్‌స్టీన్ నియమం (లేదా) కాంతి-శక్తి తుల్యతా నియమం :

కాంతి రసాయన చర్యలో పాల్గొనే ప్రతి అణువు ఒక క్వాంటం శక్తిని (లేదా) ఒక ఫోటాన్ శక్తిని శోషించుకొంటుంది." ఈ నియమాన్నే స్టార్క్ - ఐన్‌స్టీన్ నియమం అంటారు.

చర్యాపూవస్థలో శోషణం చెందిన కాంతి పౌనఃపున్యం 'ν' అనుకుంటే కాంతి రసాయన చర్యలో పాల్గొనే ఒక అణువు శోషించుకున్న శక్తి

$$E = h\nu$$
$$E = h\frac{c}{\lambda}$$

'N' అణువులు (ఒక మోల్) శోషించుకున్న శక్తి

$$E = Nh\frac{c}{\lambda}$$

$$E = \frac{2.86}{\lambda} \quad cal/mole$$

కాంతి రసాయన చర్యలో పాల్గొనే ఒక మోల్ అణువులు శోషించుకున్న ఈ శక్తిని పరిమాణాన్ని ఒక "ఐన్‌స్టీన్" అంటారు.

క్వాంటం ఫలితం (లేదా) క్వాంటం దక్షత :

ఒక కాంతి రసాయన చర్యలో పాల్గొనే అణువులు సంఖ్యకు మరియు చర్యాపూవస్థలో శోషణం చెందిన క్వాంటం ల సంఖ్యకు మధ్యగల నిష్పత్తినే, ఆ కాంతి రసాయన చర్య యొక్క క్వాంటం ఫలితం అంటారు. దీనిని ϕ అనే సంకేతం చేత సూచిస్తారు.

$$\text{క్వాంటం ఫలితం, } \phi = \frac{\text{చర్యలో పాల్గొన్న అణువుల సంఖ్య}}{\text{చర్యాపూవస్థలో శోషణం చెందిన క్వాంటంల సంఖ్య}}$$

* చర్యాపూవస్థలో శోషణం చెందిన క్వాంటం ల సంఖ్య మరియు చర్యలో పాల్గొన్న అణువుల సంఖ్య సమానంగా ఉంటే అప్పుడు ఆ చర్యకు క్వాంటం ఫలితం ϕ ఒకటి ఉంటుంది. దీనిని "సాధారణ" క్వాంటం ఫలితం" అంటారు.

ఉదా :
$$H_2S \longrightarrow H_2 + S \quad ; \phi = 1.0$$
$$SO_2 + Cl_2 \longrightarrow SO_2Cl_2 \quad ; \phi = 1.0$$

*అసాధారణ క్వాంటం ఫలితం :

చర్యలో పాల్గొనే అణువుల సంఖ్య, శోషణం చెందిన క్వాంటం సంఖ్య కన్నా ఎక్కువగా ఉంటే దానిని "అధిక క్వాంటం ఫలితం" అంటారు.

ఉదా :
$$2HI \longrightarrow H_2 + I_2 \quad ; \phi = 2$$
$$2HBr \longrightarrow H_2 + Br_2 \quad ; \phi = 2$$
$$3O_2 \longrightarrow 2O_3 \quad ; \phi = 3$$
$$CO + Cl_2 \longrightarrow COCl_2 \quad ; \phi = 10^3$$

* చర్యలో పాల్గొనే అణువుల సంఖ్య, శోషణం చెందిన క్వాంటం సంఖ్య కన్నా తక్కువగా ఉంటే దానిని "అల్ప క్వాంటం ఫలితం" అంటారు.

$$H_2 + Br_2 \longrightarrow 2HBr \quad ; \phi = 0.01$$
$$N_2 + 3H_2 \longrightarrow 2NH_3 \quad ; \phi = 0.2$$

$$CH_3 - CHO \longrightarrow CO + CH_4 \quad ; \phi = 0.5$$

అధిక క్వాంటం ఫ్రాప్తికి కారణాలు :

చర్యామిశ్రమంలోని ఒక అణువు క్వాంటం శక్తిని శోషించుకొని ఉత్తేజితం చెందుతుంది. ఈ విధంగా ఉత్తేజితం చెందిన అణువు తర్వాత శృంఖల చర్యను కొనసాగిస్తుంది. చర్యలుగా పాల్గొనే అణువుల సంఖ్యా శోషణం చెందిన క్వాంటంల సంఖ్యా కన్నా అధికంగా ఉంటుంది. ఫలితంగా క్వాంటం దక్షత "1" కన్నా ఎక్కువగా ఉంటుంది.

ఉదా:

$$2HI \longrightarrow H_2 + I_2 \quad ; \phi = 2$$

చర్యావిధానం :

$$HI + h\nu \longrightarrow H + I$$
$$HI + H \longrightarrow H_2 + I$$
$$I + I \longrightarrow I_2$$

$$2HI + h\nu \longrightarrow H_2 + I_2 \quad \phi = 2$$

పై చర్యలో ఒక క్వాంటం శోషణం చెందుట వలన రెండు అణువులు చర్యలో పాల్గొనును. కనుక ఈ చర్యా యొక్క క్వాంటం దక్షత "2"

అల్ప క్వాంటం ఫ్రాప్తికి కారణాలు :

అల్ప క్వాంటం ఫ్రాప్తికి కారణాలను క్రింది విధంగా ప్రస్తావించవచ్చును.

i). చర్యావ్యూహస్థలో శోషణం చెందిన కాంతి శక్తి దానిలో రసాయన చర్యను కలుగజేయడానికి తగినంతగా తొలపుడు క్వాంటం దక్షత తక్కువగా ఉంటుంది.

ii). చర్యావ్యూహస్థలోని ఒక అణువు క్వాంటం శక్తిని శోషించుకొని ఉత్తేజితం చెందుతుంది. ఈ విధంగా ఉత్తేజితం చెందిన అణువు చర్యలో పాల్గొనడానికి ముందే అనుత్తేజితం చెందుతుంది. అలాంటి పరిస్థితులలో క్వాంటం దక్షత "1" కన్నా తక్కువగా ఉంటుంది.

iii). చర్యావ్యూహస్థలోని ఒక అణువు క్వాంటం శక్తిని శోషించుకొని పరమాణువులుగా లేదా రాడికల్స్ గా విడిపోతుంది. ఇవి వ్యూహస్థలోని ఇతర అణువులతో చర్యా జరిపి ఉత్పన్నాలను విశ్లేషించుటకు బదులుగా ఒక దానితోనొకటి కలసిపోయి పూర్వపు అణువును విశ్లేషిస్తుంది. ఇలాంటి పరిస్థితులలో క్వాంటం దక్షత "1" కన్నా తక్కువగా ఉంటుంది.

ఉదా: H_2 మరియు Cl_2 ల మధ్య జరిగే కాంతి రసాయన చర్య.

Q. హైడ్రోజన్ మరియు క్లోరిన్‌ల మధ్యా జరిగే కాంతి రసాయన చర్య అధిక క్వాంటం ఫ్రాప్తి అని, H_2 మరియు Cl_2 ల మధ్యా జరిగే కాంతి రసాయన చర్య అల్పక్వాంటం ఫ్రాప్తి కల్గిస్తుందుటకు గల కారణాన్ని వాటి చర్యా

విధానాల పరంగా వివరించండి? (4)

Ans. H_2 మరియు Cl_2 ల మధ్య జరిగే కాంతి రసాయన చర్య అధిక క్వాంటం యోష్టిని కలిగి ఉంటుంది. $H_2 + Cl_2 \longrightarrow 2HCl$; $\phi = 10^6$

ఈ చర్య అధిక క్వాంటం యోష్టిని చర్యా విధానం పరంగా వివరించవచ్చు.

Stage: I :- ప్రాథమిక ప్రక్రియ:

చర్యామిశ్రమంలోని ఒక Cl_2 అణువు క్వాంటం శక్తిని శోషించుకొని పరమాణువులుగా వియోగం చెందును.

$$Cl_2 + h\nu \longrightarrow Cl + Cl \quad \ldots\ldots\ldots (i)$$

Stage II :- పై విధంగా విర్పడిన క్లోరిన్ పరమాణువు చర్యామిశ్రమంలో క్రింది తెలిపిన గొలుసు చర్యలను కలుగజేయును.

$$H_2 + Cl \longrightarrow HCl + H \quad \ldots\ldots\ldots (ii) \qquad \Delta = \text{ఉష్ణమోచక చర్య}$$

$$H + Cl_2 \longrightarrow HCl + Cl \quad \ldots\ldots\ldots (iii)$$

$$HCl + H \longrightarrow H_2 + Cl \quad \ldots\ldots\ldots (iv)$$

ద్వితీయ ప్రక్రియలోని (ii) & (iii) చర్యలు పలుమార్లు జరిగి చర్యామిశ్రమంలోని మొత్తం H_2 మరియు Cl_2 అణువులు ఉత్పన్నాలుగా మార్పు చెందును. ఈ విధంగా ఈ కాంతి రసాయన చర్య "శృంఖల చర్య"గా ఏర్పడుతుంది. కనుక ఇది "అధిక క్వాంటం యోష్టిని" కలిగి ఉంటుంది. చర్యామిశ్రమంలో శోషణం చెందే ఒకే క్వాంటం శక్తి వలన అనేక అణువులు చర్యలో పాల్గొనును.

ఈ చర్య శృంఖల చర్యగా కొనసాగడానికి కారణం, ద్వితీయ ప్రక్రియలోని గొలుసు చర్య (ii) ఉష్ణమోచక చర్య కావడమే.

(ii) H_2 మరియు BCl_2 ల మధ్య జరిగే కాంతి రసాయన చర్య:

ఈ చర్య అల్ప క్వాంటం యోష్టిని కలిగి ఉంటుంది.

$$H_2 + BCl_2 \longrightarrow 2HBCl \quad ; \phi = 0.01$$

ఈ చర్య అల్ప క్వాంటం యోష్టిని దాని చర్యా విధానం పరంగా వివరింపవచ్చును.

Stage I :- ప్రాథమిక ప్రక్రియ:

చర్యామిశ్రమంలోని ఒక BCl_2 అణువు క్వాంటం శక్తిని శోషించుకొని 'Bcl' పరమాణువులుగా విచ్ఛిత్తి చెందును.

$$BCl_2 + h\nu \longrightarrow Bcl + Bcl \quad \ldots\ldots\ldots (i)$$

Stage II :- ద్వితీయ ప్రక్రియ:

ప్రాథమిక ప్రక్రియలలో విర్భవించిన BCl_3 పరమాణువులు చర్యా మిశ్రమంలో క్రింద తెలిపిన గౌణ చర్యలను కలుగజేయవచ్చును.

$$H + BCl_3 \longrightarrow HBCl_3 + H \quad \ldots\ldots (iii)$$
$$BCl_2 + H \longrightarrow HBCl_3 + B \quad \ldots\ldots (iii)$$
$$HBCl_3 + H \longrightarrow H_2 + BCl_3 \quad \ldots\ldots (iv)$$
$$BCl_3 + BCl_3 \longrightarrow B_2Cl_6 \quad \ldots\ldots (v)$$

Δ = ఉష్ణ గ్రాహక చర్య

H_2 మరియు Cl_2 ల మధ్య జరిగే కాంతి రసాయన చర్యావలె ఈ చర్యాలలో కూడా గౌణ చర్యలు (iii) & (iii) పలుమార్లు జరిగి ఈ కాంతి రసాయన చర్య కూడా అత్యధిక క్వాంటం ఫ్రాప్తి వలిగి ఉండవచ్చును. కానీ ఇది అల్ప క్వాంటం ఫ్రాప్తిని కలిగి ఉండును. దీనికి కారణాన్ని ఇలా వివరించవచ్చును.

గౌణ చర్య (iii) ఉష్ణ గ్రాహక చర్య. ఈ చర్య జరుగుటకు అధిక ఉత్తేజిత శక్తి అవసరం అవుతుంది. కనుక సాధారణ పరిస్థితుల వద్ద ఈ చర్య నెమ్మదిగా జరుగుతుంది. కనుక గౌణ చర్యలు (iii) & (iv) కూడా నెమ్మదిగా జరుగును. కనుక ద్వితీయ ప్రక్రియ అనగా ఈ నాలుగు చర్యలలో చర్య V ప్రధానంగా జరుగుతుంది. ఇది ప్రాథమిక చర్యకు విరుద్ధ చర్య. అనగా ప్రాథమిక చర్యలో విర్భవించిన BCl_3 పరమాణువులు ఉత్పన్నాలను విర్భవించుటకు బదులుగా తిరిగి యదార్థపు అణువును విర్భవించును. అందుచేత ఈ కాంతి రసాయన చర్య తక్కువ క్వాంటం ఫ్రాప్తిని కలిగి ఉండును.

ఈ విధంగా ఈ కాంతి రసాయన చర్య తక్కువ క్వాంటం ఫ్రాప్తిని కలిగి ఉండుటకు ప్రధాన కారణం ద్వితీయ ప్రక్రియలలోని గౌణ చర్య (iii) ఉష్ణ గ్రాహక చర్య కావడమే.

Q. క్రింది వానికి లఘు వ్యాఖ్యలు వ్రాయండి?

1). రసాయన దీప్తి
2). యతి దీప్తి
3). స్ఫుర దీప్తి

(1). రసాయన దీప్తి :

సాధారణ ఉష్ణోగ్రత వద్ద జరిగే ఒక రసాయన చర్యలో వెలువడే కాంతిని "రసాయన దీప్తి" అని అంటారు. ఇటువంటి కాంతిని "చల్లటి కాంతి" అని కూడా అంటారు.

క్రింది వానిని రసాయన దీప్తికి ఉదాహరణగా గ్రహించవచ్చు.

i). మిణుగురు పురుగులు ఉద్గారం చేయు కాంతి, రసాయన దీప్తికి ఒక ఉదాహరణ. మిణుగురు పురుగు శరీరం నుండి "ల్యూసిఫెరిన్" అనే

ఒక phosphorescence వెలువడుతుంది. ఇది వాతావరణంలో ఆక్సిజన్ ఎంజైమ్ సమక్షంలో ఆక్సీకరణం చెందుతుంది. ఈ ఆక్సీకరణ ప్రక్రియలో కాంతి వెలువడుతుంది.

iii) రాత్రి వేళల్లో శ్మశాన వాటికలో కనిపించే వెలుపచ్చటి కాంతి రసాయన దీప్తికి ఒక ఉదాహరణ. భూమిలోపల ఎముకలు విచ్ఛిత్తి చెందుట వలన ఫాస్ఫరస్ వాతావరణంలోనికి ప్రవేశించినప్పుడు ఆక్సిజన్ సమక్షంలో P_2O_5 గా ఆక్సీకరణం చెందును. ఈ ఆక్సీకరణ ప్రక్రియలో కాంతి వెలువడుతుంది. దీనిని రసాయన దీప్తికి ఉదాహరణగా గ్రహించవచ్చు.

iii) సముద్రాలలో ఒకరకమైన "జెల్లీ చేప" చనిపోయి విచ్ఛిత్తి చెందే ప్రక్రియలో కాంతి వెలువడుతుంది.

iv) ఒక రకమైన బాక్టీరియా కలిగిన కొయ్యా విచ్ఛిత్తి చెందే ప్రక్రియలో కూడా కాంతి వెలువడుతుంది.

(ii) ప్రతిదీప్తి :

ఒక పదార్థం కాంతిని శోషించుకొని వెలువెంటనే కొంత భాగాన్ని లేదా తరంగదైర్ఘ్యం గల కాంతి రూపంలో ఉద్గారం చేయుటను "ప్రతిదీప్తి" అంటారు. ఈ దృగ్విషయాన్ని ప్రదర్శించు పదార్థాన్ని "ప్రతిదీప్తి జనకం" అంటారు. పదార్థంలో కాంతి శోషణం చెందుటను నిలుపచేసిన వెలువెంటనే ఉద్గారిత కాంతి కూడా ఆగిపోతుంది.

శోషిత కాంతి తరంగదైర్ఘ్యం ఉద్గారిత కాంతి తరంగదైర్ఘ్యం సమానంగా ఉంటే ఆ ప్రక్రియను "రెజొనెన్స్ ఫ్లోరసెన్స్" అంటారు.

ప్రతిదీప్తికి కారణాలు :

పదార్థంలో శోషణం చెందిన కాంతి శక్తి దానిలోని పరమాణువులు లేదా అణువుల నుండి e^- లను తొలగించడానికి తగినంతగా లేనప్పుడు తొలి e^- లు తక్కువశక్తి స్థాయిలనుండి అధిక శక్తి స్థాయిలలోనికి ప్రవేశించును. ఈ e^- లు తిరిగి భూస్థాయిలలోనికి ప్రవేశించినప్పుడు అవి శోషించుకొన్న శక్తినే కాంతి రూపంలో ఉద్గారం చేయును.

యురనైల్ లవణాలు, సల్ఫేట్, క్లోరోఫిల్, Hg భాష్పాలు, ఆంత్రసీన్, క్లోరైడ్ మొ॥ పదార్థాలు ప్రతిదీప్తిని ప్రదర్శించును.

ఉపయోగాలు :

i) ఆహార పదార్థాల స్థితిని తెలుసుకొనుటకు.

2). సుతిపురుగులు (Ring warming) ను గుర్తించుటకు.
3). Lightening ప్రక్రియలో ప్రతిదీప్తి ఉపయోగపడుతుంది.

③. స్ఫురదీప్తి :

ఒక పదార్థం కాంతిని శోషించుకొని దానిలో కొంతభాగాన్ని వెనువెంటనే ఉద్గారం చేస్తుంది. అయితే పదార్థం మీద కాంతి పతనం చెందుటను నిలుపుచేసిన తర్వాత కూడా ఇంతకాలం కాంతిని ఉద్గారం చేస్తూనే ఉండే ఆ దృగ్విషయాన్ని "స్ఫురదీప్తి" అందురు.

"స్ఫురదీప్తిని నిదానంగా జరిగే ప్రతిదీప్తి అనికూడా అందురు."

స్ఫురదీప్తికి కారణం :

ఒక పదార్థంలో శోషణం చెందిన కాంతిశక్తి, దానిలో రసాయన చర్యను కలుగజేయడానికి తగినంతగా ఆ లేనప్పుడు ఎలను భూ స్థాయిలలోనికి ప్రవేశించినప్పుడు అవి శోషించుకొన్న శక్తిని కాంతి రూపంలో ఉద్గారం చేయును. అయితే ఎలు అధిక స్థాయిల నుండి నిదానంగా భూస్థాయిలలోనికి ప్రవేశించును. క్లిష్టమైన నిర్మాణం కల్గిన పదార్థాలు ఈ దృగ్విషయాన్ని ప్రదర్శించును.

ఉదా :-

భారలోహ సల్ఫేటులను మలినాలుగా కల్గిన జింక్ సల్ఫైడ్, క్షారలోహ మరియు క్షారమృత్తిక లోహ సల్ఫేటులు ఈ దృగ్విషయాన్ని ప్రదర్శించును.

ఉపయోగాలు :

1). ఎలక్ట్రిక్ స్విచ్‌ల మీద, వాచ్ డయల్స్ మీద painting చేయుటకు ఉపయోగించే Paints లో స్ఫురదీప్తి పదార్థాలను ఉపయోగిస్తారు.

కిరణజన్య సంయోగక్రియ

పచ్చటి మొక్కలు CO_2 & H_2O ను ఉపయోగించుకొని క్లోరోఫిల్ మరియు సూర్యకాంతి సమక్షంలో కార్బోహైడ్రేట్లను తయారు చేసుకొను ప్రక్రియను "కిరణజన్యసంయోగక్రియ" అంటారు. ఈ ప్రక్రియలో O_2 అణుజనితంగా వెలువడుతుంది. కిరణజన్యసంయోగక్రియ ప్రక్రియలో సూర్యుని వికిరణ శక్తి జరుగుటకు రసాయన శక్తిగా మారుతుంది. ఈ కిరణజన్య-సంయోగక్రియ జరుగుటకు అవసరమైన కాంతి తరంగదైర్ఘ్యం "610 nm – 700 nm".

కిరణజన్య సంయోగక్రియలో CO_2 క్షయికరణం చెంది కార్బోహైడ్రేట్లుగాను, నీరు ఆక్సీకరణం చెంది ఆక్సిజన్‌గాను మారును.

ఈ ప్రక్రియ జరుగుటకు రెండు రకాల చర్యలు తోడ్పడును.

1). కాంతి చర్యలు
2). నిష్కాంతి చర్యలు

① కాంతి చర్యలు :

క్లోరోఫిల్‌లో ఉండే క్లోరోప్లాస్ట్‌లు ఒక క్రమమైన విధానంలో సూర్యకిరణాలను శోషించుకొని H_2O & CO_2 లకు అందించి తద్వారా కార్బోహైడ్రేట్లను O_2 & ATP లను ప్రధానం చేయును.

కిరణజన్య సంయోగక్రియను రెండు విధంగా సూచించవచ్చును.

$$H_2O + CO_2 \xrightarrow[\text{క్లోరోఫిల్}]{h\nu} \text{కర్బన అణువులు} + O_2 + \text{Heat} + \text{ATP}$$

$$\text{క్లోరోఫిల్} + h\nu \longrightarrow \text{క్లోరోఫిల్}$$

$$\text{క్లోరోఫిల్} + 6n\,CO_2 + 6n\,H_2O \longrightarrow n\,C_6H_{12}O_6 + 6n\,O_2 + \text{క్లోరోఫిల్}$$

ఇక్కడ CO_2 గ్లూకోజ్‌గాను మరియు కార్బోహైడ్రేట్‌లుగాను మారుటకు కావలసిన శక్తినే ATP అణువులు అందించును. ఒక గ్లూకోజ్ అణువు సంశ్లేషణ చెందుటకు 18 ATP అణువులు అవసరం.

HETERO CYCLIC COMPOUNDS

HETEROCYCLIC COMPOUNDS:

compounds which contain at least one atom other than carbon in their ring systems, are called heterocyclic compounds and the atom other than carbon that is involved in ring formation is known as hetero atom. The commonly found hetero atoms are N, O, S, P etc.

Ex: Furan, Pyrrole, furan etc.

classification: Based on ring size heterocyclic compounds may be classified as

(i) Five membered rings:

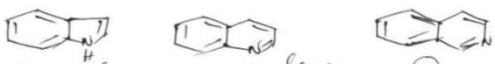
Ex: 

Furan Thiophene Pyrrole.

(ii) Six membered rings:

Pyridine Pyran Thiopyran

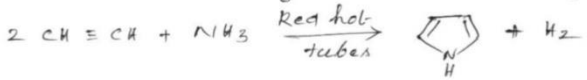
(iii) Fused heterocyclic rings:



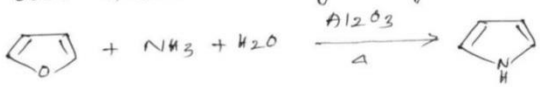
Indole Quinoline Isoquinoline

Preparation:

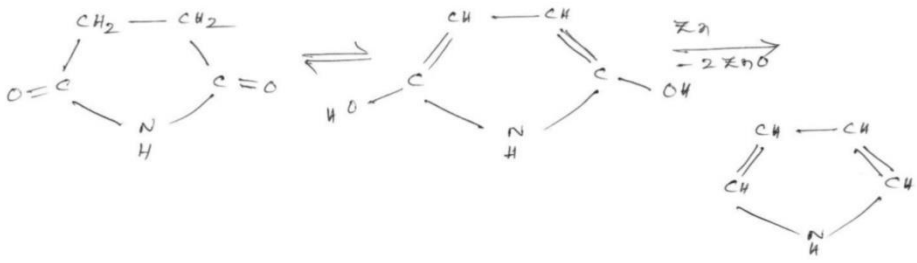
PYRROLE: (1) A mixture of acetylene and ammonia when passed through red hot tubes, pyrrole is formed.



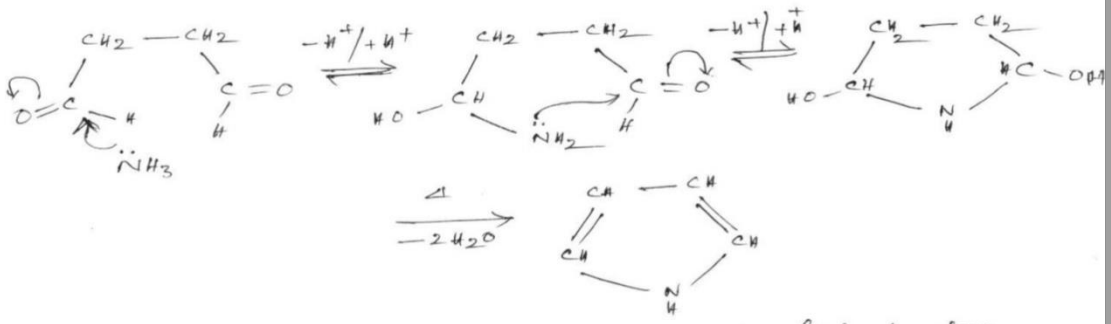
(2) A mixture of furan, ammonia and steam when passed over alumina at high temperature, pyrrole is formed.



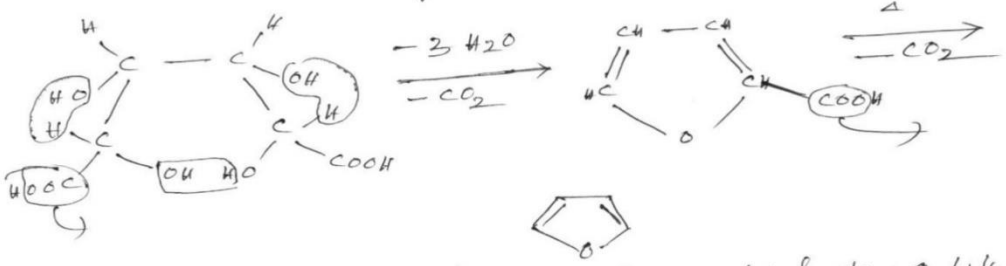
(3). Succinimide when heated with zinc dust forms pyrrole.



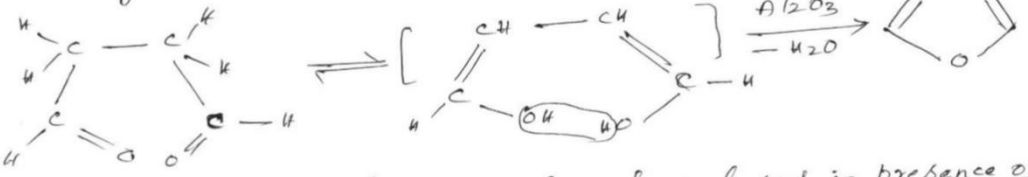
(4) By Paul - Knorr synthesis : 1,4 - Dicarbonyl compound when heated with ammonia forms Pyrrole.



FURAN : (1) Mucic acid when subjected to dry distillation forms furoic acid which on decarboxylation forms furan.

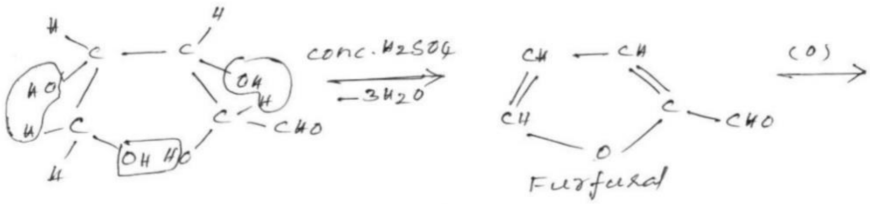


(2) By Paul - Knorr synthesis: Succinaldehyde, a 1,4 - dicarbonyl compound when heated with Al_2O_3 (or) anhydrous Zinc chloride forms furan.

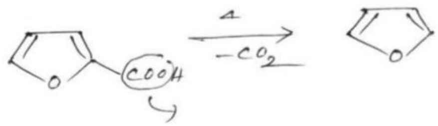


(3) A C_5 - Polysaccharide when hydrolysed in presence of an acid forms a pentose which on dehydration forms furfural. Furfural on oxidation followed by decarboxylation forms furan

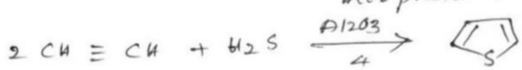
$$(C_5H_8O_4)_n \xrightarrow{H_2O} n\ C_5H_{10}O_5$$



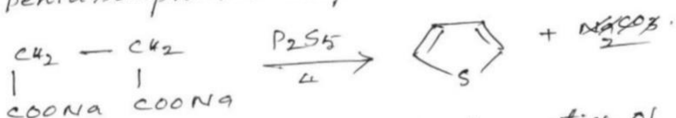
Furfural



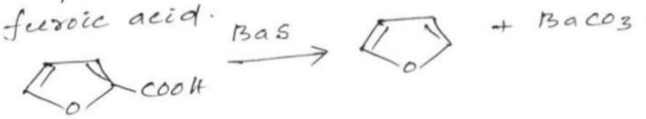
THIOPHENE : (1) a mixture of acetylene and hydrogen sulphide when passed over Al_2O_3 at high temperature, thiophene is formed.



(2). Sodium succinate when heated with phosphorous-pentasulphide, thiophene is formed.



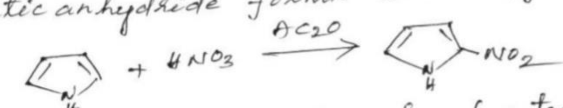
(3). Thiophene is prepared by the action of Bas on furoic acid.



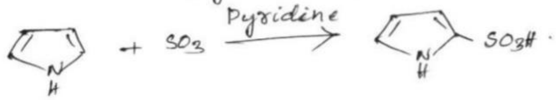
Properties :

PYRROLE : Pyrrole undergoes electrophilic substitution reactions preferably at C_2.

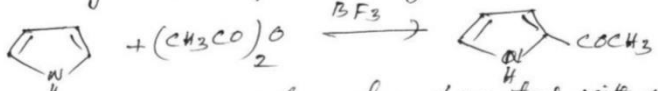
1. Nitration : pyrrole when treated with HNO_3 in presence of acetic anhydride forms 2-nitropyrrole.



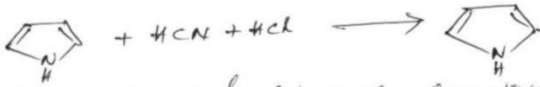
2. Sulphonation : pyrrole when treated with SO_3 in presence of pyridine forms pyrrole-2-sulphonic acid.



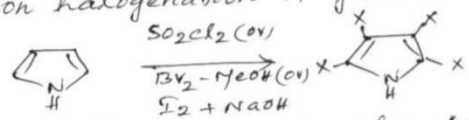
3. Friedel-Crafts acylation : pyrrole when treated with acetic anhydride in presence of BF_3 forms 2-Acetylpyrrole.



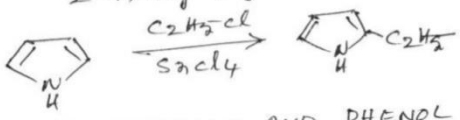
4. Formylation : Pyrrole when treated with a mixture of HCN and HCl forms 2-formyl pyrrole.



5. Halogenation : Pyrrole is more reactive towards halogenation. On halogenation it gives 2,3,4,5 - tetrahalopyrrole.

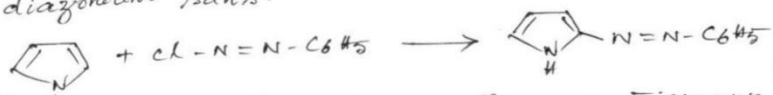


6. Alkylation : Pyrrole when treated with alkyl halide forms 2-Alkylpyrrole.



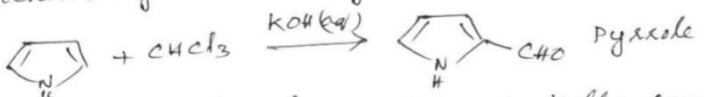
COMPARISON OF PYRROLE AND PHENOL :

(i) Pyrrole like phenol undergo coupling reaction with diazonium salts.



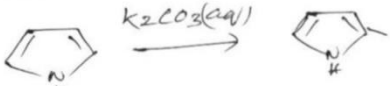
Like

(ii) Phenol, Pyrrole undergoes Reimer-Tiemann reaction



Pyrrole-2-aldehyde

(iii) Pyrrole, like phenols undergoes Kolbe reaction.

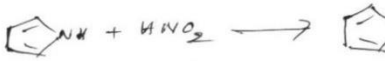


(iv) Pyrrole behaves as a weak acid. It react with strong alkalies to form salts.

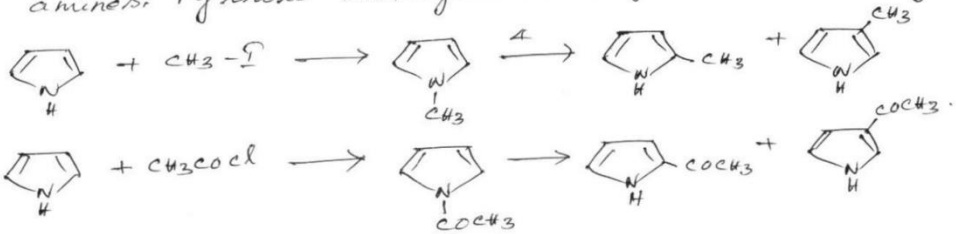
COMPARISON OF PYRROLE WITH AMINES :

1. Like aromatic secondary amines, pyrrole is a weak base. It react with strong mineral acids to form salts.

2. Like 2° - amines, it react with nitrous acid to form N-nitroso amine.

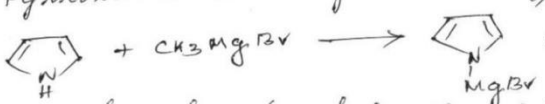


3. Like amines, pyrrole undergoes N-alkylation and N-acylation

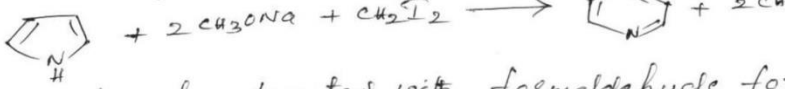


Other Properties of pyrrole :

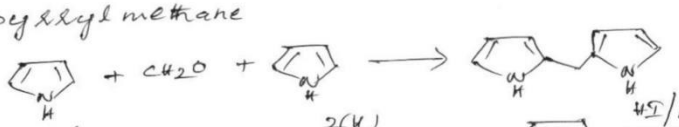
1. Reaction with Grignard reagent : pyrrole on reaction with Grignard reagent forms pyrrole magnesium bromide which on reaction with CO_2 followed by hydrolysis forms pyrrole-2-carboxylic acid.



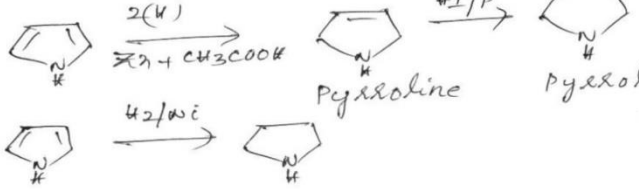
2. Pyrrole when treated with a mixture of sodium methoxide and methylene iodide gives pyridine.



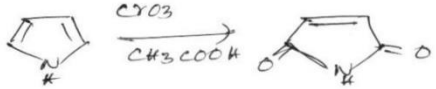
3. pyrrole when treated with formaldehyde forms dipyrrylmethane



4. Reduction :



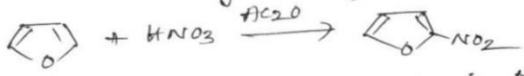
5. oxidation :



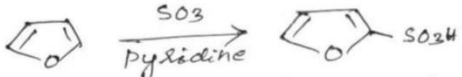
Maleimide.

Properties of Furan :

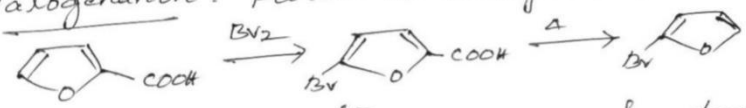
1. **Nitration** : Furan when treated nitric acid in presence of acetic anhydride forms 2-Nitrofuran.



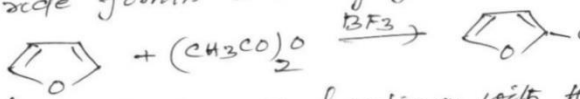
2. **Sulphonation** : Furan react with SO_3 in presence of pyridine to form Furan-2-sulphonic acid.



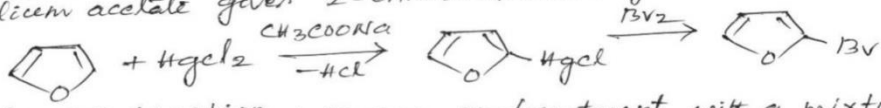
3. **Halogenation** : Furan is halogenated indirectly as follows



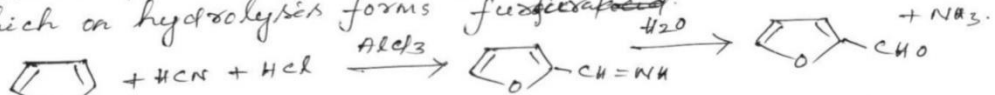
4. **Friedel craft reaction** : Furan when treated with acetic anhydride forms 2-acetyl furan.



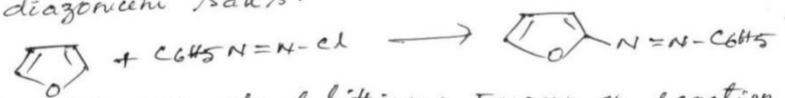
5. **Mercuration** : Furan on heating with $HgCl_2$ and aqueous sodium acetate gives 2-chloromercuric furan



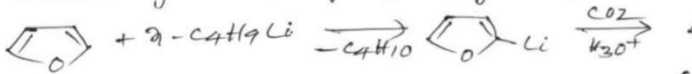
6. **Gattermann reaction** : Furan on treatment with a mixture of HCN and HCl in presence of $AlCl_3$ forms an imine which on hydrolysis forms furfural.



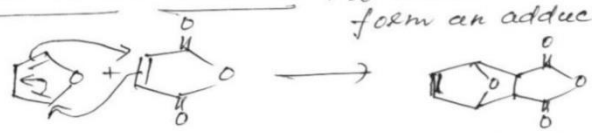
7. **Gomberg reaction** : undergoes coupling reaction with diazonium salts.



8. **Reaction with alkyl lithium**: Furan on reaction with n-butyl lithium followed by hydrolysis forms furoic acid.

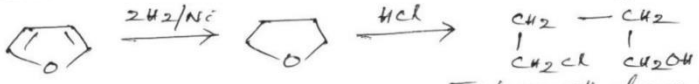


9. **Diels-Alder reaction** : It react with maleic anhydride to form an adduct.



Adduct

10. **Reduction** : Furan when reduced with Ni forms tetrahydrofuran.

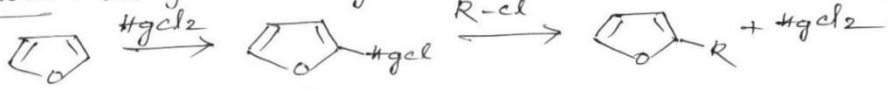


Tetramethylene chlorohydrin.

11. **Addition of oxygen**

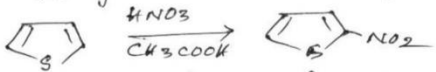
Succinaldehyde.

12. **Alkylation** : In furan alkylation is carried out indirectly

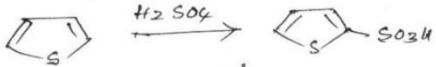


Properties of Thiophene :

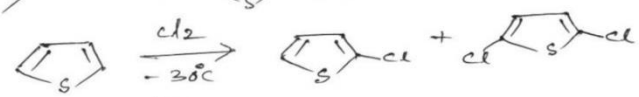
1. **Nitration** : Thiophene react with HNO_3 in presence of acetic acid to form 2-Nitrothiophene.

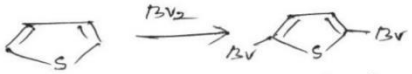


2. **Sulphonation** : Thiophene when treated with cold conc. H_2SO_4 forms 2-thiophene sulphonic acid

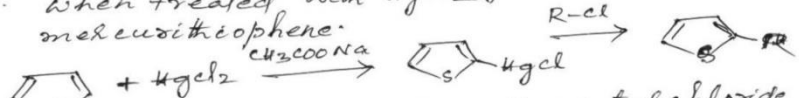


3. **Halogenation** :

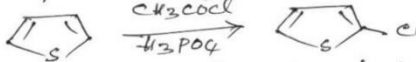




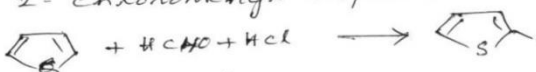
4. **Mercuration** : when treated with $HgCl_2$ forms 2-chloro-mercurithiophene.



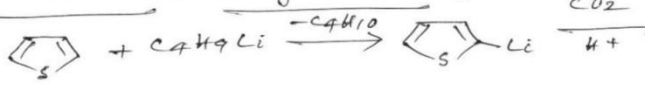
5. **Friedel craft acylation** : when treated with acetyl chloride in presence of H_3PO_4 forms 2-Acetylthiophene.



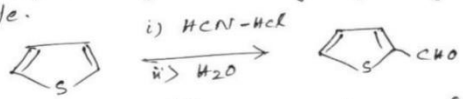
6. **chloromethylation** : when treated with a mixture of HCHO & HCl forms 2-chloromethyl thiophene.



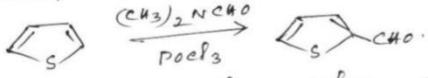
7. **Reaction with Alkyl lithium** :



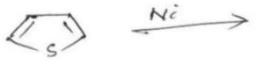
8. Gattermann reaction : Thiophene when treated with a mixture of HCN and HCl followed by hydrolysis gives thiophene-2-aldehyde.



9. Willsmayer reaction : Thiophene when treated with dimethyl-formamide forms thiophene-2-aldehyde.



10. Reduction : Thiophene when reduced with Raney Ni gives n-Butane.

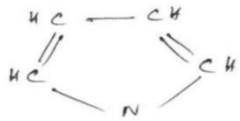


Structure Elucidation of Pyrrole :

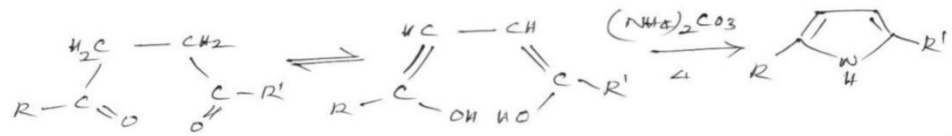
1. From analytical data the molecular formula of pyrrole was found to be C_4H_5N.
2. Pyrrole add one two moles of hydrogen to form an addition product pyrrolidine.

3. This reaction shows the presence of two double bonds in pyrrole.
3. pyrrole, like benzene undergoes electrophilic substitution reactions such as nitration, halogenation, sulphonation etc. These reactions indicate 'aromatic nature of pyrrole'.

From the above properties, the structure of pyrrole is shown as



4. conformation by synthesis : The above structure of pyrrole is conformed by its synthesis.



Limitations of diene structure : The above diene structure fails to explain the following properties.

(1) They behave like benzene and undergo electrophilic substitution reactions.

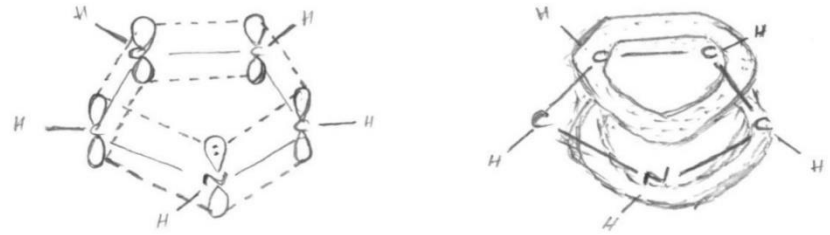
(ii) Like secondary amines, pyrrole lacks of basic character.

(iii) Pyrrole is a flat molecule.

ORBITAL STRUCTURE : The above limitations are explained interms of orbital structure of pyrrole.

In pyrrole all the four carbon atoms and nitrogen atom are sp^2 hybridised. These hybrid orbitals of 'c' and 'N' atoms form C–H, C–C, C–N & N–H sigma bonds.

The four carbon atoms contain a unhybridised p-orbital containing an unpaired electron. Nitrogen atom contains a unhybridised orbital containing a lonepair of electrons. These unhybridised p-orbitals overlap sidewise to form a cyclic delocalised π- molecular orbital containing six electrons embracing all the five atoms of the ring. The delocalised orbital would make up two π-electron clouds, one above and one below the plane of the ring.

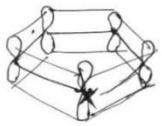


Explanation for the unique behaviour of Pyrrole :

1. As both the p-electrons on N-atom are involved in the formation of delocalised π-electron cloud, hence it lacks basic character.
2. The presence of electron cloud of six delocalised electrons makes the compound aromatic
3. Pyrrole is a flat molecule due to involvement of sp^2 hybrid orbitals
4. It undergoes addition reactions with difficulty, because the double bonds in pyrrole are not true double bonds due to resonance.
5. Pyrrole is a resonance hybrid of the following forms.

COMPARISON OF AROMATIC CHARACTER OF PYRROLE, FURAN & THIOPHENE

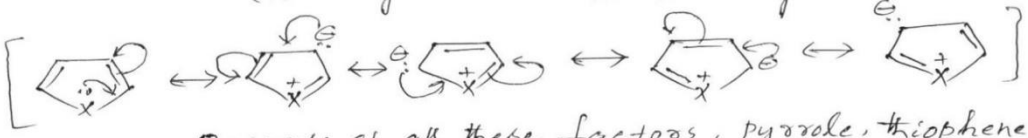
1. The three heterocyclic compounds are flat and planar rings.
2. Each one of these compounds possess 4π electrons and a lone pair of electrons on the hetero atom. These six electrons form an aromatic sextet.



X = N (or) O (or) S.

(3) The three heterocyclic compounds satisfy Huckel's rule ie they have $(4n+2)$ π electrons.

(4) They have resonance stability.



Because of all these factors, pyrrole, thiophene and furan exhibit aromatic character. But they have different degree of aromaticity.

Oxygen atom being more electronegative less readily contributes its lonepair of electrons in the formation of aromatic sextet. Hence furan is less aromatic than pyrrole and thiophene.

Sulphur being less electronegative, it can readily donates its lonepair of electrons in the formation of aromatic sextet. Hence thiophene is more aromatic than furan & pyrrole. Therefore the order of aromatic character of these three heterocyclic compounds is shown as

Thiophene > Pyrrole > Furan.

COMPARISON OF ORDER OF REACTIVITY:

The reactivity of furan, pyrrole and thiophene follows the following order: Pyrrole > Furan > Thiophene > Benzene.

The reactivity order of the above three heterocyclic compounds depends upon the relative abilities of the hetero atoms to accommodate the positive charge in the intermediate carbonium ion.

Nitrogen atom accomodates the positive charge more readily than oxygen atom and sulphur atom. Hence pyrrole is more reactive than furan and thiophene.

The sulphur atom of thiophene though less electronegative than oxygen, its 3p-orbital overlap less efficiently

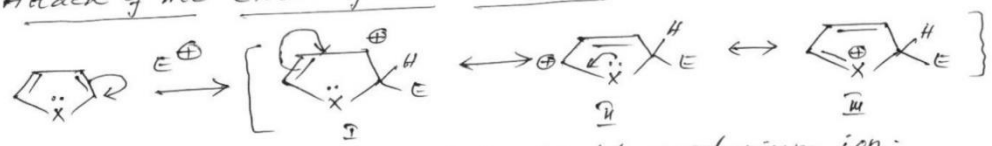
With the 2p-orbital of the aromatic π-electron system. Hence thiophene is less reactive than furan. 11

Electrophilic Aromatic Substitution:

Pyrrole, furan, thiophene are aromatic. Hence they undergo aromatic substitution reactions. They undergo electrophilic substitution predominantly in the position-2 than position-3 of the ring.

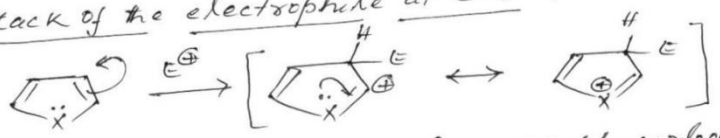
The carbonium ion formed from the attack of the electrophile at C-2 has more important resonance structures and therefore more stable, than the carbocation resulting from attack at C-3. Hammond's postulate suggests that, the reaction involving the more stable intermediate will be faster.

Attack of the electrophile at C-2 :



More stable carbonium ion.

Attack of the electrophile at C-3 :



Less stable carbonium ion.

Therefore electrophilic substitution takesplace at position-2 because it results in more stable carbocation.

PYRIDINE : Isolation from coal tar : Furan is present in the light oil fraction of coal tar. When light oil fraction is treated with dil H_2SO_4, pyridine and all other basic substances pass into solution as their sulphates. Acid layer is separated and treated with NaOH solution to regenerate pyridine and other bases. Pyridine is obtained from the mixture of bases by fractional distillation.

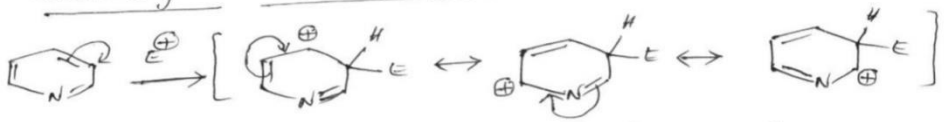
Properties of Pyridine :

1. **Basic nature** : It is a strong base. When treated with alkyl halide in presence of moist silver oxide yield N-Alkyl pyridinium hydroxide.

$$C_5H_5N + CH_3I \longrightarrow [C_5H_5-\overset{\oplus}{N}-CH_3]\,I^- \xrightarrow{AgOH} [C_5H_5-\overset{\oplus}{N}-CH_3]\,OH^-$$

2. **Electrophilic substitutions** : It undergoes electrophilic substitutions preferably at position - 3 because attack of electrophile at position - 3 result in a more stable carbonium ion.

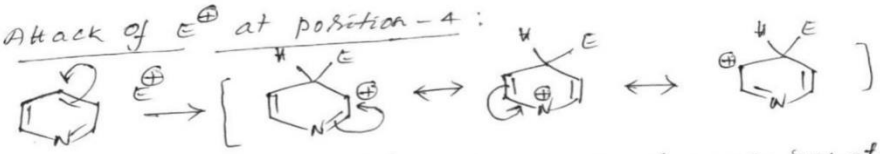
(i) Attack of $E^{\oplus}$ at position - 3



more stable carbonium ion

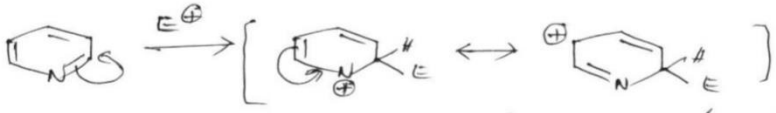
But attack of electrophile at position - 4 (or) position - 2 result in a less stable carbonium ion.

(ii) Attack of $E^{\oplus}$ at position - 4 :



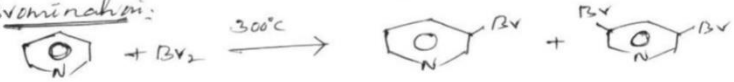
less stable carbonium ion because 'N' atom less readily accomodate +ve charge.

(iii) Attack of $E^{\oplus}$ at position - 2 :

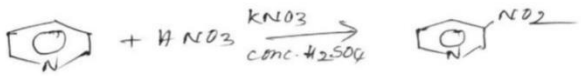


less stable carbonium ion because 'N' atom carries +ve charge.

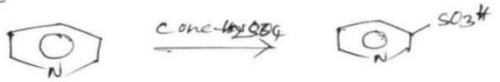
Examples : **Bromination** :



Nitration :

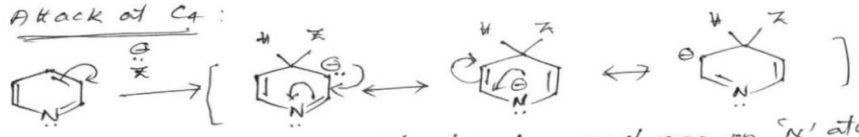


Sulphonation :



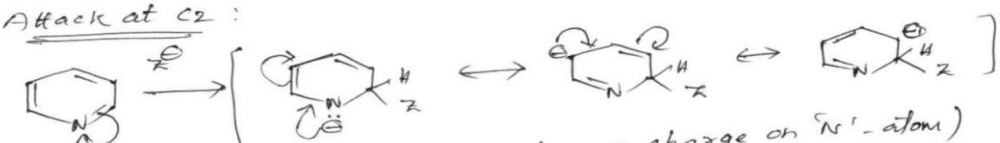
Nucleophilic substitution reactions : The electron withdrawing nature of nitrogen makes pyridine less reactive towards electrophilic substitutions and more reactive towards nucleophilic substitutions. Nucleophilic substitutions occur at positions 2 & 4 because the resultant carbanions are resonance stabilised. The negative charge on the nitrogen atom makes the carbanion extra stable.

Attack at C_4 :



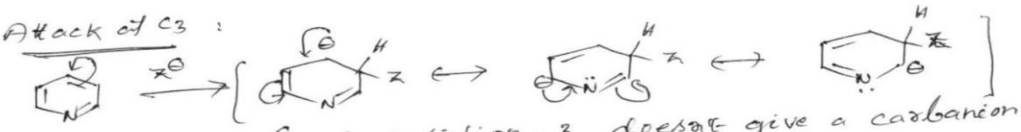
Extra stable due to -ve charge on 'N' atom.

Attack at C_2 :



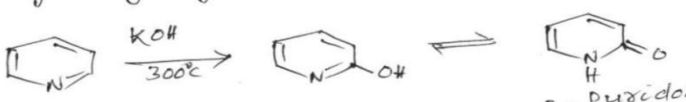
Extra stable due to -ve charge on 'N'-atom)

Attack at C_3 :



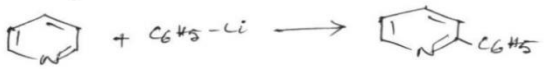
∴ Attack of $Nu^{\ominus}$ at position-3 doesn't give a carbanion that carry -ve charge on the nitrogen atom.

Examples : (1) Reaction with KOH : on heating with KOH 2-Hydroxy pyridine is formed.



2-Pyridone.

(2) Reaction with phenyl lithium : pyridine when treated with Phenyl lithium forms 2-Phenyl pyridine



3. Reaction with sodamide : pyridine when treated with sodamide forms 2-Aminopyridine. This reaction is called chichibabin reaction.

Mechanism:

4. Question bank

The respective class teachers provide the slow learners chapter wise important question bank which is useful an in-depth analysis of questions.

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

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1. **Unit wise important questions provided for slow learners**

I B.Sc – I SEMESTER - PAPER – I

(INORANIC & ORGANIC CHEMISTRY)

UNIT – I

Chemistry of p-Block Elements – I

1. Define Borane. Write the synthesis, structure, properties and uses of diborane.
2. Write the synthesis and structure of boron-nitrogen compounds
3. What are silicones? How are they classified?
4. Write the preparation and reaction of
   a. hydrazine b. hydroxyl amine
5. What are silanes? Write two preparation methods of silanes.
6. Write the general characteristics of 13 group (Boron family) elements

UNIT – II

Chemistry of p-Block Elements – II

1. Write the classification of oxides based on
   i. Chemical behaviour '
   ii. Oxygen content
2. What are inter halogen compounds? How they are classified?
   Give their general preparation
3. Write about pseudohalogens and psuedohalides and its uses?

Organo Metallic Chemistry

1. Write the Classification of Organo Metallic compounds.
2. Write the preparation, properties and applications of Alkyls of lithium
3. Write the preparation, properties and applications of Alkyls of Magnesium

UNIT – III

Structural Theory in Organic Chemistry

1. Write explanatory notes on:
   i. Inductive effect
   ii. Mesomeric effect?
2. What do you understand by the term resonance? What are the different conditions for writing resonance structures?
3. Use mesomeric effect to explain
   a. Acidity of phenol
   b. Acidity of carboxylic acid.
4. What is carbine? How we explain their stability?
5. What are the three major types of organic reactions? Explain giving examples in each case.
6. What types of bond fission possible for organic compounds?

UNIT – IV

Cyclic Hydro Carbons:

1. Suggest a possible reason why peroxide effect is for HBr and not for HCI.
2. Compare and explain the relative rates of addition HCI, HBr and HI to alkenes.
3. What do you know about Anti-Markovnikov effect?
4. Write a brief note on Diels-Alder reaction.
5. Why are conjugated dienes more stable than non-conjugated dienes?
6. Give two properties each of electophilic and nucleophilic additions to acetylene along with their mechanisms.
7. Explain: Acetylene is more acidic then ethane.
8. Acetylene forms metal acetylides but dimethyl acetylene does not form such derivatives.

Cyclo Alkanes:

1. What factors make the cyclohexane ring the most stable, abundant and important ring structure in organic chemistry? Cyclopropane undergoes ring opening reactions. Explain why?
2. Cyclopropane undergoes ring opening reactions. Explain why?

3. How will you calculate the percentage of chair and boat forms?
4. What are applications and limitations of Baeyer's stain theory?

Benzene and its reactivity

1. State Huckel's rule. Illustrate with suitable examples.
2. Write the mechanism of nitration of benzene.
3. Discuss the orbital structure of benzene.
4. Describe the Friedel-Crafts acylation. Give its mechanism
5. Define Huckel's rule. What do you understand by the term aromaticity? Give examples.
6. State Hucket's rule and explain the stability of:
   i. tropyllium cation
   ii. Cyclopentadienyl anion.
   iii. What is the cause of aromatic character of benzene in terms of Huckel's rule?
7. Discus the mechanism of electrophilic substitution of benzene.
8. What is the cause of aromatic character of benzene in terms of Huckel's rule?
9. What is aromaticity?

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S.G.S. ARTS COLLEGE : TIRUCHANOOR ROAD: TIRUPATI

DEPARTMENT OF CHEMISTRY

I B.Sc – II SEMESTER - PAPER – II

(PHYSICAL & GENERAL CHEMISTRY)

UNIT – I

Solid State:

1. Explain the laws of crystallography.
2. Define space lattice, lattice point and lattice plane.
3. What is Unit cell? Explain the types of unit cells.
4. What are Crystal systems? Explain.
5. Explain Bragg's equation.
6. State and Explain the law of symmetry.
7. What are Miller indices? Give the Miller indices of different planes in a simple cube.
8. Discus about the application of Bragg's equation in studying the crystal structure.
9. Define Crystal defect. Explain the various types of crystal defects.

UNIT- II

Gaseous State:

1. Derive vander Waal' s equation of state.
2. Define critical constants TcPc and Vc
3. Derive the relationship between critical constants and van der Waal's constants.
4. Explain the law of corresponding state.
5. Explain the structural differences between solids, liquids and gases.

Liquid State:

6. Write the properties of liquid crystal.
7. Write the applications of liquid crystals.

UNIT – III

Solutions:

1. Explain Henry's law.
2. Discuss with examples, the solutions which show positive and negative deviation from Raoult's law.

3. Define critical solution temperature (CST). Write the classification of partially miscible liquid pairs depending on their CST values.
4. Discuss the principle of stream distillation.
5. Explain Nernst distribution law.

UNIT – IV

Colloids & Surface Chemistry:

1. Write the preparation of colloids.
2. How colloids can be purified?
3. Stability of colloids.
4. Write a short note on zeta potential.
5. What are the applications of colloids.
6. Electro osmosis
7. Derive the Langmuir adsorption isotherm and mention its applications.
8. Explain Freundlisch adsorption isotherm.
9. Explain the factors influencing the adsorption of gases on solids.
10. Write the postulates of Valence bond theory.

Chemical Bonding:

1. Explain the valence bond theory.
2. Write a short note on Dipole movement.
3. Write molecular orbital diagrams for homo nuclear moleculs, Nitrogen and oxygen
4. Write molecular orbital diagrams for hetiro nuclear moleculs, Carbon monoxide and nitric oxide.
5. Write the postulates of LCAO methods

UNIT – V

Stereo Chemistry of Carbon Compounds:

1. Write to molecular representations.
2. Write about stereo isomerism.
3. Define asymmetric (or) chiral carbon. What are the conditions required for a carbon compound to show optical isomerism?
4. What are the differences between enantiomers and diastereomers?
5. Write the optical isomers of a) glyceraldehydes b) lactic acid and c) alanine.
6. Discuss the optical isomerism exhibited by tartaric acid.
7. Explain R, S configuration with example.
8. Explain D, L configuration with examples.

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DEPARTMENT OF CHEMISTRY

II B.Sc – III SEMESTER - PAPER – III

(INORGANIC & ORGANIC CHEMISTRY)

UNIT – I

Chemistry of d-Block Elements – I

1. Write briefly about characteristics of d-block elements with special reference to electronic configuration
2. Write about oxidation states of d-block elements
3. Write about magnetic properties of d-block elements
4. Write about catalytic property and complex formation of d-block

Theories of Bonding in Metals

5. Explanation of properties of metals by valence bond theory
6. Write briefly about hand theory of metals
7. Write a short note on free electron theory of metals?
8. Explain the properties of metals by VB Theory.

UNIT – II

Metal carbonyls and related compounds

1. What ate metal carbonyls write about classification of metal carbonyls
2. Write the preparation, properties and structure of nickel tetra carbonyl
3. Write the preparation, properties and structure of iron penta carbonyl
4. Define and explain effective atomic number rule with examples
5. Write the preparation, properties and structure of chromium hexa carbonyle
6. Write the preparation, properties and structure of $[Mn_2CO_{10}]$
7. Write the preparation, properties and structure of $[Fe_2CO_9]$
8. Write the preparation, properties and structure of $[Co_2CO_8]$
9. Write the preparation, properties and structure of $[Fe_3CO_{12}]$

Chemistry of f-Block Elements – I

1. What are inner transition elements? Write the electronic configuration of lanthanides and actinides?

2. What are Lanthanides? What is meant by Lanthanide contraction? What is the cause of Lanthanide contraction?
3. What are the similarities and differences of Lanthanides and Actinides? Write the comparison of Lanthanides and Actinides.

UNIT -III

I. ***Halogen Compounds***

1. SN^1 and SN^2 reaction mechanism with examples
2. Differences between SN^1 and SN^2reactions

II. ***Hydroxy Compounds***

3. Grignard synthesis of alcohols
4. *Special reactions of Phenols*
   - (i) Bromination
   - (ii) Azocoupling
   - (iii) Pinacol-Pinacoline rearrangement
   - (iv) Esterification

Name reactions

- (i) Kolbe – Schmidt
- (ii) Riemer – Tiemann reaction
- (iii) Fries rearrangement
- (iv) Schotten Baumann reaction

UNIT - IV

Carbonyl Compounds

5. Nucleophillic addition reactions
6. (i) Aldol condensation
   (ii) Cannizzaro's reaction
   (iii) Perkin reaction
   (iv) Benzoin condensation
   (v) Haloform reaction
   (vi) Knoevenagel reaction
   (vii) Baeyer-villager oxidation
   (viii) Clemmenson reduction, wolf-Kishner reduction,

MPV reduction, reduction with LiH4 and NaBH4.

7. 2, 4 DNP Test, Tollen's test, Fehling test, Schiff's test and Haloform test

UNIT - V

Carboxylic acids and derivatives

8. (i) Hydrolysis of esters (ii) Carbonation of Grignard reagents
9. (i) Hydrogen bonding (ii) Relative differences in the acidities of aromatic and aliphatic acids
10. Huns Diecker reaction, Decarbonylation by Schimdt reaction, Arrndt – Eastert synsthesis, Hell-Voll hard-Zelinsky reaction

Active Methylene Compounds.

(i) Malonic ester and Acetoacetic ester preparation and Synthetic applications

S.G.S. ARTS COLLEGE : TIRUCHANOOR ROAD : TIRUPATI

DEPARTMENT OF CHEMISTRY

II B.Sc – IV SEMESTER - PAPER – IV

SPECTOROSCOPY AND PHYSICAL CHEMISTRY

UNIT – I

Electronic Spectroscopy :

1. Explain the difference types of Molecular Spectra
2. Write the difference types of electronic transition
3. Write a short note on Chromophore auxochrome

Infrared Spectroscopy:

1. Write a note on modes of vibrations in diatomic molecules
2. Write the interpretation of Spectra-Alcohols

Proton Magnetic Resonence Spectroscopy

1. Write a note on equivalent and non equivalent crotons
2. Explain the principles nuclear magnetic resonance spectroscopy
3. Explain the chemical shift with examples
4. Write a short on the following
    a. Splitting of signals
    b. Spin spin coupling
    c. Coupling constant
5. Explain the NMR data for acetal dehyde and ethyl acetate

UNIT – II

Spectrophotometry

1. Write a short on the following
    a. Transmittance
    b. Absorbance
    c. Molar Absorptivity
2. Explain single beam spectrophotometer
3. Explain double beam spectrophotometer

4. Explain Berlambeer-lamberts law and its limitations
5. Explain the application of beer lambers law for quantitative analysis of chromium in potassium dichromate solution
6. Explain the application of beer lambers law for quantitative analysis of chromium in Manganese in mangamous sulphate

UNIT –III

Dilute Solutions :

1. Explain roults law for lowering of vapour pressure how it is useful for determination of the molecular weight of a non volatile solute
2. What are the colligative properties explain each of them with suitable examples
3. Described cottrele's method of determining molecular weight of non volatile solute by using elevation of boiling point
4. What is osmosis define osmotic pressure, explain the determination method for osmotic pressure
5. How is molecular weight of a non volatile solute determined by using Ostwald dynamic method on basis of relative lowering
6. Define an ideal solution and dilute solutions
7. Explain vanthoffs laws of osmotic pressure
8. Write a short note on
    a. Semi permeable membrane
    b. Abnormal colligative property

UNIT – IV

Electro Chemistry - I

1. What are transport numbers, describe hittarfs method for the determination of transport number
2. State and explain kohlrauschs law
3. What is Conductometric Titrations, explain various types of Conductometric Titrations
4. Write a short note on
    a. Specific resistance
    b. Specific Conductance

c. Equi valent conductance
d. Molar conductance

5. Explain Debye-Huckel-Onsagars theory of strong electrolytes
6. Discuss about Nernst equations
7. State and explain Ostwald's dilution law
8. Explain the reversible and irreversible cells
9. Explain Arhenius theory of electrolytes
10. How do you determine the EMF of a cell write any one application of EMF measurements

UNIT – V

Electro Chemistry – II

1. Discuss about Potentiometric titrations and give their advantages
2. What single electrode potential write the construction and working of standard hydrogen electrode
3. How Potentiometric titrations used to find the P^H of acid solution
4. What is standard electrode potentian explain calomel electrode with cell reactions
5. Write the principles and applications of indicator electrode
6. Discuss about electrolytes

Phase Rule

1. State and explain the gibbs phase rules , explain the phase diagram of Ag-Pb system
2. Explain the phase diagram of single component system
3. Explain the phase diagram of solid liquid system
4. Explain the phase diagram of Magnesium-Zinc system
5. Write short note on
   a. Freezing mixture
   b. Triple point
   c. Eutectic point
   d. Terms in phase rule
   e. Condensed system
   f. In congruent melting point

S.G.S. ARTS COLLEGE : TIRUCHANOOR ROAD : TIRUPATI

DEPARTMENT OF CHEMISTRY

III B.Sc – V SEMESTER - PAPER – V

UNIT - I

Coordination Chemistry

1. Explain Warner's theory of complex compounds
2. Explain valence bond theory in complex compounds
3. Explain high spin and low spin complexes with an example each
4. Discuss the salient features of crystal field theory
5. Write notes on the stereoisomerism of complex compounds with co-ordination number 4.
6. Explain the EAN rule. Give any two complexes which donot obey this rule.
7. Explain the formation of the following complexes basing on valence Bond theory.
   i. $[C_0F_6]^{3-}$
   ii. $[Zn\ (HN_3)_4]^{2+}$
   iii. $[Ni(CO)_4]$
   iv. $[Cu(NH_3)_4]$
8. What is meant by Crystal Field Stabilisation Energy
9. Define coordination number in a complex. What is the coordination number of
   i. Cobalt in $[Co(NH_3)4CIBr]$
   ii. Nickel in $[Ni(DMG)_2]$
10. What is ligand? Give examples
11. Calculate the EAN of the central Atom the following:
    i. $[Co(NH_{3)6}]^{3+}$
    ii. $[Cr(CO)_6]$
    iii. $K_6[Fe(CN)_6]$
12. Apply the concept of EAN rule on the complexes $[Fe(CN)_6]^{4+}$ and $[Fe(CN)_6]^{3-}$. What conclusion you draw about the validity of the EAN rule from these two well-known complexes?

UNIT – II

Special and magnetic properties of metal complexes stability of metal complexes

1. Describe how magnetic moment can be determined by Gouy's balance
2. What is stability constant? What are the factors that determine the stability of a complex
3. Describe Job's method of determining the composition of a metal complex.
4. Explain the term Chelates.
5. Explain the kinetic stability and thermo-dynamic stability of a complex.

UNIT – III

Nitro Hydrocarbons

1. Distinguish between alkyl nitrites and nitroalkanes.
2. How is nitrobenzene prepared? Give its reducing properties
3. Write note on:
   a. Mannich reaction
   b. Michael addition
   c. Michael Reduction
4. Explain Nef reaction
5. Explain the acidic nature of nitro alkanes.

UNIT _ IV

Nitrogen Compounds

1. a. What are amines" Give general methods of preparation of primary amines.
   b. Give a brief account of the properties of amines
2. Describe a method of separation of primary, secondary and tertiary amines in a mixture. (Hinsberg Method)
3. Write about the Basic characters of amine
4. Write about Hoffman's bromamide reaction
5. Write about the aromatic electrophilic substitution.

UNIT _ V

Thermodynamic

1. a. State the first law of thermodynamics in its different forms. Give the mathematical expression of the law.

   b. Calculate the work done in calories when 3 moles of an ideal gas at 25^0C expands isothermally and reversibly from a volume of 10 litres to a volume of 50 litres.
2. Derive the relationship between $H\triangle$ & $E.\triangle$
3. Explain Joule –Thomson effect.
4. Distinguish between thermodynamically Reversible and Irreversible process.
5. a. Derive an expression for the work done in the reversible isothermal expansion of an ideal gas.

   b. Calculate the work done when 1ogms. of an ideal gas of molecular weight 44 expands isothermally and reversible from a volume of 5 litres to 10 litres at 27^0C.
6. Which of the following is function of state. Why?
   i. Internal energy.
   ii. Work done by a gas.
7. What is a spontaneous process? Explain with examples.
8. State and explain Carnot theorem

V YEAR – SEMESTER-V – PAPER-VI

UNIT - I

Reactivity of Metal Complexes

1. What are labile and inert complexes? Explain with necessary examples.
2. How many types are nucleophilic ligand substitution reaction in complexes? Explain them with examples and necessary mechanism.
3. What are the theories and application of Trans effect in metal complexes.
4. Name the factors which influence the rate of square planar substitution reaction.

Bio-inorganic Chemistry

1. What are the elements required for the body? Name their give their percentages.
2. Write about the structure and functions of heamoglobin.
3. What is the role of Mg^{2+} ions in biochemical processes and in photosynthesis?
4. What are metal porphyrins?

UNIT – II

Chemical Kinetics

1. Derive the equations for the specific rate 'k' of a second order reactions $2A \rightarrow P$
2. What is Arrhenius equation? Explain the concept of activation energy of reaction.
3. Give the integrated equation for the rate constant of a first order reaction. Give two examples of first order reactions.
4. What is half-time ($t_{1/2}$) of a reaction? Give equations for its value for first and second order reactions.

Photo Chemistry

1. What is Grothus-Draper law? Explain with examples
2. State and explain Stark Einstein's law of photo chemical equivalence.
3. What are the factors that influence the value? Explain in detail the reasons for higher values.
4. What are flllluorescence and phosphorescence? Explain with the help of Jablonsky diagram.

5. What is Jablonsky diagram? Illustrate it briefly.
6. What is Grothus-Draper law?
7. What is Stark-Einstein's law of phot chemical equivalence? Explain.
8. What is fluorescence? Ecplain
9. What is phosphorescence? Explain.
10. What are photo-sensitized reaction? Give examples.

UNIT – III

Heterocyclic Compounds

1. Write about the aromatic character of pyrrole, furan and thiophene in detail.
2. Discuss in detail the acidic character of pyrrole.
3. Explian Diels-Alder reaction of furan with an example.
4. Give two methods of preparation of pyridine with equations.
5. What are heterocyclic compounds? Give examples.
6. Explain the aromatic character of pyrrole.
7. Explain the aromatic character of furan.

UNIT - IV

Carbohydrates

1. What are carbohydrates? How are they classified? Give examples.
2. Discuss the cyclic structure of glucose.
3. Discuss the structure of Fructose.
4. How are the following conversions effected?
    i. Aldo hexose to Aldo pertose
    ii. Aldo pentose to Aldo hexose
    iii. Keto hexose to Also hexose
    iv. Also hexose to Keto hexose
5. What are carbohydrates? Give two examples
6. What is mutarotation? Explain.
7. How do you prove that there are 5 —OH groups in Glucose? Explain.

UNIT – V

Amino Acids and Proteins

1. What are amino acids? How are they classified? Give examples.
2. Discuss the structure of proteins with necessary explanations.
3. How are amino acids separated from proteins and identified?
4. How are amino acids prepared from a-halogen acids?
5. What is Zwitter ion? Explain.
6. What are the different structures of protein?

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S.G.S. ARTS COLLEGE : TIRUCHANOOR ROAD : TIRUPATI

DEPARTMENT OF CHEMISTRY

III B.Sc – VI SEMESTER - PAPER – VI(A)

ANALYTICA METHODS IN CHEMISTRY

UNIT - I

Quantitative Analysis:

1. Write the principles of Volumetric analysis.
2. Write the Theories of Acid base Titrations.
3. Write the Theories of Redox Titrations.
4. Write the Theories of Complexo metric Titrations.
5. Write the Theories of Iodometric Titrations.
6. Write the Theories of precipitation Titrations.
7. Write a short note on
    a. Indicators
    b. Precipitation
    c. Coagulation
    d. Peptization
    e. Co-precipition
    f. Post precipitation
    g. Digestion
    h. Filtaration
    i. Drying and Ignition
8. Explain the steps involved in Chemical Analysis

UNIT – II

Treatment of Analytical Data:

1. Write the types of errors.
2. Write a short note on
    a. Significant figures.
    b. Accuracy.
    c. Precision.
    d. Standard Deviation.
    e. Confidence Limit.

UNIT – III

Separation Techniques of Chemical Analysis

1. Explain solvent extraction.
2. Explain batch extraction.
3. Explain continuation extraction.
4. Explain counter current extraction.
5. Write a short note on synergism.
6. Write the action of Ion exchange resins.
7. Write the separation of Inorganic mixtures.

UNIT – IV

Chromatography

1. Give the principle involved in chromatography.
2. Define the term 'Eluent'
3. What do you understand by the term 'R_f value'?
4. What is the advantage of descending paper chromatography over ascending paper chromatography?
5. What do you understand by the term 'Retention factor'? How does it help in the identification of components in a mixture?
6. Explain the term 'Paper' and how it acts as a chromatographic medium in paper chromatography.
7. a. Explain the principle of column chromatography.
   b. How will you proceed to separate a mixture of glycine and alanine?

UNIT – IV

Thin layer Chromatography

1. Write the basic principles and application of TLC.
2. Write the basic principles and application of HPLC
3. Write a short note on $R_f$value.
4. Adsorbedents.
5. Detection of the spots.
6. Preparation of plates.
7. Development of Chromatogram

HEAD
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5. Additional assignments and Tests

As a part of remedial work the students are given additional tests and assignments to reinforce the concept learnt in the class and improve performance in the university examinations.

Tests conducted for slow learners

Class test for slow learners conducted on 23-12-2019

SEMESTER-II

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Explain shottky and Frenkenel defects.
2. Write notes on Tyndall effect.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Derive Bragg's equation. Discuss its application in studying NaCl crystal structure.

or

b. Write a brief note on Langmuir adsorption isotherm.

4. a. What are differences between solid and liquid crystals. Write the application of liquid crystals.

or

b. Discus the application. of Nernst Distribution Law.

S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD: TIRUPATI :

DEPARTMENT OF CHEMISTRY ***ACADEMIC YEAR: 2019-20***

Tests conducted for slow learners

Class test for slow learners conducted on 02-03-2020

SEMESTER-IV

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Discuss selection rules of IR spectra.
2. Write the basic principle of NMR spectroscopy.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Explain different modules of vibrations of non linear molecules in IR with two examples.

or

b.Write a brief note on NMR splitting of signals.

4. a. i) Explain the finger print region of IR spectra.

ii) Explain the IR spectra of an harmonic oscillator.

or

b. What is chemical shift? How to measure it? What are the factors affecting chemical shift.

S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD: TIRUPATI :

DEPARTMENT OF CHEMISTRY ***ACADEMIC YEAR: 2019-20***

Tests conducted for slow learners

Class test for slow learners conducted on 02-03-2020

SEMESTER-VI

Max Marks : 25

Section – A

Answer any one question (1x5=5 Marks)

1. Briefly explain a. Precipitation and b. Co-precipitations
2. Explain the theories of Acid base Titrations.

Section – B

Answer all Questions (2x10=20 Marks)

3. a. Write the classification of errors.

or

b. Explain the following i) Accuracy ii) Precision and iii) Standard deviation

4. a. Write the principles of Volumetric analysis.

or

b. Write brief notes on i) Indicators ii) Coagulation

A sample copy of remedial class test papers

Semester: II

The appearance of dust particles in a semi-darkened room, when a sun beam enters are the well known examples of Tyndall effect. The dust particles are large enough to scatter light which renders the path of light visible.

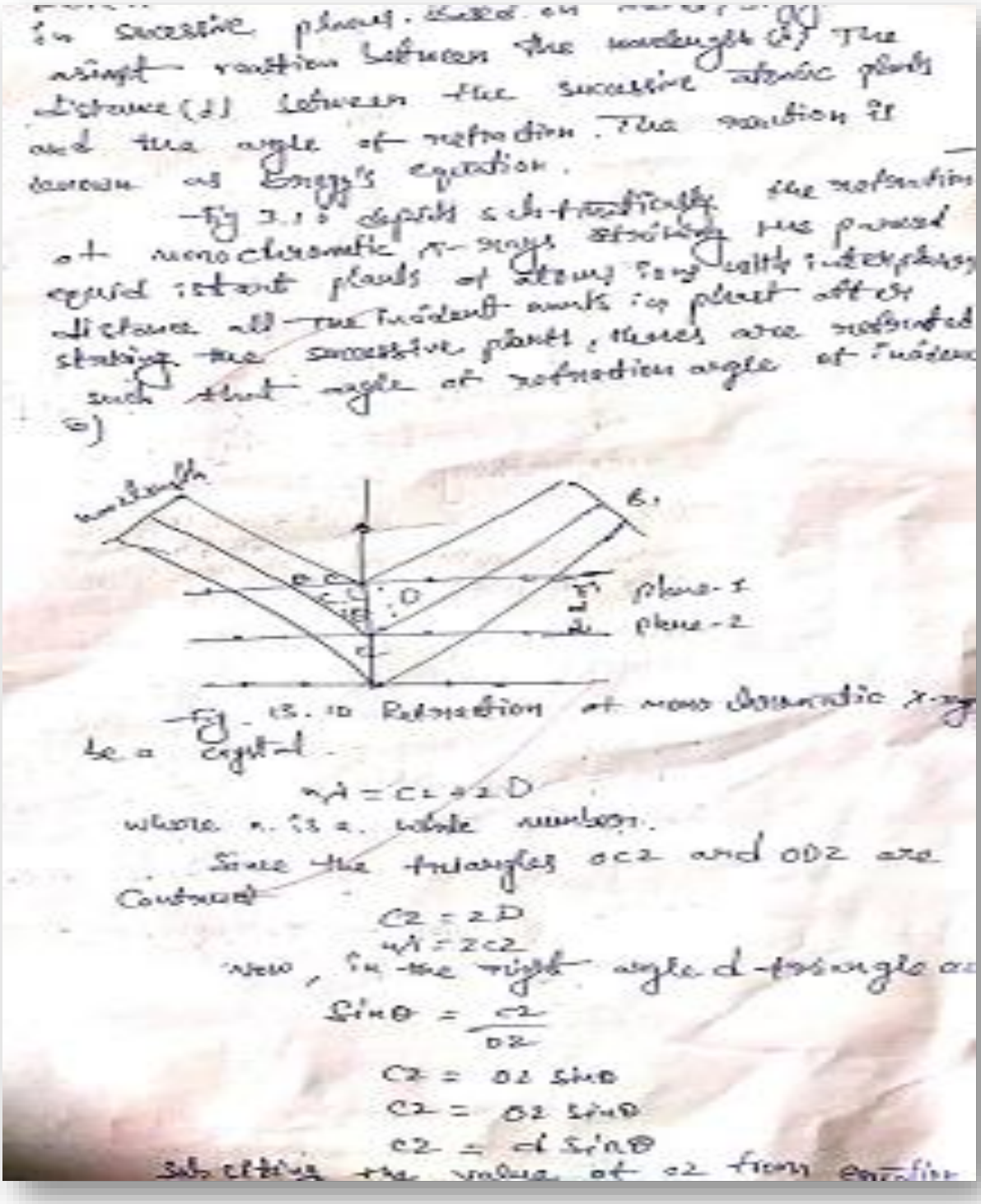
∴ There are following the of particles of diameter of the particles of the dispersed phase must be so much smaller than wavelength of the light used. The latter do not exhibit the Tyndall effect.

II. The difference b/w the refraction indices of the dispersed and the dispersion medium must be appreciably high. On the other hand in lyophilic sol' the difference of refractive indices is very small and hence the Tyndall effect is very weak.

Siedentopf and Zsigmondy applied the principles of Tyndall effect in devising the ultra microscope.

ultra - microscope :- A strong beam of light from an arc lamp powerful source is condensed by a system of cones and emits light and focussed through with the help of in a cell.

Bragg equation :- According to Bragg's a crystal could be employed not only as a transmission grating but also as a refraction grating, when X-rays are incident on a crystal face. This



PP111

$n\lambda = 2d \sin\theta$

where $n = 1, 2, 3 \ldots$

The equation is known as Bragg's equation. $n=1$ is called first order reflection. The reflection corresponding to $n=2$ is the second order reflection and so on. Thus by measuring and the incidence angle θ, we can know d/n

$$d/n = \frac{\lambda}{2\sin\theta}$$

from this d can be calculated if λ is known and viceversa. In x-ray reflection 'n' is already set as equal to 1. Thus, Bragg's equation is alternatively be written as

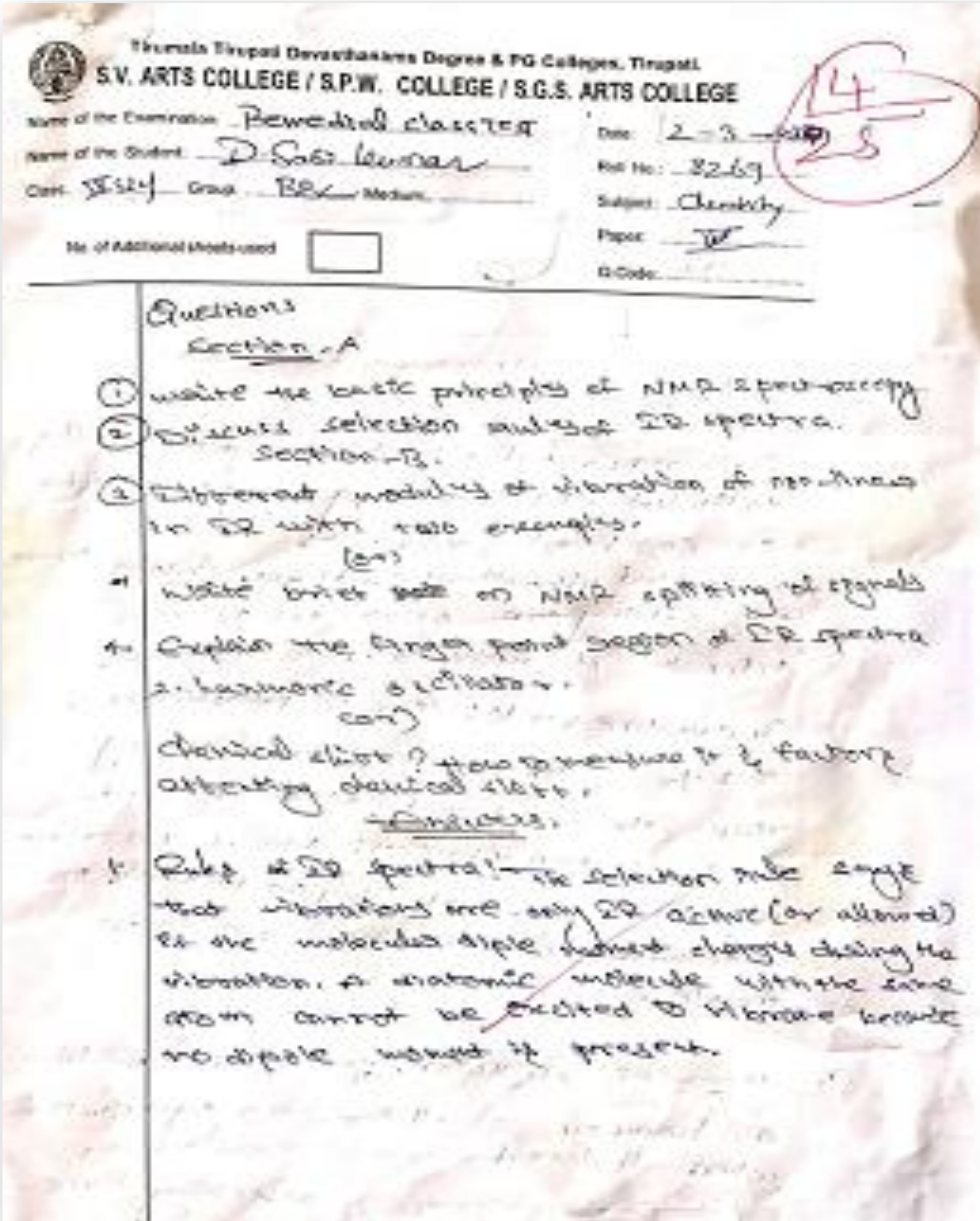
$$\lambda = 2d \sin\theta$$
$$= 2d_{hkl} \sin\theta$$

where d_{hkl} denotes the perpendicular distance b/w adjacent planes with indices hkl.

The experimental arrangement the apparatus commonly employed is Bragg's X-ray spectrometer and determination is done by two methods, one by Bragg's rotating by passing through screen and slits 2 and 'L' the monochromatic beam so details along the Bragg's rotating crystal system method are described below.

a beam of X-ray is generated from x-ray tube.

Semester-IV



Tirumala Tirupati Devasthanams Degree & PG Colleges, Tirupati.

S.V. ARTS COLLEGE / S.P.W. COLLEGE / S.G.S. ARTS COLLEGE

Name of the Examination: Remedial classes

Date: 2-3-[illegible]

Name of the Student: D. Sasi kumar

Roll No.: 8269

Class: [illegible] Group: BZC Medium:

Subject: Chemistry

Paper: IV

No. of Additional sheets used

Q.Code:

14/25

Questions

Section - A

1. Write the basic principles of NMR spectroscopy
2. Discuss selection rules of IR spectra.

Section - B.

3. Different modes of vibration of non-linear in IR with two examples.

(or)

4. Write brief note on NMR splitting of signals
5. Explain the finger print region of IR spectra & harmonic excitation.

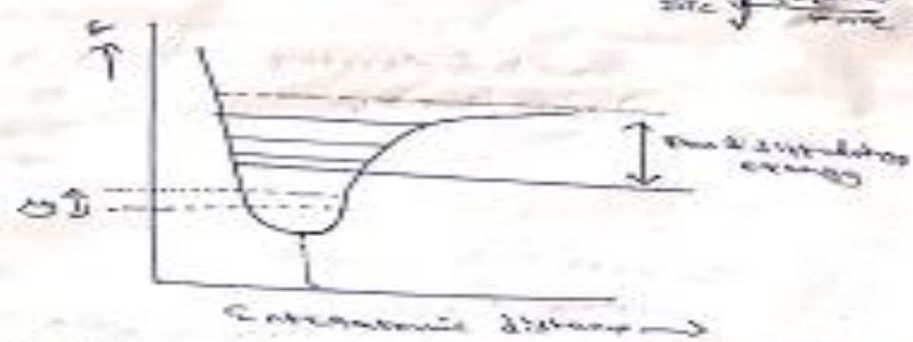
(or)

Chemical shift? How to measure it & factors affecting chemical shift.

Answers.

1. Rules of IR spectra:- The selection rule says that vibrations are only IR active (or allowed) if the molecules dipole moment changes during the vibration. A diatomic molecule with the same atom cannot be excited to vibrate because no dipole moment is present.

$$\bar{\nu} = \frac{1}{2\pi c}\sqrt{\frac{k(m_1+m_2)}{m_1 m_2}}$$



eg: CO_2

1. Degenerated vibrations: Symmetrically equivalent vibrations having same wavenumber. eg: CO_2 bending vibrations.
2. Coupled vibrations: Bonds with slightly different potential energy occur due to the network of molecules in a structure.
3. Overtone bands: Due to transition of 2 h/2, 3h/2 etc. lower intensity than the fundamental vibration band.
4. Combination vibrations: Sum or difference of two vibrations. They are helpful to determine the specification of molecule in sample..

9. Harmonic oscillator: In classical mechanics a harmonic oscillator is a system that when displaced from its equilibrium position experiences a restoring force F proportional to the displacement x.

$F = -kx$

where k is a positive constant.

→ If F is the only force acting on the system the system is called simple harmonic oscillator. It undergoes simple harmonic motion. Sinusoidal oscillation about the equilibrium point with a constant amplitude & a constant frequency (which does not depend on the amplitude).

⇒ Oscillate with a frequency lower than [illegible] the undergoes a decreasing [illegible] (underdamped oscillators).

⇒ Decay to the equilibrium position, without oscillation (overdamped oscillators)

→ If on [illegible] examples include pendulums (with small angles of displacement) masses connected to springs & acoustical [illegible]

⇒ Driven [illegible] is an external time-dependent force [illegible] the harmonic oscillator [illegible] described as given.

$$\Rightarrow \frac{d^2\theta}{dt^2} + \frac{g}{l}\sin\theta = 0$$

sin θ ≈ θ, instead consider the equation

$$\frac{d^2\theta}{dt^2} + \frac{g}{l}\theta = 0$$

$$\theta(t) = r\cos\left(\sqrt{\frac{g}{l}}\,t + \varphi\right)$$

(3) Non-linear molecule:

Acetylene, carbon dioxide, hydrogen cyanide etc.. are some examples of linear molecule whenever branched by (sec:) linked polymer such as vulcanized rubber are example of non-linear molecules.

⇒ The key difference between linear non-linear molecules is that linear molecules have a chemical structure.

⇒ non linear molecule have a zig-zag or cross-linked chemical structure.

→ All the molecules we know can be divided into two types of linear and non-linear depending on the shape of the molecules i.e the chemical structure of a molecule has linear geometry which appears to be a straight.

→ All theory molecules are categorized by non-linear molecules.

Semester-VI

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14/25

Name of the Examination: Remedial class test
Name of the Student: C. Ananasekhar
Class: VI Sem. Group: BSc Medium: Eng
No. of Additional sheets used:
Date: 22-03-20
Roll No: 8201
Subject: Chemistry
Paper: VI
Q.Code:

Section-A

Answer any One question (1x5 = 5 Marks)

1. Briefly explain a. precipitation and b. co-precipitations
2. Explain the theories of Acid base Titration

Section-B

Answer all Questions (3x10 = 20 Marks)

3. a. Write the classification of Errors

(or)

b. Explain the following (i) Accuracy (ii) Precision and (iii) Standard deviation.

4. a) Write the principles of volumetric Analysis

(or)

b) Write brief note on (i) Indicators (ii) Coagulation.

Answers

2. Acid base titration:

The chemical reaction involved in acid base titration is known as neutralisation reaction. It involves the combination of H_3O^+ ions with OH^- ions to form water. In acid-base titrations, solutions of alkali are titrated against standard acid solutions. The estimation of an alkali solution using a standard acid solution is called acidimetry. Similarly the estimation of an acid solution using a standard alkali

Solution is called alkalimetry

→ Acid base titration determines the Concentration of an acid or base by Exactly neutralising it with an acid (or) base of known Concentration.

→ Equivalence point the point at which an added titrant's moles are Stoichiometrically Equal to the moles of acid / base in the Sample

→ Analytelle unknown Solution whose Concentration is being determined in the titration

Section-B

3. a) **Classification of Errors:-**

Errors are classified into two types - Systemic (Determinate) and Random (Indeterminate) Errors.

Systemic (Determinate) Errors:-

Errors which can be avoided (or) whose magnitude can be determined is called as Systemic Errors. It can be determinable and presumably can be either avoided or corrected Systemic Errors further classified as.

→ Operational and personal Error.
→ Instrumental Error.
→ Errors of method
→ Additive or proportional Error.

Operational and Personal Error:-

Errors for which the individual analyst is Responsible and are connected with the method is called

When Errors occur during operation it is called as operational Error Eg: Transfer of Solution, effervescence, Incomplete drying.

Instrumental and Reagent Errors:-

Errors occur due to faulty instrument (or) Reagent Containing impurities Eg:- unCalibrated Weights, un-calibrated burette, pipette and measuring flasks.

Errors of Method.

When Errors occur due to method, it is difficult to correct. In gravimetric analysis, Error occurs due to Insolubility of precipitates, co precipitation, post precipitation, decomposition.

Additive (or) Proportional Errors:-

Additive Error does not depend on Instrument present in the determination Eg: loss in Weight of a crucible in which a precipitate is ignited.

Proportional Error depends on the amount of the Constituent Eg: Impurities in Standard Compound.

Random Errors:-

It occurs accidentally (or) randomly So called as Intermediate or accidental or random Error. Analyst has no Control in this Error. It follows a random distribution and a mathematical law of probability can be applied.

6. Yoga &Meditation

Regular yoga &meditation enhances mental and physical health and concentration in the work. The college organizes yoga and meditation sessions to ensure the students better performance in the academics and examinations.

Students engaged in Yoga & Meditation

7. Progress report

The various programmes organized in respective of slow learners have given positive results which can be witnessed in the improved performance of the slow learners in the end semester examinations.

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
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HOD OF CHEMISTRY

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Progress Report Of Slow Learners In Chemistry After Remedial Coaching I .BSc: I & II Semesters 2016 – 2017

S.NO	Name of The Student	STATUS OF STUDENTS (SLOW)	SUBJECT	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	A. SAI KUMAR	SLOW LEARNERS	CHEMISTY MPC(T.M)	38	54
2	B. RAJESH			35	22
3	C. CHANDU			31	52
4	C. PRAVEEN KUMAR			0	1
5	M. HARI KRISHNA			18	53
6	M .MOHAN			17	0
7	M. PRASANNA			34	65
8	N. SAI GANESH			37	53
9	P .RAM MAHESH BABU			31	0
10	R. MANOHAR			21	20
1	A. AKIL KUMAR	SLOW LEARNERS	CHEMISTRY MPC(E.M)	3	0
2	A. VEERANJANEYULU			33	37
3	B. PRAVEEN KUMAR			0	0
4	C. KIRAN KUMAR			33	50
5	C. SURI BABU			26	48
6	B. RAVI TEJA REDDY			36	AB
7	G. LALITH KISHORE			22	21
8	K. MEGHANA			35	59
9	M. VISHNU			22	51
10	P. THRIVEN KUMAR			30	42
11	P .SAI KRISHNA			4	42
12	P. LAKSHMAIAH			27	42
13	R. ASHOK KUMAR			31	66

1	B. POOJA	SLOW LEARNERS	CHEMISTRY CBZ(T.M)	25	57
2	B. NAVEEN KUMAR			35	57
3	C. UPPUTURAIAH			29	54
4	C. SAI			AB	52
5	C. NAGARJUNA			32	54
6	D. KRISHNA MOORTHY			39	24
7	G. AMMULU			30	AB
8	G. HAREESH			35	54
9	J. SANJEEVA RAYUDU			34	65
10	K. MOHAN			26	63
11	K. SANGEETHA			28	88
12	K. PRIYANKA			34	91
13	M. SUNIL			30	73
14	M .BUJJAMMA			29	86
15	M .JANARDHAN GOUD			38	75
16	R. PAVAN KALYAN			15	29
17	S. BHARGAVI			27	43
18	S. HARI KRISHNA			29	58
19	S. RAMESH			26	68
20	Y .MOULI			21	41
21	Y. GURU VISHNU			24	64
22	P. CHANDRA SEKHAR REDDY			40	70
1	B. POORNIMA	SLOW LEARNERS	CHEMISTRY MZC(E.M)	32	30
2	D. SREENATH			38	60
3	G. SIVA KUMAR			27	50
4	G. GOWTHAM SAI BABU			28	35
5	H .KEERTHI KUMAR			23	50
6	K .KOUSHIK			15	20
7	K. NAVEEN			19	51
8	K. GNANA TEJA			40	58
9	M .VENKATA SURENDRA			29	AB
1	A. UMAPATHI	SLOW LEARNERS	CHEMISTRY BBC(E.M)	34	22
2	C. PRAKASH			40	34
3	G .ANIL KUMAR			36	33
4	K. NAGARAJU			26	36
5	K. NAVEEN			25	51
6	M .HEMANTH SURYA			AB	AB

LIST OF SLOW LEARNERS: II B.Sc III &IV SEMESTERS 2016-2017					
S.NO	**NAME OF THE STUDENT**	**STATUS OF STUDENTS (SLOW)**	**SUBJECT**	**MARKS IN ODD SEMESTER**	**MARKS IN EVEN SEMESTER (After Remedial Coaching)**
1	C. ROHITH KUMAR	SLOW LEARNERS	CHEMISTRY MPC(T.M)	25	52
2	G.SUNIL KUMAR			32	52
3	M.AKHIL KUMAR			32	50
1	A. SUDHAKAR	SLOW LEARNERS	CHEMISTRY MPC(E.M)	32	68
2	B. HARISH KUMAR			31	57
3	B.V.L. SIVA			25	52
4	B.JANAKIRAMUDU			38	69
5	B. THIMMAPPA			33	54
6	C. PITHAMBERAM REDDY			28	AB
7	D.SREENIVASULU			28	AB
8	K. RAJ KUMAR			39	57
9	M.MANI			40	55
10	M. SAI KUMAR			22	50
11	M. SUNIL NAIK			31	96
12	Y. RAMANJANEYULU			24	38
13	G. CHANDU			2	20
1	A. ANIL KUMAR	SLOW LEARNERS	CHEMISTRY CBZ(T.M)	29	70
2	B. THULASI			23	AB
3	B. SUPRIYA			26	54
4	B. SAI RAM			34	25
5	B. VENKATESH			26	33
6	D. PRADEEP			28	53
7	G. SARITHA			38	36
8	K. SHOBA			31	55
9	K. KOSOGAIAH			24	31
10	P. HARI KRISHNA			25	AB
11	SARASWATHI			37	AB
12	T. SAGAR			18	AB
13	M. HARI NAIK			19	24
1	D. DAMODARAM	SLOW LEARNS	CHEMISTRY MZC(E.M)	39	20
2	M.C. VINAYAKA			29	AB
3	M. SANDEEP			39	51
4	S. PURUSHOTHAM NAIK			37	AB

5	T. NAVEEN KUMAR			24	43
6	Y. LAVANYA			37	50
7	B. VEERANJANEYULU			35	39
8	K. RAGAVENDRA			33	32
9	N. ANJI NAIK			24	50
10	T. VARUN TEJA			36	20
1	C. SWATHI	SLOW LEARNS	CHEMISTRY BBC(E.M)	40	AB
2	K. MADHU			38	38
3	M.JYOTHI REDDY			29	46
4	M.VIMAL PRABHU			29	52
5	R. MUNI PRAKASH			29	54
6	U. DENIAL			29	27
7	A. GIRI			38	26
8	K. RAJESH			40	35
9	M.CHANDRA SEKHAR NAIK			18	AB

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DEPARTMENT OF CHEMISTRY
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DEPARTMENT OF CHEMISTRY

Progress Report Of Slow Learners In Chemistry After Remedial Coaching

I .BSc : I & II SEMESTERS : 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	Programme	Marks In Odd Semester	Marks In Even Semester (After Remedial Coaching)
1	M. PRASANTH	SLOW LEARNERS	MPC	40	57
2	Y. HANUMANTHU			39	68
3	Y. HEMANTH KUMAR			39	69
4	B. HARI PRASAD			38	76
5	T.CHIRANJEEVI			38	54
6	M.RAKESH			37	51
7	V. NAGESWARA			36	52
8	P.SIVA KUMAR REDDY			32	53
9	P. SUDHEER			27	56
10	P. SRAVYA			25	AB
11	R. GUNA SEKHAR			23	52
12	K. CHANDU			22	51
13	C. GNANA SEKHAR			21	25
1	V. VENKATESH	SLOW LEARNERS	CBZ	39	55
2	B. SRAVANI			32	61
1	P. RAKESH NAIK	SLOW LEARNERS	MZC	40	41
2	G. PANDU RANGA			40	59
3	U. ANANDA			39	54
4	D. NAVEEN			38	21
5	N. PRAVEEN NAIK			34	38
6	M. SREEDHAR			29	30
7	K. VAMSI			28	56
8	T. AKASH			27	62

9	K. SANDEEP KUMAR			23	36
10	P. SRINIVASULU			22	26
11	M. RAMESH			22	33
12	G. RAJ KUMAR			21	AB
13	B. GOVARDHAN			16	18
14	A. SAMUYELL			15	AB
15	T. NAVEEN KUMAR			0	AB

1	J. VENKATESH	SLOW LEARNERS	BBC	39	52
2	A. JAGADEESWARA REDDY			38	38
3	J. VENKATESH			35	54
4	G. KIRAN KUMAR			34	74
5	S. GIRISH KUMAR			31	43
6	E. GOPINATH GOUD			30	56
7	M. GAYATHRI			30	54
8	K. SIVA			29	28
9	K. UMA			27	35
10	S. GANGADHAR			26	43
11	K. KUSUMA PRIYA			26	40
12	P. NITHISH KUMAR			21	41
13	K. AJITH			20	AB
14	T. NAVEEN KUMAR			AB	AB

LIST OF SLOW LEARNERS: II B.Sc III&IV SEMESTERS 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	SUBJECT	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	C. PRAVEEN KUMAR	Slow Learners	MPC-TM	AB	AB
2	N. SAI GANESH			39	AB
3	N. CHANDRA SEKHAR			38	61
4	M. SUDHAKAR			29	27
5	A.B. LOKESH			22	AB
6	R. MANOHAR			22	20
7	B. RAJESH			19	AB
1	C. SURI BABU	Slow Learners	MPC-EM	35	39
2	A. VEERANJANEYULU			32	39
3	G. LALITH KISHORE			28	40
4	R. ASHOK KUMAR			25	AB
5	P. SAI KRISHNA			18	18
6	M. VISHNU			17	18
7	T.B. BALARAMUDU			55	74
8	P. LAKSHMAIAH			39	20
9	K. VINAY			37	57
10	A. AKIL KUMAR			6	AB
1	K. SANGEETHA	Slow Learners	CBZ-TM	40	
2	B. GAJENDRA			39	54
3	D. KRISHNA MOORTHY			37	19
4	R. INDRASENA REDDY			36	21
5	C. UPPUTURAIAH			34	26
6	Y. MOULI			30	65
7	M. SUNIL			28	49
8	C. SAI			26	17
9	B. POOJA			24	26
10	M. VAMSI			24	21
11	R. PAVAN KALYAN			22	25
1	S. VENKATA NAGATEJASREE	SLOW LEARNERS	BBC	39	38
2	B. POORNIMA			36	21

3	G. Gowtham Sai Babu			28	21
4	D. SREENATH			24	56
5	K. GNANA TEJA			23	21
6	M. RAVI NAIK			23	74
7	K.KOUSHIK			20	AB
1	T.Y. MANOJ KUMAR	SLOW LEARNERS	MZC	39	68
2	P. LAKSHMI PRASAD			33	AB
3	K. REDDI BABU			28	39
4	M. HEMANTH SURYA			28	62
5	G. ANIL KUMAR			26	63
6	T.SAI KIRAN			24	41
7	C. PRAKASH			22	30

LIST OF SLOW /ADVANCED LEARNERS: III B.Sc V & VI SEMESTERS: 2017 – 2018

S.NO	Name Of The Student	Status Of Students (SLOW)	Subject	Marks In Odd Semester	Marks In Even Semester (After Remedial Coaching)
1	C. ROHITH KUMAR	SLOW LEARNERS	MPC-TM	37	51
2	R. VINOD KUMAR			22	23
3	O. ASHOK KUMAR			20	23
1	B. HARISH KUMAR	SLOW LEARNERS	MPC-EM	33	34
2	K. SUJITH KUMAR			32	41
3	K.ASHOK KUMAR REDDY			29	51
4	M. SAI KUMAR			22	18
5	K. YELLAPA			21	33
6	Y. RAMANJANEYULU			18	7
7	D. SREENIVASULU			0	0
1	B. VENKATESH	SLOW LEARNERS	CBZ-TM	37	54
2	K. KOSOGAIAH			27	53
3	M. HARI NAIK			22	23
4	HARI KRISHNA			20	23
1	M.C. VINAYAKA	SLOW LEARNERS	MZC	40	42
2	T. AMARNADH			38	AB
3	K. RAGAVENDRA			38	40
4	R. JAGADESH BABU			37	61
5	M. SANDEEP			29	AB
6	VARUN TEJA			23	31
7	N. ANJI NAIK			20	41
1	Y. VIKAS REDDY	SLOW LEARNERS	BBC	39	35
2	K.RAJESH			38	AB
3	M. JYOTHI REDDY			37	63
4	U. DENIAL			37	54
5	A. GIRI			31	AB
6	S.S. ANOOF			23	53

DEPARTMENT OF CHEMISTRY

Progress Report Of Slow Learners In Chemistry After Remedial Coaching

I .BSc: I & II Semesters: 2018 – 2019

S.NO	NAME OF THE STUDENT	Status Of Students (SLOW)	Subject	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER (After Remedial Coaching)
1	O. ARAVIND	SLOW LERNERS	CHEMISTRY (MPC)	3.7	5.4
2	A. HEMANTH KUMAR			3.3	2.2
3	B.SIVARAMUDU			3.3	5.2
4	B. DATHATHREYUDU			3.1	5.4
5	K.NITHISH			3.1	
6	P. MUNI SEKHAR			2.8	5.6
7	S. NANDA KUMAR			2.1	6.6
1	B. GANGADHAR	SLOW LERNERS	CHEMISTRY (CBZ)	3.7	5.6
2	A. MANENDRA			3.7	5.4
3	K. HIMA BINDU			3.3.	
4	K. VENKATESH			3.2	
5	B. MOHAN NAIK			3.1	6.1
6	D. SAI KUMAR			2.8	
7	M. HARI PRAKASH NAIK			2.7	3
8	M. NARESH BABU			2.7	
9	K. KALPANA			2.5	AB
10	C. CHANDU			2.4	AB
11	B. CHANDRA SEKHAR NAIK			2.2	5.7
12	K.CHENCHAIAH			2.2	AB
13	G. NAGARAJU			2	5.5
14	G. MAHENDRA			1.9	
1	K. SANTHOSH KUMAR	SLOW LERNERS	CHEMISTRY (MZC)	AB	AB
2	S. SOMA SEKHAR			4	5.3
3	E. NAVEEN			3.9	3.3
4	G. PRANAY KUMAR REDDY			3.5	5.7
5	E. REDDY SEKHAR			3.5	5.2
6	C. RAJA SEKHAR			3.2	
7	B. RANGADU			3.2	2.5
8	N. VINOD KUMAR			3.2	3.9

9	P. SAI THARUN			2.8	
10	A. BALAJI			2.7	6.3
11	T.GOWTHAM			2.7	5.5
12	G. SWARNA LATHA			2.7	2
13	B. GANESH			2.6	3.2
14	B. VIJAY KUMAR			2.6	2.5
15	G. BHEEMUDU			2.5	5.3
16	T. MOUNIKA			2.4	5.3
17	Y. DEESNESH BABU			2.3	5.8
1	M. NARASIMHULU	SLOW LERNERS	CHEMISTRY (BBC)	3.7	6.4
2	K. BHARGAV			4	5.6
3	B. CHINNA THIMMAPPA			4	5.7
4	S. SAI SREE			4	4.2
5	P. VANDHANA			4	6.2
6	M. SATISH KUMAR			3.6	6.4
7	K. RAMANJINEYULU			3.5	5.3
8	P. NARASHIMHA REDDY			3.2	3.3
9	U. SOMESWARA RAO			3.1	5.9
10	R. TEJASREE			2.8	3.2
11	K.DEENA			2.6	3.2
12	M.NAVEEN			2.5	6.1
13	T.VIJAY			2.4	3.2
14	B. HAREESH			2.2	3.4
15	P. JAYCHANDRA			1.6	5.4
16	A. PEDDA OBULESU			AB	AB
17	T.SUJAN			AB	AB

LIST OF SLOW LEARNERS: II B.Sc : III & IV SEMESTERS : 2018 - 2019

S.No	Name of The Student	Status Of Students (SLOW)	Subject	Marks in odd semester	Marks In Even Semester (After Remedial Coaching)
1	P. SIVA KUMAR REDDY	SLOW LERNERS	CHEMISTRY (MPC)	21	5.1
2	P. MAHESH			31	2.4
3	P. SUDHEER			33	6.3
4	N. SUMAN			37	5
5	M. RAKESH			39	AB
1	B. SRAVANI	SLOW LERNERS	CHEMISTRY (CBZ)	44	6.5
2	S. BALAJI NAIK			43	3.4
3	T. REVATHI			47	4.4
1	T. AKASH	SLOW LERNERS	CHEMISTRY (MZC)	AB	6.5
2	P. SRINIVASULU			17	AB
3	SK. MAZEED BABU			19	AB
4	M. SREEDHAR			20	6.5
5	P. RAKASH NAIK			25	6.8
6	B. YUGANDHAR			25	5.2
7	N. PRAVEEN NAIK			27	6.2
8	V. PARSARAM			30	5.5
9	A. BALAJI			35	3.7
1	P. NITHISH KUMAR	SLOW LERNERS	CHEMISTRY (BBC)	19	2.2
2	G. KIRAN KUMAR			29	7.4
3	S. GANGADHAR			33	5.4
4	C. LOKESH			35	6.6
5	M. GAYATHRI			36	6.2
6	E. GOPINATH GOUD			38	8
7	J. VENKATESH			39	2.7

LIST OF SLOW LEARNERS: III B.Sc V & VI SEMESTER 2018 - 2019

S.NO	NAME OF THE STUDENT	Status Of Students (SLOW)	Subject	Marks in odd semester	Marks In Even Semester (After Remedial Coaching)
1	M. HARI KRISHNA	SLOW LERNERS	CHEMISTRY (MPC)	17	68
2	M. SUDHAKAR			19	68
3	M. VISHNU			21	AB
4	P. LAKSHMAIAH			21	22
5	K. SACHIN CHOWDARY			30	71
6	G. MIDHILESH			39	65
		SLOW LERNERS	CHEMISTRY (CBZ)	NIL	
	NIL				
1	G. GOWTHAM SAI BABA	SLOW LERNERS	CHEMISTRY (MZC)	17	22
2	K. NAVEEN			22	37
3	S. VENKATANAGA TEJSREE			34	53
4	M. BHANU PRAKASH			40	59
		SLOW LERNERS	CHEMISTRY (BBC)		
1	C. PRAKASH			20	27
2	B. SUNIL KUMAR NAIK			21	50
3	P. LAKSHMI PRASAD			21	53
4	A. HARI KRISHNA			28	33
5	N. MAHESH KUMAR			30	52
6	G. ANIL KUMAR			33	AB
7	T. SAI KIRAN			33	55
8	M. HEMANTH SURYA			36	35

Progress report of slow learners after remedial coaching

LIST OF SLOW LEARNERS: I B.Sc I & II SEMESTERS - 2019 – 2020

S.No	Name of the Student	Status Of Students (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	A. RAGHU KUMAR	SLOW	MPC.T.M	3.8	2.5
2	A. VENKATAARJUN	LEARNERS	&E.M		3.2
3	A. CHANDU PRAKASH				3.6
1	A M. SAMEERA	SLOW		1.9	2.7
2	B.DINESH NAIK	LEARNERS		2	3.3
3	C. LAKSHMI NARAYANA			2.2	3.5
4	C. HARINAD			2.2	5.3
5	C. HEMANTH KUMAR			2.4	5.3
6	C. CHANDRA SEKHAR			2.5	5.4
7	D. SURYA		CBZ.T.M	2.7	5.4
8	D. CHANDRA			2.7	5.4
9	D.MUNI LATHA			2.8	5.4
10	E. SREE KANTH			3.1	5.5
11	VEERANJINEYULU			3.2	5.5
12	G. RAJASEKHAR			3.3	5.5
13	K. SASI KUMAR			3.7	5.9
14	K. USHA			3.7	5.9
1	A. GIRIRAMESH	SLOW	MZC E.M	2.2	5.1
2	A. GIRISHMA	LEARNERS		3.1	5.4
1	A. CHANDINI	SLOW		2.4	2.5
2	A. POOJITHA	LEARNERS		2.8	5.2
3	D. GOUTHAM KRISHNA			2.8	5.4
4	VENKATASAI KUMAR			3.2	5.4
5	V. MANJUNADH NAIK		BBC E.M	3.5	5.5
6	D. GOVARDHAN			3.7	5.6
7	V. JANSI LAKSHMI			3.8	5.6
8	C. BABU			3.8	5.8
9	C. KARUNAKAR			3.8	6
10	C. RAMAKANTH			3.9	6

LIST OF SLOW /ADVANCED LEARNERS: II B.Sc : III & IV SEMESTERS - 2019 - 2020

S.No	Name of the Student	STATUS OF STUDENTS (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	NIL	NIL	MPC E.M &T.M	NIL	
1	A. BALAJI	SLOW LEARNERS	MZC(EM)	NIL	2.6
2	G.BEEMUDU				3
3	G.CHANDRA SEKAR				3.6
4	Y.DINESH BABU				3.8
1	T.ASHA	SLOW LEARNERS	CBZ(EM)	2.9	5.4
2	B.CHANDRASEKAHR			3.4	6
3	C.CHANDU			3.7	
1	T.BARATH KUMAR	SLOW	BBC.E.M	2.1	3.3
2	K. BARGAV	LEARNERS		2.1	4.1
3	P. CHANDRA SHEKAR			2.1	5
4	B. CHINNA			2.2	5
5	S. CHOUDAIAH			2.2	5.2
6	K.DEENA			2.4	5.3
7	B.DILIP			2.5	5.3
8	B.HARI BABU			2.7	5.5
9	K.JAGADESWARA			2.7	5.6
10	M. CHANDRA			2.7	5.6
11	T.JAYARAM			2.8	6
12	B. LAKSHMI KANTH			2.9	6.2
13	K. LAKSHMI NARASIMHA			3	7
14	G. MALLIKARJUNA			3.1	7
15	S. MANJUNDH			3.1	7.1
16	M. MAREPPA			3.3	7.2

LIST OF SLOW LEARNERS: III B.Sc V&VI SEMESTERS - 2019 - 2020

S.No	Name of the Student	STATUS OF STUDENTS (Slow)	SUBJECT	MARKS IN ODD SEM	MARKS IN EVEN SEM
1	T.AVINASH	SLOW LEARNERS	MPC E.M	2.9	0
2	K.BAGYA LAKSHMI		&T.M	3.2	2.3
3	S.GANGADHAR			3.5	2.6
1	NIL	SLOW LEARNERS	CBZ(T.M)	NIL	
1	A.BALAJI	SLOW LEARNERS	MZC(E.M)	3.7	4.1
1	NIL	SLOW LEARNERS	BBC(E.M)	NIL	

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TIRUPATHI

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD,TIRUPATI

DEPARTMENT OF CHEMISTRY

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HOD OF CHEMISTRY

SEMESTER WISE NUMBER OF SLOW LEARNERS Batch .. 2016-19				
PROGRAMME NAME	PROGRAMME CODE	SEMESTER	TOTAL NUMBER OF STUDENTS	TOTAL NUMBER OF SLOW LEARNERS
B.Sc	MPC	II SEMESTER	44	23
		III SEMESTER	44	15
		IV SEMESTER	39	21
		V SEMESTER	35	14
		VI SEMESTER	31	8

PROGRAMME WISE SLOW LEARNERS PERFORMANCE Batch .. 2016-19				
PROGRAMME NAME	PROGRAMME CODE	SEMESTER	TOTAL NUMBER OF STUDENTS	TOTAL NUMBER OF SLOW LEARNERS
B.Sc	CBZ	II SEMESTER	24	22
		III SEMESTER	27	4
		IV SEMESTER	25	11
		V SEMESTER	26	12
		VI SEMESTER	24	2

The above data is a direct eye witness for the positive outcome of remedial coaching.

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DEPARTMENT OF COMPUTER SCIENCE					
PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING					
LIST OF SLOW LEARNERS					
S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS · (SLOW)	Group & Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	SANAMBATLA HIMAVANTH	SLOW LEARNERS	II YEAR MSCS-EM	42	44

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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

PROGRESS REPORT OF SLOW LEARNERS IN COMPUTER SCIENCE AFTER REMEDIAL COACHING

LIST OF SLOW LEARNERS

S.NO	NAME OF THE STUDENT	STATUS OF STUDENTS (SLOW)	Group& Medium	MARKS IN ODD SEMESTER	MARKS IN EVEN SEMESTER
1	G VAMSI KRISHNA	SLOW LEARNERS	II YEAR IV SEM MSCS	20	66
2	KESANI SAI KUMAR			18	41
3	SANAMBATLA HIMAVANTH			22	AB

HEAD
DEPARTMENT OF COMPUTER SC
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Students with varied degree of learning capabilities and diversified back ground get admission in the college. Therefore the level of challenge faced by them is also different. The college is empathetic to this aspect and takes maximum steps to optimize the learning outcomes of the students. In this direction the college has undertaken a proactive step of identifying slow learners and advanced learners.

Students who scored 40% or <40% marks in the university examinations are categorized as slow learners and those scored >70% are advanced learners**.**

B.Sc: Batch-2016-2019

Department of chemistry

A sample copy of advanced learners belonging to chemistry department is presented here. Complete details are available in college website.

Programme code	Semester wise number of advanced learners					
	Semester-II	Semester-III	Semester-IV	Semester-V		Semester-VI
				Paper-5	Paper-6	
MPC - E.M&TM	4	4	4	7	11	5
BBC-E.M	5	10	8	10	11	14
MZC-E.M	3	7	11	6	9	14
CBZ-TM	8	NIL	1	3	4	2
Total	**20**	**21**	**24**	**26**	**35**	**35**

B.Sc: Batch-2017-2020						
Programme code	Semester wise number of advanced learners					
	SEMESTER-II	SEMESTER-III	SEMESTER-IV	SEMESTER-V		SEMESTER-VI
				Paper-5	Paper-6	
MPC -E.M&TM	5	8	5	16	15	27
BBC-E.M	2	5	NIL	NIL	3	2
MZC-E.M	7	6	5	6	7	25
CBZ-TM	9	6	14	11	14	20
Total	23	25	24	33	39	74

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S.ARTS COLLEGE, TTD, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

BATCH-2016-2019

Programme code	SEMESTER WISE NUMBER FOR ADVANCED LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	**SEMESTER-III (BASED ON II SEMESTER)**	**SEMESTER-IV (BASED ON III SEMESTER)**	**SEMESTER-V (BASED ON IV SEMESTER)**	**SEMESTER-VI (BASED ON V SEMESTER) PAPER-5**	**SEMESTER-VI (BASED ON V SEMESTER) PAPER-6**
MPCS	9	9	20	18	23	27
MSCS	12	16	30	29	24	27
Total	21	25	50	47	47	54

BATCH-2017-2020

Programme code	SEMESTER WISE NUMBER FOR ADVANCED LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	SEMESTER-III (BASED ON II SEMESTER)	SEMESTER-IV (BASED ON III SEMESTER)	SEMESTER-V (BASED ON IV SEMESTER)	SEMESTER-VI (BASED ON V SEMESTER) PAPER-5	SEMESTER-VI (BASED ON V SEMESTER) PAPER-6
MPCS	24	29	34	39	37	38
MSCS	27	30	28	38	34	34
Total	51	59	62	77	71	72

BATCH-2018-2021

Programme code	SEMESTER WISE NUMBER FOR ADVANCED LEARNERS					
	SEMESTER-II (BASED ON I SEMESTER)	SEMESTER-III (BASED ON II SEMESTER)	SEMESTER-IV (BASED ON III SEMESTER)	SEMESTER-V (BASED ON IV SEMESTER)	SEMESTER-VI (BASED ON V SEMESTER) PAPER-5	SEMESTER-VI (BASED ON V SEMESTER) PAPER-6
MPCS	31	26	36	45	44	46
MSCS	26	31	43	46	43	44
Total	57	57	79	91	87	90

II. A list of representative activities conducted for advanced learners is shown below

1. Quiz competition

In order to boost self esteem and confidence in the students the college organized quiz competitions. They encourage students self awareness of progress and self assessment. This can help the students to identify the areas they need to develop. A sample copy is presented here.

Tirupati
Date: 26-08-2019

From
The HOD of CHEMISTRY
S.G.S. Arts College
Tirupati

To
The Principal
S.G.S. Arts College
Tirupati

Respected Sir

Sub:- Permission for conducting Quiz programme – Request – Regarding.

-:oo0oo:-

We are conducting Quiz programme for I, II and III year students of B.Sc. (MPC, MZC, BBC & CBZ) in Room No.225. So, kindly grant us permission on 29-08-2019.

Thanking you Sir

Yours faithfully

HOD of Chemistry

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE,TIRUCHANOOR ROAD,TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No:9490728655
Mail.ID: parameswara1964@gmail.com

COMMITTEE FOR QUIZ PROGRAMME

ACADEMIC YEAR – 2019-2020

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Dr. V. Venkata Lakshmi, Coordinator	
2	Dr. P. Suguna, Member	
3	Smt. K. Revathi, Member	

SIGNATURE OF THE HOD

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

chem

"OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

TOPIC: PHARMACEUTICAL USES OF MEDICINAL CHEMISTRY **DATE: 29/08/2019**

S.No	Programme name	Programme Code	Name Of The Student	Signature Of The Student	BATCH	Marks Awarded (30)	Remarks
1	II B.Sc	MPC	M.Tharunkumar	M.Tharunkumar	A	28	I PLACE
2			B.Sharmila	B. Sharmila			
3			B.Varaprasad	B. Varap			
4			M.SandeepReddy	M.Sandeep			
5			B.Nagaraju naik	B. Nagaraju			
6			B.sri vijay	B. Srivijay			
7		MPC	C.Varaprasad	C. Varaprasad	B	25	II PLACE
8			C.Chandra uday kumar	C. Chandra uday kumar			
9			T.Vishnuvardhan	T. Vishnu			
10			M.Raja	M. Raja			
11			N.KanthaReddy	Kanthareddy			
12			D.Karthik	D. karthik			
13		MZC	C.Chandu	C. Chandu	C	24	
14			P.Lokesh	P. Lokesh			
15			P.Vaishnavi	P. Vaishnavi			
16			R.Pavithra	R. Pavithra			
17			M.Kumar	M. Kumar			
18			N.Usman	N. Usman			
19		BBC	K.Venkatasuresh	K. Venkata suresh	D	27	II PLACE
20			P. Praddeep Kumar	P. Praddeep kumar			
21			K.Sunetha	K. Sunetha			
22			V. Manjunadh Naik	V. Manjunath			
23			K.RAJA	K. Raja			
24			A.POOJITHA	A. Poojitha			
25		BBC	M.Devendra	M. Devendra	E	23	
26			D. Goutham Krishna	D. Goutham			
27			B.Venkatasai Kumar Naik	B. Venkatasai			
28			P. Govardhan	P. Govardhan			
29			S.Pravallika	S. pravallika			
30			T. Sreelatha	T. Sreelatha			

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI
S.G.S ARTS COLLEGE,TIRUCHANOOR ROAD,TIRUPATI
DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No:9490728655
Mail.ID: parameswara1964@gmail.com

QUIZ - DEPARTMENT OF CHEMISTRY

ACADEMIC YEAR ::2019-20						
S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION	TOTAL NO. OF BATCHES	WINNING BATCH
1	II B.Sc.	MPC	PHARMACEUTICAL USES OF MEDICINAL CHEMISTRY	29/8/2019	5 (30 Nos.)	1(6Nos.) (MPC)
		CBZ				
		MZC				
		BBC				

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of CHEMISTRY
EVENT ORGANISED REPORT: 2019 – 20

1	Name of the Department	**Chemistry**
2	Name of Event Organized	**QUIZ COMPETITION**
3	Title of Event Organized	Pharmalogical uses of medicinal plants
4	Name of Co-ordinator of Event	**Sri.D.Parameswara**
5	Class of the Participants	**B.Sc,MPC,BBC,MZC&CBZ**
6	No. of Participants	**30**
7	Objective of the Event	➢ To encourage students to look beyond their textual knowledge and establish a relationship between theory and application of the learnt concepts
8	Outcome of the Event	• **Interactive platform** • **Changes scope of learning** • **Encourages team work** • **Bridges academia-industry gap**

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Presentation of certificate to the quiz winner A.Raghu kumar &M. Sandeep Reddy of IIB.Sc : MPC -EM by HOD Sri.D.PARAMESWARA

QUIZ FORMAT

TIRUPATI,
DATE : 24-08-2019

FROM,
The HOD of MICROBIOLOGY
S.G.S. ARTS COLLEGE,
TIRUPATI.

TO ,
THE PRINCIPAL
S.G.S. ARTS COLLEGE,
TIRUPATI.

Respected Sir/ Madam,

Sub : Request for permission to conduct Quiz programme Regarding

***

We are conducting Quiz programme for III year students of **MICROBIOLOGY** in Room No 316So, kindly grant us permission to conduct Quiz programme on **26-08-2019**.

Thanking you Sir / Madam

Yours faithfully

HOD

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

DEPARTMENT OF MICROBIOLOGY

COMMITTEE FOR QUIZ PROGRAMME

ACADEMIC YEAR – 2019 - 2020

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Dr. K. Sridevi	K. Sridevi
2		

SIGNATURE OF THE HOD

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF MICROBIOLOGY

ACADEMIC YEAR 2019 -2020

S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION	TOTAL NO. OF BATCHES	WINNING BATCH
1	QUIZ	III MZC	Food Microbiology	28-08- 2018	4	B

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DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: MICROBIOLOGY

Academic year: 2019 - 2020

S.No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	Quiz Competition	26-08-2019	III MZC	12	

HOD

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of MICROBIOLOGY
EVENT ORGANISED REPORT: 2019 - 2020

1	Name of the Department	**MICROBIOLOGY**
2	Name of Event Organized	**QUIZ**
3	Title of Event Organized	**Food microbiology**
4	Name of Coordinator of Event	**Dr. K. Sridevi**
5	Class of the Participants	**MZC**
6	No. of Participants	**12**
7	Objective of the Event	**To know the pathogenic and non pathogenic micro organisms that found in food and their importance in Fermented food , food borne disease etc.,**
8	Outcome of the Event	**It is useful for methods of isolation , detection, identification of food borne and spoiled food in food Microbiology**

HOD

HEAD
DEPT. OF MICROBIOLOGY
SGS ARTS COLLEGE
TIRUPATI-517501 (A.P.)

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of MICROBIOLOGY

EVENT ORGANISED REPORT: (Year)

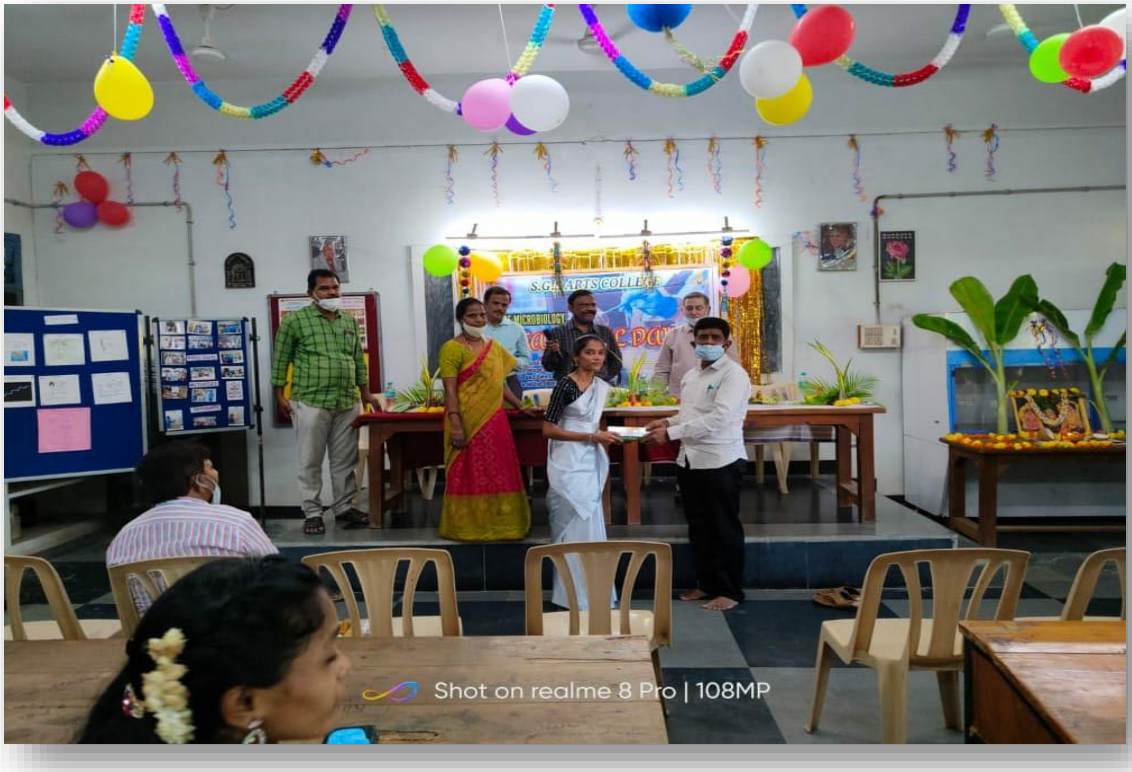
A quiz competion is conducted by the Department of **MICROBIOLOGY** , in the view of Competitive awareness on **26-08-2019** .

Batch. No	Name of the Student	Class	Roll No	Student Signature	Marks Awarded	Total	Remarks
1	**Chakrapani Tharun**	**MZC**	18704	Chakrapani Tharun	**20**	8	
	D.Gawtham gowda	**MZC**	18711	D.Gawthamgowda		7	
	K.E.pavan kumar	**MZC**	18718	K.E.Pavankumar		5	
2	**E.Reddy sekhar**	**MZC**	18724	E.Reddysekhar	**23**	8	
	K.Sai krishna	**MZC**	18727	K.Sai krishna		8	
	P.Subramanyam	**MZC**	18730	P.Subramanyam		7	
3	R.Gayathri	**MZC**	18708	R.Gayathri	**25**	8	**WIN**
	M.Gurramma	**MZC**	18712	M.Gurramma		9	
	K.Kalyani	**MZC**	18713	K.Kalyani		8	
4	**G.Nagarjuna**	**MZC**	18716	G.Nagarjuna	**19**	7	
	E.Naveen	**MZC**	18717	E.Naveen		7	
	B.N.Rangadu	**MZC**	18722	Rangadu		5	

Organizing Committee:

1. Dr. K. Sri devi K. Sridevi

*** Photo Evidence**



Tirupati
Date: 02-09-2019

From

The HOD of Computer Science
S.G.S. Arts College,
Tirupati.

To

The Principal,
S.G.S. Arts College,
Tirupati.

Respected Sir/Madam,

Sub: - Permission for conducting Quiz programme – Request – Regarding.

***

We are conducting Quiz programme for II & III year students of B.Sc(Mpcs & Mscs)

Em in Room No 214. So, kindly grant us permission on 04-09-2019.

Thanking you Sir/Madam,

Yours faithfully

HOD
HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

COMMITTEE FOR QUIZ PROGRAMME

ACADEMIC YEAR (2019-2020)

S.NO.	NAME OF THE COMMITTEE MEMBERS	SIGNATURE
1	Smt. B.Triveni, Coordinator	
2	Smt N.Jayalakshmi	
3	Sri D.Chaithanya Kumar	

SIGNATURE OF THE HOD

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: Computer Science

Academic year: (2019-2020)

S.No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	QUIZ COMPETITION	04-09-2019	III B.Sc (Mpcs & Mscs)	20 (4 Batches)	

HOD

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of Computer Science
EVENT ORGANISED REPORT: (2019-2020)

1	Name of the Department	**Computer Science**
2	Name of Event Organized	**QUIZ COMPETITION**
3	Title of Event Organized	**Object Oriented Programming using Java**
4	Name of Coordinator of Event	**Smt B.Triveni**
5	Class of the Participants	**II & III B.Sc MPCS, MSCS [EM]**
6	No. of Participants	**20**
7	Objective of the Event	• To evaluate the knowledge of the participants within academics as well as beyond academics and to make them familiar with the prospects of quizzes and the objectivity of the questions.
8	Outcome of the Event	1. There is clear evidence from psychological experiments that practicing retrieval of something after learning it, for instance by taking a quiz or test, makes us more likely to retain it for the long term. 2. Testing identifies gaps in knowledge. 3. Essentially it reduces forgetting which makes the next related study area more productive.

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF COMPUTER SCIENCE

ACADEMIC YEAR (2019-2020)

S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME	DATE OF CONDUCTION
	II & III B.Sc	MPCS	Object Oriented Programming using Java	04-09-2019
		MSCS		

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

QUIZ - DEPARTMENT OF COMPUTER SCIENCE

ACADEMIC YEAR (2019-2020)

S.NO.	PROGRAMME NAME	PROGRAMME CODE	TITLE OF THE QUIZ PROGRAMME		TOTAL NO. OF BATCHES	WINNING BATCH
1	III B.Sc	MPCS	17495	Bhargav Embeti	A	1ST PRIZE
2			17496	Brahmaiah Sandu		
3			17505	Mahesh Avisineni		
4			17513	Prabhakar Badikela		
5			17526	Surekha Pulluru		
6		MSCS	17541	Allipeera Dudekula	B	
7			17545	Bhuvanesh Ragam		
8			17550	Ganesh Kumar Gujjala		
9			17513	Prabhakar Badikela		
10			17518	Ravali Baduru		
11	II B.Sc	MPCS	18522	Hemasri Kuppa	C	
12			18524	Jayalakshmi Kailasam		
13			18525	Mabbu Sharif Gajuguduru		
14			18543	Sandhya Putturu		
15			18558	Varshini Annem		
16		MSCS	18577	Deepa Kadirimangalam	D	2ND PRIZE
17			18586	Indhumathi Lalapeta		
18			18589	Lakshmi Manne		
19			18609	Venkata Chanukya Kethineni		
20			18610	Venkat Nisanth Kanupuru		

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

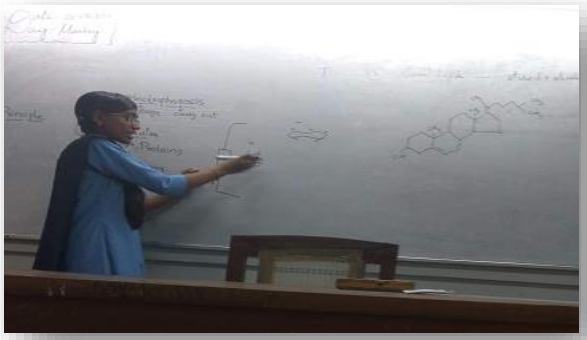
Photo Gallery

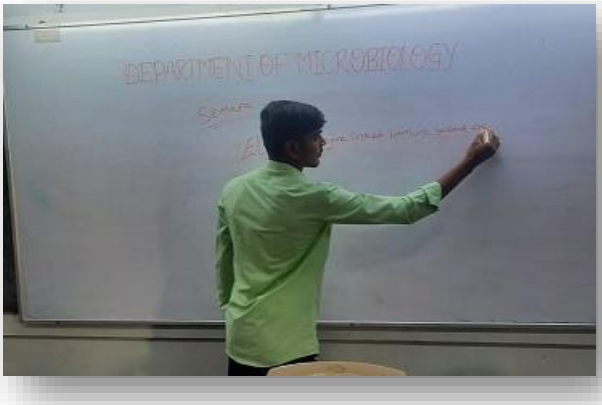
2. STUDENT SEMINAR

The college promotes the learning capabilities of the students by making them engaged in seminars on the topic of their choice. This ensures the students to learn about the latest information and new skills related to the concerned subjects.

Student seminars conducted in the department of sankrit for the academic year 2017-18 on the topic “UNNATHI” .

Student seminars conducted in the department of chemistry for the academic year 2019-20 on the topic “Factors affecting distributionlaw” .





Student seminars conducted in the department of Microbiology for the academic year 2019-20 on the topic “General properties of liquids” .

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

Sri Govindaraja Swamy Arts College

Department of Computer Science

STUDENT SEMINAR

Date : 25-09-2018

OBJECTIVES:

1. To progress the communication skills by motivating them to give seminars in different topics.
2. To make students Expertise in Verbal Communication.
3. To remove stage fear in students.
4. To explore expression of the students and also to gauge their awareness quotient.

CONVENER

- Sri P.Udaya, HOD, Department of Computer Science

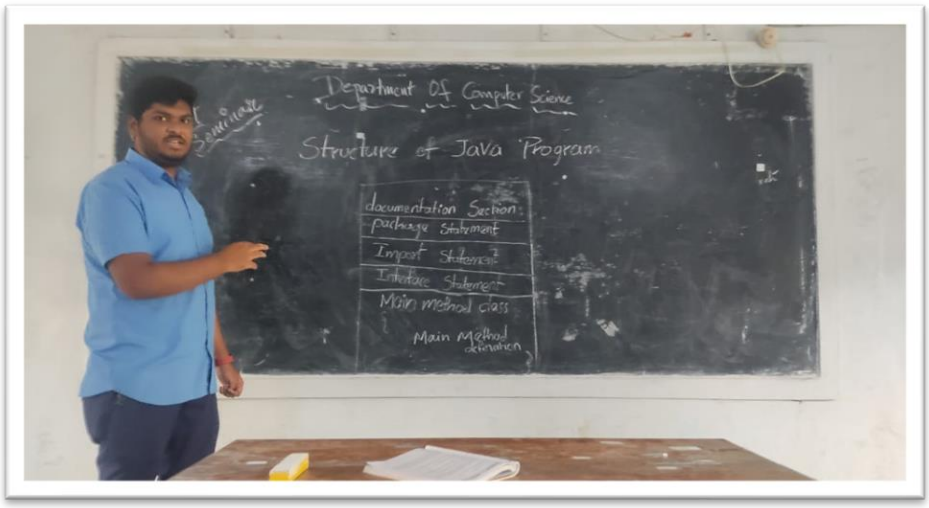
ORGANIZING MEMBERS

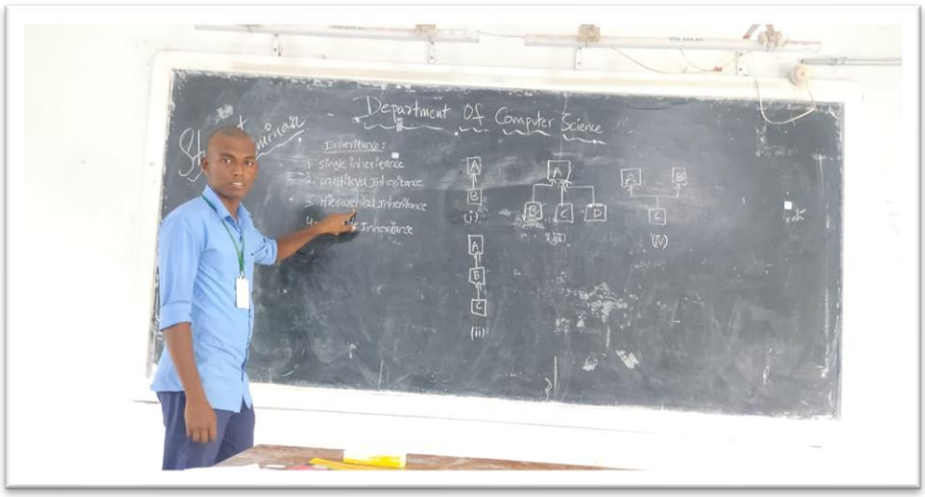
- SmtB.Triveni
- SmtN.Jayalakshmi
- Sri D.Chaitanya Kumar

TOPIC: **Fundamentals of Java**

- Fundamentals of OOP
- Structure of Java Program
- Control Structures in Java
- Types of Inheritance
- Interface
- Thread Life Cycle
- Applet Life Cycle

PHOTO GALLERY:





Attendance Sheet for IIB.Sc. (MPCS) EM

Sno	Roll No	Name of the Student	Signature
1	17491	Anjan KumarPottiramigalla	P. Anjan Kumar
2	17492	ArjunKarli	K. Arjun
3	17493	AshokKanimeni	K Ashok
4	17500	JosephMerugumolu	M. Joseph
5	17501	JyothsnaKondru	K. Jyothsna
6	17503	LokeshKommala	K. Lokesh
7	17504	MadhuUndela	U. Madhu
8	17505	MaheshAvisineni	A. Mahesh
9	17506	ManoharBandari	B. Manohar
10	17509	MohiniIrri	I. Mohini
11	17510	Om PrakashAaramadaka	A. Om Prakash
12	17511	Pavan KumarAnderi	A. Pavan Kumar
13	17513	PrabhakarBadikela	B. Prabhakar
14	17514	Praneeth KumarUppuluru	U. Praneeth Kumar
15	17515	PrasanthMuppuri	M. Prasanth
16	17517	Ram PrasadKorlakunta	K. Ram Prasad
17	17518	RavaliBaduru	B. Ravali
18	17519	Ravi Durga Prasad Donti	D. Ravi Durga Prasad
19	17520	Ravi TejaMekala	M. Ravi Teja
20	17521	Sai PrakashPenchili	P. Sai Prakash
21	17522	SatishVeligiram	V. Satish
22	17523	Shaik Adam Basha	Shaik Adam Basha
23	17524	SravaniPidugu	P. Sravani
24	17525	SreekanthChejarla	C. Sreekanth
25	17526	SurekhaPulluru	P. Surekha
26	17527	Suri BabuMotikala	M. Suri Babu
27	17528	TejaswiniManyam	M. Tejaswini
28	17531	UrukunduMaliga	M. Urukundu
29	17532	VedhavathiPamula	P. Vedhavathi
30	17533	VeerannaBoya	B. Veeranna
31	17534	VeereshBoya	B. Veeresh
32	17535	VeereshNallagunti	N. Veeresh
33	17537	Vignesh KumarRaja	R. Vignesh Kumar
34	17538	YaminiThatithoti	T. Yamini
35	17540	YeswanthPangaluri	P. Yeswanth

HEAD
DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Attendance Sheet for II B.Sc. (MSCS) EM

Sno	Roll No	Name of the Student	Signature
1	17541	AllipeeraDudekula	D. Allipeera
2	17542	AnjaliBodannagari	B. Anjali
3	17544	AnushaMinuku	M. Anusha
4	17545	BhuvaneshRagam	R. Bhuvanesh
5	17546	ChalapathiArava	A. Chalapathi
6	17547	Chandra SekharNama	N. Chandra Sekhar
7	17548	Deva.G	G. Deva
8	17550	Ganesh KumarGujjala	G. Ganesh kumar
9	17551	HemanthMekala	M. Hemanth
10	17552	HimavanthSanambatla	S. Himavanth
11	17553	JagadeswariSakam	S. Jagadeswari
12	17554	Kiran KumarChakali	C. Kiran Kumar
13	17555	LokeshChitteni	C. Lokesh
14	17557	LoknathUpendram	V. Loknath
15	17558	MadhuPeta	B. [illegible]
16	17559	ManikantaBalagani	B. Manik
17	17560	MeghanaMaddela	M. Meghana
18	17561	Muni Sai KumarPonna	P. M[illegible]
19	17562	NarsappaPandikona	P. Narasappa
20	17563	Raja Dhannasi	D. Raja
21	17564	RajuSingadi	S. Raju
22	17565	Rethish Sai.JM	M. Rethish Sai
23	17566	RupeshKanakappagari	K. Rupesh
24	17567	Sai KumarAnthati	A. Sai kumar
25	17571	Santhosh KumarKantineni	K. Santhosh
26	17573	Sathish KumarGurrappagari	G. Sathish Kumar
27	17574	SrinivasuluBapana	B. Srinivasulu

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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

Sno	Roll No	Name of the Student	Signature
28	17575	Sumalatha.B	B. Sumalatha
29	17576	SunilPasupula	P. Sunil
30	17577	SupriyaDidi	D. Supriya
31	17578	SureshGonimadathala	G. Suresh
32	17579	SuryaGuggilla	G. Surya
33	17584	VenugopalPutakala	P. Venu
34	17585	VinayJadapalli	J. Vinay
35	17586	Vishnu VardhanMamuduru	M. Vishnu Vardhan
36	17587	YeswanthGownipalli	G. Yeswanth

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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

REPORT:

The department of Computer Science has magnificently organised a seminar for II BSc MSCs-EM and MPCs-EM students on the above listed topics .Students of II BSc MSCs-EM and MPCs-EM have actively participated in this event and given away seminars on their respective topics .This Seminar was able to Encourage, motivate and Improve Communication skills for the better performance of students. Also this seminar enriched the knowledge in Fundamentals of Java Programming related topics. The Students and Lecturers of the Department of computer science participated the event and made this programme grand success.

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DEPARTMENT OF COMPUTER SCIENCE
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department: Statistics

Academic year:2019-20

S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	SEMINAR	25/02/2020	B.Sc: (MSCs MPS) BA(MES)	50	-

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of STATISTICS

EVENT ORGANISED REPORT: 2019-20

1	Name of the Department	**Statistics**
2	Name of Event Organized	**Seminar**
3	Title of Event Organized	**Role of Statisticians – Census survey and Analysis**
4	Name of Co-ordinator of Event	**Dr.D.Chandra Kesavulu Naidu**
5	Resource Person	**Prof.B.Sarojamma,Board of Studies Chair person, Dept of Statistics, S.V.University.**
6	Class of the Participants	**B.Sc:MSCs MPS BA: MES**
7	No. of Participants	**50**
8	Objective of the Event	➢ To create awareness in students about **Role of Statisticians – Census survey and Analysis**
9	Outcome of the Event	➢ Became aware of process of Census survey in India ➢ Understood the role of Statisticians in the process

HOD

D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

A seminar was organised and Conducted by the Head Of Dept; Statistics Dr. D. Chandra Resavulu Naidu, for the students of Statistics (BA, BSc).

Topic: "Role of Statisticians - Census Survey and Analysis".

Resource Person: Prof. B. Sarojamma
Board of Studies - Chairperson
Dept of Statistics
S.V. University, Tirupati.

All the staff members and students of the dept attended the seminar. Students were made aware of the process of Census Survey in India, its importance and role of Statisticians.

Staff: 1) Dr. K. VISHAKUMAR
2) Dr. K. VENKATESH
3) Dr. M. VASUDEVA REDDY

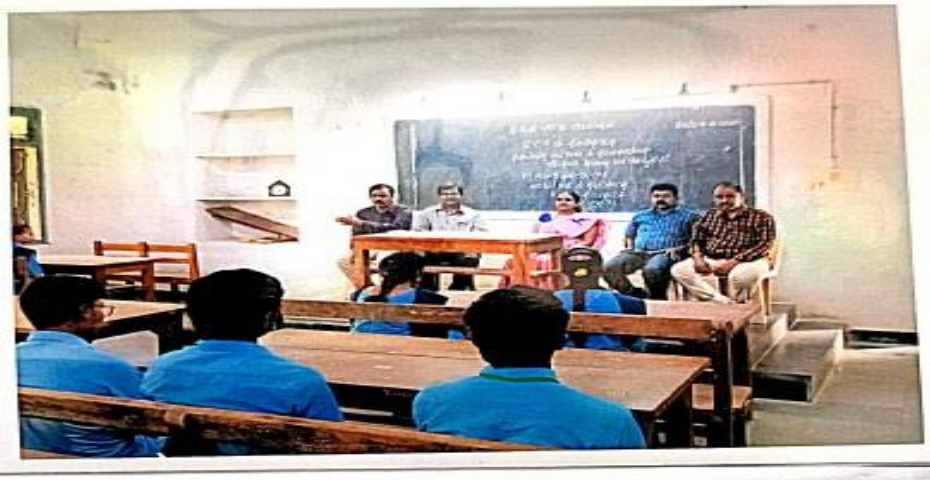
D. CHANDRA RESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

50

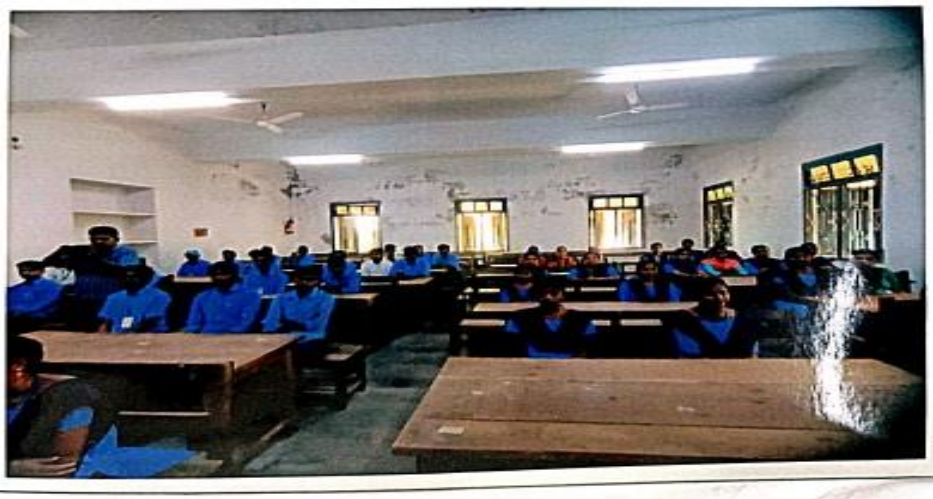
1	B. Mahesh	B. Mahesh.
2	D. Mounika	D. Mounika
3	B. Lavanya	B. Lavanya
4	K. Akhil	K. Akhil
5	M. Ashok	M. Ashok
6	P. Bhadri	P. Bhadri
7	M. varma teja	M.
8	U. Naveen	U. Naveen
9	V. Venu Gopal	V. Venu gopal
10	P. sambayya	P. Sambayya
11	K. Anil kumar	K.
12	D. Lakshmini	D. Lakshmini
13	O. vignesh	O.

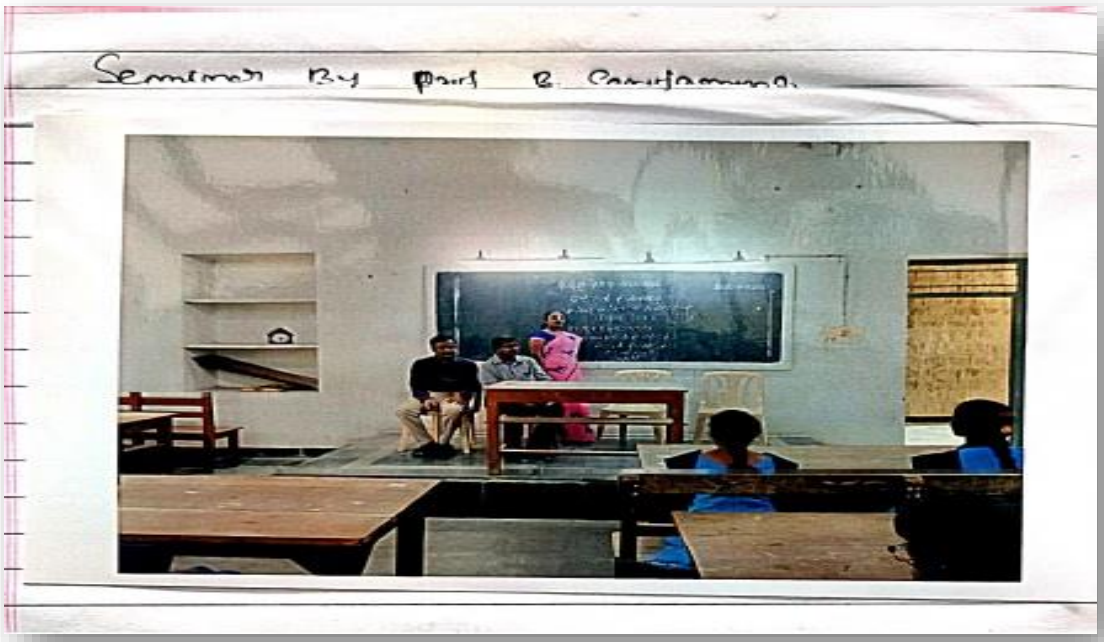
14. T. Uma Maheswari
15. B. Siva Jyothi
16. M. Keerthi
17. N. Sireesha
18. T. Raju
19. P. Vamsi
20. S. Manoj Kumar
21. B. Anand
22. R. Chandu
23. P. Devi
24. C. Latha
25. G. Bhumika
26. Gr. Lokeshwari
27. E. Aruna
28. K. Pooja
29. S. Sunil Kumar
30. P. Suresh
31. N. Suresh Babu
32. P. Pavan Kalyan
33. T. Suresh
34. M. Sandhya
35. V. Vamsi
36. V. Mohan
37. Y. Govardhana
38. Y. Venkata Subbaiah
39. P. Sidhayya
40. G. Sai Teja
41. M. Narendra
42. V. Leela Prasadh
43. TGS Kishore
44. P. Jashwanth
45. S. Abdul Samadh
46. P. Bhanu Prakash
47. N. Muni Chaitanya
48. M. Dinesh
49. G. Mohan Krishna
50. GD. Gopal

Seminar on "Role of Statisticians in Crop Survey and Analysis" by Prof. B. Sarojamma.



Staff and Students participated in the Seminar.




Seminar By Prof. B. Ravikumar.

3. Elocution & Essay writing competition

Elocution competition helps the students to develop their self expression and confidence. Elocution competition enable the students to improve their articulation and pronunciation of speech and oratory skills.

Writing encourages and enhances critical thinking and helps in organizing thoughts and improve communication skills. It provides an opportunity to the students to express their thoughts ,feelings,ideas and beliefs. In this direction the college organized elocution and essay competitions.

Students participating in Elocution competitions organized by the department of sanskrit for the academic year 2020-21 on the topic “Importance of Upanishad”.

4. Coaching for competitive examinations

The college has conducted coaching for competitive examinations to promote the employability of the students.

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY, M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

Sno	Academic Year	No of Students benefitted by Guidance for Competitive Examination and career counseling
1	2020-2021	NIL
2	2019-2020	363
3	2018-2019	296
4	2017-2018	290
5	2016-2017	282

S.G.S. ARTS COLLEGE Co-ordinator NAAC

PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web: sgsac.edu.in

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5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

ACADEMIC YEAR 2019-2020

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	104
2	PG CET Coaching	Department of Statistics	30
3	PG CET Coaching	Department of Mathematics	23
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	44
6	PG CET Coaching	Department of Commerce	88
7	PG CET Coaching	Department of Microbiology	7
8	PG CET Coaching	Department of Biotechnology	27
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			363

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5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

ACADEMIC YEAR 2018-2019

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	100
2	PG CET Coaching	Department of Statistics	33
3	PG CET Coaching	Department of Mathematics	20
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	22
6	PG CET Coaching	Department of Commerce	46
7	PG CET Coaching	Department of Microbiology	10
8	PG CET Coaching	Department of Biotechnology	25
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			296

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ACADEMIC YEAR 2017-2018

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	95
2	PG CET Coaching	Department of Statistics	29
3	PG CET Coaching	Department of Mathematics	26
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	23
6	PG CET Coaching	Department of Commerce	41
7	PG CET Coaching	Department of Microbiology	14
8	PG CET Coaching	Department of Biotechnology	22
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			290

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PRINCIPAL (FAC)

5.1.4.1: Number of students benefitted by guidance for competitive examinations and career counseling offered by the institution year wise during last five years

ACADEMIC YEAR 2016-2017

Sno	competitive examinations and career counseling	Name of the organizer	No of Students Participated
1	Coaching for Competitive Examinations	Career Guidance and Placement Cell, S.G.S. Arts College	105
2	PG CET Coaching	Department of Statistics	28
3	PG CET Coaching	Department of Mathematics	19
4	PG CET Coaching	Department of Chemistry	20
5	PG CET Coaching	Department of Physics	23
6	PG CET Coaching	Department of Commerce	48
7	PG CET Coaching	Department of Microbiology	4
8	PG CET Coaching	Department of Biotechnology	15
9	PG CET Coaching	Department of Economics	20
Total Number of Students Participated in competitive examinations and career counseling			282

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S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the Department: Chemistry

Academic year:2019-20

S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	PGCET	13/12/2019	B.Sc : MPC,BBC,MZC&CBZ	20	-

HOD
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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S. ARTS COLLEGE, TTD, TIRUPATI
Department of CHEMISTRY
EVENT ORGANISED REPORT: 2019 – 20

1	Name of the Department	**Chemistry**
2	Name of Event Organized	**PGCET Coaching**
3	Title of Event Organized	**PGCET Coaching – Chemistry**
4	Name of Co-coordinator of Event	**Sri.D.Parameswara**
5	Class of the Participants	**B.Sc,MPC,BBC,MZC&CBZ**
6	No. of Participants	**20**
7	Objective of the Event	An ability to develop critical thinking and efficient in problem solving skills
8	Outcome of the Event	**To provide maximum scope for succeeding higher education**

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S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY, M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

DEPARTMENT OF CHEMISTRY

CIRCULAR

31-10-2019

S.G.S A.C / Chem / 2019 / DM / 04

The faculty members of department of chemistry are informed that there will be a staff meeting at 9.00 A.M in the HOD chamber on 01-11-2019 to discuss about P.G CET coaching to be conducted for the final year students for the academic year 2019-20.

S.No.	Name of the Faculty	Mobile No.	Signature
1	Sri D. Parameswara	9490728655	
2	Dr. V. Venkata Lakshmi	9492075161	
3	Dr. P. Suguna	7989925972	
4	Sri. K. Revati	9177721917	

AGENDA

- To discuss about of P.G CET coaching to be conducted for the final year students for the academic year 2019-20.
- To finalize schedule for the P.G CET coaching
- To finalize timetable for the P.G CET coaching
- To discuss workload distribution for the P.G CET coaching

MINUTES

- The faculty members discussed about the schedule and time table for the P.G CET coaching and made the following resolutions unanimously.
- It was planned to conduct P.G CET coaching for the final year students in the second week of December 2019.
- Sri.D.Parameswara was made Coordinator for the P.G CET coaching.
- The workload distribution for the P.G CET was as shown bellow.
- It was planned to conduct P.G CET coaching classes at the rate of 3hrs per week on every Friday from 8.30 to 10 AM & 4.00 to 5.30 PM.

S.No	Name of the Faculty	Subject	No of Hours
1	Sri D. Parameswara	Physical chemistry	21
2	Dr. V. Venkata Lakshmi	Inorganic chemistry	20
3	Dr. P. Suguna	Organic chemistry	9
4	Smt.K.Revati	Organic chemistry	10

Signatures

1.
2.
3.
4.

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Permission letter

02-11-2019

From

D. Parameswara,

HOD, Dept of chemistry,

S.G.S Arts College,

Tirupati.

To

The Principal,

S.G.S Arts College,

Tirupati.

Respected Sir,

Sub: Request for permission to conduct coaching classes for P.G CET –Reg.

We have planned to conduct coaching classes for P.G CET in chemistry for the final year students of B.Sc with chemistry combination from the second week of December 2019. The coaching classes enable the students to perform better in P.G CET and other entrance examinations conducted for admission into M.Sc. So, please permit us to conduct the side coaching classes.

Thanking you

Yours faithfully

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

INTIMATION

SGSAC/Chem dept/PGCET 2019/HOD/01 04-11-2019

All the final year students of B.Sc with chemistry combination are informed that coaching for P.G CET has been scheduled to commence from 13-12-2019.The classes are conducted at the rate of 3hrs per week and every Friday from8.30 to 10 AM & 4.00 to 6.00 PM. in room number 129. So, the interested students are instructed to enroll their name on or before 07-11-2019.

H.O.D of the Department

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

To

1. All teaching staff of chemistry Department for circulation.
2. Department notice board.

DEPARTMENT OF CHEMISTRY
PGCET COACHING PLAN

PROGRAMME SCHEDULE : (WEEKLY 3HRS) VENUE : Room No.129

week	Date	Time	Topic	Faculty incharge
week-1	13/12/2019	8.30-10.30	S-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Structural theory in organic chemistry	Dr.B.Sathynarayana
	14/12/2019	8.30-10.00	P-Block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
week-2	20/12/2019	8.30-10.00	Organo metallic compunds	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
	21/12/2019	8.30-10.00	d-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Hydro carbons	Dr.B.Sathynarayana
week-3	27/12/2019	8.30-10.00	f-block elements	Dr.V.Venkata lakshmi
		4.00-5.30	Halogen compounds ,Hydroxy compounds	Dr.B.Sathynarayana
	28/12/2019	8.30-10.00	Bonding in metals	Dr.V.Venkata lakshmi
		4.00-5.30	Carbonyle compounds	Dr.B.Sathynarayana
week-4	03/01/2020	8.30-10.00	Metal carbonyles	Dr.V.Venkata lakshmi
		4.00-5.00	Carboxyllic acids	Dr.B.Sathynarayana
	04/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Carbohydrates	Dr.P.Suguna
week-5	10/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Amino acids &Nitro compounds	Dr.P.Suguna
	11/01/2020	8.30-10.00	Co-ordination chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Amino acids &Nitro compounds	Dr.P.Suguna
week-6	24/01/2020	8.30-10.00	Analytical chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Hetero cyclic compounds	Dr.P.Suguna
	25/01/2020	8.30-10.00	Metallurgy	Dr.V.Venkata lakshmi
		4.00-5.30	Stereo chemistry	Dr.P.Suguna

week-7	07/02/2020	8.30-10.00	Bioinorganic chemistry	Dr.V.Venkata lakshmi
		4.00-5.30	Stereo chemistry	Dr.P.Suguna
	08/02/2020	8.30-10.00	Gaseous state	Sri.D.Parameswara
		4.00-5.30	Solid state	Sri.D.Parameswara
week-8	14/02/2020	8.30-10.00	Solutions	Sri.D.Parameswara
		4.00-5.30	Surface chemistry& colloids	Sri.D.Parameswara
	15/02/2020	8.30-10.00	Electro chemistry	Sri.D.Parameswara
		4.00-5.30	Electro chemistry	Sri.D.Parameswara
week-9	21/02/2020	8.30-10.00	Phase rule	Sri.D.Parameswara
		4.00-5.30	Colligative properties of dilute solutions	Sri.D.Parameswara
	22/02/2020	8.30-10.00	Chemical kinetics	Sri.D.Parameswara
		4.00-6.00	Chemical kinetics	Sri.D.Parameswara
week-10	27/02/2020	8.30-10.00	Thermo dynamics	Sri.D.Parameswara
		4.00-5.30	Thermo dynamics	Sri.D.Parameswara
	29/02/2020	8.30-10.00	Photo chemistry	Sri.D.Parameswara
		4.00-5.30	Photo chemistry	Sri.D.Parameswara

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DEPARTMENT OF CHEMISTRY
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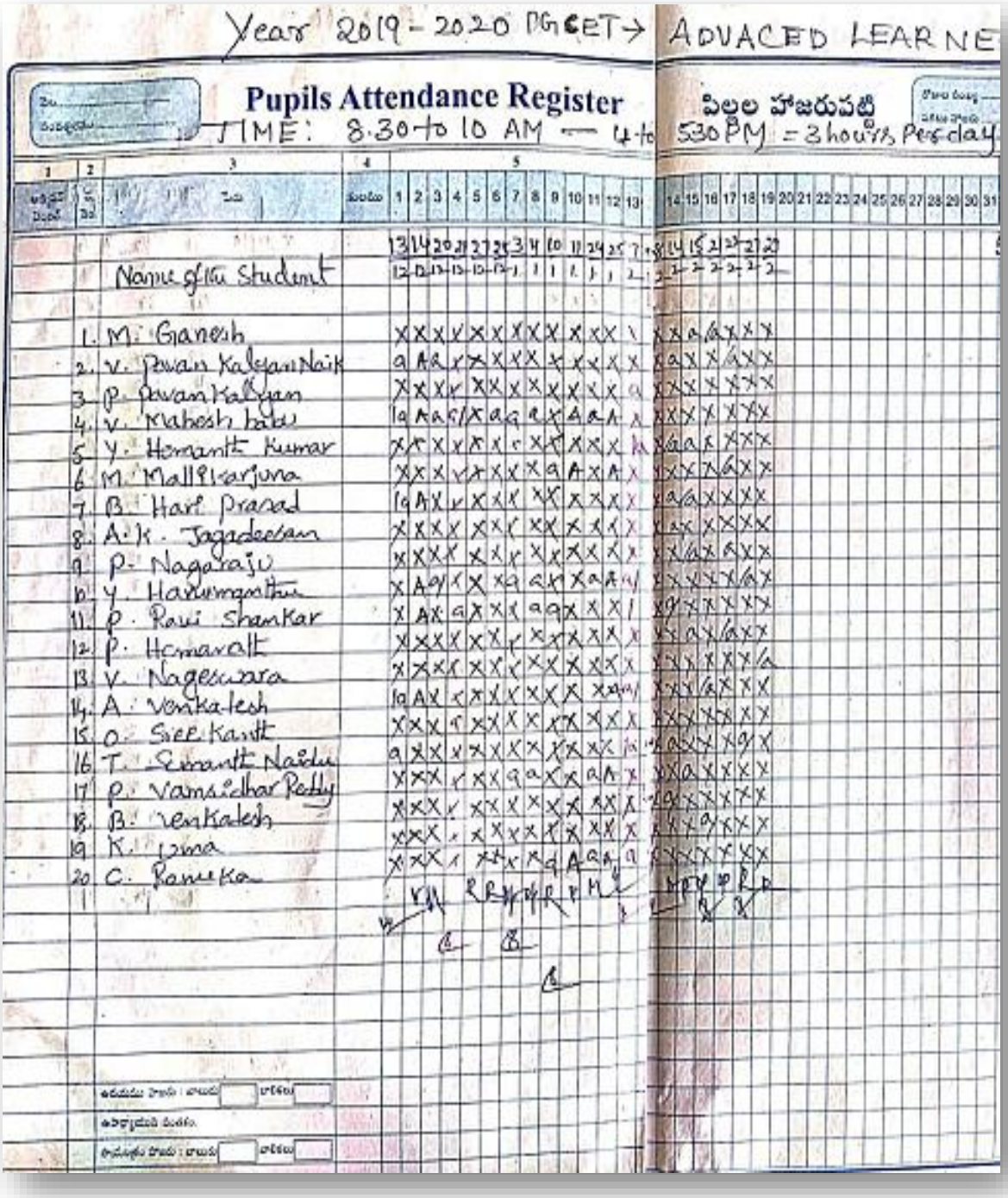

Year 2019 - 2020 PG CET → ADVACED LEARNE

Pupils Attendance Register

TIME: 8.30 to 10 AM — 4 to 5.30 PM = 3 hours Per day

పిల్లల హాజరుపట్టి

		Name of the Student
1		M. Ganesh
2		V. Pawan Kalyan Naik
3		P. Pawan Kalyan
4		V. Mahesh babu
5		Y. Hemanth Kumar
6		M. Mallikarjuna
7		B. Hari Prasad
8		A. K. Jagadeesam
9		P. Nagaraju
10		Y. Harikrishna
11		P. Ravi Shankar
12		P. Hemavathi
13		V. Nageswara
14		A. Venkatesh
15		O. Sree Kanth
16		T. Seemanth Naidu
17		P. Vamsidhar Reddy
18		B. Venkatesh
19		K. Uma
20		C. Renuka

DEPARTMENT OF CHEMISTRY

PGCET COACHING , STUDENTS MARKS LIST

S.NO	Name of the student	Test-1 05/01/2020 (Maximum marks)	Test-2 09/02/2020 (Maximum marks)	Grand Test 08/03/2020 (Maximum marks)
1	MANDADHI GANESH	46	45	69
2	V PAVAN KALYAN NAIK	45	39	71
3	P PAVAN KALYAN	45	42	68
4	VAKITI MAHESH BABU	39	41	AB
5	Y HEMANTHKUMAR	43	45	80
6	MALLELA MALLIKARJUNA	42	49	73
7	BYRICETTY HARI PRASAD	48	47	82
8	A K JAGADEESAN	47	45	81
9	Pathina nagaraju	49	43	79
10	YERUKALA HANUMANTHU	Ab	43	75
11	PEDDAULITHI RAVI SHANKAR	41	45	68
12	POOJARI HEMANATH	45	46	67
13	VELPULA NAGESWARA	47	49	83
14	Arigela Venkatesh	48	AB	85
15	ODUGURU SREEKANTH	43	45	71
16	TURRA SUMANTH NAIDU	41	43	85
17	P.VAMSIDHAR REDDY	42	45	66
18	B.VENKATESH	47	43	65
19	K.UMA	42	43	80
20	C.RENUKA	41	40	81

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S.G.S ARTS COLLEGE
STUDENT FEED BACK REPORT

STUDENT FEED BACK FORM:

Date: **Academic year: 2019-20**

Semester: VI **Name of the programme : Coaching for PGCET**

Department :Chemistry

If you have attended our PGCET coaching classes we encourage you to submit feedback for us. This will help us to improve our future sessions and better meet the needs of our faculty.

S.NO	CRITERIA	RATING				
1	Content of the programme	Very poor	Poor	Average	Good	Excellent
2	Presentation style &delivery	Very poor	Poor	Average	Good	Excellent
3	Was it motivating	Very little	Less	Average	Lots	Too much
4	was the time given for questions	Very little	Less	Average	Lots	Too much
5	Was it useful for future	Very useless	Useless	Average	Useful	Very much
6	To What extent you have learnt from this programme	Very little	Less	Average	Lots	Too much
7	What is your all rating to this programme	Very poor	Poor	Average	Good	Excellent

General Comments: It has motivated me to study well.

Students Signature:

S.G.S ARTS COLLEGE :TIRUPATI
DEPARTMENT OF CHEMISTRY
PGCET COACHING FOR THE ACADEMIC YEAR 2019-20

NOTICE

14-12-2019

Students of PGCET coaching are informed that the test schedule for PGCET coaching is as shown bellow.

Test-1	**: 05-01-2020**
Test-2	**: 09-02-2020**
Grand test	**: 08-03-2020**

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P.G. CET QUESTION PAPER :: 2019-20

TEST-I

Name of the Student :

Group **:** **Maximum Marks : 50**

Answer all the question

1. The compound the exhibits highest dipole moment is ( )
   a. Nitrobenzene b. O-nitrophenol
   c. m-nitrophenol d. p-nitrophenol
2. Nitrobenzene on reduction with tin and HCl gives ( )
   a. Nitrobenzene b. Aminobenzene
   c. Phenyl Hydroxyl Amine d. Hydrazo benzene
3. Coupling reaction is shown by ( )
   a. Pyrrole b. Furan c. Thiophene d. All the above
4. In Anthracene the electrophilic substitution occurs at ( )
   a. C-1 position b. C-6 position
   c. C-3 position d. C-9 position
5. The relative lowering of vapour pressure of a solution containing Non-volatile solute is ( )
   a. Proportional to mole fraction of the solvent
   b. Proportional to mole fraction of the solute
   c. Equal to the mole fraction of the solvent'
   d. Equal to the mole fraction of the solute
6. In vander Waal's equation of state for a real gas, the term which accounts for The molecular forces are ( )
   a. (V-b) b. R c.$[P + a/V^2]$ d. $(RT)^{-1}$
7. Azeotropic mixture of HCL and H_2O has _____HCl ( )
   a. 48.4% b. 22.2% c. 36.5% d. 20.24%

8. In a single – component system, if degree of freedom is zero, maximum number of phases that can co-exist is ( )
   a. 0 b. 1 c. 2 d. 3
9. Specific conductivity (k), and equivalent conductivity (y) of an electrolyte are related as ( )
   a. $k = \Lambda V$ b. $k = V / \Lambda$ c. $k = \Lambda / V$ d. $\Lambda = kV$

10. In the process of photosynthesis, the chlorophyll acts as ( )
    a. Catalyst b.Intiator

c. Photosensitizer d, Promotor

11. In *Joule – Thomson* expansion ( )

a. dS = 0 b. dG = 0 c. dE = 0 D. dH = 0

12. The instrument used to measure optical rotation is ( )

a. Polari meter b. Spectrometer c. Calorimeter d. Voltmeter

13. $A1_4C_3$, belongs to the category of ( )

a. Acetyl ides b. Allylides c. Mathanides d. Interstitial

14. The order of reactivity towards electrophilic substitution among benzene, pyrrole, furan and thiophene is ( )

a. Pyrrol>benzene>furan>thiophene
b. Thiophene>benzene>pyrrol>furan
c. Pyrrol>furan>thiophene>benzene
d. Pyrrol>thiophene>furan>benzene

15. Which of the following cycle alkanes is the most stable? ( )

a. Cyclo propane b. Cyclo butane
c. Cycle pentane d. Cycle hexane

16. Which of the following is the strongest acid? ( )

a. Phenol b. 2-Nitro phenol
c. 2, 4 Dinitro phenol d. Picric acid

17. The trade name of trichloroethylene is ( )

a. Teron b. Westron c. Westrosol d. DDT

18. What metal ion is the oxidizing agent in tollens' reagent? ( )

a.Cu^+ b. Ag^+ c. Fe^{2+} d. Fe^{3+}

19. Which of the following compounds most acidic? ( )

a. CH_3CI b. CH_3OH c. $H_2C = CH_2$ d. H-C = C-H

20. Which of the following cycloalkanes is most stable? ( )

a. Cyclopropane b. Cyclobutene
c. Cyclopentane d. Cyclohexane

21. Which of the following compounds form only a mono substituted?
Derivative on bromination ? ( )

a. N-butane b. Propane
c.2-methyl butane d. 2, 2-demithyl propane

22.The number of possible stereoisomers for 2, 3 dibromo butane are ( )

a. Two b. Four c. Three d. Six

23. Which of the following do not respond to 'Iodoform' test? ( )

a. Acetaldehyde b. Benzaldehyde
c. Acetophenone d. Methylethyl ketone

24. Which of the following is a rearrangement reaction? ( )

a. Wurtz reaction b. Claisen condensation
c. Pinacol-pinacoline reaction d. Sulphonation

25. Which of the following prove the wave nature of electron? ( )

a. Uncertainty principle b. Photoelectric effect

c. Electron diffraction d. Zeeman Effect

26. When sodium chloride is added to a mixture of oil and caustic soda, the soap gets precipitated because ( )
    a. NaCl is ionic compound
    a. Of an increase in concentration of Na^+ ion
    b. Of the insolubility of the soap in presence of Cl^- ions
    c. Of decrease in the solubility product of Nacl in presence of the soap

27. For a gas the deviation from ideal behaviour is maximum at ( )
    a. 0^o C, 1 atmos b. -10^o C, 1 atmos
    c. 0^o C, 5 atmos d. -10^oC, 5 atmos

28. The degrees of freedom at eutectic point ( )
    a. 0 b. 1 c. 2 d.3

29. In which of the following processes, the entropy will decrease? ( )
    a. Boiling of water b. Freezing of a liquid
    c. Expansion of a gas d. Melting of ice

30. The equivalent conductance at infinite dilution is equal to ( )
    a. The sum of the ionic conductances of cation and anion
    b. The sum of the transport numbers of the cation and anion
    c. The sum of the specific conductivity and cell constant
    d. The sum of equivalent and specific conductivities

31. Inversion temperature is ( )
    a. The temperature at which a gas liquefies
    b. The temperature to which a gas should be cooled before applying pressure
    c. The temperature above which the liquid cannot exist
    d. The temperature at which both gas and liquid can coexist

32. Gold number means ( )
    a. The number of gold atoms present in a standard gold sol
    b. The minimum amount of an electrolyte which should be added to a Standard gold sol which just prevents its coagulation
    c. The minimum amount of gold chloride which should be dispersed in water to give a stable gold sol
    d. The number of moles of a protective colloid which should be added to a standard gold sol to prevent its coagulation

33. The two effects associated with the Debye-Huckel theory are: ( )
    a. Relaxation effect and asymmetry effect
    b. Relaxation effect and mesomeric effect
    c. Relaxation effect and electrophoretic effect
    d. Only mesomeric effect

34. The light emitted by a glowworm is an example of ( )
    a. Fluorescence b. Phosphorescence
    c. Photosynthesis d. Chemiluminescence

35. Which of the following ion has a magnetic moment of 3,87 B.M.? ( )

a. Mn^{2+} b. Fe^{3+} c. Cr^{3+} d. Sc^{3+}

36. The molecular shape of XeF_4 is ( )

a. Linear b. Square

c. Square planar d. Pentagonal bipyramid

37. The strong field ligand is ( )

a. en b. NH_3 c. NO_2^{θ} d. CN^{θ}

38. The most important air pollutants responsible for depletion of the Ozone layer are ( )

a. Oxides of nitrogen b. Oxides of sulphur

C. Oxides of halogens d. Chlorofluoro hydrocarbons

39. Which of the following ions is colourless? ( )

a. Sc^{3+} b. Ti^{3+} c. Fe^{3+} d. V^{3+}

40. The molarity of pure water ( )

a. 55.6 b. 5 c. 100 d. 18

41. The half life of radon is 3.8 days. The time (in days) by which its 1/20 of the Amount will be left behind is ( )

a. 9 approx. b. 16.45 c. 7.6 d. 1.9

42. The function of haemoglobin in red blood cells is ( )

a. To transport dioxygen from lungs to tissues

b. To convert CO_2 of the tissues (absorbed from environment) to HCO_2 ions

c. The O_2 transforms haemoglobin to oxyhaemoglobin which is facilitated PH Fe^{2+} ions for storage of energy in lungs, but not in muscles

e. The affinity for haemoglobin towards O_2 decreases with pH and ligands like CO, CN etc. reduce it to deoxyhaemoglobin, which is not good O^2 carrier

43. The most stable oxidation state of lanthanides is : ( )

a. +4 b. +2 c. +3 d. +1

44. The indicator employed in the complexometric titration of magnesium using E.D.T.A. is : ( )

a. Phenolphthalein b. Methyl Orange

c. Erichrome Black – T d. Catechol Violet

45. According to Fajan's rules, the bond will acquire covalent character, if the cation is ( )

a. Large Size b. Small Size

c. Largest Size d. Independent of its size

46. Which of the following compounds undergo cannizaro's reaction when treated with an alkali? ( )

(a) $CH_3CH_2CH_2CH_2CHO$ (b) $(CH_3)_3CHCHCH_2CHO$

(c) $CH_3CH_2CH(CH_3)CHO$ (d) $(CH_3)_3C\text{-}CHO$

47. Which of the following is six membered ring ( )

(a) Pyrrole (b) Pyridine

(c) Galactose (d) Maltose

48. The order of molecularity of reaction of inversion of cane sugar respectively

is ( )

(a) 1,2 (b) 2,1 (c) 1,1 (d)2,2

49. The quantum yield of H_2 and Br_2 reaction in the presence of light is ( )

(a)1.0 (b) 0.1 (c) 0.01 (d) 10

50. During fermentation, the mixture is stirred well because ( )

(a) Of heating (b) co_2 is liberated

(c)Yeast is present (d) Alcohol is obtained

ANSWERS									
1	d	11	d	21	d	31	b	41	b
2	b	12	a	22	C	32	d	42	c
3	d	13	c	23	D	33	d	43	c
4	d	14	c	24	C	34	d	44	c
5	b	15	d	25	C	35	c	45	b
6	a	16	d	26	B	36	c	46	b
7	d	17	c	27	a	37	d	47	b
8	c	18	b	28	a	38	d	48	a
9	c	19	d	29	b	39	a	49	c
10	a	20	d	30	a	40	a	50	b

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TEST-II

1. Nitrobengene combines with hydrogen in the presence of platinum to yield (c)
   (a) Aniline (b)Benzenze
   (c)Toluene (d) Azobenzenze
2. Aromatic compounds undergo (d)
   (a) Nucleophilic addition reactions
   (b) Nucleophilic substitution reactions
   (c) Electrophilic addition reactions
   (d) Electrophilic substitution reactions
3. Which of the following bases is most likely to favour SN^1 reaction rather than SN^2 reaction in the hydrolysis of an alkyl halide? (c)
   (a)OH^{---} (b) OEt^{---} (c)H_2O (d)NH_2^{--}
4. The natural amino acids are (a)
   (a) D – rotatory (b) L - rotatory
   (c) L-Configuration (d) D- Configuration
5. Camphor can be easily purified by the process of (c)
   (a) Sublimation (b) Distillation
   (c)Crystallisation (d) Sedimentation
6. The number of acidic hydrogen atoms in 1-butyne and 2-butyne respectively are (b)
   (a)1,0 (b)0,1 (c)1,1 (d)1,2
7. HNC and HCN are (a)
   (a) Functional isomers (b) Tautomers
   (c)Chain isomers (d) Position isomers
8. HNC and HCN are (b)
   (a)Functional isomers (b)Tautomers
   (c)Chain isomers (d)Position isomers

9. Grignard's reagent react with ketones by addition or hydrolysis forming (c)
   (a)Primary alcohol (b)Secondary alcohol
   (c)Tertiary alcohol (d)Aldehyde
10. The reagent used for distinguishing alcohols is (b)
    (a)Br_2/CCl_4 (b)Lucas reagent
    (C)Schiffs reagent d)$Fecl_3$
11. The product of destructive distillation of wood is (b)
    (a)Acetic acid (b)Ethyl alcohol
    (c)Soap (d)Kerosene
12. Ammonolysisof an ester leads to the formation of (b)

(a)Uride (b)Amide
(c)Amine (d)Nitrite

13.Phenyl isocyanide test is used to identify (c)
(a)Aromatic secondary amines
(b)Aromatic teritiary amines
(c)Aromatic primary amines
(d)Quarternary ammonium compounds

14.The vanthoff factor i for a dilute solution of NaHSO4 IS (d)
(a)1/2 (b)1 (c)2 (d)3

15.Real gas exist at (b)
(a)high pressure and high temperature
(b)high pressure and lowtemperature
(c)low pressure and high temperature
(d)low pressure and low temperature

16.The electrolyte which is effective in coagulating gold sol is (c)
(a)Sodium chloride (b)Barium chloride
(c)Aluminium chloride (d)Ferric chloride

17. EDTA is a (d)
(a)bidentate ligand (b)tridentate ligand
(c)pentadentate ligand (d)hexadentate ligand

18.Chalcogens can exhibit oxidation states of (c)
(a)odd number (b)even number
(c)odd and even numbers (d)fractional numbers

19.The oxidation state of oxygen in O2F2 is (a)
(a)+1 (b)+2 (c)+4 (d)-2

20. An ester on hydrolysis gives a carboxylic acid which on kolbe's electrolysis gives ethane (d)
(a) Ethyl methonoate (b) Ethyl propanoate
(c) Ethyl ethanoate (d) Methyl methanoate

21. Identify the alkane which gives only one product on chlorination (b)
(a) n - Pentane (b) Neopentane
(c) 2-Methyl butane (d) n – Butane

22. Reaction of benzene with n-propyl chloride in presence of $AICI_3$ gives the following compound as the major product (b)
(a) n - Propyl benzene (b) Isopropyl benzene
(c) Chlorobenzen (d) Benzyl chloride

23. The carbon content present in Sesquiterpenes is (d)
(A) C_{10} (B) C_{20}
(C) C_{30} (D) C_{15}

24. The addition of H_2O to acetylene in presence of H_2SO_4+H_gSO_4 is an example for (A) Electrophilic addition (B) Nucleophilic addition

(C) Free radical addition (D) Electrophilic substitution

25. Aniline on reaction with bromine water gives (d)
(A) o-Bromoaniline (B) p-Bromoaniline
(C) 2,4-Dibromoaniline (D) 2,4,6- Tribromoaniline

26. The number of asymmetric carbons present in glucose is (d)
(A) 1 (B) 2 (C) 3 (D) 4

27. Which of the following statements is correct in case of diastereomers? (c)
(A) Identical physical and chemical properties
(B) Mirror image relationship
(C) No mirror image relationship
(D) Optically inactive

28. The compound that reacts fastest with Leucas reagent is (d)
(A) I-Butanol (B) 2- Butanol
(C) 3- Methyl-2-butanol (D) 2- Methyl-2- butanol

29. Isocyanide test is used for the identification of (d)
(A) Alcohols (B) Nitroalkanes
(C) Aldehydes (D) Amines

30. Sugar gets charred with con.H_2SO_4. It is because of (b)
(A) Dehydrogenation (B) Dehydration
(C) Hydration (D) Hydrolysis

31. An aqueous solution containg Ni^{2+} and Cd^{2+} has been acidified and thenH_2S gas passed into it. In this process. (b)
(A) Both Ni^{2+} and Cd^{2+} get precipitated
(B) Only Cd^{2+} gets precipitated
(C) Only Ni^{2+} gets precipitated
(D) Neither Ni^{2+} nor Cd^{2+} gets precipitated

32. The mean free path of the molecule of a gas is (b)
(A) Directly proportional to pressure
(B) Inversely proportional to pressure
(C) Independent of pressure
(D) Inversely proportional to square root of pressure

33. Which of the solution has the highest vapour pressure at a specified temperature? (b)
(A) 0.1 m glucose (B) 0.1m NaCl
(C) 0.1m $CaCl_2$ (D)0.1m $FeCl_3$

34. The process in which an excited molecule in its triplet state returns to its ground state in its single state by emitting the radiation is known as (b)
(A) Fluorescence (B) Phosphorescence
(C) Photosensitization (D) Chemiluminescence

35. Which of the following is a double salt? (d)

(A) $CuSO_4.4NH_3.H_2O$ (B) $Fe(CN)_2.4KCN$
(C) $CoCl_3.6NH_3$ (D) $KCl.MgCl_2.6H_2O$

36. $K_4[Fe(CN)_6]$ is named according to IUPAC system as (c)
(A) Potassium hexacyanoferrum (II) (B) Potassium ferrocyanide
(C) Potassium hexacyanoferrate (II) (D) Potassium hexacyanoferrate

37. Which among the following is not coloured? (a)
(A) TiO_2 (B) $VOSO_4$ (C) $TiCl_3$ (D) $CrCl_3$

38. The oxidation state of S inperoxomonosulphuric acid (H_2SO_5) is (c)
(A) +2 (B) +4 (C) +6 (D) +8

39. The shape of $Ni(C0)_4$ is (a)
(A) Tetrahedral (B) Trigonal
(C) Square planar (D) Trigonalbipyramidal

40. Cu^{2+} and Cd^{2+} ions can be separated from each other using (b)
(A) H S (B) KCN (C)HCl (D) None of these

41. The reaction of formaldehyde with organolithium compoundfollowed by hydrolysis gives (a)
(A) Primary alcohol (B) Secondary alcohol
(C) Tertiary alcohol (D) None of these

42. A diatomic molecule has eight electrons in bonding orbitals, four electrons in non-bonding orbitals and two electrons in antibonding orbitals. The bond order present in it is (b)
(A) 1 (B) 2 (C) 3 (D) 4

43. Detergent have common group (d)
a. RNA b. RONa c. RCOONa d. $RC_6H_4SO_3Na$

44. Ethyl alcohol is denatured by (d)
a. Methanol and Fomic acid
b.KCN
c. CH_3OH and C_6H_6
d. CH_3OH and Pyridine

45. Which is strongest acid? (d)
a. CCl_3COOH b. CH_3COOH
c. CBr_3COOH d. CF_3COOH

46. Product formed when benzaldehyde is refluxed with alcoholic KCN Solution (a)
a. Benzoin b. Phenyl cyanide
c. Phenyl isocyanate d. Phenyl isocyanide

47. The correct structure of H_2O_2 is (d)
a. H_2O_2 b. H-O-O-H
c. H – O d. H - O
I I

H – O O – H

48. German silver is made of (b)

a. Cu, Cr. Sn b. Cu, Zn, Ni

c. Cu, Ge, Ag d. Zn, Ge, Ag

49. K_2CS_3 can be called potassium? (c)

a. Sulfocyanide b. Thiocarbide

c. Thiocarnonate d. Thiocyanate

50. The metal that can be extracted from sea water (c)

a. Cr b. Ca c. Mg d. all

ANSWERS									
1	C	11	D	21	D	31	B	41	A
2	D	12	B	22	B	32	B	42	B
3	C	13	C	23	D	33	B	43	D
4	A	14	D	24	B	34	B	44	D
5	C	15	B	25	D	35	D	45	D
6	B	16	C	26	D	36	C	46	A
7	A	17	D	27	C	37	A	47	D
8	B	18	C	28	D	38	C	48	B
9	C	19	A	29	D	39	A	49	C
10	D	20	D	30	B	40	B	50	C

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GRAND TEST

SECTION – A (100M)

1. Which of the following is not a chiral compound? [b]
   (O) a. 2-bromopentane b. 3-chloropentane
   c. 2-chlorohexane d. 3-bromohexane

2. A tri-substituted benzene derivative will have [a]
   a. 2 isomers b. 4 isomers
   c. 3 isomers d. 5 isomer

3. The most stable carbocation is [c]
   a. CH_3^+ b. $CH_3CH_2^+$ c. $C_6H_5-CH_2^+$ d. $C_3H_2CH_2^+$

4. Select the antiaromatic species from the following: cyclopropenyl cation, cyclopentadiene, naphthalene, benzene, pyridine, cyclobutadiene, cyclopentadienyl cation [d]
   a. cyclopentadiene
   b. cyclopentadiene & pyridine
   c. cyclobutadiene & pyridine
   d. cyclopentadienyl cation & cyclobutadiene

5. Identify the chiral α – amino acid from the following: [b]
   a. glycine b. valine c. naproxen d. quinidine

6. Identify the most appropriate reaction that proceeds through carbine intermediate (from the following) [a]
   a. Reimer – Tiemann reaction b. Diels – Alder reaction
   c. Witting reaction d. Grignard reaction

7. How many stereoisomers are possible for butane-2,3-dicarboxylic acid? [b]
   a. 1 b. 2 c. 3 d. 4

8. Cannizarro reaction requires [b]
   a. formaldehyde and HCl b. benzaldehyde and NaOH
   c. benzaldehyde and aniline d. benzaldehyde and Na_2SO_3

9. Trimethylbenzene will have the following number of isomers. [b]
   a. Aldol condensation b. Cannizzaro reaction
   c. Claisen condensation d. Beckmann rearrangement

10. Michael reaction is [c]
   a. 1,2-addition of nucleophiles to ketones
   b. 1,2-addition of nucleophiles to olefins
   c. 1,4-addition of nucleophiles to α, β-unsaturated carbonyl compounds

d. 1,4-electrophilic addition to α, β-unsaturated carbonyl compounds

11 Indentify the most appropriate reagent for conversion of RCOOH to RCH_2OH from the following [a]

a. N_aBH_4 b. N_aH c. BH_3 THF d. H_2 [Pd] C

12. Identify the reaction that most appropriately proceeds through a carbocation intermediate [b]

a. Diels-Alder reaction b. Friedel-Crafts reaction

c. Witting reaction d. Hydroboration

13. Conversion of benzaldehyde to cinnamic acid is an example of [b]

a. Grignard reaction b. Knoevenagel condensation

c. Wolff rearrangement d. Mannich reaction

14. Identify the most appropriate reagent for conversion of cyclopentanone to cyclohexanone [c]

a. Methyl iodide b. methane sulfonic acid

c. diazomethane d. dimethyl sulphate

15. The structure of well-known pesticide DDT is [c]

a. 1,1,1-trichloro-2,2-bis(*p*-chlorophenyl) ethane

b. trimethylaluminium

c. 1,1-dichloro-2,2-difluoroethane

d. tetraethyl read

16. The intermediate in the anti-Markonikov addition of HBr to an alkene is a [d]

a. carbanion b. carbocation

c. carbene d. free radical

17. Identify the reaction in which two C-C bonds are formed in a single step [b]

a. Grignard reaction b. Diels-Alder reaction

c. Michanel reaction d. Friedel-Crafts reaction

18. Which of the following is a chiral molecule? [a]

a. 3-hexanol b. 1-hexonal

c. 1,6 hexanediol c. hex-3-en-l-ol

19. Identify the molecule that has the highest boiling point [d]

a. n-decane b. l-decane c. l-decyne d. l-decanol

20. Beckmann rearrangement is the conversion of [c]

a. oxime to nitrile b. amine to acid

c. nitro group to amine d. oxime to amide

21. The compound that exhibits geometrical isomerism is : [b]

a. 1-butene b. 2-butene c. 1-butanol d. 2-butanol

22. A terpene among the following is : [b]

a. D-xylose b. geraniol c. nicotine d. L-proline

23. At pH 7.0, amino acids exist as : [d]

a. negatively charges species b. positively charges species

c. uncharged species d. zwitterions

24. Molecular formula of an olefinic ketone containing a ring will be [c]
a. $C_nH_{2n}O$ b. $C_nH_{2n+2}O$ c. $C_nH_{2n-4}O$ d. $C_nH_{2n-6}O$

25. Which of the following molecules has the lowest pK_8? [c]
a. phenol b. *p*-nitro phenol
c. *p*-methyl phenol d. *p*-amino phenol

26. Which of the following statements is NOT true about S_N2 reactions: [b]
a. The transition state is linear
b. The configuration is inverted
c. The intermediate is planar
d. The reaction is second order

27. The maximum number of isomers that a molecule with 5 asymmetric centres can have is : [b]
a. 5 b. 32 c. 64 d. 16

28. Which of the following properties is characteristic of only chiral molecules : [c]
a. a sharp melting point b. strong UV-VIS absorption
c. non-zero optical rotation D. high dipole moment

29. 19. Identify the terpene from the following [b]
a. quinine b. menthol c. glucose d. cholesterol

30. The compound which would exhibit intramolecular hydrogen bonding is [a]
a. o-hydroxybenzaldehyde b. p-hydroxybenzaldehyde
c. m-hydroxybenzaldehyde d. o-hydroxytolune

31. R-OH have higher boiling points than R-Cl because of [b]
a. molecular shape b. hydrogen bonding
c. molecular size d. dipole moment

32. Which is the correct order of basicity? [c]
a. pyrrole > pyridine > Et_3N
b. Et_3N > pyrrole > pyridine
c. pyridine > Et_3N > pyrrole
d. Et_3N > pyridine > pyrrole

33. Which of the following forms intramolecular hydrogen bond? [d]
a. benzoic acid b. salicylic acid
c. phenol d. *p*-nitrobenzoic acid

34. α-D-glucose and β-D-glucose are best described as [a]
a. anomers b. epimers
c. enantiomers d. conformers

35. Which of the following molecules does not have 4n+2 π-electrons [c]
a. benzene b. phenanthrene
c. cyclobutdiene d. chlorobenzene

36. The most appropriate reagent for conversion of RCOOH to RCH_2OH is [d]
a. $NaBH_4$ b. Na/EtOH c. $LiBH_4$ d. $LiAlH_4$

37. Which of the isomers of pentane has the highest boiling point? [a]
a. $CH_3(CH_2)_3CH_3$ b. $(CH_3)_2CHCH_2CH_3$
c. $(CH_3)_2$ d. All isomers have the same b.p.

38. A mixture of benzoic acid and naphthalene can be separated most easily by [d]

a. sublimation b. distillation
c. fractional distillation d. crystallization

39. An ester RCOOR' is hydrolyzed. The products are RCOOH and R'OH. Which of the following is true? [b]

a. Both O-atoms in the acid were present in the ester
b. Only the carborryl O-atom of the ester is retained in the acid
c. Only the alkoxy O-atom of the ester is retained in the acid
d. O-atoms in the ester are scrambled in the reaction

40. When phenyl magnesium bromide is reacted with methanol, the product of the reaction is [c]

a. phenol b. anisole c. benzene d. bromobenzene

SECTION – B

41. The hydrogen halide with the maximum dipole moment is [a]

a. HF b. HCl c. HBr d. HI

42. Which property among the following indicates weak intermolecular forces? [b]

a. small molar volume b. low boiling point
c. large dipole moment d. high electrical conductivity

43. An octahedral complex MA_4B_2 [b]

a. will have two constitutional isomers
b. will have two stereoisomers
c. can not show isomerism
d. will be optically active

44. Which of the following is true about the S_N2 displacement reaction? [d]

a. always proceeds with racemisation
b. always proceeds with retention of configuration
c. sometimes inversion, sometimes retention
d. always proceeds with inversion of ion

45. The structure of X_4F_4 is [c]

a. square pyramid b. tetrahedral
c. square planar d. octahedral

46. Among the ions, Sc^{3+}, Ti^{3+}, Cr^{3+} and Fe^{3+}, which one is expected to give a colourless aqueous solution? [a]

a. Sc^{3+} b. Ti^{3+} c. Cr^{3+} d. Fe^{3+}

47. Which of the following molecules is linear? [a]

a. CO_2 b. NO_2 c. SO_2 d. OF_2

48. The hydrogen atom transitions involving absorption from n = 0 to all other n correspond to [a]

a. Lyman series b. Balmer series
c. Brackett series d. Paschen-Beck series

49. The oxidation state of oxygen in H_2O_2 is [c]
a. 0 b. +1 c. -1 d. -2

50. Water gas is an equimolar mixture of [c]
a. CO_2 and N_2 b. CO and N_2
c. CO and H_2 d. CO and H_2O

51. The IUPAC name for $[Fe(CN)_6]^{3-}$ is [a]
a. hexacyano ferrate (II) ion b. hexacyano ferrate (III) ion
c. hexacyano iron (II) ion d. iron (III) hexacyanide ion

52. The catalyst used in Haber's process for the manufacture of ammonia contains [d]
a. Pt b. Pd c. Ni d. Fe

53. Which of the following is a by-product of soap manufacture? [d]
a. Phenol b. glucose c. glycol d. glycerol

54. Identify the most appropriate reagent for the reduction of benzoic acid to benzyl alcohol [a]
a. N_aBH_4 b. $LiBH_4$ c. $LiAlH_4$ d. DiBAL-H

55. Identify the most polar molecule from the following [d]
a. dec-1-ene b. decanal
c. decan-2-one d. decan-1-ol

56. Among the following, which is the most appropriate description for the characteristic of an electron? [c]
a. particle b. wave
c. both particle and wave d. neither particle nor wave

57. An example of benzene mechanism is : [a]
a. $PhCl + NH_3/ NH_2^- \longrightarrow PhNH_2$
b. $PhH + HNO_3 + H_2SO4 \longrightarrow PhNO_2$
c. $PhNH_2 + HONO \longrightarrow PhN_2$
d. $PhI + CuCN \longrightarrow PhCN$

58. The splitting of nuclear energy levels in a magnetic field is known as [a]
a. Zeeman effect b. Stark effect
c. Mossbauer effect d. Cotton effect

59. Which of the following pairs of chemicals cannot be used to make hydrogen gas? [c]
a. HPO_3 b. H_3PO_3 c. H_3PO_4 d. $H_4P_2O_7$

60. The transition state in S_N2 displacement reaction has the geometry [d]
a. trigonal bipyramid b. tetrahedral
c. octahedral d. planar

61. Which of the following chemicals are responsible for depletion of ozone layer [b]
a. hydrocarbons b. chlorofluorocarbons

c. sulphur compounds d. lead compounds

62. Which of the following is a polyamide? [b]
a. polystyrene b. nylon c. teflon d. rubber

63. Which of the following molecules will have zero dipole moment?
(i). *o*-dichlorobenzene (ii). *m*-dichlorobenzene
(iii). *p*-dichlorobenzene (iv). 1,3,5-trichlorobenzene [d]
a. (i) and (ii) b. (ii) and (iii)
b. (iii) and (iv) d. (i) and (iv)

64. Addition of NH_4Cl and NH_4OH to a solution containing Fe^{3+} and Zn^{2+} gives a [c]
a. brown precipitate of $Fe(OH)_3$
b. white precipitate of $Zn(OH)_2$
c. green precipitate of $Fe(OH)_3$
d. mixture of $Fe(OH)_3$ and $Zn(OH)_2$

65. Complete the sentence : *Werner proposed his theory to explain* [a]
a. bonding in transition metal complexes
b. bonding in benzene
c. structure of silicates
d. optical activity of α-amino acids

66. Which one of the following is false? [a]
According to Fajan's rules, covalent bonding is favoured by
a. small positive ion b. large negative ion
c. large positive ion d. large charges on either ion

67. A white power, which dissolves readily in water, reacts with dil.HCl producing a colourless gas. The powder is likely to be [c]
a. sodium chloride b. calcium chloride
c. calcium carbonate d. sodium carbonate

68. H_2S is a weak acid. Addition of a few drops of HCl to an aqueous solution of H_2S will [b]
a. lower the concentration of S^{2-} ions
b. raise the concentration of S^{2-} ions
c. have no effect on the concentration of S^{2-} ions
d. increase the pH value

69. The color of anhydrous $FeSO_4$ is [b]
a. Green b. Blue c. Violet d. White

70. The compound having the lowest oxidation state of iron is [d]
a. HF b. HCl c. HBr d. HI

71. In qualitative analysis, Ag is detected in the first group, while pb is detected in both first and second groups. This is because, [d]
a. AgCl is much less soluble than $PbCl_2$
b. AgCl is much less soluble than $PbCl_2$
c. The solubilities of the chlorides are same, but traces of PbS are easily seen due to its black colour
d. AgS is soluble, but PbS is insoluble

72. Which metal atom is part of a vitamin molecule? [c]
a. iron b. bent c. cobalt d. zinc

73. In diborance, there are [c]
a. two 3c-4e BHB bonds b. four 3c-4e BHB bonds
c. two 3c-2e BHB bonds d. four 3c-2e BHB bonds

74. The number of unpaired electrons in gaseous species of Mn^{3+}, Cr^{3+} and V^{3+} respectively are: [c]
a. 4, 4 and 2
b. 3, 3 and 2
c. 4, 3 and 2
d. 3, 3 and 3

75. In the standard Daniel cell, the flow is from [a]
a. Zn electrode to the Cu electrode
b. Cu electrode to the Zn electrode
c. Zn electrode to the H_2 electrode
d. Pt electrode to the Cu electrode

76. Which one of the following compounds shows an electron spin resonance (ESR) spectrum? [d]
a. benzene b. zinc oxide
c. methyl blue d. copper sulphate

77. The number of degrees of freedom of water at its triple point is [c]
a. 3 b. 1 c. 0 d. 2

78. In a spontaneous process [a]
a. both entropy and free energy increase
b. both entropy and free energy decrease
c. entropy decreases and free energy increases

79. Work done in an isothermal reversible expansion of a perfect gas from a volume V_i volume V_f is [a]
a. greater than that in an adiabatic reversible expansion from the same initial to final state
b. equal to that in an adiabatic reversible expansion from the same initial to final state
c. less than that in an adiabatic reversible expansion from the same initial to final state
d. such data cannot be compared δ

80. The lock-and-key hypothesis was proposed by Emil Fischer to explain [a]
a. enzyme catalysis b. photosynthesis
c. respiration d. heredity

81. Entropy of activation of the Diels-Alder reaction will be [c]
a. positive b. negative c. zero d. positive or negative

82. The total entropy change in a Carnot cycle is [a]
a. zero b. positive c. negative d. infinite

83. When a non-volatile solute is added to a solvent what is the difference in vapor pressure expressed as a faction of original vapor pressure equal to? [c]
a. Mole fraction of solute in vapor phase
b. Mole fraction of solvent in vapor phase
c. Mole fraction of solute in liquid phase
d. Mole fraction of solvent in liquid phase

84. The relation dG = Vdp – SdT leads to which of the following Maxwell's equation? [d]

a. $\left(\frac{\delta T}{\delta V}\right)_s = -\left(\frac{\delta p}{\delta S}\right)_v$

b. $\left(\frac{\delta T}{\delta p}\right)_s = -\left(\frac{\delta p}{\delta S}\right)_p$

c. $\left(\frac{\delta p}{\delta T}\right)_v = -\left(\frac{\delta S}{\delta V}\right)_T$

d. $\left(\frac{\delta V}{\delta T}\right)_p = -\left(\frac{\delta S}{\delta p}\right)_T$

85. In an exothermic chemical reaction as the temperature rises, the equilibrium [a]

a. shifts towards the reactants b. shifts towards the products
c. remain unchanged d. exhibits an anomalous behaviour

86. An emulsion is [b]

a. liquid dispersed in gas b. liquid dispersed in liquid
c. solid dispersed in liquid d. solid dispersed in gas

87. The van der Waals equation for 'n' mole of a gas is given by: [c]

a. $(P + a/V^2)(V-b) = nRT$
b. $(P + na/V^2)(V-nb) = nRT$
c. $(P + n^2a/V^2)(V-nb) = nRT$
d. $(P + na/V^2)(V-nb) = nRT$

88. If a solid sphere is reshaped into a solid cube, the surface area [a]

a. increases b. decreases
c. does not change d. becomes half

89. Viscosity of liquid increases with rise tin temperature [c]

a. Decreases b. Constant
c. Increases d. Independent

90. A water drops is spherical in shape due to [a]

a. Surface tension b. poise
d. Viscosity d. Reflection

91. At critical temperature, the surface tension will be [d]

a. 1 b. 2 c. 3 d. 0

92. The unit of rate constant of Zero-order reaction is [b]

a. 5^{-5} b. mol L^{-1} c. L2 mol^{-2} 5^{-1} d. L mol^{-1} 5^{-1}

93. Chemical dissociation is [b]

a. exothermic b. reversible by endothermic
c. reversible d. endothermic

94. Spin multiplicity can be calculated by using [b]

a. 2S b. 2S+1 c. 2S-1 d. 2J+1

95. The properties of a real gas coincide with those of an ideal gas at [c]

a. $p \longrightarrow 0$ at any T b. $p \longrightarrow 0$ at any Boyle temperature T_β
c. T_β and at any pressure d. $T \longrightarrow \infty$ and $p \longrightarrow 0$

96. Chemiluminescence involves the emission of [c]
a. UV radiation b. Infra-red radiation
c. Visible radiation c. All the above

97. The main function of the salt bridge is [d]
a. to allow ions to go from one cell to another
b. to provide link between the half cells
c. to keep the emf of the cell positive
d. to maintain electric neutrality of the solution in two half cells

98. Transport number of CH_3COO^- ions is greater than that of $C_6H_5COO^-$ because CH_3COO^- ion [a]
a. has lower ionic mass b. is smaller in size
c. has CH_3 in place of C_6H_5 d. all the above

99. An acidic buffer solution can be prepared by mixing solutions by [b]
a. ammonium acetate and acetic acid
b. ammonium chloride and ammonium hydroxide
c. sulphuric acid and sodium sulphate
d. sodium chloride and sodium hydroxide

100. If the relative lowering of pressure of o-xylene is 0.005 due to addition of 0.5 grams of non-volatile solute in 500 grams of solvent, what is the molecular weight of the solute? [a]
a. 21.3 g/mole b. 23.1 g/mole c. 32.1 g/mole d. 1.23 g/mole

5. Group discussions

The institute encourage the advance learners to participate in group discussions which enhance their critical thinking, Understanding ability, Communication skills management skills through the use of roles within the groups and assessment skills involved in assessing options to make decisions about their group's final answers.

Students engaging in group discussion on TMV (Tobacco Mosaic virus) conducted by Department of Microbiology on 09-12-2021.

S G S ARTS COLLEGE, TIRUPATI

DEPARTMENT OF **ZOOLOGY**

2018-19

GROUP DISCUSSION CONDUCTED

ON

EVOLUTIONARY CHANGES IN HEART IN CHORDATES

Group Discussion		
Academic Year 2018-2019		
S No	**Names**	**Details**
1	Name of the Department	Zoology
2	Date of the Group Discussion conducted	Sep 29,2018
3	Name of Organizers	S N Shameer, B Ravi
4	Class of the Students	B.Sc., (CBZ & MZC)
6	Topic	Evolutionary changes in heart in chordates
7	No of Students attended	30
8	No of Faculty attended	2
9	Feedback	To know how the heart has evolved in the chordates.

HEAD
Department of Zoology
S.G.S. Arts College
TIRUPATI-517 501

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S G S ARTS COLLEGE, TIRUPATI

DEPARTMENT OF **ZOOLOGY**

GROUP DISCUSSIONS

Date: 14-9-18

CIRCULAR ISSUED TO THE STUDENTS

This is to inform that the department is intended to conduct group discussion on faculty preferred topic. In this regard, I hereby request all the B.Sc. CBZ/MZC Zoology students that those who are interested to participate in Discussion can enrol their names in the Zoology staff room.

HEAD OF THE DEPARTMENT,
ZOOLOGY.

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Department of Zoology
S.G.S. Arts College
TIRUPATI-517 502

GROUP DISCUSSION CONDUCTED ON

EVOLUTIONARY CHANGES IN HEART IN CHORDATES

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S.G.S ARTS COLLEGE ,T.T.D,TIRUPATI.

DEPARTMENT OF ZOOLOGY 2018-19

The students who have participate in the Group Discussion

S.NO	ROLL NO	NAME OF THE STUDENT	SIGNATURE
1	18701	A.BALAJI	Balaji
2	18702	C.BHANU VARA PRASAD	Bhanu vara prasad
3	18703	G.BHEEMUDU	Bheemudu
4	18704	T.CHAKRAPANI	chakrapani
5	187-5	G.CHANDR SEKHAR	chandr sekhar
6	18706	Y.DINESH BABU	Dinesh babu
7	18707	B.GANESH	Ganesh
8	187-8	R.GAYATHRI	Gayathri
9	18709	K.GOPAL	Gopal
10	18710	T.GOWTHAM	Gowtham
11	18711	B.GOWTHAM GOWDA	Gowtham gowda
12	18712	M.GURRAMMA	Gurramma
13	18713	K.KALYANI	kalyani
14	18714	K.LIKITHA	Likitha
15	18715	T.MOUNIKA	mounika

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TIRUPATI-517 501

16	19673	M.NAVEEN NAIK	Naveen Naik
17	19674	G.RAJASEKHAR	Raja sekhar
18	19675	A.M.SAMEERA	sameera
19	19676	P.SANDEEP KUMAR	Sandeep kumar.
20	19677	P.SANTOSH KUMAR	Santosh kumar.
21	19678	K.SASI KUMAR	Sasi kumar.
22	19679	E.SREEKANTH	Sreekanth.
23	19680	D.SURYA	Surya.
24	19681	K.SWATHI	Swathi
25	19682	K.USHA	usha.
26	19683	G.VEERANJINEYULU	Veeranjineyulu.

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35	19773	C.RAMA KRISHNA	Rama krishna
36	19774	A.RAMESH	Ramesh
37	19775	P.REDDY PRAKASH	Reddy prakash
38	19776	G.REDDI SEKHAR NAIDU	Reddishakar naidu
39	19777	REENA SHAIK	Reena shaik
40	19778	V.SAI CHAITANYA	sai chaitanya
41	19779	G.SIVA KUMAR	Siva kumar
42	19780	K.SOHITH KUMAR	sohith kumar
43	19781	C.SURESH BABU	Suresh babu
44	19782	N.USMAN	Usman
45	19783	P.VAISHNAVI	Vaishnavi
46	19784	C.VENKATA RAMANA	Venkata ramana
47	19785	K.VINOD	Vinod

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S.G.S. Arts Colleg
TIRUPATI-517 501

Department of Computer Science

Acamedic Year : 2019 -2020

Activity : Poster Making Competition organized by Department of Computer Science

Date:14-01-2019

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S. ARTSCOLLEGE, TIRUPATI

(FOUNDED IN 1969)

DEPARTMENT OF COMPUTER SCIENCE

CIRCULAR

Date: 11/ 01 /2019

All the interested first year BSc students with **Computer Science Combination groups** are hereby informed to participate Poster Making Competition on Topic **"Internet Safety/Cyber Safety"** , which is to be organized in Room no 216 on 14/01/2019 by 11:00AM.

HOD

DEPARTMENT OF COMPUTER ...
S.G.S. ARTS COLLEGE
TIRUPATI

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

Sri Govindaraja Swamy Arts College,TTD,Tirupati

Department of Computer Science

POSTER MAKING COMPETITION

Date:14-01-2019

OBJECTIVES:

1. To bring out the hidden talents in Poster designing .
2. To explore creative expression of the students and also to gauge their awareness quotient.
3. To improve overall performance of students by using different pedagogical techniques.
4. Ability to write simple tips for staying safe online in the form of poster thereby creating awareness to all on specified topic .

CONVENOR

- Sri P.Udaya, HOD, Department of Computer Science

ORGANIZING MEMBERS

- Smt B.Triveni
- Smt N.Jayalakshmi
- Smt D.Chaitanya Kumar

TOPIC: Internet Safety/Cyber Safety

ABSTRACT:

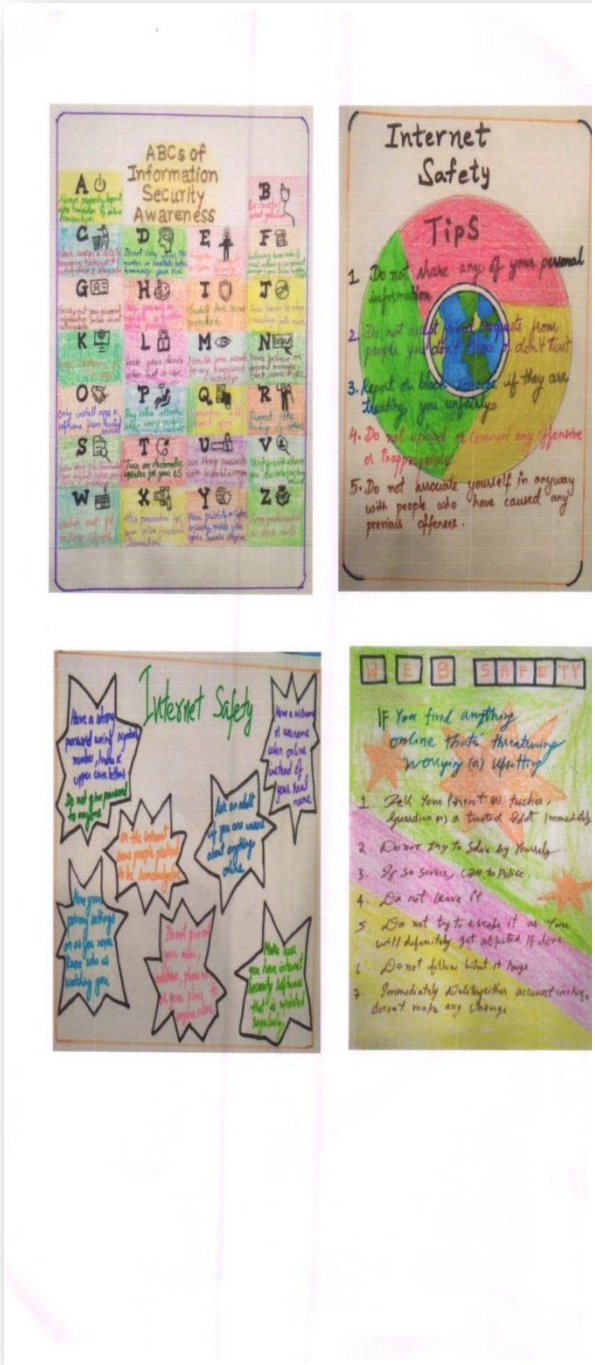
The department of computer science thought to organize Poster Making Competition programme for I BSc. MSCs-EM and MPCs-EM on the Topic **"Internet Safety/Cyber Safety"** to create awareness of personal safety and security risks to private information and property associated with using the internet, and the self-protection from computer crime.

PHOTO GALLERY


Tirupati.


POSTER
MAKING
S.G.S. ARTS COLLEGE
Tirupati.
Department of
Computer Science

Posters Gallery



Sri Govindaraja Swamy Arts College,TTD,Tirupati

Department of Computer Science

POSTER MAKING COMPETITION

ATTENDANCE SHEET

Sno	Roll No	Group & Year	Name of the Student	Signature
1	18520	MPCs I[st] year	B. Hashavardhan Reddy	B. Hashavardhan Reddy
2	18565	MSCs I[st] year	B. Ajay kumar	B. Ajay kumar
3	18568	MSCs I[st] year	T. Anitha	T. Anitha
4	18513	MPCs I[st] year	E. Balaji	E. Balaji
5	18569	MSCs I[st] year	B. Aravind Naidu	B. Aravind Naidu
6	18575	MSCs I[st] year	V. chanidini	V. chanidini
7	18515	MPCs I[st] year	A. Balarama Krishna	A. Balarama Krishna
8	18514	MPCs I[st] year	M. Balanjineyulu	M Balan
9	18576	MSCs I[st] year	M. charan Teja	M. charan Teja
10	18581	MSCs I[st] year	V. Govardhan Krishna	V. Govardhan Krishna
11	18577	MSCs I[st] year	K. Deepa	K. Deepa
12	18516	MPCs I[st] year	M. charan kumar	M. Charan Kumar
13	18519	MPCs I[st] year	P. venna guru sai Kumar	P. Venna Guru
14	18521	MPCs I[st] year	P. Hemanth kumar	P. Hemanth Kumar
15	18584	MSCs I[st] year	P. Haritha	P. Haritha
16	18586	MSCs I[st] year	L. Indhumathi	L. Indhumathi
17	18588	MSCs I[st] year	K. Kishore	K. Kishore
18	18589	MSCs I[st] year	M. Lakshmi	M. Lakshmi
19	18523	MPCs I[st] year	N. Jagan Goud	N. Jagan Goud
20	18524	MPCs I[st] year	K. Jayalakshmi	K. Jayalakshmi
21	18595	MSCs I[st] year	A. Penchala Narasimhulu	A. Penchala Narasimhulu
22	18596	MSCs I[st] year	G. Ramesh	G. Ramesh
23	18598	MSCs I[st] year	K. M. Rithik	K. M. Rithik
24	18602	MSCs I[st] year	C. Sreedevi	C. Sreedevi
25	18525	MPCs I[st] year	G. Habbu Sharif	G. Habbu Sharif

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Sno	Roll No	Group & Year	Name of the Student	Signature
26	18605	MSCs I[st] year	S. Thimmaiah	S. Thimmaiah
27	18606	MSCs I[st] year	K. Vamsi Krishna	K. Vamsi Krishna
28	18607	MSCs I[st] year	M. Venkata Amarnath	M. Venkata Amarnath
29	18529	MPCs 1[st] year	G. MOUNIKA	G. Mounika
30	18533	MPCs 1[st] year	K. Naveen kumar	Naveen Kumar
31	18541	MPCs 1[st] year	Saiteja	Saiteja
32	18546	MPCs I[st] year	Shaik Yasmin	Shaik Yasmin
33	18563	MPCs I[st]	G. Surendra	G. Surendra
34	18542	MPCs I[st] year	D. Sandhya	D Sandhya
35	18613	MSCs 1st	E. VIJAY GOUD	E.VIJAY GOUD
36	18553	MPCs 1st Year	Mangali Swasth	Swasth
37	18615	MSCs 1st	P. Geethanjali	P.Geetha
38	18554	MPCs 1st	V. SWAPNA	Swapna
39	18557	MPCs 1st	B Vamsi	Vamsi
40	18558	MPCs 1st	K Dhanakumar	K. Dhanakumar
41	18558	MPCs 1st	V Vanishini	Vanishini
42	18559	MPCs 1st	K. Venkata Jagadeesh Kumar	K. Venkata Jagadeesh
43	18547	MPCs I[st] year	G. Shireesha	G. Shireesha
44	18560	MPCs I[st] year	M. Venkatasubramanyam	M. Venkatasubramanyam
45	18561	MPCs 1st	M. VIGAY KUMAR	Vijay
46	18562	MPCs I[st] year	A. Venkata chakravarthi	A. Venkata Chakravarthi

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DEPARTMENT OF COMPUTER SCIENCE
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TIRUPATI

Prize Winners:

1) K. Jayalakshmi - I[st] Prize
2) B. Hashavardhan Reddy - 2[nd] Prize
3) G. Mounika - 3[rd] Prize

6. Expert Talks /Guest lecturers/Seminars.

Dept. of Physics organized a seminar on date 21.11.2019 on Analytical Techniques for material characterization, delivered by Prof. P.Sreedhar Reddy, Registrar, S.V.University, Tirupati.

Dept. of Physics organized a seminar on date 15.06.2016 on Nano Science and Nano Technology, delivered by Prof. OMD Hussain, Dept. of Physics, S.V.University, Tirupati.

Dept. of Physics organized a seminar on date 13.10.2016 on Atmospheric observations on Troposphere using Radar system, delivered by Prof. D. Narayana Rao, Former Director, GST Radar Centre, Gadanki.

7. STUDENTS' OFFICE BEARERS (STUDENTS' COUNCIL)

The student office bearers work under the guidance of the college. The main function of the students office bearers is to organize various activities to provide opportunities to the students of the college to exhibit their talents.

Objectives:

True education aims at holistic and integrated development of an individuals personality including intellectual , physical, emotional, moral and ethical aspects in search of truth.

Since the stude3nts are the central force of any educational institution, their participation in its activities undoubtedly essential within the overall perspective stated above.Student office bearers for various association are nominated on merit basis as per S.V.University guidelines,

Chairman, Secretary and Joint secretary are selected on rotation basis from B.A., B.Com and B.Sc courses . Secretary and Joint secretary for Maths and Science, Arts, Commerce, Language Biological Sciences, Fine arts and Sports are nominated on merit basis.

Institution facilitates students' representation and engagement in various administrative, co-curricular and extracurricular activities following duly established processes and norms (student council, students representation on various bodies)

7. STUDENT UNION OFFICE BEARERS: 2020-2021

2020 - 2021 – STUDENTS UNION

Union Chairman : **M. Gurramma,** III B.Sc., M.Z.C., (E.M.)

Union Secretary: **A. Anusha** III B.A., M.E.S., (E.M.)

MATHS & SCIENCE ASSOCIATION

Secretary : **B. Harshavardhan Reddy,** III B.Sc., M.P.Cs. (E.M.)

Joint Secretary : **P. Sakeel,** II B.Sc.,M.P.Cs., (E.M.)

LIFE SCIENCE ASSOCIATION

Secretary : **R. Gayatri,** III B.Sc., M.Z.C. (E.M.)

Joint Secretary : **A. Chandini,** II B.Sc., B.B.C. (E.M.)

COMMERCE ASSOCIATION

Secretary : **D. Srikanth,** II B.Com., CA (E.M.)

Joint Secretary : **V. Yugandhar,** II B.Com., CA (E.M.)

ARTS ASSOCIATION

Secretary : **P.Ganesh Achary,** III B.A.,A.S.C.A., (E.M.)

Joint Secretary : **V. Pavan Kumar,** II B.A., H.P.T. (T.M.)

LANGUAGE ASSOCIATION

Secretary : **E. Vijay Goud,** III B.Sc., M.S.Cs. (E.M.)

Joint Secretary : **D. Divyasri,** II B.Com., CA (E.M.)

FINE ARTS ASSOCIATION

Secretary : **M. Prabhu Kumar,** III B.A., H.E.P. (E.M.)

Joint Secretary : **A. Karunya,** II B.A., H.E.P. (E.M.)

SPORTS ASSOCIATION

Sports & Games_Secretary: **P. Veera Guru Saikumar,** III B.Sc., M.P.Cs (E.M.)

* * *

2019-2020

Sno	Class	Name of the Representative	Roll No
1	III Bcom-CA-E.M	Kum. K.Deepa(Union Chairman)	17167
2	III B.Sc-MSCs-EM	Mr. A.Venkatesh(Union Secretary)	17582
3	III B.Sc-MPCs-EM	Mr. B.Mohana Krishna(Maths & Science Association Secretary)	17508
4	II B.Sc-MSCs-EM	Kum. K.Sireesha(Maths & Science Association Joint Secretary)	18599
5	III BSC-BBC-E.M	Kum. Jahnavi Takkolu (Life Sciences Association Secretary)	17672
6	II BSc-MZC-EM	Kum. M.Guramma(Life Sciences Association Joint Secretary)	18712
7	III BCom-CA-E.M	Mr. V.Sravan Kumar(Commerce Association Secretary)	17208
8	II BCom-CA-EM	Mr. D.Sreekanth(Commerce Association Joint Secretary)	18178
9	III BA-HEP-TM	Mr. K.Rajendra Prasad(Arts Association Secretary)	17307
10	II BA-MES-EM	Kum. A.Anusha(Arts Association Joint Secretary)	18373
11	III BSc-MPCs-EM	Kum. K.JyotsnaLanguage Association Secretary)	17501
12	II BSc-MPCs-EM	Kum. Shaik Yashmin(Language Association Joint Secretary)	18546
13	III BA-MES-EM	Mr. Y.Anil(Fine Arts Association Secretary)	17391
14	II BCom-GEN-EM	Kum. P.Yamini(Fine Arts Association Joint Secretary)	18101
15	I BA-HEP-EM	Kum. K.Karunya(Fine Arts Association Joint Secretary)	HT: 0120008001
16	III BCom-GEN-E.M	Mr. Y.Prudviraj(Sports&Games Association Secretary)	17112
17	III B.Sc-MZC-EM	Mr. D.Sandeep(Sports&Games Association Secretary)	17741

BSc Class Representatives			
Sno	**Class**	**Name of the Representative**	**Roll No**
1	III B.Sc-MSCs-EM	Mr. A.Venkatesu	17582
2	III B.Sc-MPCs-EM	Mr. M.Prasanth	17515
3	III B.Sc-MPC-EM	Mr. P.Pavan Kalyan	17473
4	III B.Sc-MPC-TM	Mr. Y. Hemanth Kumar	17428
5	III B.Sc-MPS-TM	Mr. O.Jayakumar	17598
6	III B.Sc-BBC-EM	Mr. B.Mahesh Babu	17681
7	III B.Sc-MZC-EM	Mr. P.Devanadh Goud	17719
8	III B.Sc-CT&HM-EM	Kum.Pavani Krishna Benguluru	17782
9	III B.Sc-CBZ-TM	Mr. V.Venkatesh	17644
10	II B.Sc-MSCs-EM	Mr. Bharath kumar	18571
11	II B.Sc-MPCs-EM	Mr. B.Harshavardhan Reddy	18520
12	II B.Sc-MPC-EM	Mr. Obul Reddy	18453
13	II B.Sc-MPC-TM	Mr. P.Chandra Sekhar	18405
14	II B.Sc-MPS-TM	Mr. O.Jayakumar	HT: 0318008134
15	II B.Sc-BBC-EM	Mr. M.Mareepa	18670
16	II B.Sc-MZC-EM	Mr. V.Pavan Kumar Goud	18718
17	II B.Sc-CT&HM-EM	Mr. Chandu Chainuru	18756
18	II B.Sc-CBZ-TM	Mr. A.Ashok	18616
19	I B.Sc-MSCs-EM	Mr. G. Dasari Gopal	HT: 0320008161
20	I B.Sc-MPCs-EM	Mr. K.Adimuni Peddaiah	HT: 0320008099
21	I B.Sc-MPC-EM	Mr. A.M.Raghuram Kumar	HT: 0320008026

22	I B.Sc-MPC-TM	Mr. A.M. RAGHURAM KUMAR	HT: 0320008040
23	I B.Sc-MPS-TM	Mr. G.Reddi Prasad	HT: 0320008057
24	IB.Sc-BBC-EM	Mr. M.Vivek Amruth Narayana	HT: 0320008282
25	I B.Sc-MZC-EM	Mr. A..Karthik	HT: 0320008214
26	I B.Sc-CT&HM-EM	Mr. V. Durga Nagendra	HT: 0320008346
27	I B.Sc-CBZ-TM	Mr. G. Rajasekhar	HT: 0320008012
BCom Class Representatives			
28	III B.Com-CA-EM	Mr. M.Anji Naik	17162
29	III B.Com-GEN-EM	Mr. C.Ramesh	17041
30	III B.Com-GEN-TM	Mr. Murali Chinthala	17032
31	II B.Com-CA-EM	Mr. M. Venkata Suresh	18187
32	II B.Com-GEN-EM	Miss . P.Yamini	18101
33	II B.Com-GEN-TM	Mr. Y.Sai Kumar	18042
34	I B.Com-CA-EM	Mr. G.Tarun	HT: 0220008143
35	IB.Com-GEN-EM	Mr. E.Mohan	HT: 0220008031
36	I B.Com-GEN-TM	Mr. M.Ravi	HT:0220008070
BA Class Representatives			
37	III B.A-HEP-EM	Mr. Hemanth Vuranduru	17345
38	III B.A-HEP-TM	Mr. Venkatanarayana Reddy Devireddy	17315
39	III B.A-HPT-TM	Mr. T.Eswaraiah	17367
40	III B.A-ASCA-EM	Mr. Vamsi Gundlapalli	17253
41	III B.A-MES-EM	Pradeep Madaka	17399
42	II B.A-HEP-EM	Mr. Y.Surya Mahesh	18336

43	II B.A-HEP-TM	Mr. A.Kiran Kumar	18273
44	II B.A-HPT-TM	Mr. M.Eswar Naik	18346
45	II B.A-ASCA-EM	Mr. A.Adisankar	18201
46	II B.A-MES-EM	Dilhi Babu.G	18374
47	I B.A-HEP-EM	Mr. A.Hari Prasad Reddy	0120008002
48	I B.A-HEP-TM	Mr. N.Siddeswaraiah	0120008041
49	I B.A-HPT-TM	Mr. G.Nagaraju	0120008069
50	I B.A-ASCA-EM	Mr. J.Vinay Kumar	0120008086

HOSTEL BEARERS: 2019-2020

Sno	Committee Members	Designation	Role
1	Dr. C.Ramesh	Principal	Chief Warden
2	Dr.Y.Mallikarjun Rao	HOD,Commerce	Warden
3	Sri K. Rajesh	Lecturer,Economics	Deputy Warden
4	Mr. Prasanth	Student,III BSc-MPCs-EM	Mess Secretary
5	H.Mahadeva	Student,III BA-HEP	Store Secretary

MESS COMMITTEE

Sno	Name	Group	Role
1	P.Pavan Kalyan	III BSc-MPC-EM	Associate Members
2	D.Naveen	III Bcom-General	
3	G.Vamshi	III BA-ASCA	
4	G.Varaprasad	III BSc-CT&HM	
5	B.Harsha Vardhan	II BSc-MPCs	
6	B.Srikanth	II Bcom –General	
7	Ganesh Achari	II BA-ASCA	

S.G.S. ARTS COLLEGE ★ TIRUPATI ★ Co-ordinator NAAC

PRINCIPAL S.G.S. ARTS COLLEGE TIRUPATI

2. Bridge Course

Abstract of the Bridge Course conducted is shown here complete details are available in college Website

Phone: (0877) 2264599 E-mail:sgsartscollegettd@gmail.com Web: sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Academic year 2016-17

Summary of Bridge courses

The Institution has conducted Bridge courses with an objective to Bridge the gap between subjects studied at their higher secondary level and the subjects they would be studying in their Graduation level. The syllabus for the course is framed in such way that they get basic knowledge and the subjects they would be learning through Graduation.

S.NO	Department	Date of conduction	Total Number of hours taught	Students to whom Bridge course was given	Number of students benefited
1	Chemistry	18-07-2016 to 23-07-2016	30	Students who have no chemistry back ground in their 10+2 level(Vocational study)	3
2	Mathematics	18-07-2016 to 23-07-2016	30	Students with biological background in their 10+2 level	96
3	Statistics	11-07-20-16 to 16-07-2016	30	Students with biological background in their 10+2 level	96

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

S.G.S. ARTS COLLEGE Co-ordinator NAAC TIRUPATI

Tirupathi
06.03.2021

From
Sri. D. Chandrakesavulu Naidu,
HOD of Statistics,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.

Sir,

Sub:- Permission for conducting Statistics Bridge course to students of first year Bi.P.C Degree – request – Regarding.

&&&&&

Kindly permit the Department of Statistics to conduct Bridge Course to the students of first year, who have not studied statistics in intermediate level, for one week i.e. from 15-03-2021 to 16- 03-2021. Time table and syllabus are enclosed.

Thanking you sir

Yours faithfully
D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

2020-2021
SGS Arts College,Tirupati
Department of Statistics

Circular

First-year degree students who studied Bi.P.C Course at intermediate level are informed to attend the Statistics bridge course which will be held at Room No. 222, Department of Statistcs, Tirupati from 15-03-2021 to 16-03-2021. All Biology Student should attend the with out fail

D. CHANDRA KESAVULU NAIDU, M.Phil., Ph.D
LECTURER & HEAD
Dept. of Statistics
S.G.S. ARTS COLLEGE
TIRUPATI

Head,
Department of Statistics

2020-2021
Bridge Course-Summary

Name of the programme	Bridge course in Statistics
Duration of Course	One Week
Date	15-03-2021 to 16-03-2021
Venue	Room No. 222

S.No	Subject name	Faculty Name	No. of Students
1	Meaning Scope and Limitations of Statistics	Dr. K Vijaya Kumar	116
2	*Measures of Central Tendency*	*Dr. K Venkatesh*	*116*
3	Measures of Dispersion	Dr. M Vasudeva Reddy	116

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TIRUPATI

S.G.S. ARTS COLLEGE
TIRUPATI

2020-2021
Department of Statistics
Bridge Course Timetable

Day	10AM-11:30 AM	11:30AM-1PM	2PM-3PM	3PM-4PM
Wednesday	Meaning Scope and Limitations of Statistics	Measures of Dispersion	Measures of Central Tendency	Q&A Session
Thursday	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	Measures of Dispersion	
Friday	Measures of Dispersion	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	
Saturday	Meaning Scope and Limitations of Statistics	Measures of Dispersion	Measures of Central Tendency	
Monday	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	Measures of Dispersion	
Tuesday	Measures of Dispersion	Measures of Central Tendency	Meaning Scope and Limitations of Statistics	

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TIRUPATI

2020-2021

Bridge Course in Statistics

Syllabus

S.No	Subject	Syllabus
1	**Meaning Scope and Limitations of Statistics**	Primary and Secondary, Classification and Tabulation, Construction of frequency distribution. Graphical Representation : Histogram, Bar, Pie and Frequency polygon.
2	**Measures of Central Tendency**	Features of good average, Arithmetic Mean, Median, Mode. Empirical relationship between Mean,central Median and Mode and skewness based on values.
3	**Measures of Dispersion**	Range, Quartile Deviation (QD), Mean Deviation (MD), Variance, Standard Deviation (SD), relationship between QD, MD and SD. Familiarization of the concepts relating to Correlation and Linear Regression

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TIRUPATI

SGS ARTS COLLEGE::TTD::TPT
DEPARTMENT OF STATISTICS
BBC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Marks
1	20919	Chakradhar Goud E	22
2	20920	Chan Basha P	19
3	20921	Dheeraj M	21
4	20922	Divya Bharathi Reddy G	19
5	20923	Ganesh B	23
6	20924	Harshavardhan K G	25
7	20925	Hemalatha M	ABA
8	20926	Hemanth K	23
9	20927	Indira Mythri D	17
10	20928	Janaki Raman R S	20
11	20929	Kalyan M	22
12	20930	Kiran N	21
13	20931	Krishna D	19
14	20932	Kushwanth T	23
15	20933	Lakhsmanna T	19
16	20934	Leena P	25
17	20935	Mamatha A	25
18	20936	Manjusha K	24
19	20937	Mohana A	23
20	20938	Naga Roshini N	19
21	20939	Nataraj D	17
22	20940	Naveen Kumar K	21
23	20941	Pavan Kumar Raju O	22
24	20942	Pavan Kumar K	18
25	20943	Phanendra Kumar B	23
26	20944	Praveen S	24
27	20945	Ravi Kumar T	19
28	20946	Sai Praathap K	21
29	20947	Sai Sree Harsha P V	17
30	20948	Sai Sreenivas K	19
31	20949	Subba Reddy B	22
32	20950	Surendra Babu S	24
33	20951	Tejasree R	25
34	20952	Usha Sree P	25
35	20953	Vara Prasad M	19
36	20954	Vidhyavathi B	20
37	20955	Vignesh Yadav T	17
38	20956	Vijaya Brahma C	18
39	20957	Yamini B	23

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TIRUPATI

SGS ARTS COLLEGE::TTD::TPT
DEPARTMENT OF STATISTICS
BBC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	15.03.2021	16.03.2021
1	20919	Chakradhar Goud.E	P	P
2	20920	Chan Basha.P	P	P
3	20921	Dheeraj.M	P	P
4	20922	Divya Bharathi Reddy.G	P	P
5	20923	Ganesh.B	P	P
6	20924	Harshavardhan.K.G	P	P
7	20925	Hemalatha.M	A	A
8	20926	Hemanth.K	P	P
9	20927	Indira Mythri.D		P
10	20928	Janaki Raman.RS	P	P
11	20929	Kalyan.M	A	P
12	20930	Kiran.N	P	P
13	20931	Krishna.D	P	A
14	20932	Kushwanth.T	P	P
15	20933	Lakhsmanna.T	P	P
16	20934	Leena.P	P	P
17	20935	Mamatha.A	P	P
18	20936	Manjusha.K	P	P
19	20937	Mohana.A	P	P
20	20938	Naga Roshini.N	P	P
21	20939	Nataraj.D	P	P
22	20940	Naveen Kumar.K	A	P
23	20941	Pavan Kumar Raju.O	P	P
24	20942	Pavan Kumar.K	P	P
25	20943	Phanendra Kumar.B	P	P
26	20944	Praveen.S	P	P
27	20945	Ravi Kumar.T	P	P
28	20946	Sai Praathap.K	P	P
29	20947	Sai Sree Harsha.P.V	P	P
30	20948	Sai Sreenivas.K	P	P
31	20949	Subba Reddy.B	P	P
32	20950	Surendra Babu.S	P	P
33	20951	Tejasree.R	P	P
34	20952	Usha Sree.P	P	P
35	20953	Vara Prasad.M	P	P
36	20954	Vidhyavathi.B	P	P
37	20955	Vignesh Yadav.T	P	P
38	20956	Vijaya Brahma.C	P	P
39	20957	Yamini.B	P	P

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TIRUPATI

SGS ARTS COLLEGE::TTD::TPT
DEPARTMENT OF STATISTICS

BBC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Signature
1	20919	Chakradhar Goud.E	E. chakradhar goud
2	20920	Chan Basha.P	Chand Basha
3	20921	Dheeraj.M	M.Dheeraj
4	20922	Divya Bharathi Reddy.G	Divya Bharathi
5	20923	Ganesh.B	
6	20924	Harshavardhan.K.G	K. Harshavardhan
7	20925	Heenalatha.M - TC	
8	20926	Hemanth.K	K. Hemanth Kumar
9	20927	Indira Mythri.D	D. Indira Mythri
10	20928	Janaki Raman.RS	R.S. Janaki
11	20929	Kalyan.M	M. kalyan
12	20930	Kiran.N	N. Kiran
13	20931	Krishna.D	D. Krishna
14	20932	Kushwanth.T	T. Kushwanth
15	20933	Lakhsmanna.T	T. Lakshmanna
16	20934	Leena.P	P. Leena
17	20935	Mamatha.A	A. Mamatha
18	20936	Manjusha.K	Manjusha
19	20937	Mohana.A	A. Mohana
20	20938	Naga Roshini.N	N. N. Roshini
21	20939	Nataraj.D	D. Nataraj
22	20940	Naveen Kumar.K	K. Naveen Kumar
23	20941	Pavan Kumar Raju.O	O. PK Raju
24	20942	Pavan Kumar.K	K. Pavan Kumar
25	20943	Phanendra Kumar.B	B. Phanendra Kumar
26	20944	Praveen.S	S. Praveen
27	20945	Ravi Kumar.T	T. Ravi Kumar
28	20946	Sai Praathap.K	sai prathap
29	20947	Sai Sree Harsha.P.V	P. V. Sree Harsha
30	20948	Sai Sreenivas.K	
31	20949	Subba Reddy.B	B. Subba Reddy
32	20950	Surendra Babu.S	S. Surendra Babu
33	20951	Tejasree.R	R. Tejasree
34	20952	Usha Sree.P	P. Usha Sree
35	20953	Vara Prasad.M	M. Vara Prasad
36	20954	Vidhyavathi.B	B. Vidhyavathi
37	20955	Vignesh Yadav.T	T. Vignesh Yadav
38	20956	Vijaya Brahma.C	
39	20957	Yamini.B	B. Yamini

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SGS ARTS COLLEGE::TTD::TPT
DEPARTMENT OF STATISTICS
MZC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Marks
1	20969	ARAVIND .K	23
2	20970	ASHOK DUPPANAGURTHI	24
3	20971	AVINASH SEDAM	19
4	20972	AYESHA SHAIK	21
5	20973	BALAKULLEYAPPA AAKULEDU	22
6	20974	CHAITANYA BALINENI	19
7	20975	DEEPAK ROYAL BANDLA	17
8	20976	Dora Babu. Kaputhota	16
9	20977	DHANUSH CHINTHALA	22
10	20978	DHEERAJ KUMAR SODEM	24
11	20979	GOWTHAM NAIK MUDE	25
12	20980	GURU RAJU KURUVA	24
13	20981	HANISH BABU BODANAPU	25
14	20982	HARSHITA DAKSHINADI	25
15	20983	HEMANTH KARNAM	17
16	20984	JEEVAN KUMAR KEELAPATLA	19
17	20985	KISHORE NAIK MUDE	18
18	20986	LOKANADHAM GUTHI	20
19	20987	KUSHWANTH KUMAR.M	21
20	20988	LAKSHMIKANTH NAIK.B	24
21	20989	LOKESH ANKE	22
22	20990	LOKESH PUTTURU	19
23	20991	MADHU BANDILA	17
24	20992	MANESH KUMAR JAGGALI	22
25	20993	MANOJ KUMAR REDDY .V	19
26	20994	PAVAN KUMAR JANGAM	19
27	20995	PRBHAKARA CHINTAKAYALA	20
28	20996	RAMAKRISHNA NAIK BUKKE	21
29	20997	RISHITHA POKALA	22
30	20998	SAI PRAKASH VEMULA	24
31	20999	SANDHYA YERRABALLI	21
32	21000	SATHEESH GOLLAPALLI	23
33	21001	Seetharamanjaneyulu.K	19
34	21002	SHARATH KUMAR.Y	17
35	21003	SHIK TASNEEM MALLELA	16
36	21004	SIVANI BOMMI	19
37	21005	THULASI NALLIAHGARI	21
38	21006	VENKATESH AADHULA	24
39	21007	VENKATESH.T	25

D. CHANDRA RESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

SGS ARTS COLLEGE::TTD::TPT
DEPARTMENT OF STATISTICS
MZC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	15.03.2021	16.03.2021
1	20969	ARAVIND .K	1	1
2	20970	ASHOK DUPPANAGURTHI	1	1
3	20971	AVINASH SEDAM	1	1
4	20972	AYESHA SHAIK	1	1
5	20973	BALAKULLEYAPPA AAKULEDU	1	1
6	20974	CHAITANYA BALINENI	1	1
7	20975	DEEPAK ROYAL BANDLA	1	1
8	20976	Dora Babu. Kaputhota	A	1
9	20977	DHANUSH CHINTHALA	1	1
10	20978	DHEERAJ KUMAR SODEM	1	1
11	20979	GOWTHAM NAIK MUDE	1	1
12	20980	GURU RAJU KURUVA	1	1
13	20981	HANISH BABU BODANAPU	1	1
14	20982	HARSHITA DAKSHINADI	1	1
15	20983	HEMANTH KARNAM	1	1
16	20984	JEEVAN KUMAR KEELAPATLA	1	1
17	20985	KISHORE NAIK MUDE	1	1
18	20986	LOKANADHAM GUTHI	1	1
19	20987	KUSHWANTH KUMAR.M	1	1
20	20988	LAKSHMIKANTH NAIK.B	1	1
21	20989	LOKESH ANKE	1	1
22	20990	LOKESH PUTTURU	1	1
23	20991	MADHU BANDILA	1	1
24	20992	MANESH KUMAR JAGGALI	1	1
25	20993	MANOJ KUMAR REDDY .V	1	1
26	20994	PAVAN KUMAR JANGAM	A	1
27	20995	PRBHAKARA CHINTAKAYALA	1	1
28	20996	RAMAKRISHNA NAIK BUKKE	1	1
29	20997	RISHITHA POKALA	1	1
30	20998	SAI PRAKASH VEMULA	1	1
31	20999	SANDHYA YERRABALLI	1	1
32	21000	SATHEESH GOLLAPALLI	1	A
33	21001	Seetharamanjaneyulu.K	1	1
34	21002	SHARATH KUMAR.Y	1	1
35	21003	SHIK TASNEEM MALLELA	1	1
36	21004	SIVANI BOMMI	1	1
37	21005	THULASI NALLIAHGARI	1	1
38	21006	VENKATESH AADHULA	1	1
39	21007	VENKATESH.T	1	1

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SGS ARTS COLLEGE::TTD::TPT
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MZC :: 2020-21 Bridge Course : Basic Statistics

S.NO	Roll number	Name of the student	Signature
1	20969	ARAVIND .K	K. Aravind
2	20970	ASHOK DUPPANAGURTHI	D. Ashok
3	20971	AVINASH SEDAM	S. Avinash
4	20972	AYESHA SHAIK	Sk. Ayesha
5	20973	BALAKULLEYAPPA AAKULEDU	A. Bala Kulleyappa
6	20974	CHAITANYA BALINENI	B. Chaitanya
7	20975	DEEPAK ROYAL BANDLA	B. Deepak
8	20976	Dora Babu. Kaputhota	K. Dora babu.
9	20977	DHANUSH CHINTHALA	C. Dhanush
10	20978	DHEERAJ KUMAR SODEM	S. Dheeraj Kumar
11	20979	GOWTHAM NAIK MUDE	M. Gowtham Naik
12	20980	GURU RAJU KURUVA	K. Gururaju
13	20981	HANISH BABU BODANAPU	B. Hanish Babu
14	20982	HARSHITA DAKSHINADI	D. Harshitha
15	20983	HEMANTH KARNAM	K. Hemanth.
16	20984	JEEVAN KUMAR KEELAPATLA	K. Jeevan Kumar
17	20985	KISHORE NAIK MUDE	M. Kishore Naik
18	20986	LOKANADHAM GUTHI	G. Lokanadham
19	20987	KUSHWANTH KUMAR.M	M. Kushwanth Kumar
20	20988	LAKSHMIKANTH NAIK.B	Lakshmi Kanth
21	20989	LOKESH ANKE	A. Lokesh
22	20990	LOKESH PUTTURU	P. Lokesh
23	20991	MADHU BANDILA	B. Madhu
24	20992	MANESH KUMAR JAGGALI	J. Manesh Kumar
25	20993	MANOJ KUMAR REDDY .V	Manoj Kumar
26	20994	PAVAN KUMAR JANGAM	J. Pavan Kumar
27	20995	PRBHAKARA CHINTAKAYALA	C. Prabhakar
28	20996	RAMAKRISHNA NAIK BUKKE	Ramakrishna Naik
29	20997	RISHITHA POKALA	P. Rishitha
30	20998	SAI PRAKASH VEMULA	Sai Prakash
31	20999	SANDHYA YERRABALLI	Sandhya
32	21000	SATHEESH GOLLAPALLI	Satheesh
33	21001	Seetharamanjaneyulu.K	Seetharamanjaneyulu
34	21002	SHARATH KUMAR.Y	Y. Sharath Kumar
35	21003	SHIK TASNEEM MALLELA	Tasneem
36	21004	SIVANI BOMMI	B. Sivani
37	21005	THULASI NALLIAHGARI	Thulasi
38	21006	VENKATESH AADHULA	A. Venkatesh
39	21007	VENKATESH.T	T. Venkatesh

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2	20870	BHANU KIRAN YADALA	Y. Bhanukiran
3	20871	BHANU PRASAD NALAVALI	N. Bhanu Prasad
4	20872	BHARATH KUMAR .K	K. Bharath kumar
5	20873	BHARGAVA DAKKA	D. Bhargava
6	20874	BHUVANESH CHERUKURI	C. Bhuvanesh
7	20875	CHAITANYA MADDURI	M Chaithanya
8	20876	CHANDRA VAMSI AREKANTI	A. Chandra Vamsi
9	20877	CHANDU KOPPALA	K. Chandu
10	20878	GANESH ERRABALLI	E. Ganesh
11	20879	GOWTHAMI KATINENI	K. Gowthami
12	20880	GUNATEJA RAYAPATI	R. Guna teja
13	20881	HARI OM MADURI	M. Hari Om
14	20882	JANARDHAN KOTHAKOTA	Janardhan. K
15	20883	JAVEED SHAIK	Javeed. Shaik
16	20884	JHANSI BAI KELUTH	K. Jhansi Bai
17	20885	KAMAL KATUPAKAM	
18	20886	KIRAN KUMAR VADIGILA	V. Kiran Kumar
19	20887	MAHESH ABBIGANDLA	A. Mahesh
20	20888	MEGHANA ALWAR	Meghana. A
21	20889	MOHAN CHEEMALA	C. MOHAN
22	20890	NARASIMHA VRADAN DAMAA	
23	20891	NAVEEN KUMAR MEKALA	
24	20892	PAVAN KALYAN BURRI	B. Pavan kalyan
25	20893	PRAVEEN LOKASANI	L. praveen
26	20894	PRAVEEN VALLURU	V. Praveen Kumar
27	20895	PREMSAGAR NAKKA	N. Sagar
28	20896	RADHIKA JANAM	Radhika. Janam
29	20897	RAGHU BHEEMANNAGARI	Raghu Bharmanagari
30	20898	Sai Yaswanth Kumar.C	C.
31	20899	SARAT CHANDRA .S	S. Sarath Chandra
32	20900	SIVA KUMAR BOLLA	Siva Kumar Golla

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33	20901	SIVANATH REDDY MUDAMALA	Siva nathreddy
34	20902	SWATHI GUNJI	G. Swathi
35	20903	UMESH CHANDRA MANGALI	Mumesh chandra
36	20904	VAMSI BANDILA	Vamsi
37	20905	VINEELA PARICHERLA	P. Veneela
38	20906	VINOD KUMAR MALLEPOGU	M. Vinod Kumar

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S.NO	Roll number	Name of the student	15.03.2021	16.03.2021
1	20869	ANUPRIYA KUPPALA	1	1
2	20870	BHANU KIRAN YADALA	1	1
3	20871	BHANU PRASAD NALAVALI	1	1
4	20872	BHARATH KUMAR .K	1	1
5	20873	BHARGAVA DAKKA	1	1
6	20874	BHUVANESH CHERUKURI	1	1
7	20875	CHAITANYA MADDURI	1	1
8	20876	CHANDRA VAMSI AREKANTI	1	1
9	20877	CHANDU KOPPALA	1	1
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11	20879	GOWTHAMI KATINENI	1	1
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13	20881	HARI OM MADURI	1	1
14	20882	JANARDHAN KOTHAKOTA	1	1
15	20883	JAVEED SHAIK	1	1
16	20884	JHANSI BAI KELUTH	1	1
17	20885	KAMAL KATUPAKAM	1	1
18	20886	KIRAN KUMAR VADIGILA	1	1
19	20887	MAHESH ABBIGANDLA	1	A
20	20888	MEGHANA ALWAR	1	1
21	20889	MOHAN CHEEMALA	1	1
22	20890	NARASIMHA VRADAN DAMAA	1	1
23	20891	NAVEEN KUMAR MEKALA	1	1
24	20892	PAVAN KALYAN BURRI	1	1
25	20893	PRAVEEN LOKASANI	1	1
26	20894	PRAVEEN VALLURU	1	1
27	20895	PREMSAGAR NAKKA	1	1
28	20896	RADHIKA JANAM	1	1
29	20897	RAGHU BHEEMANNAGARI	1	1
30	20898	Sai Yaswanth Kumar.C	1	1
31	20899	SARAT CHANDRA .S	1	1
32	20900	SIVA KUMAR BOLLA	1	1
33	20901	SIVANATH REDDY MUDAMALA	1	1
34	20902	SWATHI GUNJI	1	1
35	20903	UMESH CHANDRA MANGALI	1	1
36	20904	VAMSI BANDILA	1	1
37	20905	VINEELA PARICHERLA	1	1
38	20906	VINOD KUMAR MALLEPOGU	1	1

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1	20869	ANUPRIYA KUPPALA	22
2	20870	BHANU KIRAN YADALA	22
3	20871	BHANU PRASAD NALAVALI	16
4	20872	BHARATH KUMAR .K	19
5	20873	BHARGAVA DAKKA	20
6	20874	BHUVANESH CHERUKURI	17
7	20875	CHAITANYA MADDURI	23
8	20876	CHANDRA VAMSI AREKANTI	24
9	20877	CHANDU KOPPALA	22
10	20878	GANESH ERRABALLI	20
11	20879	GOWTHAMI KATINENI	19
12	20880	GUNATEJA RAYAPATI	19
13	20881	HARI OM MADURI	12
14	20882	JANARDHAN KOTHAKOTA	20
15	20883	JAVEED SHAIK	22
16	20884	JHANSI BAI KELUTH	23
17	20885	KAMAL KATUPAKAM	17
18	20886	KIRAN KUMAR VADIGILA	20
19	20887	MAHESH ABBIGANDLA	23
20	20888	MEGHANA ALWAR	24
21	20889	MOHAN CHEEMALA	25
22	20890	NARASIMHA VRADAN DAMAA	25
23	20891	NAVEEN KUMAR MEKALA	20
24	20892	PAVAN KALYAN BURRI	19
25	20893	PRAVEEN LOKASANI	12
26	20894	PRAVEEN VALLURU	18
27	20895	PREMSAGAR NAKKA	17
28	20896	RADHIKA JANAM	20
29	20897	RAGHU BHEEMANNAGARI	22
30	20898	Sai Yaswanth Kumar.C	20
31	20899	SARAT CHANDRA .S	19
32	20900	SIVA KUMAR BOLLA	21
33	20901	SIVANATH REDDY MUDAMALA	22
34	20902	SWATHI GUNJI	23
35	20903	UMESH CHANDRA MANGALI	19
36	20904	VAMSI BANDILA	17
37	20905	VINEELA PARICHERLA	20
38	20906	VINOD KUMAR MALLEPOGU	19

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DEPARTMENT of STATISTICS

- SGS ARTS COLLEGE

MODEL EXAM PAPER - 2020-21

BRIDGE COURSE - BIOLOGY STUDENTS

21/25

NAME:

Group:

Total Marks: 25

Time: 1hr

SECTION - A

All questions carry equal marks (10×1 = 10)

Fill in the blanks

(1) Data obtained by conducting a survey is called primary data

(2) The data collected from published reports is known Secondary data

(3) statistics deals with only Quantitative data

(4) The difference between the upper and lower limit of a class is called class interval

(5) An arrangement of data in rows and columns is known as tabulation.

(6) Measures of central tendency are also known as typical value or _____.

(7) Median is same as fifteeth percentile.

(8) Best measure of dispersion is Mean deviation

(9) The term regression was introduced by Sir. Francis Galton

(10) The significance of regression parameters can be tested by S-test

SECTION-B

Multiple choice questions (15×1 = 15)

(11) The word 'statistics' is used as: [c]

a) singular b) plural

c) singular and plural both d) none of the above

(12) which of the following represents data? [c]

a) a single value b) only two values in a set

c) a group of values in a set d) none of the above

(13) In a grouped data, the number of classes preferred are [a]

a) minimum possible b) adequate

c) maximum possible d) any arbitrarily chosen number

(14) Trilinear chart is used to portray simultaneously: [b]

a) two variables b) three variables

c) four variables d) any number of variables.

(15) Non-dimensional diagrams are also known as: [a]

a) cubes b) spheres

c) pictograms d) All the above

(16) ogive curve occur for, [c]

a) more than type distribution

b) less than type distribution

c) both (a) and (b)

d) none of (a) and (b)

(17) which of the following is not a measure of dispersion? [d]

a) mean deviation b) quartile deviation

c) standard deviation d) average deviation from mean.

which one of the given measures of dispersion is considered best? [a]

a) standard deviation b) quartile deviation

c) range d) variation

(19) which of the following is a unitless measure of dispersion [c]

a) standard deviation b) mean deviation

c) coefficient of variation d) range

(20) The term regression was introduced by: [b]

a) R.A. Fisher b) Sir Francis Galton

c) Karl pearson d) none of the above

(21) Regression equation is also named as: [d]

a) prediction equation b) estimating equation

c) line of average relationship d) all the above

(22) In the regression line $Y = \alpha + \beta X$, β is called the [a]

a) slope of the line b) intercept of the line

c) neither (a) nor (b) d) both (a) and (b)

(23) If $\beta_{yx} > 1$, then β_{xy} is [a]

a) less than 1 b) greater than 1

c) equal to 1 d) equal to 0

(24) The shape of trilinear charts is that of a: [c]

a) cone b) cube c) equilateral triangle d) pyramid

(25) Ogives for more than type and less than type distributions intersect at: [b]

a) mean b) median

c) mode d) origin

Key — 2020-2021 model exam paper.

SECTION - A

① primary, ② secondary, ③ Quantitative data,
④ class interval, ⑤ tabulation, ⑥ typical value, ⑦ fiftieth,
⑧ standard deviation, ⑨ Sir Francis Galton,
⑩ t-test

SECTION - B

⑪ c, ⑫ c, ⑬ b, ⑭ b, ⑮ c, ⑯ c, ⑰ d, ⑱ a, ⑲ c,
⑳ b, ㉑ d, ㉒ a, ㉓ a, ㉔ c, ㉕ b

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TIRUMALA TIRUPATI DEVASTHANAMS
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Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY,
M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

Tiruprti
Date:22.07.2019

From
Sri. D. Parameswara,
HOD of Chemistry,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.

Sir,

Sub:- Permission for conducting Chemistry Bridge course to students of first year degree – request – Regarding.

&&&&&

Kindly permit the Department of Chemistry to conduct Bridge Course to the students of first year degree who studied Vocational Course in intermediate for one week i.e. from 24-07-2019 to 30-07-2019. Time table and syllabus are enclosed.

Thanking you sir

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE

Yours faithfully

S.G.S ARTS COLLEGE,TIRUCHANOOR ROAD,TIRUPATI
Department of Chemistry

Circular

First-year Life Sciences students who studied vocational courses at the intermediate level are informed to attend the bridge course which will be held at seminar hall, SGS Degree College, Tirupati from 24-07-2019 to 30-07-2019.

Head,
Department of Chemistry

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DEPARTMENT OF CHEMISTRY
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Bridge Course-Summary

Name of the programme	Bridge course in chemistry for vocational students
Duration of Course	One Week
Date	24-07-2019 to 30-07-2019
Venue	Seminar Hall

S.No	Subject name	Faculty Name	No. of Students
1	Inorganic Chemistry	P Suguna	4
2	Physical Chemistry	V Venkata Lakshmi	4
3	Organic Chemistry	D Parameswara	4
4	General Chemistry	K. Revathi	4

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

2019-2020
Bridge Course in Chemistry
Students List

S.No	Name of the student	Signature
1	P.Munendra	P. Munendra
2	M.Naveen Naik	M. Naveen Naik
3	N.Lokesh	N Lokesh.
4	Shaik Reena	Shaik Reena
5	S.Mary	S. Mary
6	M.Pandiyan	M. Pandiyan

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Department of Chemistry
Bridge Course Time Table

Day	10AM-11AM	11AM-12NOON	12NOON-1PM	2PM-3PM	3PM-4PM
Wednesday	Inorganic	Physical	Organic	General	Q&A Session
Thursday	Physical	Organic	General	Inorganic	
Friday	Organic	General	Inorganic	Physical	
Saturday	General	Inorganic	Physical	Organic	
Monday	Inorganic	Physical	Organic	General	
Tuesday	Physical	Organic	General	Inorganic	

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Bridge Course in Chemistry

Syllabus

S.No	Subject	Syllabus
1	Inorganic Chemistry	Electronic Configuration of elements Classification of elements in to s,p,d,f blocks. Periodic properties
2	Physical Chemistry	States of matter- Gaseous state,solid state Solutions Ionic Equilibria
3	Organic Chemistry	Nomenclature of Organic Compounds Types of Reactions Types of Reagents Bond polarising Effects Reaction Intermediates
4	General Chemistry	Stereo Chemistry Colloids Adsorption Analytical Chemistry

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S. G. S. ARTS COLLEGE
TIRUPATHI

Department of Chemistry
Bridge Course Student Attendance List

S.No	Name of the student	Course Studied in Intermedate	Programme Joined in UG	24/7/19	25/7/19	26/7/19	27/7/19	29/7/19	30/7/19
1	P.Munendra	Dairy Science	CBZ	X	X	X	a	X	X
2	M.Naveen Na	Dairy Science	CBZ	X	X	X	X	X	X
3	N.Lokesh	Agl.Polytechnic	CBZ	X	a	X	X	X	X
4	Shaik Reena	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X
5	S.Mary	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X
6	M.Pandiyan	MLT (Medical Lab Technician Course)	MZC	X	X	X	X	X	X

1.

Signature of the faculty

2.

3.

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Bridge Course in Chemistry
Students Marks List

S.No	Name of the student	Marks
1	P.Munendra	13 - Thirteen only
2	M.Naveen Naik	12 - Twelve only
3	N.Lokesh	15 - fifteen only
4	Shaik Reena	13 - Thirteen only
5	S.Mary	14 - fourteen only
6	M.Pandiyan	13 - Thirteen only

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S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD,TIRUPATI
DEPARTMENT OF CHEMISTRY
OBJECTIVE TEST - BRIDGE COURSE 2019-2020

Answer all questions Maximum Marks: 20

---

1. The physical and chemical properties of elements are the periodic function of their atomic weights is ( )
a) Mendeleev's periodic law
b) Lother Meyer's periodic law
c) Moseley's periodic law
d) Bohr's periodic law

2. Calcium has an atomic number of 20. A stable calcium atom has an electronic configuration of: ( )
a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
b) $1s^2 1p^6 1d^{10} 1f^2$
c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2$
d) $1s^2 2s^2 2p^6 3s^2 3p^6$

3. The correct order of second ionization potential of C, N, O and F is ( )
a) C > N > O > F b) O > N > F > C c) O > F > N > C d) F > O > N > C

4. Of the different allotropic forms of phosphorus, most reactive is ( )
a) Violet phosphorus
b) White phosphorus
c) Scarlet phosphorus
d) Red Phosphorus

5.Which of the following is not an element of first transition series? ( )
a) Iron
b) Chromium
c) Magnesium
d) Nickel

6. Which group contains coloured ions out of the following? ( )
1. Cu^-. 2. Ti^{4-} 3. Co^{2+} 4. Fe^{2-}
a) 1, 2, 3, 4
b) 3, 4
c) 2, 3
4) 1, 2

7.In the equation of sate of an ideal gas PV = nRT, the value of the universal gas constant would depend only on ( )
a) The pressure of the gas
b) The units of the measurement

c) The nature of the gas
d) None of these

8. At the same temperature and pressure, which of the following gases ( )
will have the highest kinetic energy per mole

a) Oxygen
b) Methane
c) Hydrogen
d) All the same

9. On dissolving sugar in water at room temperature solution feels cool ( )
to touch.Under which of the following cases dissolution of sugar will be
most rapid?

a)Sugar crystals in cold water.
b) Sugar crystals in hot water.
c) Powdered sugar in cold water.
d) Powdered sugar in hot water.

10. Which one of the following is non-crystalline or amorphous? ( )

a) Diamond
b) Graphite
c) Glass
d) Common Salt

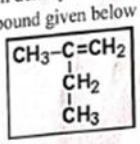
11. Which of the following is a characteristic property of equilibrium? ( )

a) Number of moles of reactants and products is always equal.
b) Catalyst affects the equilibrium state.
c) It never proceeds to completion.
d) Rate of forward and backward reactions are not equal.

12.Find the incorrect statement for a nucleophile ( )

a) A nucleophile is a Lewis acid
b) Nucleophiles do not seek electron
c) Ammonia is a nucleophile
d) Nucleophiles attack low electron density sites

13. The correct name for the compound given below is: ( )



a) 2-methyl-1-butene
b) 2-ethyl-1-propene
c) 2-ethyl-1-pentane
d) 3-methyl-2-butene

14. Alkene under high temperature and high-pressure forms ____ ( )
a) Alcohol
b) Polyalkyne
c) Polyalkane
d) Polyalkene

15. The phenomenon in which 2 or more structures, involving identical position of atoms can be written for a particular molecule, is called ___ ( )
a) Conjugation
b) Resonance
c) Hyper conjugation
d) Vibration

16. Which of the following is a permanent electron displacement effect? ( )
a) Inductomeric
b) Electromeric
c) Inductive
d) All of the mentioned

17. Tyndall effect confirms the ( )
(a) gravity effect on the sol. particles
(b) light scattering by the sol. particles
(c) heterogeneous nature of sols.
(d) Brownian motion of the sol. particles

18. Which of the following can adsorb larger volume of hydrogen gas? ( )
(a) Finely divided platinum
(b) Colloidal solution of palladium
(c) Small pieces of palladium
(d) A single metal surface of platinum

19. Which of the following aqua regia is. ( )
a 1 part conc. HCl and 3 parts conc. HNO
b) 3 part conc. HCl and 1 part conc. HNO
c) 2 parts conc. HCl and 1 part conc. HNO
d) 3 parts conc. HCl and 2 parts conc. HNO2

20. Which of the following can make difference in optical isomers? ( )
a) heat
b) temperature
c) polarized light
d) pressure

Answers:

1	A	11	C
2	A	12	A
3	C	13	A
4	D	14	C
5	C	15	B
6	B	16	C
7	B	17	C
8	D	18	B
9	D	19	B
10	C	20	C

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Tiruprti
Date:05.03.2021

From
Sri. D. Parameswara,
HOD of Chemistry,
SGS Arts College,
Tirupathi.

To
The Principal,
SGS Arts College,
Tirupathi.

Sir,

Sub:- Permission for conducting Chemistry Bridge course to students of first year degree – request – Regarding.

&&&&&

Kindly permit the Department of Chemistry to conduct Bridge Course to the students of first year degree who studied Vocational Course in intermediate for two days i.e. 08-03-2021 and 09-03-2021. Time table and syllabus are enclosed.

Thanking you sir

Yours faithfully

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S Arts College,Tirupati
Department of Chemistry

Circular

First-year Life Sciences students who studied vocational courses at the intermediate level are informed to attend the bridge course which will be held at seminar hall, SGS Degree College, Tirupati from 08-03-2021 to 09-03-2021.

Head,
Department of Chemistry

DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Bridge Course in Chemistry
Students List

S.No	Name of the student	Signature
1	T.Radhika	T. Radhika
2	N.Bhanu prasad	N.T.Bhanu Prasad
3	K.Kamal	K.Kamal
4	K.Chandu	K. Chandu
5	T.Vignesh yadav	T. Vignesh Yadav
6	P.Chand basha	P. Chand basha
7	M.Kishor naik	M. Kishor naik
8	G. Satheesh	G. Satheesh

SIGNATURE

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Department of Chemistry
Bridge Course Time Table

Day	10AM-11AM	11AM-12NOON	12NOON-1PM	2PM-3PM	3PM-4PM
Monday	Inorganic	Physical	Organic	General	Q&A Session
Tuesday	Physical	Organic	General	Inorganic	

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TIRUPATHI

Bridge Course in Chemistry

Syllabus

S.No	Subject	Syllabus
1	Inorganic Chemistry	Electronic Configuration of elements Classification of elements in to s,p,d,f blocks. Periodic properties
2	Physical Chemistry	States of matter- Gaseous state,solid state Solutions Ionic Equilibria
3	Organic Chemistry	Nomenclature of Organic Compounds Types of Reactions Types of Reagents Bond polarising Effects Reaction Intermediates
4	General Chemistry	Stereo Chemistry Colloids Adsorption Analytical Chemistry

Department of Chemistry
Bridge Course Student Attendance List

S.No	Name of the student	Course Studied in Intermedate	Programme Joined in UG	8/3/21	9/3/21
1	j.Radhika	Diploma in agriculture	CBZ	X	X
2	N.Bhanu prasad	Diploma in agriculture	CBZ	X	X
3	K.Kamal	(Medical Lab Technician	CBZ	X	X
4	K.Chandu	Diploma in seed	CBZ	X	X
5	T.Vignesh yadav	Diploma in agriculture	MZC	X	X
6	P.Chand basha	Diploma in agriculture	BBC	X	a
7	M.Kishor naik	Diploa in organic farming	MZC	X	X
8	G.sathesh	MLT (Medical Lab Technician Course)	MZC	X	X

Signature of the faculty

1)

2)

3)

Head,Department of Chemistry

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DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

Bridge Course in Chemistry
Students Marks List

S.No	Name of the student	Marks
1	j.Radhika	13 - Thirteen only
2	N.Bhanu prasad	16 - Sixteen only
3	K.Kamal	14 - fourteen only
4	K.Chandu	15 - fifteen only
5	T.Vignesh yadav	14 - fourteen only
6	P.Chand basha	17 - seventeen only
7	M.Kishor naik	16 - six teen only
8	G.Satheesh	13 - Thirteen only

SIGNATURE

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S. G. S. ARTS COLLEGE
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Bridge Course-Summary

Name of the programme	Bridge course in chemistry for vocational students
Duration of Course	Two days
Date	08-03-2021 to 09-03-2021.
Venue	Seminar Hall

S.No	Subject name	Faculty Name	No. of Students
1	Inorganic Chemistry	P Suguna	4
2	Physical Chemistry	V Venkata Lakshmi	4
3	Organic Chemistry	D Parameswara	4
4	General Chemistry	K. Revathi	4

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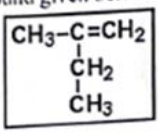
S. G. S. ARTS COLLEGE
TIRUPATHI

S.G.S ARTS COLLEGE: TIRUCHANOOR ROAD,TIRUPATI
DEPARTMENT OF CHEMISTRY
OBJECTIVE TEST - BRIDGE COURSE 2020-2021

Answer all questions Maximum Marks: 20

1. Which of the following is a characteristic property of equilibrium? ( )
   a) Number of moles of reactants and products is always equal.
   b) Catalyst affects the equilibrium state.
   c) It never proceeds to completion.
   d) Rate of forward and backward reactions are not equal.
2. Find the incorrect statement for a nucleophile ( )
   a) A nucleophile is a Lewis acid
   b) Nucleophiles do not seek electron
   c) Ammonia is a nucleophile
   d) Nucleophiles attack low electron density sites
3. The correct name for the compound given below is: ( )



   a) 2-methyl-1-butene
   b) 2-ethyl-1-propene
   c) 2-ethyl-1-pentane
   d) 3-methyl-2-butene
4. Alkene under high temperature and high-pressure forms ____ ( )
   a) Alcohol
   b) Polyalkyne
   c) Polyalkane
   d) Polyalkene
5. The phenomenon in which 2 or more structures, involving identical position of atoms can be written for a particular molecule, is called ___ ( )
   a) Conjugation
   b) Resonance
   c) Hyper conjugation
   d) Vibration
6. Which of the following is a permanent electron displacement effect? ( )
   a) Inductomeric

b) Electromeric
c) Inductive
d) All of the mentioned

7. Tyndall effect confirms the ( )
(a) gravity effect on the sol. particles
(b) light scattering by the sol. particles
(c) heterogeneous nature of sols.
(d) Brownian motion of the sol. particles

8.Which of the following can adsorb larger volume of hydrogen gas? ( )
(a) Finely divided platinum
(b) Colloidal solution of palladium
(c) Small pieces of palladium
(d) A single metal surface of platinum

9. Which of the following aqua regia is. ( )
a 1 part conc. HCl and 3 parts conc. HNO
b) 3 part conc. HCl and 1 part conc. HNO
c) 2 parts conc. HCl and 1 part conc. HNO
d) 3 parts conc. HCl and 2 parts conc. HNO2

10. Which of the following can make difference in optical isomers? ( )
a) heat
b) temperature
c) polarized light
d) pressure

11. The physical and chemical properties of elements are the periodic ( )
Function of their atomic weights is
a) Mendeleev's periodic law
b) Luther Meyer's periodic law
c) Moseley's periodic law
d) Bohr's periodic law

12. Calcium has an atomic number of 20. A stable calcium atom ( )
has an electronic configuration of:
a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
b) $1s^2 1p^6 1d^{10} 1f^2$
c) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2$
d) $1s^2 2s^2 2p^6 3s^2 3p^6$

13. The correct order of second ionization potential of C, N, O and F is ( )
a) C > N > O > F b) O > N > F > C c) O > F > N > C d) F > O > N > C

14. Of the different allotropic forms of phosphorus, most reactive is ( )
a) Violet phosphorus

b) White phosphorus
c) Scarlet phosphorus
d) Red Phosphorus

15. Which of the following is not an element of first transition series? ( )
a) Iron
b) Chromium
c) Magnesium
d) Nickel

16. Which group contains coloured ions out of the following? ( )
1. Cu^{-}. 2. Ti^{4-} 3. Co^{2+} 4. Fe^{2-}
a) 1, 2, 3, 4
b) 3, 4
c) 2, 3
4) 1, 2

17. In the equation of sate of an ideal gas PV = nRT, the value of the ( ) universal gas constant would depend only on
a) The pressure of the gas
b) The units of the measurement
c) The nature of the gas
d) None of these

18. At the same temperature and pressure, which of the following gases ( ) will have the highest kinetic energy per mole
a) Oxygen
b) Methane
c) Hydrogen
d) All the same

19. On dissolving sugar in water at room temperature solution feels cool ( ) to touch. Under which of the following cases dissolution of sugar will be Most rapid?
a)Sugar crystals in cold water.
b) Sugar crystals in hot water.
c) Powdered sugar in cold water.
d) Powdered sugar in hot water.

20. Which one of the following is non-crystalline or amorphous? ( )
a) Diamond
b) Graphite
c) Glass
d) Common Salt

Answers:

1	C	11	A
2	A	12	A
3	A	13	C
4	C	14	D
5	B	15	C
6	C	16	B
7	C	17	B
8	B	18	D
9	B	19	D
10	C	20	C

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATI

3. ENDOWMENT PRIZES

With an objective to identify the best performing students, to motivate a students for achieving excellence in both curricular and extra-curricular activities our college has an institutionalized endowment prizes.

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

Students benefited by institutional level Prizes/Awards/ Medals for Academic achievements

Fund received from the donors for the ward of endowment cash prizes, awards, medals etc:

Existing total capital amount Rs 1,63,928 ,amount has been deposited as Fixed Deposit in Union Bank of India, the amount received as interest and the above deposit will be for the award of cash prize or award or medal who stood as Toppers.

1. **Ramesh Memorial** cash award: This awards instituted by his father Sri.G.Ramachandra Reddy Retired Chief Engineer .A.P to be awarded to the 3rd year BCom student who has secures highest aggregate marks in all semesters in Ist and 2nd years.
2. Prof. **G. Leela Krishna** Endowment Cash award: This award is instituted by Prof. **K.V.Naidu**, Former, Director of SVIMS, Tirupati as a token of Respect on his teacher, Prof. **G. Leela Krishna** to be awarded to the student of 3rdB.Scwho has secures the highest marks in Botany in all semesters of Ist year and 2nd year.

3 . **SmtPaluvai. Vijayalakshmi** Memorial cash award: This cash award of was instituted by her husband **Sri P. Subbarao** , Retd. HOD. Of Chemistry, S.V.Arts College, and former HOD. Of Chemistry of our College to be

awarded to (the student of III BSc with Chemistry combination of either medium scoring the highest total marks in all the semesters of Ist and 2nd year) to top rankers in Chemistry , one each at 2nd year and 3rd year who have secures the highest aggregate marks .

3. Sri **T. Surendranath Reddy** cash award: This prize is awarded to the topper in commerce who secures the highest marks in I & II year B.Com

4. **Dr. Dasa Raju, Rama Raju** endowment cash award: This prize is awarded to the student of 3rd year who got highest marks in Statistics I & II .

5. **Sri K.V.Gopalakrishna Murthy** and **Smt. K.Saraswathy** endowment cash award: This prize is awarded to the topper in English who has secured the highest aggregate marks in I and II Year.

6. **Mr. A. Goutham, Endownment**cash award: This prize is awarded to (The Topper in Mathematics who secures the highest marks in 2nd year BSc) to Top rankers in Mathematics, one each at Ist year and 2nd year, who have secured highest marks.

7. **Late Sri.P.Purushotham Naidu** Memorial endowment cash awards: The awards are instituted by his affectionate friends BA(HEP) 1992-1995 batch .This prize is awarded to EIGHT Top rankers one each at first year BA(HEP) and 2nd year BA(HEP).
   First year BA (HPT, MES, ASA)
   Second year BA (HEP, HPT, MES).

8. **Prof K.Sri Ramarao** Rtd HOD of Botany endowment prize: This prize is Top rankers one in Botany and the other in Microbiology who secures highest marks in all the semesters of First year & Second Year in subjects of Botany and Microbiology.

9. **S.G.S. Arts College** Science cash award: This prize is awarded to the student who secures the highest aggregate marks in all the subjects in all the semesters of BSc.
10. **Sri S.RangaRajan** cash award: This prize is awarded to the student of Maths combinations who secures the highest marks in all the group subjects of First and second year students.

Winners of the above prizes, Awards and Medals for their Academic Achievements.

2019-2020

Sno	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	K.Deepa	III Bcom-CA	87.3%	0218008014	**Ramesh Memorial** cash award, **T. Surendranath Reddy** cash award
2	S.Munikumar	III Bsc-BBC	551/600	0318008166	**G. Leela Krishna** Endowment Cash award
3	T.Avinash	III BSc-BBC	87.0%	0318008143	**Smt Paluvai. Vijayalakshmi** Memorial cash award
4	T.Avinash	III BSc-BBC	279/300	0318008143	**Smt Paluvai. Vijayalakshmi** Memorial cash award
5	A.Venkatesh	III BSc-MSCs	590/600	0318008121	**Dasa Raju, Rama Raju** endowment cash award
6	A.Venkatesh	III BSc-MSCs	89.3%	0318008121	**S.RangaRajan** cash award
7	B.Mohana Krishna	III BSc-MSCs	239/300	0318008311	**K.V.Gopalakrishna Murthy** and **Smt. K.Saraswathy** endowment cash ward
8	B.Mohana Krishna	III BSc-MSCs	197/200	0318008311	**A. Goutham, Endownment** cash ward
9	B.Srinivasulu	II BSC-MPCs	200/200	0319008675	**A. Goutham, Endownment** cash ward
10	K.Sireesha	II BSc-MPC	200/200	0319008172	**A. Goutham, Endownment** cash ward
11	G.Surendra	II BSc-MPC	200/200	0319008120	**A. Goutham, Endownment** cash ward

2018-2019

Sno	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	N.Purushotham	III Bcom-CA	86%	0217008579	**Ramesh Memorial** cash award, **T. Surendranath Reddy** cash award
2	S.Snehalatha	III Bsc-BBC	559/600	0317008681	**G. Leela Krishna** Endowment Cash award
3	P.Bharath Naveen	III BSc-BBC	550/600	0317008652	**Smt Paluvai. Vijayalakshmi** Memorial cash award
4	P.Guru sai	III BSc-MSCs	560/600	0317008675	**Dasa Raju, Rama Raju** endowment cash award
5	P.Gurusai	III BSc-MSCs	88%	0317008675	**S.RangaRajan** cash award
6	P.Shilpa	III BSc-MZC	245/300	0317008668	**K.V.Gopalakrishna Murthy** and **Smt. K.Saraswathy** endowment cash Award
7	S.Yerramala	II BA-HPT ™	84%	0117008019	**Late.Purushottam Naidu Endowment** cash ward
8	B.Pavan Kumar	II BA-HEP (EM)	84%	0117008505	**Late.Purushottam Naidu Endowment** cash ward
9	G.C.Gunasekhar	II BA-ASCA (EM)	83%	0117008513	**Late.Purushottam Naidu Endowment** cash ward
10	C.Sreenu	II BA-MES (EM)	87%	0117008510	**Late.Purushottam Naidu Endowment** cash ward
11	N.Purushotham	III Bcom-CA	86%	0217008579	**T. Surendranath Reddy** cash award

2017-2018

S. No	Name of the Student	Group & Year	Marks/ Percentage	Hall Ticket No	Winner of the prize
1	P.Gowtham Kumar	III BCom-CA	87.84%	0216008592	**Ramesh Memorial** cash award
2	N.Asthvadath	III BSc-BBC	475/550	0316008653	**G. Leela Krishna** Endowment Cash award
3	G.Venkateswarulu	III BSc-MPC(EM)	86.72%	0316008515	**A. Goutham, Endownment** cash ward
4	M.Sunil Kumar	III BSc-MPC(EM)	255/300	0316008523	**A. Goutham, Endownment** cash ward
5	N.Giridhar	III Bom	75.3%	0216008584	**T. Surendranath Reddy** cash award
6	B.Yuva Kishore	III BSc-MSCs	375/400	0316008709	**Dasa Raju, Rama Raju** endowment cash award
7	B.Muthyalu	II Bsc-MScs	81.0%	0316008572	**K.V.Gopalakrishna Murthy** and **Smt. K.Saraswathy** endowment cash ward
8	G.Venkateswarulu	III BSc-MPC(EM)	200/200	0316008515	**A. Goutham, Endownment** cash ward
9	P.Gurusai	II BSc-MSCs	195/200	0317008675	**A. Goutham, Endownment** cash ward
10	R.Kalairasi	II BSc-MSCs	195/200	0317008678	**A. Goutham, Endownment** cash ward
11	A.Ganesh	II BSc-MPCs	86.18%	0316008533	**S.RangaRajan** cash award

2016-2017

Sno	Name of the Student	Group & Year	Marks	Hall Ticket No	Winner of the prize
1	B.Bharath	III BCom-CA	1186/1300	215042814	**Ramesh Memorial** cash award
2	B.Bharath	III BCom-CA	355/500	215042814	**T. Surendranath Reddy** cash award
3	B.SuryaPrathap	III BSc- BBC	265/300	315042844	**Ramesh Memorial** cash award
4	B.SuryaPrathap	III BSc- BBC	118/150	315042844	**SmtPaluvai. Vijayalakshmi** Memorial cash award
5	A.Chandra Sekhar Naidu	III BSc-MZC	377/450	315042812	**SmtPaluvai. Vijayalakshmi** Memorial cash award
6	B.Masthan	III BSc-MSCs	187/300	315042771	**Dasa Raju, Rama Raju** endowment cash award
7	B.Masthan	III BSc--MSCs	831/900	315042771	**S.G.S. Arts College** Science cash award
8	G.Moulali	III BSc-MSCs	157/200	315042780	**K.V.Gopalakrishna Murthy** and **Smt. K.Saraswathy** endowment cash ward
9	R.Prathap	III BSc-MPC	300/300	315042726	**A. Goutham, Endownment** cash ward
10	B.Narendra Reddy	II BSc-MPCs	200/200	316008541	**A. Goutham, Endownment** cash ward
11	B.Masthan	II BSc-MScs	1099/1300	315042771	**S.RangaRajan** cash award

S.G.S. ARTS COLLEGE Co-ordinator NAAC TIRUPATI

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF CHEMISTRY

Endowment Award Winners

The students of chemistry who excelled in their academics and award winners for their outstanding performance in S.V.U Examinations.

ACADEMIC YEAR 2016-17					
S.No.	Name of the Student	Group & Year	Marks	Hall Ticket No	Prize Winner
1	B.SuryaPrathap	III BSc- BBC	118/150	315042844	Smt. Paluvai. Vijayalakshmi Memorial cash award
2	A.Chandra Sekhar Naidu	III BSc-MZC	377/450	315042812	SmtPaluvai. Vijayalakshmi Memorial cash award
ACADEMIC YEAR 2018-19					
1	P.Bharath Naveen	III BSc-BBC	550/600	0317008652	Smt. Paluvai. Vijayalakshmi Memorial cash award
ACADEMIC YEAR 2019-20					
1	T.Avinash	III BSc-BBC	87.0%	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award
2	T.Avinash	III BSc-BBC	279/300	0318008143	Smt Paluvai. Vijayalakshmi Memorial cash award

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

4. Certificate Courses

No University curriculum can adequately cover all areas of importance and relevance. It is important for higher education institutions to supplements the curriculum to make students prepared better to meet the demands of industry and develop their own interest and attitudes .The institution offers a wide variety of short term certificate courses which are conducted out of college hours and semester breaks. The institution offers two types of courses 1. General interest courses which are open to all 2. Subject specific courses. The student can pursue these courses as per his/her required. The courses are conducted by many departments.

Phone: (0877) 2264599
Web: sgsac.edu.in

e-mail:sgsactd9@gmail.com
sgsartscollegetd@gmail.com

TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S. ARTS COLLEGE
Tiruchanur Road, TIRUPATI-517501

Prof N. VENUGOPAL REDDY M.Sc., M.Phil., Ph.D
PRINCIPAL (FAC)

1.2.2.1: How many Add on /Certificate programs are offered within the last 5 years.

S. No	Academic Year	Total Number of Certificate Courses
1	2020-21	Nil
2	2019-20	10
3	2018-19	10
4	2017-18	10
5	2016-17	10

S.G.S. ARTS COLLEGE * TIRUPATI * Co-ordinator NAAC

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

SGS ARTS COLLEGE, TIRUPATI
CERTIFICATE COURSES
Academic Year: 2019-20
BROUCHURE

No University Curriculum can adequately cover at areas of importance and relevance. It is important for higher education institutions to supplement the curriculum to make students prepared better to meet the demands of industry and develop their own interests and aptitudes.

Our college offers a wide variety of short term certificate courses which are conducted out of college hours and semester breaks.

Our institution offers two types of courses.

1. General interest courses which are open to all.
2. Subject specific courses.

The student can pursue these courses as per his requirement. The courses are organized by many departments. Each course has a coordinator whose phone number is prominently displayed in the brochure.

Each course has its own assessment pattern, both practical and theoretical. Successful completion of the course earns you certification issued by the college.

S. No	Certificate Course	Organized by	Coordinator	Contact No
1	Photoshop	Computer Science	Sri P. Udaya	9440257300
2	Tally	Commerce	Dr. Y. Mallikarjuna Rao	9848533623
3	Archaeology	History	Sri A. Madhusudana Rao	9440506830
4	Floriculture	Botany	Smt T. Sakunthala	9052060327
5	Hotel & Hospitality Management	HMCT	Dr. N. Jagadeesh Kumar	8499840536
6	Soft skills	English	Dr. K. Madhu Kumar	9440257300
7	E-Commerce	Commerce	Dr. Y. Mallikarjuna Rao	9848533623
8	Maintenance of House hold Electrical Appliance	Physics	Dr. K. Ravindranath Reddy	9440164872
9	MS-Office	Computer Science	Sri P. Udaya	9440257300
10	Vedic Mathematics	Mathematics	Sri P. Munibhaskar Rao	9989305142

Note: For admission contact coordinator for the respective course.

As the basic idea is to make ample opportunities available to students, there are no restrictions on how many courses a student can join and one student can acquire several certifications during the period of degree course.

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S.G.S. ARTS COLLEGE
TIRUPATI

SGS ARTS COLLEGE, TIRUPATI
Certificate Courses
Academic Year: 2019-20

The following departments are conducting the certificate courses. The students those who are interested to study the certificate courses are requested to consult the HOD of the following department.

S. No	Department	Name of the Course	Eligibility of the Students
1	Computer Science	Photoshop	All Graduate students
2	Commerce	Tally	B.Com & B.A ASCA students
3	History	Archaeology	B.A H.E.P & H.P.T students
4	Botany	Floriculture	All Life Science Group students
5	HMCT	Hotel & Hospitality Management	HMCT students
6	English	Soft skills	All Graduate students
7	Commerce	E-Commerce	B.Com & B.A ASCA students
8	Physics	Maintenance of House hold Electrical Appliance	All B. Sc students
9	Computer Science	MS-Office	All B. Sc students
10	Mathematics	Vedic Mathematics	B. Sc M.S.Cs, M.P.Cs, M.P.S, M.P.C & M.E.S students

PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

CERTIFICATE COURSE : MS-OFFICE

ACADEMIC YEAR : 2019-2020

Permission Letter

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

Date: 14-01-2020.

From

HOD,
Department of Computer Science,
S.G.S Arts College,
Tirupati.

To

The Principal,
S.G.S Arts College,
Tirupati.

Respected Sir,

Sub: Requesting for permission to conduct Certificate Course in **MS-Office** – Regarding.

We have planned to conduct a forty hours Certificate Course on **MS-Office** for two batches of Second Year students from **20-01-2020 to 13-02-2020**. Therefore kindly permit us to conduct this Certificate Course.

Thanking You,

Yours Faithfully,

CIRCULAR

It is to inform all the faculty of **Department of Computer Science** to attend staff meeting at **11:00 PM** in the HOD chamber to discuss the conduct of Certificate Courses for the academic year **2019-20.**

Faculty Members Attended:

HOD	Name of the Faculty	Signature
Sri P.Udaya	1. Smt B.Triveni 2. Smt N.Jayalakshmi 3. Sri D.Chaitanya Kumar	

Agenda: **Certificate courses to be conducted for the Academic Year 2019-20.**

All the staff members of the department discussed the agenda. The following resolutions were passed unanimously.

- It was decided to conduct a certificate course on **MS-Office** to the Second Year students in the first week of **September** month.
- It was decided to conduct a certificate course on **MS-Office** to the Second year students in the first week of **February** month.
- HOD offered vote of thanks.

Faculty: **Signature:**

1. Smt B.Triveni
2. Smt N.Jayalakshmi
1. Sri D.Chaitanya Kumar

NOTICE

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

It is to inform to all the **II Year B.Sc MPCS, MSCS, MPC, MPS** students that we are going to conduct a Certificate Course. The details of the course offered are as follows.

Course name	**MS-Office**
Course date & duration	**20/January/2020 : 40 Hours**
Timings	**Batch 1:** Theory: 8:00 AM to 9:00 AM Practical: 4:00 PM to 5:00 PM **Batch 2:** Theory: 5:00 PM to 6:00 PM Practical: 8:00 AM to 9:00 AM
Course Mentors	**Batch 1:** B.Triveni, D.Chaitanya Kumar **Batch 2:** N.Jayalakshmi, D.Chaitanya Kumar

Students who are interested to join the course can contact the Coordinator of Certificate Course Sri P.Udaya for further details of course content, study material and prerequisites of the course.

HOD

Course Coordinator

Details of the Course

Name of the Certificate Course	:	MS-Office
Batch	:	1
Timing	:	Theory: 8:00 AM to 9:00 AM
		Practical: 4:00 PM to 5:00 PM
No. of students/ Dept/Year/ Course Duration	:	60/ Comp.Science/ 2019-20/ 40 Hours
Venue	:	Theory: Room No 215
		Practical: Comp.Science Lab
Eligible Groups	:	II Year B.Sc MPCS, MSCS, MPC, MPS
Trainers	:	B.Triveni, D.Chaithanya Kumar
Certificate	:	Yes
Course Coordinator	:	Sri P.Udaya

Details of the Course

Name of the Certificate Course	:	MS-Office
Batch	:	2
Timing	:	Theory: 4:00 PM to 5:00 PM
		Practical: 8:00AM to 9:00AM
No. of students/ Dept/Year/ Course Duration	:	60/ Comp.Science/ 2019-20/ 40 Hours
Venue	:	Theory: Room No 215
		Practical: Comp.Science Lab
Eligible Groups	:	II Year B.Sc MPCS, MSCS, MPC, MPS
Trainers	:	N.Jayalakshmi, D.Chaithanya Kumar
Certificate	:	Yes
Course Coordinator	:	Sri P.Udaya

S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office

Course Contents

Theory (20 Hours)

MS WORD

Module 1: Text Basics, Text Formatting and saving file, Working with Objects

- Typing the text, Alignment of text
- Editing Text: Cut, Copy, Paste, Select All, Clear
- Find & Replace
- New, Open, Close, Save, Save As
- Formatting Text: Font Size, Font Style
- Font Colour, Use the Bold, Italic, and Underline
- Change the Text Case
- Line spacing, Paragraph spacing
- Shading text and paragraph
- Working with Tabs and Indents
- Shapes, Clipart and Picture, Word Art, Smart Art
- Columns and Orderings - To Add Columns to a Document
- Change the Order of Objects
- Page Number, Date & Time
- Inserting Text boxes
- Inserting Word art
- Inserting symbols
- Inserting Chart

Module 2: Header & Footers, Working with bullets and numbered lists

- Inserting custom Header and Footer
- Inserting objects in the header and footer
- Add section break to a document
- Multilevel numbering and Bulleting
- Creating List
- Customizing List style
- Page bordering
- Page background

Module 3: Tables, Styles and Content

- Working with Tables, Table Formatting
- Table Styles
- Alignment option
- Merge and split option
- Using Build- in Styles, Modifying Styles
- Creating Styles, Creating a list style
- Table of contents and references
- Adding internal references

- Inserting Column, Pie chart etc.
- Create an effective chart with Chart Tool
- Design, Format, and Layout options
- Adding chart title
- Changing layouts
- Chart styles
- Editing chart data range
- Editing data series
- Changing chart

Module 10: Proofing and Printing

- Page setup, Setting print area, Print titles
- Inserting custom Header and Footer
- Inserting objects in the header and footer
- Page Setup, Setting margins, Print Preview, Print
- Enable back ground error checking
- Setting AutoCorrect Options

MS POWERPOINT

Module 11: Setting Up PowerPoint Environment:

- New, Open, Close, Save, Save As
- Typing the text, Alignment of text
- Formatting Text: Font Size, Font Style
- Font Color, Use the Bold, Italic, and Underline
- Cut, Copy, Paste, Select All, Clear text
- Find & Replace
- Working with Tabs and Indents

Module 12: Creating slides and applying themes

- Inserting new slide
- Changing layout of slides
- Duplicating slides
- Copying and pasting slide
- Applying themes to the slide layout
- Changing theme color
- Slide background
- Formatting slide background
- Using slide views

Module 13: Working with bullets and numbering

- Multilevel numbering and Bulleting
- Creating List
- Page bordering
- Page background
- Aligning text
- Text directions
- Columns option

Module 14: Working with Objects

- Adding a Footnote
- Adding Endnote

Module 4: Merging Documents

- Typing new address list
- Importing address list from Excel file
- Write and insert field
- Merging with outlook contact
- Preview Result
- Merging to envelopes
- Merging to label
- Setting rules for merges
- Finish & Merge options

Module 5: Proofing the document, Printing

- Check Spelling As You Type.
- Mark Grammar Errors As You Type.
- Setting AutoCorrect Options
- Page Setup, Setting margins
- Print Preview, Print

MS EXCEL

Module 6: Introduction to Excel, Formatting excel work book:

- Introduction to Excel interface
- Understanding rows and columns, Naming Cells
- Working with excel workbook and sheets
- New, Open, Close, Save, Save As
- Formatting Text: Font Size, Font Style
- Font Color, Use the Bold, Italic, and Underline
- Wrap text, Merge and Centre
- Currency, Accounting and other formats
- Modifying Columns, Rows & Cells

Module 7: Perform Calculations with Functions

- Creating Simple Formulas
- Setting up your own formula
- Date and Time Functions, Financial Functions
- Logical Functions, Lookup and Reference
- Functions Mathematical Functions
- Statistical Functions, Text Functions.

Module 8: Sort and Filter Data with Excel

- Sort and filtering data
- Using number filter, Text filter
- Custom filtering
- Removing filters from columns
- Conditional formatting

Module 9: Create Effective Charts to Present Data Visually

- Shapes, Clipart and Picture, Word Art, Smart Art
- Change the Order of Objects
- Inserting slide header and footer
- Inserting Text boxes
- Inserting shapes, using quick styles
- Inserting Word art
- Inserting symbols
- Inserting Chart

Module 15: Hyperlinks and Action Buttons

- Inserting Hyperlinks and Action Buttons
- Edit Hyperlinks and Action Button
- Word Art and Shapes

Module 16: Working With Movies and Sounds

- Inserting Movie From a Computer File
- Inserting Audio file
- Audio Video playback and format options
- Video options, Adjust options
- Reshaping and bordering Video

Module 17: Using SmartArt and Tables

- Working with Tables, Table Formatting
- Table Styles
- Alignment option
- Merge and split option
- Converting text to smart art

Module 18: Animation and Slide Transition

- Default Animation, Custom Animation
- Modify a Default or Custom Animation
- Reorder Animation Using Transitions
- Apply a Slide Transition, Modifying a
- Transition, Advancing to the Next Slide

Module 19: Using slide Master

- Using slide master
- Inserting layout option
- Creating custom layout
- Inserting place holders
- Formatting place holders

Module 20: Slide show option

- Start slide show
- Start show from the current slide
- Rehearse timing
- Creating custom slide show
- Check Spelling As You Type
- Setting AutoCorrect Options
- Save as video
- Save as JPEG files
- Save as PowerPoint Show file
- Print Preview, Print

Practical (20 Hrs): Based upon above content

Duration of the Course: 40 Hours

Theory : 20 Hours

Practical : 20 Hours

Eligibility : II Year BSc MPCS, MSCS, MPC, MPS Groups

Intake : 60 Students /Batch

Head of the Department : Sri P.Udaya

Course Mentors :

1. Smt B.Triveni
2. Smt N.Jayalakshmi
3. Sri D.Chaithanya Kumar

S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office
Academic Year: 2019 -2020

BATCH 1: TIME TABLE

20 Jan 2020 to 13 Feb 2020

DAY/TIME	THEORY	PRACTICAL
	8.00 – 9.00 AM	4.00 – 5.00 PM
MONDAY	BT	DCK
TUESDAY	DCK	BT
WEDNESDAY	BT	DCK
THURSDAY	DCK	BT
FRIDAY	BT	DCK
SATURDAY	DCK	BT

HOD

BT: B.Triveni
DCK: D.Chaitanya Kumar

S.G.S.ARTS COLLEGE, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN MS-Office
Academic Year: 2019 -2020

BATCH 2: TIME TABLE

20 Jan 2020 to 13 Feb 2020

DAY/TIME	THEORY	PRACTICAL
	4.00 – 5.00 PM	8.00 – 9.00 AM
MONDAY	NJ	DCK
TUESDAY	DCK	NJ
WEDNESDAY	NJ	DCK
THURSDAY	DCK	NJ
FRIDAY	NJ	DCK
SATURDAY	DCK	NJ

HOD

NJ : N.Jayalakshmi
DCK: D.Chaitanya Kumar

Department of Computer Science:

MS-Office Certificate Course Students Name List Batch 1

S.No	Name of the Student	Class	Roll No
1	Anil Kumar Mukku	II B.Sc(MPCS)-EM	18511
2	Anil Kumar Talari	II B.Sc(MPCS)-EM	18512
3	Balaji.E	II B.Sc(MPCS)-EM	18513
4	Balanjineyulu Mangala	II B.Sc(MPCS)-EM	18514
5	Balarama Krishna Yadav A	II B.Sc(MPCS)-EM	18515
6	Charan Kuamar Mandati	II B.Sc(MPCS)-EM	18516
7	Chenchaiah Gunji	II B.Sc(MPCS)-EM	18517
8	Dinesh Kumar Mangali	II B.Sc(MPCS)-EM	18518
9	Guru Sai Kumar.P.V	II B.Sc(MPCS)-EM	18519
10	Harshavardhan Reddy B	II B.Sc(MPCS)-EM	18520
11	Hemanth Kumar Paipuru	II B.Sc(MPCS)-EM	18521
12	Hemasri Kuppa	II B.Sc(MPCS)-EM	18522
13	Jagan Goud Nandimalla	II B.Sc(MPCS)-EM	18523
14	Jayalakshmi Kailasam	II B.Sc(MPCS)-EM	18524
15	Mabbu Sharif Gajuguduru	II B.Sc(MPCS)-EM	18525
16	Mahendra Dasari	II B.Sc(MPCS)-EM	18526
17	Manendra Kumar R	II B.Sc(MPCS)-EM	18527
18	Mohan Krishna Kotekallu	II B.Sc(MPCS)-EM	18528
19	Mounika Gali	II B.Sc(MPCS)-EM	18529
20	Mounika Sanambatla	II B.Sc(MPCS)-EM	18530
21	Nagadinesh Nagisetty	II B.Sc(MPCS)-EM	18531
22	Narasimhudu Kukka	II B.Sc(MPCS)-EM	18532
23	Naveen Kumar Kanchi	II B.Sc(MPCS)-EM	18533
24	Purushotham Reddy B	II B.Sc(MPCS)-EM	18534
25	Raj Kumar Giddaiahgari	II B.Sc(MPCS)-EM	18535
26	Rajesh Golla	II B.Sc(MPCS)-EM	18536
27	Rajesh Pannamareddi	II B.Sc(MPCS)-EM	18537
28	Rakesh Gajjala	II B.Sc(MPCS)-EM	18538
29	Rasool Yendhoti	II B.Sc(MPCS)-EM	18539

30	Ravi Teja Kamineni	II B.Sc(MPCS)-EM	18540
31	Saiteja Padma	II B.Sc(MPCS)-EM	18541
32	Sandhya Devarakonda	II B.Sc(MPCS)-EM	18542
33	Sandhya Putturu	II B.Sc(MPCS)-EM	18543
34	Sasidhar Bottla	II B.Sc(MPCS)-EM	18544
35	Shaik Yasmin	II B.Sc(MPCS)-EM	18546
36	Shireesha Gounori	II B.Sc(MPCS)-EM	18547
37	Siva Sankar Thati	II B.Sc(MPCS)-EM	18548
38	Sravan Kumar Chagalamarri	II B.Sc(MPCS)-EM	18549
39	Sreenivasulu Boya	II B.Sc(MPCS)-EM	18550
40	Subbarayudu Bayanaboina	II B.Sc(MPCS)-EM	18551
41	Suresh Kota	II B.Sc(MPCS)-EM	18552
42	Suresh Mangali	II B.Sc(MPCS)-EM	18553
43	Aravind Obalavaram	II B.Sc(MPC)-TM	18404
44	Chandrasekhar.B	II B.Sc(MPC)-TM	18405
45	Dathathreyudu Bandarupalli	II B.Sc(MPC)-TM	18406
46	Hareesh Derangi	II B.Sc(MPC)-TM	18407
47	Munisekhar Pamanji	II B.Sc(MPC)-TM	18408
48	Sunil Chemuru	II B.Sc(MPC)-TM	18409
49	Suresh Jangamreddy	II B.Sc(MPC)-TM	18411
50	Vishnuvardhan Chalampalem	II B.Sc(MPC)-TM	18412
51	Bharadhwaz Yannam	II B.Sc(MPC)-EM	18441
52	Giri Kondari	II B.Sc(MPC)-EM	18442
53	Govardhan Dattam	II B.Sc(MPC)-EM	18443
54	Hemadri.P	II B.Sc(MPC)-EM	18444
55	Hemantha Kumar Aare	II B.Sc(MPC)-EM	18445
56	Kalathur Vetti Nithish	II B.Sc(MPC)-EM	18446
57	Karimulla Pinjari	II B.Sc(MPC)-EM	18447
58	Mallikarjuna Avula	II B.Sc(MPC)-EM	18448
59	Mahesh Vanka	II B.Sc(MPC)-EM	18449
60	Manjunath Bhagathagalla	II B.Sc(MPC)-EM	18450

Department of Computer Science:

MS-Office Certificate Course Students Name List Batch 2

S.No	Name of the Student	Class	Roll No
1	Ajay Kumar Bandar	II B.Sc(MSCS)-EM	18565
2	Ajay Kumar Medikonda	II B.Sc(MSCS)-EM	18566
3	Adiseshu Mabbu	II B.Sc(MSCS)-EM	18567
4	Anitha Thalari	II B.Sc(MSCS)-EM	18568
5	Aravind Naidu Buraga	II B.Sc(MSCS)-EM	18569
6	Balaji Ballikura	II B.Sc(MSCS)-EM	18570
7	Bharath Kumar Paidikalva	II B.Sc(MSCS)-EM	18571
8	Bhaskar Kamatham	II B.Sc(MSCS)-EM	18572
9	Chakri Paipuri	II B.Sc(MSCS)-EM	18573
10	Charan Ulasala	II B.Sc(MSCS)-EM	18574
11	Chandini Velagambeti	II B.Sc(MSCS)-EM	18575
12	Charan Tej Markapuram	II B.Sc(MSCS)-EM	18576
13	Deepa Kadirimangalam	II B.Sc(MSCS)-EM	18577
14	Dhana Kumar Kota	II B.Sc(MSCS)-EM	18578
15	Eswar Kanakappagari	II B.Sc(MSCS)-EM	18579
16	Eswara Narasimha Vara Prasad	II B.Sc(MSCS)-EM	18580
17	Govardhan Krishna Vallepu	II B.Sc(MSCS)-EM	18581
18	Guru Swamy Velampati	II B.Sc(MSCS)-EM	18582
19	Hari Prasad Mala	II B.Sc(MSCS)-EM	18583
20	Haritha Pujari	II B.Sc(MSCS)-EM	18584
21	Hemanth Rajolu	II B.Sc(MSCS)-EM	18585
22	Indhumathi Lalapeta	II B.Sc(MSCS)-EM	18586
23	Jyothi Kuruva	II B.Sc(MSCS)-EM	18587
24	Kishore Kovuru	II B.Sc(MSCS)-EM	18588
25	Lakshmi Manne	II B.Sc(MSCS)-EM	18589
26	Leelakrishna Gadiboina	II B.Sc(MSCS)-EM	18590
27	Manoj Kumar Pallipattu	II B.Sc(MSCS)-EM	18591
28	Manoj Naramala	II B.Sc(MSCS)-EM	18592

29	Nanda Kumar Kala	II B.Sc(MSCS)-EM	18593
30	Pavan Kumar Reddy Vengareddy	II B.Sc(MSCS)-EM	18594
31	Penchala Narasimhulu Amooru	II B.Sc(MSCS)-EM	18595
32	Ramesh Gunda	II B.Sc(MSCS)-EM	18596
33	Ramu Goud Ramagouni	II B.Sc(MSCS)-EM	18597
34	Rithik.K.M	II B.Sc(MSCS)-EM	18598
35	Shireesha Kuruva	II B.Sc(MSCS)-EM	18599
36	Shyam Kumar Pokuru	II B.Sc(MSCS)-EM	18600
37	Siva Sankar Bijigam	II B.Sc(MSCS)-EM	18601
38	Sreedevi Chittoor	II B.Sc(MSCS)-EM	18602
39	Sreekanth.K	II B.Sc(MSCS)-EM	18603
40	Suresh Madigundu	II B.Sc(MSCS)-EM	18604
41	Thimmaiah Siruvella	II B.Sc(MSCS)-EM	18605
42	Vamsi Krishna Naidu Kommineni	II B.Sc(MSCS)-EM	18606
43	Venkata Amaranath Malasani	II B.Sc(MSCS)-EM	18607
44	Venkata Anirudh Reddy Venna	II B.Sc(MSCS)-EM	18608
45	Venkata Chanukya Kethineni	II B.Sc(MSCS)-EM	18609
46	Venkat Nisanth Kanupuru	II B.Sc(MSCS)-EM	18610
47	Venkateswarlu Bingi	II B.Sc(MSCS)-EM	18611
48	Venu Bharath.R.T	II B.Sc(MSCS)-EM	18612
49	Vijya Goud Ediga	II B.Sc(MSCS)-EM	18613
50	Vinay Sada	II B.Sc(MSCS)-EM	18614
51	Aravind Chakravarthy	II B.Sc(MPS)-TM	18474
52	Dastagiri Boya	II B.Sc(MPS)-TM	18476
53	Erri Swamy Getu	II B.Sc(MPS)-TM	18477
54	Marimeni Niranjan	II B.Sc(MPS)-TM	18478
55	Ramesh Boya	II B.Sc(MPS)-TM	18480
56	Yuva raj Kante	II B.Sc(MPS)-TM	18481
57	Nanda Kumar.S	II B.Sc(MPC)-EM	18451
58	Narendra Mattem	II B.Sc(MPC)-EM	18452
59	Obul Reddy Boggala	II B.Sc(MPC)-EM	18453
60	Pavan Kumar.N	II B.Sc(MPC)-EM	18454

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS ATTENDANCE LIST BATCH 1 (THEORY) 8:00 AM to 9:00 AM

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
1	18511	Anil Kumar Mukku	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P
2	18512	Anil Kumar Talari	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P
3	18513	Balaji.E	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P
4	18514	Balanjineyulu Mangala	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P
5	18515	Balarama Krishna Yadav A	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P
6	18516	Charan Kuamar Mandati	P	A	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P
7	18517	Chenchaiah Gunji	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
8	18518	Dinesh Kumar Mangali	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P
9	18519	Guru Sai Kumar.P.V	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
10	18520	Harshavardhan Reddy B	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
11	18521	Hemanth Kumar Paipuru	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P
12	18522	Hemasri Kuppa	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
13	18523	Jagan Goud Nandimalla	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
14	18524	Jayalakshmi Kailasam	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
15	18525	Mabbu Sharif Gajuguduru	P	P	P	A	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P
16	18526	Mahendra Dasari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	A
17	18527	Manendra Kumar R	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A
18	18528	Mohan Krishna Kotekallu	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
19	18529	Mounika Gali	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P
20	18530	Mounika Sanambatla	P	A	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
21	18531	Nagadinesh Nagisetty	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	A

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
22	18532	Narasimhudu Kukka	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P
23	18533	Naveen Kumar Kanchi	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
24	18534	Purushotham Reddy B	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P
25	18535	Raj Kumar Giddaiahgari	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P
26	18536	Rajesh Golla	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P
27	18537	Rajesh Pannamareddi	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P
28	18538	Rakesh Gajjala	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	A
29	18539	Rasool Yendhoti	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P
30	18540	Ravi Teja Kamineni	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P	P	P	P
31	18541	Saiteja Padma	P	P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P
32	18542	Sandhya Devarakonda	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A
33	18543	Sandhya Putturu	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
34	18544	Sasidhar Bottla	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	A	P	P	P	P
35	18546	Shaik Yasmin	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P
36	18547	Shireesha Gounori	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
37	18548	Siva Sankar Thati	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
38	18549	Sravan Kumar Chagalamarri	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P	P	P	P	A	P
39	18550	Sreenivasulu Boya	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P
40	18551	Subbarayudu Bayanaboina	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
41	18552	Suresh Kota	P	P	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P
42	18553	Suresh Mangali	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P
43	18404	Aravind Obalavaram	P	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
44	18405	Chandrasekhar.B	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	A	P	P
45	18406	Dathathreyudu Bandarupalli	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	A	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
46	18407	Hareesh Derangi	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P
47	18408	Munisekhar Pamanji	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P
48	18409	Sunil Chemuru	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
49	18411	Suresh Jangamreddy	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P
50	18412	Vishnuvardhan Chalampalem	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P
51	18441	Bharadhwaz Yannam	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P
52	18442	Giri Kondari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P
53	18443	Govardhan Dattam	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P
54	18444	Hemadri.P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
55	18445	Hemantha Kumar Aare	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P
56	18446	Kalathur Vetti Nithish	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P
57	18447	Karimulla Pinjari	P	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P
58	18448	Mallikarjuna Avula	P	P	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P
59	18449	Mahesh Vanka	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A
60	18450	Manjunath Bhagathagalla	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS ATTENDANCE LIST BATCH 1 (PRACTICAL) 4:00 PM to 5:00 PM

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
1	18511	Anil Kumar Mukku	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P
2	18512	Anil Kumar Talari	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P
3	18513	Balaji.E	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P
4	18514	Balanjineyulu Mangala	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
5	18515	Balarama Krishna Yadav A	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
6	18516	Charan Kuamar Mandati	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
7	18517	Chenchaiah Gunji	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P
8	18518	Dinesh Kumar Mangali	P	P	P	P	A	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P
9	18519	Guru Sai Kumar.P.V	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
10	18520	Harshavardhan Reddy B	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P
11	18521	Hemanth Kumar Paipuru	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A
12	18522	Hemasri Kuppa	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P
13	18523	Jagan Goud Nandimalla	P	P	P	P	P	P	A	P	P	A	P	P	P	P	P	P	P	P	P	P
14	18524	Jayalakshmi Kailasam	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P
15	18525	Mabbu Sharif Gajuguduru	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
16	18526	Mahendra Dasari	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P
17	18527	Manendra Kumar R	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A
18	18528	Mohan Krishna Kotekallu	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A
19	18529	Mounika Gali	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
20	18530	Mounika Sanambatla	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P
21	18531	Nagadinesh Nagisetty	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P
22	18532	Narasimhudu Kukka	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
23	18533	Naveen Kumar Kanchi	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P
24	18534	Purushotham Reddy B	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
25	18535	Raj Kumar Giddaiahgari	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P
26	18536	Rajesh Golla	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P
27	18537	Rajesh Pannamareddi	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P	A
28	18538	Rakesh Gajjala	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
29	18539	Rasool Yendhoti	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P
30	18540	Ravi Teja Kamineni	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	P
31	18541	Saiteja Padma	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	A	P	P	P
32	18542	Sandhya Devarakonda	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A	P	P
33	18543	Sandhya Putturu	P	P	P	P	P	P	P	A	P	P	P	P	P	P	A	P	P	P	P	P
34	18544	Sasidhar Bottla	P	P	P	A	P	A	P	P	P	P	P	P	A	P	P	P	P	P	P	P
35	18546	Shaik Yasmin	P	A	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P
36	18547	Shireesha Gounori	P	P	P	A	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
37	18548	Siva Sankar Thati	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	P	P	P	P
38	18549	Sravan Kumar Chagalamarri	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P	A	P	P
39	18550	Sreenivasulu Boya	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	A	P	P	P	P
40	18551	Subbarayudu Bayanaboina	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P
41	18552	Suresh Kota	P	P	P	P	P	P	P	P	P	A	P	A	P	P	P	P	P	P	P	P

S.No	Roll No	Name of the Student	20-01-2020	21-01-2020	22-01-2020	23-01-2020	24-01-2020	25-01-2020	27-01-2020	28-01-2020	29-01-2020	30-01-2020	31-01-2020	03-02-2020	04-02-2020	05-02-2020	06-02-2020	07-02-2020	10-02-2020	11-02-2020	12-02-2020	13-02-2020
42	18553	Suresh Mangali	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
43	18404	Aravind Obalavaram	P	P	P	P	P	P	A	P	P	P	P	A	P	P	A	P	P	P	A	P
44	18405	Chandrasekhar.B	P	P	P	P	P	P	P	P	P	P	A	P	P	A	P	P	A	P	D	A
45	18406	Dathathreyudu Bandarupalli	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P
46	18407	Hareesh Derangi	P	A	P	A	P	P	P	P	P	A	P	P	P	P	A	P	P	A	P	P
47	18408	Munisekhar Pamanji	A	P	A	P	P	P	P	P	P	P	P	A	P	P	P	P	A	P	P	P
48	18409	Sunil Chemuru	P	P	P	P	P	P	P	A	P	P	P	P	P	P	P	A	P	P	P	P
49	18411	Suresh Jangamreddy	P	P	P	P	A	A	P	P	A	P	P	P	P	P	A	P	P	P	P	P
50	18412	Vishnuvardhan Chalampalem	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P
51	18441	Bharadhwaz Yannam	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P	P	P
52	18442	Giri Kondari	P	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	A	P	P
53	18443	Govardhan Dattam	P	A	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P
54	18444	Hemadri.P	P	P	A	P	A	P	P	A	P	A	P	P	P	P	P	P	P	P	P	A
55	18445	Hemantha Kumar Aare	A	P	P	A	P	A	P	P	A	P	P	P	A	P	P	P	P	P	P	P
56	18446	Kalathur Vetti Nithish	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P
57	18447	Karimulla Pinjari	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P	A	P	P
58	18448	Mallikarjuna Avula	P	P	A	P	P	P	A	P	P	P	P	P	P	P	P	P	A	P	P	P
59	18449	Mahesh Vanka	P	A	P	P	P	P	P	A	P	P	P	P	P	P	P	P	P	P	P	P
60	18450	Manjunath Bhagathagalla	A	P	P	P	P	P	P	P	A	P	P	P	A	P	P	P	P	P	P	A

NOTICE

SGSAC/Dept.of Computer Science/HOD/Certificate Course/2019-20

All the Second Year B.Sc MPCS, MSCS, MPC(EM & TM), MPS Batch 1 and Batch 2enrolled students of Certificate Course in **MS-Office** are hereby informed that the **Final** Examination will be held on 15-02-2020.

Date: 13-02-2020

HOD

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS TEST ATTENDANCE BATCH 1

Date: 15-02-2020 **Timinig: 4:00 to 5:00 PM**

S.No	Roll No	Name of the Student	Signature
1	18511	Anil Kumar Mukku	M. Anil Kumar
2	18512	Anil Kumar Talari	T. Anil Kumar
3	18513	Balaji.E	Balaji.E
4	18514	Balanjineyulu Mangala	M. Balanjineyulu
5	18515	Balarama Krishna Yadav A	A. Balakrishnayadav
6	18516	Charan Kuamar Mandati	Charan Kumar Mandati
7	18517	Chenchaiah Gunji	G. Chenchaiah
8	18518	Dinesh Kumar Mangali	M. Dinesh Kumar
9	18519	Guru Sai Kumar.P.V	P.V. Guru Sai Kumar
10	18520	Harshavardhan Reddy B	B. Harshavardhan Reddy
11	18521	Hemanth Kumar Paipuru	P. Hemanth Kumar
12	18522	Hemasri Kuppa	K. Hemasri
13	18523	Jagan Goud Nandimalla	N. Jagan Goud
14	18524	Jayalakshmi Kailasam	K. Jayalakshmi
15	18525	Mabbu Sharif Gajuguduru	G. Mabbu Sharif
16	18526	Mahendra Dasari	Mahendra Dasari
17	18527	Manendra Kumar R	R. Manendra Kumar
18	18528	Mohan Krishna Kotekallu	Mohan Krishna Kotekallu.
19	18529	Mounika Gali	G. Mounika
20	18530	Mounika Sanambatla	S. Mounika
21	18531	Nagadinesh Nagisetty	N. Nagadinesh
22	18532	Narasimhudu Kukka	K. Narasimhudu.
23	18533	Naveen Kumar Kanchi	K. Naveen Kumar
24	18534	Purushotham Reddy B	Purushotham Reddy B.
25	18535	Raj Kumar Giddaiahgari	Raj Kumar Giddaiahgari
26	18536	Rajesh Golla	G. Rajesh
27	18537	Rajesh Pannamareddi	P. Rajesh
28	18538	Rakesh Gajjala	G. Rakesh
29	18539	Rasool Yendhoti	Y. Rasool
30	18540	Ravi Teja Kamineni	K. Ravi Teja
31	18541	Saiteja Padma	P. saiteja
32	18542	Sandhya Devarakonda	D. Sandya
33	18543	Sandhya Putturu	Sandhya putturu.
34	18544	Sasidhar Bottla	Sasidhar Bottala.
35	18546	Shaik Yasmin	Yasmin. S
36	18547	Shireesha Gounori	G. Shireesha

S.No	Roll No	Name of the Student	Signature
37	18548	Siva Sankar Thati	T. Siva Sankar
38	18549	Sravan Kumar Chagalamarri	C. Sravan Kumar.
39	18550	Sreenivasulu Boya	B. Srinivasulu
40	18551	Subbarayudu Bayanaboina	B. Subbarayudu
41	18552	Suresh Kota	Suresh Kota
42	18553	Suresh Mangali	M. Suresh
43	18404	Aravind Obalavaram	O. Aravind
44	18405	Chandrasekhar.B	B. Chandrsekhar.
45	18406	Dathathreyudu Bandarupalli	B. Dathathreyudu
46	18407	Hareesh Derangi	D. Hareesh
47	18408	Munisekhar Pamanji	Munisekhar Pamanji.
48	18409	Sunil Chemuru	Sunil Chemuru
49	18411	Suresh Jangamreddy	Suresh.J
50	18412	Vishnuvardhan Chalampalem	C. Vishnuvardhan
51	18441	Bharadhwaz Yannam	Y. Bharadhwaz
52	18442	Giri Kondari	Giri Kondari
53	18443	Govardhan Dattam	D. Govardhan
54	18444	Hemadri.P	P. Hemadri
55	18445	Hemantha Kumar Aare	A. Hemanth kumar
56	18446	Kalathur Vetti Nithish	Kalathur Vetti Nithish.
57	18447	Karimulla Pinjari	P. Karimulla
58	18448	Mallikarjuna Avula	A. Mallikarjuna
59	18449	Mahesh Vanka	V. Mahesh
60	18450	Manjunath Bhagathagalla	B. Manjunath

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
STUDENTS TEST ATTENDANCE BATCH 2

Date: 15-02-2020 **Timinig: 4:00 to 5:00 PM**

S.No	Roll No	Name of the Student	Signature
1	18565	Ajay Kumar Bandar	B. Ajay kumar
2	18566	Ajay Kumar Medikonda	M. Ajay kumar
3	18567	Adiseshu Mabbu	Adiseshu mabbu
4	18568	Anitha Thalari	T. Anitha
5	18569	Aravind Naidu Buraga	B. Aravind Naidu
6	18570	Balaji Ballikura	B. Balaji
7	18571	Bharath Kumar Paidikalva	Bharath kumar Paidikalva
8	18572	Bhaskar Kamatham	K. Bhaskar
9	18573	Chakri Paipuri	P. Chakri
10	18574	Charan Ulasala	U. Charan
11	18575	Chandini Velagambeti	V. Chandini
12	18576	Charan Tej Markapuram	M. charan Tej
13	18577	Deepa Kadirimangalam	K. Deepa
14	18578	Dhana Kumar Kota	K. Dhana kumar
15	18579	Eswar Kanakappagari	Eswar. K.
16	18580	Eswara Narasimha Vara Prasad	E. Narasimha vara prasad
17	18581	Govardhan Krishna Vallepu	V. Govardhan krishna.
18	18582	Guru Swamy Velampati	V. Guru swamy
19	18583	Hari Prasad Mala	M. Hariprasad
20	18584	Haritha Pujari	P. Haritha
21	18585	Hemanth Rajolu	Hemanth Rajolu.
22	18586	Indhumathi Lalapeta	L. Indhumathi
23	18587	Jyothi Kuruva	K. Jyothi
24	18588	Kishore Kovuru	K. Kishore
25	18589	Lakshmi Manne	M. Lakshmi
26	18590	Leelakrishna Gadiboina	G. Leelakrishna
27	18591	Manoj Kumar Pallipattu	P. Manoj kumar
28	18592	Manoj Naramala	N. Manoj
29	18593	Nanda Kumar Kala	K. Nanda kumar

S.No	Roll No	Name of the Student	Signature
30	18594	Pavan Kumar Reddy Vengareddy	Pavan kumar Reddy. V
31	18595	Penchala Narasimhulu Amooru	Penchala Narasimhulu Amooru
32	18596	Ramesh Gunda	G. Ramesh
33	18597	Ramu Goud Ramagouni	R. Ramu Goud
34	18598	Rithik.K.M	K.M. Rithik
35	18599	Shireesha Kuruva	Shireesha Kuruva.
36	18600	Shyam Kumar Pokuru	P. Shyamkumar
37	18601	Siva Sankar Bijigam	Siva Shankar Bijigam
38	18602	Sreedevi Chittoor	C. Sreedevi
39	18603	Sreekanth.K	K. Sree Kanth
40	18604	Suresh Madigundu	Suresh Madigundu.
41	18605	Thimmaiah Siruvella	S. Thimmaiah
42	18606	Vamsi Krishna Naidu Kommineni	K. Vamsi krishna Naidu
43	18607	Venkata Amaranath Malasani	M. Venkata Amaranath
44	18608	Venkata Anirudh Reddy Venna	V. Venka Anirudh Reddy
45	18609	Venkata Chanukya Kethineni	K. Venkata chanukya
46	18610	Venkat Nisanth Kanupuru	K. Venkat Nisanth
47	18611	Venkateswarlu Bingi	Bingi venkateswarlu
48	18612	Venu Bharath.R.T	Venu Bharath. R.T
49	18613	Vijya Goud Ediga	E. Vijya Goud
50	18614	Vinay Sada	S. Vinay
51	18474	Aravind Chakravarthy	Aravind Chakravarthy.
52	18476	Dastagiri Boya	B. Dastagiri
53	18477	Erri Swamy Getu	G. Erriswamy
54	18478	Marimeni Niranjan	Niranjan Marimeni
55	18480	Ramesh Boya	B. Ramesh
56	18481	Yuva raj Kante	K. Yuvaraj
57	18451	Nanda Kumar.S	S. Nanda kumar
58	18452	Narendra Mattem	M. Narendra
59	18453	Obul Reddy Boggala	B. Obul Reddy
60	18454	Pavan Kumar.N	N. Pavan kumar

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15/02/2020

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
Batch 1 & 2
ACADEMIC YEAR: 2019-20
Final Examination

Max Marks: 50 **Duration: 2 Hours**

Name of the Student : N. Nagadinesh, II. B.Sc. MPCs-EM

MULTIPLE CHOICE QUESTIONS EACH QUESTION CARRY 2 MARKS

1. Right clicking on something in word
A. Delete the object
B. Nothing the right mouse button is there for left handed people
C. Opens the short-cut list
D. Insert the picture

2. We can create the new shortcut key from
A. Options from Tools menu
B. Customize from Tools menu
C. Properties from File menu
D. Auto Correct from Toolbars menu

3. The autocorrect tool
A. Provides statistics about your document
B. Displays words with the same or similar meaning
C. Checks the grammar in the document
D. Checks for misspelled words as you type and correct them as defined

4. Which key or key combination will move the insertion point to the bottom of your document?
A. Ctrl+End
B. PageDown
C. Ctrl+Home
D. End

5. To display hyperlink fields in a Word document, you can press the
A. Ctrl+Shift+F9 key
B. Shift + F9
C. Ctrl+Alt key
D. None of above

6. To center worksheet titles across a range of cells, you must
a. Select the cells containing the title text plus the range over which the title text is to be centered
b. Widen the columns
c. Select the cells containing the title text plus the range over which the title text is to be enfettered
d. Format the cells with the comma style

7. How do you delete a column?

a. Select the column heading you want to delete and select the Delete Row button on the standard toolbar
b. Select the column heading you want to delete and select Insert Delete from the menu
c. Select the row heading you want to delete and select Edit>Delete from the menu
d. Right click the column heading you want to delete and select delete from the shortcut menu

8. How can you find specific information in a list?
a. Select Tools > Finder from the menu
b. Click the Find button on the standard toolbar
c. Select Insert > Find from the menu
d. Select Data > Form from the menu to open the Data Form dialog box and click the Criteria button

9. When integrating word and excel, word is usually the
a. Server
b. Destination
c. Client
d. Both b and c

10. When a label is too long to fit within a worksheet cell, you typically must
a. Shorten the label
b. Increase the column width
c. Decrease the column width
d. Adjust the row height

11. When you have completely finished working with a document you should
A. Close it
B. Save it
C. Type it
D. Edit it

12. To cut or copy text you must first
A. Click the remove/duplicate button
B. First highlight the text and click the Copy or Cut command from edit menu
C. Click on File menu
D. None of above

13. How many items can you copy to the Office Clipboard
A. 24
B. 20
C. 12
D. 10

14. The command cross reference is available in
A. View menu
B. Format menu
C. Insert menu
D. Tools menu

15. We can send the word documents to
A. Microsoft PowerPoint
B. Microsoft Excel

c. Microsoft Access
d. All of above

16. Which of the following are types of sound files?
a. LOG files
b. DAT files
c. WAV files
d. DRV files

17. Which of the following can you use to add times to the slides in a presentation? a. Microsoft graph
b. Microsoft Table
c. Microsoft Excel
d. Microsoft Word

18. Which view in PowerPoint can be used to enter speaker comments? a. Normal
b. Slide show
c. Slide sorter
d. Notes page view

19. The best way to insert a new slide in a presentation is to use the
a. Normal view
b. Special view
c. Slide show view
d. Slide sorter view

20. Which is the best view for getting your thoughts for a presentation out on the computer?
a. Outline view
b. Notes page view
c. Slide sorter view
d. Slide view

21. The default page orientation of word document is
A. Landscape
B. Portrait
C. Long sides
D. Double long side

22. You can print the documents
A. Print icon from Standard toolbar
B. Print command from File menu
C. Ctrl+P
D. All of the above

23. The thesaurus tool
A. Finds repeated words in document
B. Check for synonyms and antonyms
C. Identifies words with capitalization problems
D. All of the above

24. To view smaller text on the screen you can
A. Decrease the font points
B. Decrease the zoom percentage
C. Increase the view percentage
D. Increase the editing percentage

25. Which of the following is not option for changing the case of the text?
A. Indent case
B. Sentence case
C. Toggle case
D. Lower case

******All the Best******

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
MARKS SHEET BATCH 1

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
1	18511	Anil Kumar Mukku	46	four Six
2	18512	Anil Kumar Talari	48	four Eight
3	18513	Balaji.E	42	four Two
4	18514	Balanjineyulu Mangala	44	four four
5	18515	Balarama Krishna Yadav A	44	four four
6	18516	Charan Kuamar Mandati	48	four Eight
7	18517	Chenchaiah Gunji	46	four Six
8	18518	Dinesh Kumar Mangali	44	four four
9	18519	Guru Sai Kumar.P.V	42	four Two
10	18520	Harshavardhan Reddy B	42	four Two
11	18521	Hemanth Kumar Paipuru	44	four four
12	18522	Hemasri Kuppa	46	four Six
13	18523	Jagan Goud Nandimalla	48	four Eight
14	18524	Jayalakshmi Kailasam	50	fifty Only
15	18525	Mabbu Sharif Gajuguduru	42	four Two
16	18526	Mahendra Dasari	44	four four
17	18527	Manendra Kumar R	46	four Six
18	18528	Mohan Krishna Kotekallu	48	four Eight
19	18529	Mounika Gali	42	four Two
20	18530	Mounika Sanambatla	44	four four
21	18531	Nagadinesh Nagisetty	48	four Eight
22	18532	Narasimhudu Kukka	46	four six
23	18533	Naveen Kumar Kanchi	44	four four
24	18534	Purushotham Reddy B	42	four Two
25	18535	Raj Kumar Giddaiahgari	44	four four
26	18536	Rajesh Golla	46	four Six
27	18537	Rajesh Pannamareddi	48	four Eight
28	18538	Rakesh Gajjala	44	four four
29	18539	Rasool Yendhoti	48	four Eight
30	18540	Ravi Teja Kamineni	42	four Two
31	18541	Saiteja Padma	42	four Two
32	18542	Sandhya Devarakonda	44	four four
33	18543	Sandhya Putturu	46	four Six
34	18544	Sasidhar Bottla	48	four Eight
35	18546	Shaik Yasmin	42	four Two
36	18547	Shireesha Gounori	44	four four
37	18548	Siva Sankar Thati	46	four Six
38	18549	Sravan Kumar Chagalamarri	48	four Eight
39	18550	Sreenivasulu Boya	42	four Two
40	18551	Subbarayudu Bayanaboina	44	four four

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
41	18552	Suresh Kota	44	four four
42	18553	Suresh Mangali	46	four six
43	18404	Aravind Obalavaram	42	four Two
44	18405	Chandrasekhar.B	44	four four
45	18406	Dathathreyudu Bandarupalli	48	four Eight
46	18407	Hareesh Derangi	46	four six
47	18408	Munisekhar Pamanji	46	four six
48	18409	Sunil Chemuru	42	four Two
49	18411	Suresh Jangamreddy	44	four four
50	18412	Vishnuvardhan Chalampalem	48	four Eight
51	18441	Bharadhwaz Yannam	44	four four
52	18442	Giri Kondari	42	four Two
53	18443	Govardhan Dattam	44	four four
54	18444	Hemadri.P	46	four six
55	18445	Hemantha Kumar Aare	48	four Eight
56	18446	Kalathur Vetti Nithish	46	four six
57	18447	Karimulla Pinjari	48	four Eight
58	18448	Mallikarjuna Avula	44	four four
59	18449	Mahesh Vanka	46	four six
60	18450	Manjunath Bhagathagalla	46	four six

SGS ARTS COLLEGE, TTD, TIRUPATI
DEPARTMENT OF COMPUTER SCIENCE
CERTIFICATE COURSE IN "MS-Office"
MARKS SHEET BATCH 2

S.No	Roll No	Name of the Student	Marks in Figure	Marks in Words
1	18565	Ajay Kumar Bandar	46	four six
2	18566	Ajay Kumar Medikonda	48	four Eight
3	18567	Adiseshu Mabbu	44	four four
4	18568	Anitha Thalari	42	four Two
5	18569	Aravind Naidu Buraga	50	fifty Only
6	18570	Balaji Ballikura	46	four six
7	18571	Bharath Kumar Paidikalva	42	four Two
8	18572	Bhaskar Kamatham	48	four Eight
9	18573	Chakri Paipuri	44	four four
10	18574	Charan Ulasala	46	four six
11	18575	Chandini Velagambeti	44	four four
12	18576	Charan Tej Markapuram	48	four Eight
13	18577	Deepa Kadirimangalam	46	four six
14	18578	Dhana Kumar Kota	44	four four
15	18579	Eswar Kanakappagari	46	four six
16	18580	Eswara Narasimha Vara Prasad	48	four Eight
17	18581	Govardhan Krishna Vallepu	40	forty Only
18	18582	Guru Swamy Velampati	42	four Two
19	18583	Hari Prasad Mala	44	four four
20	18584	Haritha Pujari	46	four six
21	18585	Hemanth Rajolu	48	four Eight
22	18586	Indhumathi Lalapeta	42	four Two
23	18587	Jyothi Kuruva	42	four Two
24	18588	Kishore Kovuru	44	four four
25	18589	Lakshmi Manne	46	four six
26	18590	Leelakrishna Gadiboina	46	four six
27	18591	Manoj Kumar Pallipattu	44	four four
28	18592	Manoj Naramala	48	four Eight
29	18593	Nanda Kumar Kala	44	four four
30	18594	Pavan Kumar Reddy Vengareddy	48	four Eight
31	18595	Penchala Narasimhulu Amooru	46	four six

32	18596	Ramesh Gunda	42	four Two
33	18597	Ramu Goud Ramagouni	44	four four
34	18598	Rithik.K.M	46	four six
35	18599	Shireesha Kuruva	44	four four
36	18600	Shyam Kumar Pokuru	46	four six
37	18601	Siva Sankar Bijigam	48	four Eight
38	18602	Sreedevi Chittoor	46	four Six
39	18603	Sreekanth.K	44	four four
40	18604	Suresh Madigundu	42	four Two
41	18605	Thimmaiah Siruvella	42	four Two
42	18606	Vamsi Krishna Naidu Kommineni	44	four four
43	18607	Venkata Amaranath Malasani	46	four six
44	18608	Venkata Anirudh Reddy Venna	48	four Eight
45	18609	Venkata Chanukya Kethineni	42	four Two
46	18610	Venkat Nisanth Kanupuru	44	four four
47	18611	Venkateswarlu Bingi	46	four Six
48	18612	Venu Bharath.R.T	48	four Eight
49	18613	Vijya Goud Ediga	44	four four
50	18614	Vinay Sada	48	four Eight
51	18474	Aravind Chakravarthy	42	four Two
52	18476	Dastagiri Boya	44	four four
53	18477	Erri Swamy Getu	46	four six
54	18478	Marimeni Niranjan	48	four Eight
55	18480	Ramesh Boya	44	four four
56	18481	Yuva raj Kante	48	four Eight
57	18451	Nanda Kumar.S	46	four Six
58	18452	Narendra Mattem	44	four four
59	18453	Obul Reddy Boggala	42	four Two
60	18454	Pavan Kumar.N	44	four four

S G S ARTS COLLEGE, TIRUPATI

DEPARTMENT OF COMPUTER SCIENCE

CERTIFICATE COURSE IN MS-OFFICE

Academic year : 2019 - 20

REPORT OF THE COURSE:

1. Name of the Course : MS-OFFICE
2. Name of the Course Coordinator : P Udaya
3. No of Students Admitted : 100 (60 per batch)
4. Faculty Members : 03

S.NO	Name of the Teacher	Paper	No of Hours (Classes) Engaged	
			Theory	Problem Solving / Practicles
1.	Smt. B.Triveni	I	6	6
2.	Smt. N.Jayalakshmi	I	6	6
3.	Sri D.Chaithanya Kumar	I	8	8

5. Number of Students Appeared for Exam : Batch I : 60
   Batch II : 60
6. Number of Students Passed : Batch I : 60
   Batch II : 60
7. Percentage of Results : 100%
8. Date of Theory Examination : 15-02-2020
9. Name of the Examiner : B.Triveni, N.Jayalakshmi
10. Date of Declaration of Results : 16-02-2020

HoD

Signature of Course Coordinator

PHOTO GALLERY

Distribution of Certificates to B.S.c MPCs and MSCs Second year students for completing the Certificate Course on the Topic "MS Office " by Guest of Honour Dr. N. Venugopal Reddy, H.O.D, Physics, S.V. Arts College, Tirupati.

Guest of Honor Dr. N. Venugopal Reddy & Staff of Computer Science Department

STUDENT NAME : P. SANDHYA II BSc MPCs EM

STUDENT NAME : T.ANIL KUMAR II BSc MPCs EM

Student Name: P.Manoj Kumar, II B.S.c-MSCs

Student Name: C.Sreedevi, II B.S.c-MSCs

Certificates of B.S.c MPCs and MSCs Second year students for completing the Certificate Course on the Topic "MS Office" .


OM NAMO VENKATESAYA
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S ARTS COLLEGE:TIRUPATI

CERTIFICATE OF COURSE COMPLETION

This is to certify that Mr/Kum. N.Nagadinesh , II B.S.c. MPCS-EM

group, has attended the certificate course in MS Office conducted by the department of Computer Science from 20-01-2020 to 13-02-2020 He/She has successfully completed the course and fulfilled all the requirements for the award of the certificate in MS Office.

Course Coordinator

PRINCIPAL
PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.


OM NAMO VENKATESAYA
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S ARTS COLLEGE:TIRUPATI

CERTIFICATE OF COURSE COMPLETION

This is to certify that Mr/Kum. P.Sardhya , II B.S.c. MPCs-EM

group, has attended the certificate course in MS Office conducted by the department of Computer Science from 20-01-2020 to 13-02-2020 He/She has successfully completed the course and fulfilled all the requirements for the award of the certificate in MS Office.

Course Coordinator

PRINCIPAL
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S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.


OM NAMO VENKATESAYA
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S ARTS COLLEGE:TIRUPATI

CERTIFICATE OF COURSE COMPLETION

This is to certify that Mr/Kum. R.Hemanth , II B.S.c. MSCs-EM

group, has attended the certificate course in MS Office conducted by the department of Computer Science from 20-01-2020 to 13-02-2020 He/She has successfully completed the course and fulfilled all the requirements for the award of the certificate in MS Office.

Course Coordinator

PRINCIPAL
PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.


OM NAMO VENKATESAYA
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S ARTS COLLEGE:TIRUPATI

CERTIFICATE OF COURSE COMPLETION

This is to certify that Mr/Kum. C.Sravan Kumar , II B.S.c. MPCs-EM

group, has attended the certificate course in MS Office conducted by the department of Computer Science from 20-01-2020 to 13-02-2020 He/She has successfully completed the course and fulfilled all the requirements for the award of the certificate in MS Office.

Course Coordinator

PRINCIPAL
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T.T.D., TIRUPATI.


OM NAMO VENKATESAYA
TIRUMALA TIRUPATI DEVASTHANAMS
S.G.S ARTS COLLEGE:TIRUPATI

CERTIFICATE OF COURSE COMPLETION

This is to certify that Mr/Kum.B.Aravind Naidu, II B.S.c. MSCS-EM

group, has attended the certificate course in MS Office conducted by the department of Computer Science from 20-01-2020 to 13-02-2020 He/She has successfully completed the course and fulfilled all the requirements for the award of the certificate in MS Office.

Course Coordinator

PRINCIPAL
PRINCIPAL
S.G.S. ARTS COLLEGE
T.T.D., TIRUPATI.

OM NAMO VENKATESAYA"

TIRUMALA TIRUPATI DEVASTHANAMS, TIRUPATI

S.G.S ARTS COLLEGE, TIRUCHANOOR ROAD,TIRUPATI

DEPARTMENT OF CHEMISTRY

Sri D.PARAMESWARA
M.Sc.,Mphil.,(Ph.D)
HOD OF CHEMISTRY

Phone No: 9490728655
Mail.ID: parameswara1964@gmail.com

Department of Chemistry

BATCH :: 2016-19		GROUP::MPC		
SEMESTER WISE PROGRESSION OF STUDENT RESULTS				
SEMESTER	**TOTAL NUMBER OF STUDENTS IN THE CLASS**	**TOTAL NUMBER OF SLOW LEARNERS IN THE CLASS**	**TOTAL NUMBER OF ADVANCED LEARNERS IN THE CLASS**	**TOTAL NUMBER OF AVERAGE LEARNERS IN THE CLASS**
II SEMESTER (Based on the results of I semester (2016-17)	44	23	2	19
III SEMESTER (Based on the results of II semester (2016-17)	44	15	4	25
IV SEMESTER (Based on the results of III semester (2017-18)	39	21	3	15
V SEMESTER (Based on the results of IV semester (2017-18)	35	14	4	17
VI SEMESTER (Based on the results of V semester (2018-19)	31	8	5	18

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

BATCH :: 2016-19	GROUP: CBZ			
SEMESTER WISE PROGRESSION OF STUDENT RESULTS				
SEMESTER	TOTAL NUMBER OF STUDENTS IN THE CLASS	TOTAL NUMBER OF SLOW LEARNERS IN THE CLASS	TOTAL NUMBER OF ADVANCED LEARNERS IN THE CLASS	TOTAL NUMBER OF AVERAGE LEARNERS IN THE CLASS
II SEMESTER (Based on the results of I semester (2016-17)	24	22	NIL	2
III SEMESTER (Based on the results of II semester (2016-17)	27	4	7	16
IV SEMESTER (Based on the results of III semester (2017-18)	25	11	NIL	14
V SEMESTER (Based on the results of IV semester (2017-18)	26	12	1	13
VI SEMESTER (Based on the results of V semester (2018-19)	24	2	3	19

HEAD
DEPARTMENT OF CHEMISTRY
S. G. S. ARTS COLLEGE
TIRUPATHI

III. Other programme

INDEX		
S.No	**Content**	**Page number**
1	Orientation programme	
2	Bridge course	
3	Endowment prizes	
4	Certificate courses	

1. Student Orientation Programme

Objectives:

To help new students to adjust and feel comfortable in new environment.

- To inculcate the culture of the college
- To make them aware of the curriculum
- To explain discipline regulations of the college
- To make the students to get familiarization with the college, departments
- To make the students to know extra-curricular activities in the college

2019-2020

Brief Report of Programme

Name of the Programme	Student Induction Programme for B.Sc. First Year
Date of the Programme	19-06-2019
Name of the Organizing Committee	Admission Committee
Objectives for organising the Programme	To create awareness among the students about academic curriculum, College discipline and activities
Nature of Programme	Lectures and Q & A
Brief Report	Principal **Dr. C. Ramesh** delivered a speech on College activities and discipline
No. of participants	102

S.G.S. ARTS COLLEGE Co-ordinator NAAC TIRUPATI

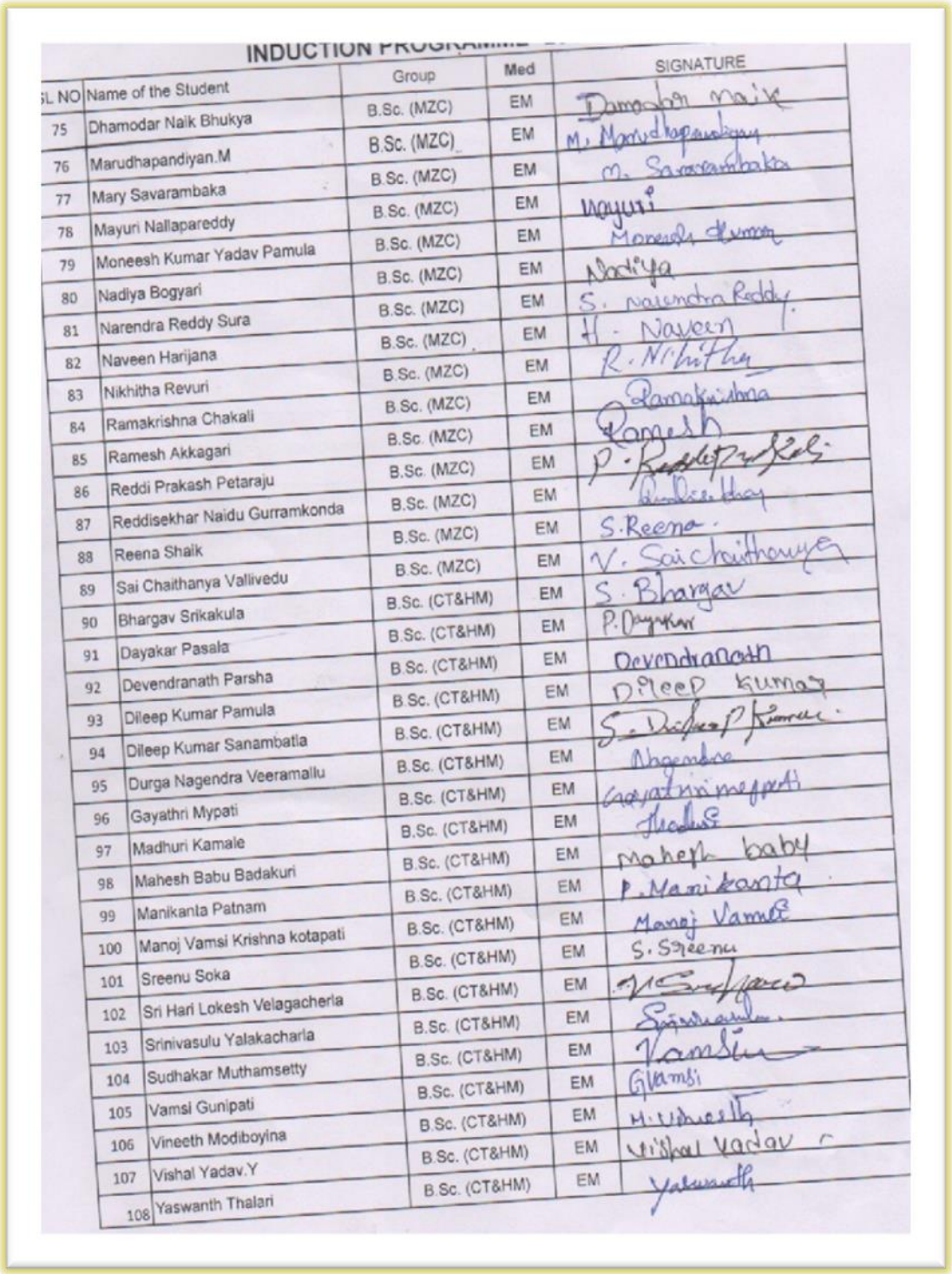
PRINCIPAL
S.G.S. ARTS COLLEGE
TIRUPATI

INDUCTION PROGRAMME -2019-20

SL NO	Name of the Student	Group	Med	SIGNATURE
1	Balasubramanyam Palli	B.Sc. (MPC)	EM	P. Balasubramanyam
2	Balasundaram Parampogu	B.Sc. (MPC)	EM	P. Bala Sundaram
3	Chandana Priya Esukapalli	B.Sc. (MPC)	EM	E. Chandana Priya
4	Chandu Prakash Avulapalli	B.Sc. (MPC)	EM	A. Chandu Prakash
5	Chinthalapattadai Chandran Uday Ku	B.Sc. (MPC)	EM	C. Uday Kumar
6	Tharun Kumar Mallay	B.Sc. (MPC)	EM	Tharun Kumar
7	Varaprasad Boya	B.Sc. (MPC)	EM	Varaprasad
8	Varaprasad Cherukuri	B.Sc. (MPC)	EM	C. Varaprasad
9	Venkata Arjun Amirisetty	B.Sc. (MPC)	EM	A. Venkata Arjun
10	Venkatesh Yanati	B.Sc. (MPC)	EM	Y. Venkatesh
11	Vishnu Prasad Kammapatti	B.Sc. (MPC)	EM	K. Vishnu Prasad
12	Vishnu Vardhan Talari	B.Sc. (MPC)	EM	Vishnu Vardan
13	Akhil Kothapalli	B.Sc. (MPS)	EM	K. Akhil
14	Anil Kumar Kedri	B.Sc. (MPS)	EM	K. Anil Kumar
15	Ashok Mundla	B.Sc. (MPS)	EM	Ashok.
16	Badri Penchala	B.Sc. (MPS)	EM	P. Badri
17	Bharath KumarAmilineni	B.Sc. (MPS)	EM	Bharath Kumar
18	Chethan.S	B.Sc. (MPS)	EM	S Chethan
19	Teja Muneswar Vulasala	B.Sc. (MPS)	EM	Teja muneswar vulasala
20	Varma Teja Mondla	B.Sc. (MPS)	EM	M. Varma Teja
21	Venu Gopal Varikunta	B.Sc. (MPS)	EM	Venu gopal
22	Vignesh Obalavaram	B.Sc. (MPS)	EM	vignesh
23	Adimuni Peddaiah Kattubadi	B.Sc. (MPCS)	EM	K. Adimuni peddaiah
24	Govardhan Manchala	B.Sc. (MPCS)	EM	M. Govardhan
25	Harathi Kalyani	B.Sc. (MPCS)	EM	Harathi kalyani
26	Hari Krishna Singamala	B.Sc. (MPCS)	EM	S. Harikrishna
27	Hema Neerugatti	B.Sc. (MPCS)	EM	N. Hema
28	Hemalatha Aruru	B.Sc. (MPCS)	EM	Hemalatha Aruru
29	Pavan Kalyan Junuguru	B.Sc. (MPCS)	EM	J. Pavan Kalyan
30	Pavan Kumar Pulimadugu	B.Sc. (MPCS)	EM	P. Pavan Kumar
31	Pavan Kumar Reddy Katam	B.Sc. (MPCS)	EM	P. pavan kumar
32	Pedda Swamy Dudekula	B.Sc. (MPCS)	EM	Pedda Swamy
33	Praveen Kolle	B.Sc. (MPCS)	EM	M. Praveen
34	Veerendrachari Kammari	B.Sc. (MPCS)	EM	K. Veerendrachari
35	Vemana Mangali	B.Sc. (MPCS)	EM	M Vemana
36	Venkateswara Rao Dumpala	B.Sc. (MPCS)	EM	Venkateswara Rao
37	Vidhyadhar Madapuri	B.Sc. (MPCS)	EM	Vidhyadhar

INDUCTION PROGRAMME -2019-20

SL NO	Name of the Student	Group	Med	SIGNATURE
38	Vijayarathnam Basavamgari	B.Sc. (MPCS)	EM	Vijayrathnam
39	Mahesh Naik Bukke	B.Sc. (MSCS)	EM	B Mahesh Naik
40	Mallikarjuna Ganta	B.Sc. (MSCS)	EM	mahesh naik
41	Manoj Kumar Sreeram	B.Sc. (MSCS)	EM	Manojkumar
42	Mohan Vatambeti	B.Sc. (MSCS)	EM	Mohan
43	Mohan Krishna Galam	B.Sc. (MSCS)	EM	G Mohan krishna
44	Sireesha Narukuru	B.Sc. (MSCS)	EM	N. Sreesha
45	Siva Jyothi Bodicherla	B.Sc. (MSCS)	EM	Siva Jyothi. B.
46	Siva Shankar Reddy Gundre	B.Sc. (MSCS)	EM	G. Siva Shankar Reddy
47	Somasekhar Kukati	B.Sc. (MSCS)	EM	kSomasekhar
48	Srinivasulu Punamalli	B.Sc. (MSCS)	EM	P. Srinivasulu
49	Adi Lakshmi Bukya	B.Sc. (CBZ)	TM	B Adi Lakshmi
50	Balachandra Devula	B.Sc. (CBZ)	TM	D. Bala Chandra
51	Chandra Sekhar Chowdam	B.Sc. (CBZ)	TM	Balachandra
52	Deepa Marimani	B.Sc. (CBZ)	TM	Deepa M
53	Dinesh Naik Bukke	B.Sc. (CBZ)	TM	Dinesh Naik Bukke
54	Sandeep Kumar Pachigalla	B.Sc. (CBZ)	TM	P. Sandeep Kumar
55	Santhosh Kumar Pasala	B.Sc. (CBZ)	TM	P. Santhosh Kumar
56	Sasi Kumar K	B.Sc. (CBZ)	TM	K. Sasikumar
57	Bhavani Jeepalem	B.Sc. (BBC)	EM	Bhavani
58	Chandini Adimulam	B.Sc. (BBC)	EM	Chandini
59	Devendra Mala	B.Sc. (BBC)	EM	M. Devendra
60	Devi Chitta	B.Sc. (BBC)	EM	C. Devi
61	Goutham Krishna Bandila	B.Sc. (BBC)	EM	Goutham krishna
62	Mahesh Reddy Thippathatigari	B.Sc. (BBC)	EM	Mahesh Reddy T.
63	Mallikarjuna Karenna	B.Sc. (BBC)	EM	K. Mallikarjuna
64	Manjunath Naik Bhukya	B.Sc. (BBC)	EM	Manjunath Naik
65	Poojitha Anthati	B.Sc. (BBC)	EM	A Poojitha
66	Poornima Nelavoy	B.Sc. (BBC)	EM	N. Poornima
67	Pradeep Kumar Pemara	B.Sc. (BBC)	EM	Pradeep kumar
68	Prakash Naik Mude	B.Sc. (BBC)	EM	Prakash Naik
69	Siva Kumar Muga	B.Sc. (BBC)	EM	Sivakumar
70	Sowmya Gangalapudi	B.Sc. (BBC)	EM	G. Sowmya
71	Sreelatha Thamarla	B.Sc. (BBC)	EM	T. Sreelatha
72	Brahma Reddy Maruprolu	B.Sc. (MZC)	EM	Brahma reddy
73	Chandu Chinthagunta	B.Sc. (MZC)	EM	C. Chandu
74	Charan Teja Bathalapalli	B.Sc. (MZC)	EM	B. Charan Teja



INDUCTION PROGRAMME

SL NO	Name of the Student	Group	Med	SIGNATURE
75	Dhamodar Naik Bhukya	B.Sc. (MZC)	EM	Damodhar naik
76	Marudhapandiyan.M	B.Sc. (MZC)	EM	M. Marudhapandiyan
77	Mary Savarambaka	B.Sc. (MZC)	EM	M. Savarambaka
78	Mayuri Nallapareddy	B.Sc. (MZC)	EM	Mayuri
79	Moneesh Kumar Yadav Pamula	B.Sc. (MZC)	EM	Moneesh Kumar
80	Nadiya Bogyari	B.Sc. (MZC)	EM	Nadiya
81	Narendra Reddy Sura	B.Sc. (MZC)	EM	S. Narendra Reddy
82	Naveen Harijana	B.Sc. (MZC)	EM	H. Naveen
83	Nikhitha Revuri	B.Sc. (MZC)	EM	R. Nikhitha
84	Ramakrishna Chakali	B.Sc. (MZC)	EM	Ramakrishna
85	Ramesh Akkagari	B.Sc. (MZC)	EM	Ramesh
86	Reddi Prakash Petaraju	B.Sc. (MZC)	EM	P. Reddi Prakash
87	Reddisekhar Naidu Gurramkonda	B.Sc. (MZC)	EM	Reddisekhar
88	Reena Shaik	B.Sc. (MZC)	EM	S. Reena
89	Sai Chaithanya Vallivedu	B.Sc. (MZC)	EM	V. Sai chaithanya
90	Bhargav Srikakula	B.Sc. (CT&HM)	EM	S. Bhargav
91	Dayakar Pasala	B.Sc. (CT&HM)	EM	P. Dayakar
92	Devendranath Parsha	B.Sc. (CT&HM)	EM	Devendranath
93	Dileep Kumar Pamula	B.Sc. (CT&HM)	EM	Dileep Kumar
94	Dileep Kumar Sanambatla	B.Sc. (CT&HM)	EM	S. Dileep Kumar
95	Durga Nagendra Veeramallu	B.Sc. (CT&HM)	EM	Nagendra
96	Gayathri Mypati	B.Sc. (CT&HM)	EM	Gayathri mypati
97	Madhuri Kamale	B.Sc. (CT&HM)	EM	Madhuri
98	Mahesh Babu Badakuri	B.Sc. (CT&HM)	EM	Mahesh baby
99	Manikanta Patnam	B.Sc. (CT&HM)	EM	P. Manikanta
100	Manoj Vamsi Krishna kotapati	B.Sc. (CT&HM)	EM	Manoj Vamsi
101	Sreenu Soka	B.Sc. (CT&HM)	EM	S. Sreenu
102	Sri Hari Lokesh Velagacherla	B.Sc. (CT&HM)	EM	[illegible]
103	Srinivasulu Yalakacharla	B.Sc. (CT&HM)	EM	Srinivasulu
104	Sudhakar Muthamsetty	B.Sc. (CT&HM)	EM	Vamsi
105	Vamsi Gunipati	B.Sc. (CT&HM)	EM	Gvamsi
106	Vineeth Modiboyina	B.Sc. (CT&HM)	EM	M. Vineeth
107	Vishal Yadav.Y	B.Sc. (CT&HM)	EM	Vishal Yadav
108	Yaswanth Thalari	B.Sc. (CT&HM)	EM	Yaswanth


S.G.S.ARTS COLLEGE
TTD, Tirupati
THE MANAGEMENT, PRINCIPAL, STAFF & STUDE
CORDIALLY INVITE YOU TO THE
ORIENTATION PROGRAMM
for the First Year BA, Bcom, Bsc St
9th August 2019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ORIENTATION PROGRAMME
for the First Year BA, Bcom, Bsc Students
9th August 2019

Dr.C.Ramesh, Principal delivering key- note address during the orientation programme conducted for first year on 09-08-2019.

S.G.S. ARTS COLLEGE, TIRUCHANOOR ROAD, TIRUPATI

Student supporting programs / Events conducted

Name of the department. Statistics

Academic year.2019-20

S. No	Name of the Program/ Event	Date	Class	No. of Students Participated	Remark
1	ORIENTATION PROGRAMME	06/07/2019	B.Sc:MSCs MPS BA: MES,ASCA	43	-

S.G.S. ARTS COLLEGE, TTD, TIRUPATI

Department of STATISTICS

EVENT ORGANISED REPORT: 2019-20

1	Name of the Department	Statistics
2	Name of Event Organized	Orientation programme for I year students
3	Title of Event Organized	Awareness program on Statistics in real life.
4	Name of Co-ordinator of Event	Dr.D.Chandra kesavulu Naidu
5	Resource Person	Dr P.Maheswari, Assistant Professor in Statistics , SVA Degree College form Women, Srikalahasthi.
6	Class of the Participants	B.Sc:MSCs MPS BA: MES,ASCA
7	No. of Participants	43
8	Objective of the Event	➢ To create awareness on Statistics in real life. Career opportunities and higher education.
9	Outcome of the Event	➢ Appreciated use of Statistics in real life ➢ Got to know Career Opportunities for Statistics graduates and Higher education Opportunities.

D. CHANDRA KESAVULU NAIDU,
M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

HOD

Orientation Programme

06/07/2019

The department of Statistics has conducted Orientation programme for the I year Students of BA, B.Sc (MSCs, MPS, MES & ASCA).

Objective: To bring awareness among the student about Statistics in real life applications, Career opportunities and higher studies.

Resource Person: Dr. P. Maheswari,
Asst. Professor in Statistics
S.V.A. Govt. Degree college for women,
Srikalahasti.

Keypoints: 1. Statistics - in Real life
2. Career opportunities after UG program
3. Institutions offering PG program in Statistics

The Head of dept Dr. D. Chandra Kesavulu Naidu and faculty members addressed students to enlighten about rules & regulations, discipline to be followed in the college for better future.

D. CHANDRA KESAVULU NAIDU, M.Sc., M.Phil., Ph.D
LECTURER & HEAD
DEPT. OF STATISTICS
S.G.S. ARTS COLLEGE
TIRUPATI

Faculty: 1) Dr. K. Vijaya Kumar – K Vijaya
2) Dr. K. Venkatesh
3) Dr. M. Vasudeva Reddy – M.V.S

1	P. Chakri	P. Chakri
2	K. Bhaskar	K. Bhaskar
3	U. Charan	U. Charan
4	K. Deepa	K. Deepa
5	B. Aravindh Naik	B. Aravindh
6	M. Ajay Naidu	M. Ajay Naidu
7	M. Suresh	M. Suresh
8	K. Venkata Chanikya	K. Venkata Chanikya
9	S. Vinay	S. Vinay

10	C. Vijay Goud.	44	M. Meghana.
11	P. Haritha		P. Haritha
12	R. Hemanth.		R. Hemanth
13	K. Jyothi		Jyothi K
14	M. Lakshmi		M. Lakshmi.
15	N. Manoj		N. manoj
16	G. Ramesh		G.
17	K. Sireesha		K. Sireesha
18	C. Sridevi		C.
19	B. Vijay Vardhan		B. Vijay.
20	B. Viswanath		B. Viswanath
21	J. Chandrudu		
22	S. Gurunath		S. Gurunath
23	B. Vasu		B.
24	P. Rakesh		P. Rakesh
25	G. Surya		P. Rakesh
26	M. Hemanth		M. Hemanth
27	P. Lokesh		P. Lokesh
28	D. Raja		D.
29	S. Jagadeeswari		S. Jagadeeswari
30	J. Vinay		J. Vinay.
31	R. Bhuvanesh		R. Bhuvanesh
32	D. Supriya		D. Supriya
33	M. Anusha.		M. Anusha
34	T. Sunil		T. Sunil.
35	C. Lokesh		C. Lokesh
36	G. Suresh		G. Suresh
37	B. Sumalatha		B. Sumalatha
38	M. Anusha.		M. Anusha
39	K. Rupesh		K. Rupesh
40	D. Allipeera.		D. Allipeera
41	C. Sandeep		
42	U. Sai kumar		U. Sai kumar
43	U. Lokanath		U. Lokanath

Orientation Programme on to bring awareness among the students about statistics in real life applications, career opportunities and higher Studies by DRP/Mahesewari.

All three year statistics Students participated in Orientation Programme.

STUDENT UNION OFFICE BEARERS:

Phone: (0877) 2264599 **E-mail:**sgsartscollegettd@gmail.com **Web:** sgsac.edu.in

TIRUMALA TIRUPATI DEVASTHANAMS

S.G.S. ARTS COLLEGE

Tiruchanur Road, TIRUPATI-517501

2018-2019

Sno	Class	Name of the Representative	Roll No
1	III B.A-HPT-TM	Mr. S.Yerramala(Union Chairman)	16284
2	III B.Sc-BBC-EM	Mr. S.Snehalatha(Union Secretary)	16736
3	III B.Sc-MSCs-EM	Mr. P.Gurusai(Maths & Science Association Secretary)	16576
4	II B.Sc-BBC-EM	Kum. T.Jahanavi(Maths & Science Association Joint Secretary)	17672
5	III B.A-HEP-TM	Mr. B.Pavan Kumar(Arts Association Secretary)	16358
6	II B.A - HEP-EM	Mr. K.Rajendra Prasad(Arts Association Joint Secretary)	17307
7	III B.Com-CA-EM	Mr. N.Purushottam(Commerce Association Secretary)	0217008579
8	II B.Com-CA-EM	Kum. K.Deepa(Commerce Association Joint Secretary)	17167
9	III B.Sc-MZC-EM	Kum. P.Shilpa(Language Association Secretary)	16675
10	II B.Sc-MPCs-EM	Mr. M.Joseph(Language Association Joint Secretary)	17500
11	III B.Sc-MPS-TM	Mr. B.Sumanth Naik(Sports&Games Association Secretary)	317008004
BSc Class Representatives			
Sno	**Class**	**Name of the Representative**	**Roll No**
1	III B.Sc-MSCs-EM	Mr. D.Venkata Rami Reddy	16612

2	III B.Sc-MPCs-EM	Mr. T.Lokesh	16534
3	III B.Sc-MPC-EM	Mr.M.Sandeep Reddy	0317008027
4	III B.Sc-MPC-TM	Mr. D.L.Viswanath	0317008020
5	III B.Sc-MPS-TM	Mr. M.Baby	0317008037
6	III B.Sc-BBC-EM	Mr.B.Lokesh Reddy	0317008715
7	III B.Sc-MZC-EM	V.Naveen Sahit	0317008589
8	III B.Sc-CT&HM-EM	Mr. J.Hemanth	16760
9	III B.Sc-CBZ-TM	Mr,Chandra sekhar reddy	0317008664
10	II B.Sc-MSCs-EM	Mr. P.Lokesh	17556
11	II B.Sc-MPCs-EM	Mr. M.Prasanth	17515
12	II B.Sc-MPC-EM	Mr. A.K.Jagadeesh	17465
13	II B.Sc-MPC-TM	K.Chandu	17457
14	II B.Sc-MPS-TM	Mr. A.Jayakumar	17598
15	II B.Sc-BBC-EM	Mr. B.Mahesh	17681
16	II B.Sc-MZC-EM	Mr. P.Devananda Goud Mr.G.Jaya Kumar	17719 0318008263
17	II B.Sc-CT&HM-EM	Mr Aravindhan.V	17768
18	II B.Sc-CBZ-TM	Mr. M.Venkatesh	17643
19	I B.Sc-MSCs-EM	Mr. G.Ramesh	18596
20	I B.Sc-MPCs-EM	Mr. Natteti Venkateswarulu	18560
21	I B.Sc-MPC-EM	SurendraGuvvala M.Raghu Ram Kumr	18462
22	I B.Sc-MPC-TM	Chandrasekhar.B	18405
23	I B.Sc-MPS-TM	Mr. B.Dastagiri	18476

24	IB.Sc-BBC-EM	M.Marappa Mr. T.Bharat Kumar	18652
25	I B.Sc-MZC-EM	Chakrapani Tharun M.Gurramma	18704
26	I B.Sc-CT&HM-EM	Mr. M.Naveen	18772
27	I B.Sc-CBZ-TM	K.Durga ManendraAvula	18628
BCom Class Representatives			
28	III B.Com-CA-EM	Mr. P.Seshari Naidu	0217008503
29	III B.Com-GEN-EM	Mr. D.Rajasekhar	217008582
30	III B.Com-GEN-TM	Mr A.Koteswar Rao	217008501
31	II B.Com-CA-EM	K.Deepa Mr. V.Harish	17176
32	II B.Com-GEN-EM	K.Ramesh KishoreKullanivagaira	17095
33	II B.Com-GEN-TM	B.Babu Naik Giri Babu.K	17010
34	I B.Com-CA-EM	Mr D.Mahanandi Mr. M.Venkatsuresh	18187
35	IB.Com-GEN-EM	Mr. Y.Sasi Kumar Mr. C.Jaswanth	18069
36	I B.Com-GEN-TM	Mr. Y.Sai Kumar G.Kranthi Kiran	18042

BA Class Representatives			
37	III B.A-HEP-EM	Mr. B.Sunil	16361
38	III B.A-HEP-TM	Mr. A.Jayaram	16307
39	III B.A-HPT-TM	Mr. M.Venkataramana	16289
40	III B.A-ASCA-EM	Bhanu Prakash.B	16223
41	III B.A-MES-EM	B.Ashok	116008524
42	II B.A-HEP-EM	Dhana SekharAllavala	17343
43	II B.A-HEP-TM	NaveenPoola	17302
44	II B.A-HPT-TM	Mr.S.Yerramala	117008019
45	II B.A-ASCA-EM	Mr.P.Giddaiah	117008506
46	II B.A-MES-EM	Mr.G.Venkata Ratnam	17408
47	I B.A-HEP-EM	Mr. Surya Mahesh Bala	0119008075
48	I B.A-HEP-TM	Mr. K.Venkatesh	0119008072
49	I B.A-HPT-TM	Mr.M.Eswar Naik	0119008093
50	I B.A-ASCA-EM	Mr.P.Nagendra	0119008015
51	I B.A-MES-EM	S.Lalith Madhav	0119008082

HOSTEL COMMITTEE			
Sno	**Committee Members**	**Designation**	**Role**
1	Dr. C.Ramesh	Principal	Chief Warden
2	Sri P.Muni Bhaskar Rao	HOD,Mathematics	Warden
3	Sri K. Rajesh	Lecturer,Economics	Dy.Warden
4	E.Anil Kumar	Student,III BSc-MPCs-EM	Store Secretary
5	B.Sumanth Naik	Student,III BSc-MPS-TM	Sports Secretary
MESS COMMITTEE			
Sno	**Name**	**Group**	**Role**
1	G.Hariprasad	III BSc-MPCs-EM	Associate Members
2	P.Bharath Naveen	III BBC-EM	
3	C.Sreenu	III BA-MES-EM	
4	N.Raghu	III Bcom-CA-EM	
5	Subramanyam	II Bcom-Gen-EM	
6	A.Venkatesu	II BSc-MSCs-EM	
7	N.Udaya Kumar Reddy	II BA-ASCA-EM	

STUDENT UNION OFFICE BEARERS:

Department of Chemistry

To foster leadership qualities among the students, the institution has constituted a Student Council and the toppers in the academics are selected as representatives of student Council. The following are the students of chemistry who represent the student council.

2020-2021

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Kum. M.Guramma	III BSc-MZC-EM	18712	UNION Chairmen
2	R.Gayathri	III BSc-MZC-EM		Life Sciences Association Secretary
3	A.Chandini	II BSc-BBC-EM		Life Sciences Association Joint Secretary

2019-2020

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Kum. Jahnavi Takkolu	III BSC-BBC-E.M	17672	Life Sciences Association Secretary
2	Kum. M.Guramma)	II BSc-MZC-EM	18712	Life Sciences Association Joint Secretary
3	Mr. D.Sandeep	III B.Sc-MZC-EM	17741	Sports&Games Association Secretary

2018-2019

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	Mr. S.Snehalatha	III B.Sc-BBC-EM	16736	Union Secretary)
2	Kum. T.Jahanavi	II B.Sc-BBC-EM	17672	Maths & Science Association Joint Secretary
3	Kum. P.Shilpa	III B.Sc-MZC-EM	16675	Language Association Secretary

2017-2018

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	G.Venkateswarulu	III B.Sc-MPC-EM		Maths & Science Association Joint Secretary
2	S. Snehalatha	II B.Sc-BBC-EM		Language Association Joint Secretary

2016-2017

SNO	Name of the student	Class	Roll Number	Designation in Student Council
1	G.Venkateswarulu	II B.Sc-MPC-EM		Maths & Science Association Joint Secretary